

Article

# CTAB-Assisted Fabrication of Bi<sub>2</sub>WO<sub>6</sub> Thin Nanoplates with High Adsorption and Enhanced Visible Light-Driven Photocatalytic Performance

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**Abstract:** Two-dimensional thin Bi<sub>2</sub>WO<sub>6</sub> nanoplates have been fabricated using a cetyltrimethylammonium bromide (CTAB)-assisted hydrothermal method. We investigated the proposed formation mechanism based on the crystalline structures of the thin Bi<sub>2</sub>WO<sub>6</sub> nanoplates. The high adsorption ability and excellent visible-light driven photocatalytic activities of the Bi<sub>2</sub>WO<sub>6</sub> nanoplates were illustrated, in view of exposed (001) facets of nanoplates possessing faster separation of photo-generated charge carriers and increased catalytically active sites. Such a cost-effective way to obtain Bi<sub>2</sub>WO<sub>6</sub> nanoplates offers new possibilities for the design of adsorptive semiconductor photocatalysts with strengthened photocatalytic activities.

**Keywords:** Bi<sub>2</sub>WO<sub>6</sub> nanoplates; semiconductor photocatalyst; visible light-driven

## 1. Introduction

The semiconductor photocatalysis technique is reported to be an efficient and environmentally friendly method for the treatment of organic environment pollutants [1]. Due to the fact visible light accounts for 43% of sunlight compared to 4% from UV light, visible light responsive semiconductor photocatalysts have become a major focus of water treatment research in order to realize the full use of solar power and conserve energy [2]. As one of the simplest members of the aurivillius oxide family of layered perovskites, bismuth tungstate (Bi<sub>2</sub>WO<sub>6</sub>), with its narrow bandgap (ca. 2.6–2.7 eV) has aroused significant attention in view of its good photocatalytic performance in the degradation of chloroform, acetaldehyde and organic dyes under visible light irradiation [3–7]. Considering the close dependence of the photocatalytic performance of Bi<sub>2</sub>WO<sub>6</sub> species on their morphology, Bi<sub>2</sub>WO<sub>6</sub> photocatalysts of various sizes and shapes have been widely studied (such as Bi<sub>2</sub>WO<sub>6</sub> nanoparticles [8], hierarchical Bi<sub>2</sub>WO<sub>6</sub> microspheres [9], self-assembled microdiscs [10], etc.). Particularly, increasing attention has been recently focused on the two-dimensional Bi<sub>2</sub>WO<sub>6</sub> micro- and nanostructures represented by mono-layered nanosheets [11] and ultrathin nanoplates [12], in view of their increased surface-to-volume ratio, faster separation of photogenerated electrons and holes with plenty of catalytically active sites and excellent structural stability [13], which are attractive and important features for improving the photocatalytic activities of Bi<sub>2</sub>WO<sub>6</sub> under visible light irradiation.

Herein, two-dimensional thin Bi<sub>2</sub>WO<sub>6</sub> nanoplates with round edges were firstly obtained using a simple hydrothermal reaction system employing Bi(NO<sub>3</sub>)<sub>3</sub>·5H<sub>2</sub>O and Na<sub>2</sub>WO<sub>4</sub>·2H<sub>2</sub>O as raw materials and cetyltrimethylammonium bromide (CTAB) as surfactant. High adsorption and enhanced visible light-driven photocatalytic performance of these Bi<sub>2</sub>WO<sub>6</sub> thin nanoplates for the decomposition of rhodamine B (RhB) aqueous solution was demonstrated in detail.

## 2. Materials and Methods

### 2.1. Synthesis

$\text{Bi}(\text{NO}_3)_3 \cdot 5\text{H}_2\text{O}$ ,  $\text{Na}_2\text{WO}_4 \cdot 2\text{H}_2\text{O}$  and CTAB chemical reagents used in this work were of analytical grade, bought from Sinopharm Chemical Reagent Company (Shanghai, China) and were used without further purification. In a typical procedure for the synthesis of thin  $\text{Bi}_2\text{WO}_6$  nanoplates, 2 mmol  $\text{Bi}(\text{NO}_3)_3 \cdot 5\text{H}_2\text{O}$ , 0.5 g CTAB, and 1 mmol  $\text{Na}_2\text{WO}_4 \cdot 2\text{H}_2\text{O}$  were respectively dissolved in 10 mL distilled water. Firstly, 10 mL CTAB aqueous solution was added dropwise to 10 mL  $\text{Bi}(\text{NO}_3)_3 \cdot 5\text{H}_2\text{O}$  aqueous solution and stirred for 30 min. Then 10 mL  $\text{Na}_2\text{WO}_4 \cdot 2\text{H}_2\text{O}$  aqueous solution was added to the above solution and the mixture solution was stirred for 1 h at room temperature until a precursor suspension was obtained. Finally, the resulting solution was transferred to a 50 mL Teflon-line autoclave, and maintained at 180 °C for 20 h. After the autoclave cooled down to room temperature naturally, the prepared samples were collected by centrifugation, washed with distilled water and absolute ethanol for several times, and then dried at 80 °C in a drying oven for 6 h. Commercial nano- $\text{TiO}_2$  with size of 5–10 nm and phenol were purchased from Shanghai Aladdin Biochem Technology Company (Shanghai, China).

### 2.2. Characterization

The obtained products were characterized by X-ray diffraction (XRD, Philips X' Pert Pro Super, Amsterdam, The Netherlands), field emission scanning electron microscopy (FESEM, Hitachi S-4800, Tokyo, Japan), transmission electron microscope (TEM, Phillips Tecnai-12, Amsterdam, The Netherlands), high-resolution transmission electron microscopy (HRTEM, FEI Tecnai G2 F30 S-TWIN, Hillsboro, OR, USA) and energy-dispersive X-ray (EDX) analysis was obtained with an EDAX detector installed on the same HRTEM. UV-visible diffuse reflectance spectra (DRS) were recorded on a UV-vis diffuse reflectance spectrum (Varian Cary 5000, Palo Alto, CA, USA). The Brunauer-Emmett-Teller (BET) surface area was measured using a Micrometrics ASAP 2020 (Micrometrics, Norcross, GA, USA).

### 2.3. Photocatalytic Activity Test

The photocatalytic activities of the  $\text{Bi}_2\text{WO}_6$  samples were evaluated by the degradation of rhodamine B (RhB) or phenol under a simulated sunlight irradiation using a 300 W Xe lamp with a 420 nm cut-off filter as a light source. In each experiment, 50 mg of photocatalyst was added into 100 mL RhB of solution ( $1 \times 10^{-5}$  mol/L) or phenol solution (15 mg/L). Prior to visible light illumination, the suspensions were magnetically stirred in the dark for 60 min to reach adsorption-desorption equilibrium between thin  $\text{Bi}_2\text{WO}_6$  nanoplates and RhB aqueous solution, after that the solution was exposed to visible light irradiation with non-stop stirring. During visible light irradiation, about 4 mL of suspension was removed from the reactor at certain time intervals and centrifuged to get stop the deposition on the  $\text{Bi}_2\text{WO}_6$  nanoplates photocatalyst. The filtrates were analyzed by recording changes of the absorption band maximum (553 nm) in the UV-vis spectra of RhB by using a spectrophotometer (UV-2700, Shimadzu, Suzhou, China). The photocatalytic efficiency was calculated according to the following equation:

$$\text{Removal efficiency (\%)} = \frac{C_0 - C}{C_0} \times 100\%$$

where  $C_0$  and  $C$  represent the original concentration of RhB and phenol in dark, and the reaction concentration of RhB, phenol after a certain time visible-light irradiation, separately.

## 3. Results and Discussion

### 3.1. Structure and Morphology

Two-dimensional  $\text{Bi}_2\text{WO}_6$  thin nanoplates was firstly obtained by a CTAB-assisted hydrothermal method under 180 °C for 20 h. The XRD pattern of the  $\text{Bi}_2\text{WO}_6$  thin nanoplate product is shown

in Figure 1 where eight major sharp reflection peaks located at  $2\theta = 28.21^\circ$ ,  $32.80^\circ$ ,  $47.06^\circ$ ,  $56.05^\circ$ ,  $58.62^\circ$ ,  $68.84^\circ$ ,  $76.24^\circ$ ,  $78.50^\circ$  showed the well-crystallized orthorhombic phase of  $\text{Bi}_2\text{WO}_6$  with lattice parameters of  $a = 5.457 \text{ \AA}$ ,  $b = 5.436 \text{ \AA}$ , and  $c = 16.427 \text{ \AA}$ , which matches very well with the standard values (Powder Diffraction File No: 73-1126). No peaks of any other phases or impurities were detected in this pattern. The intensity ratio of the (200) or (020) peak to the (113) peak is 0.80, evidently higher than the standard value of 0.185, while the full width at half maximum (FWHM) of the (200) or (020) Bragg peak is narrower than that of the (113) peak and indicates a higher grain size of  $\text{Bi}_2\text{WO}_6$  along the (100) and (010) directions compared to the (001) direction. Therefore both the above analysis indicate that  $\text{Bi}_2\text{WO}_6$  nanoplates with special anisotropic features will grow along the (001) plane [12,14].

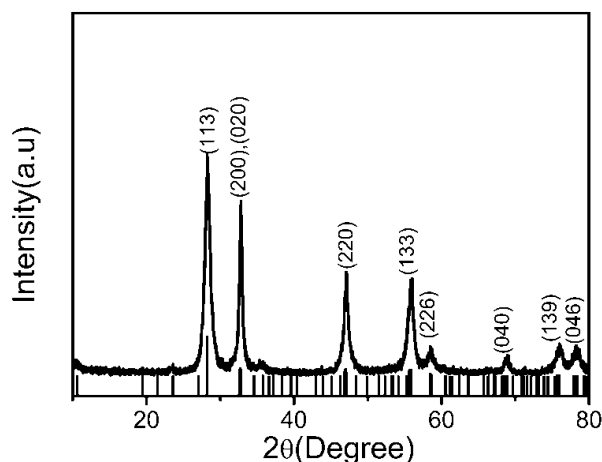


Figure 1. XRD pattern of  $\text{Bi}_2\text{WO}_6$  thin nanoplates prepared at  $180^\circ\text{C}$  for 20 h.

An X-ray photoelectron spectroscopy (XPS) analysis on our obtained two-dimensional thin  $\text{Bi}_2\text{WO}_6$  nanoplates was performed in order to study the chemical states and surface composition of the elements existing in the  $\text{Bi}_2\text{WO}_6$  thin nanoplates (Figure 2). Figure 2a is the XPS full survey spectra on  $\text{Bi}_2\text{WO}_6$  thin nanoplates, where we can see the existence of Bi, W and O elements in  $\text{Bi}_2\text{WO}_6$  nanoplates (C element belongs to the binder), and a tiny peak at 68.38 eV can be attributed to the presence of Br element [15] (the EDS pattern in Figure S1 also indicates the existence of Br element), which will possibly increase the adsorption and photocatalytic activity of  $\text{Bi}_2\text{WO}_6$  nanoplates according to the reported literature results [15].

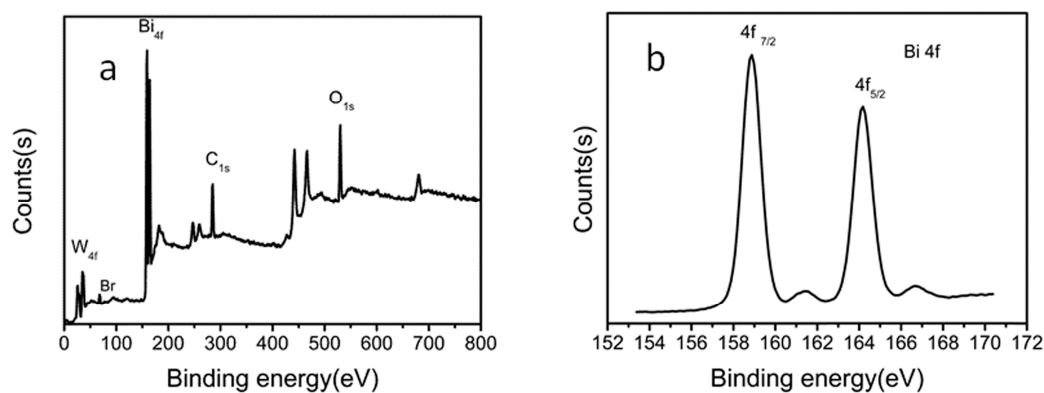
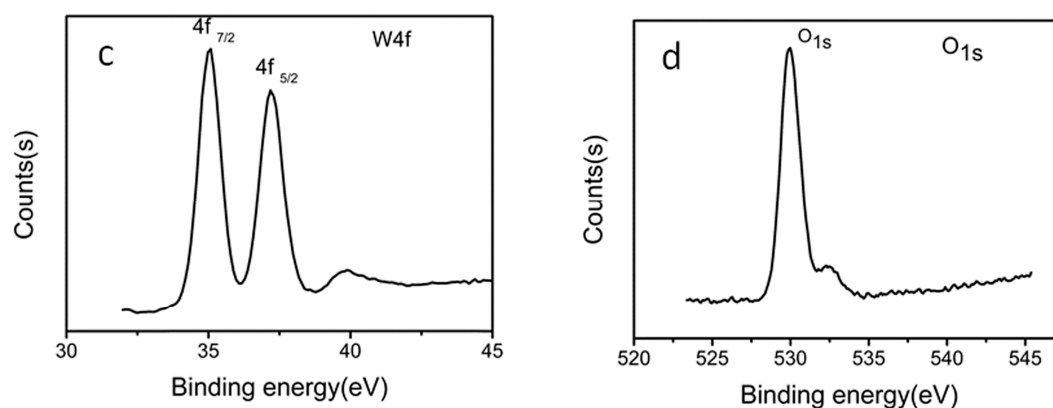


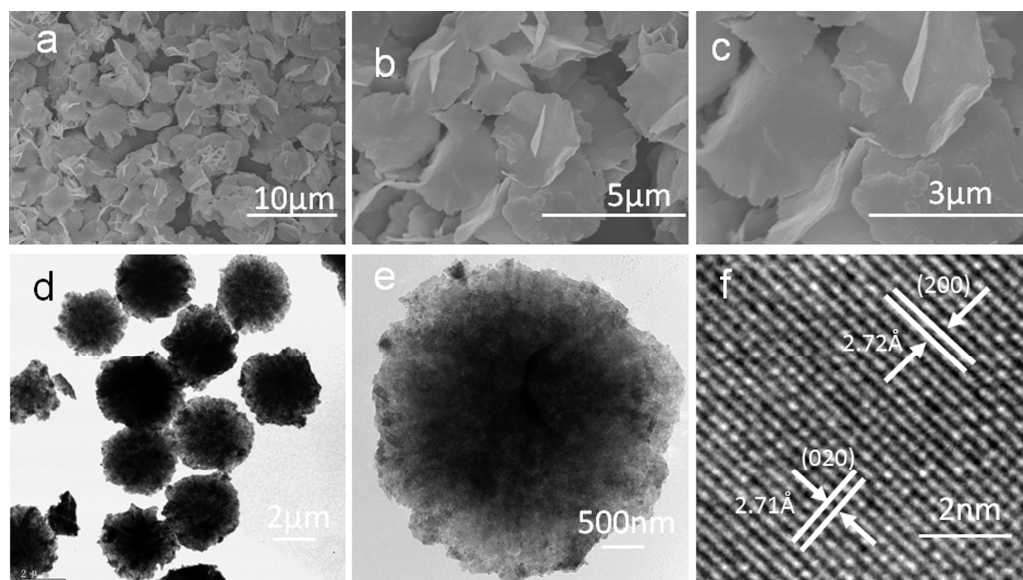
Figure 2. Cont.



**Figure 2.** XPS spectra of the  $\text{Bi}_2\text{WO}_6$  nanoplates: (a) the typical survey, the high-resolution spectra of (b)  $\text{Bi}_{4f}$ ; (c)  $\text{W}_{4f}$  and (d)  $\text{O}_{1s}$ .

The high-resolution spectra survey of separate Bi, W and O elements were displayed respectively as follows: two peaks centered at 158.8 eV and 164.2 eV (Figure 2b) could be attributed to the relevant  $4f_{7/2}$  and  $4f_{5/2}$  of  $\text{Bi}^{3+}$  element, the binding energies of 35.1 eV and 37.2 eV (Figure 2c) are separately assigned to  $4f_{7/2}$  and  $4f_{5/2}$  from  $\text{W}^{6+}$  element and the non-symmetric  $\text{O}_{1s}$  peak located at 529.8 eV should be due to the contributions from crystal lattice oxygen and adsorbed oxygen existent in the oxides (Figure 2d) [11,16].

The formation and morphologies of two-dimensional  $\text{Bi}_2\text{WO}_6$  thin nanoplates can be demonstrated by their FESEM and TEM images. Figure 3a shows a panoramic FESEM image of a representative  $\text{Bi}_2\text{WO}_6$  sample made up of abundant dispersive thin nanoplates with sizes of ca. 3.0  $\mu\text{m}$ . The higher FESEM magnification images in Figure 3b,c clearly reveal  $\text{Bi}_2\text{WO}_6$  nanoplates displaying thin and two-dimensional microstructures, from which the thickness of the  $\text{Bi}_2\text{WO}_6$  nanoplates was measured as around 60 nm.



**Figure 3.** (a) General and (b,c) higher magnification SEM images of  $\text{Bi}_2\text{WO}_6$  thin nanoplates; (d,e) TEM images of  $\text{Bi}_2\text{WO}_6$  thin nanoplates; (f) HRTEM image taken on a certain part of  $\text{Bi}_2\text{WO}_6$  nanoplate.

Additionally, Panels (d) and (e) of Figure 3 show the bright-field TEM images of  $\text{Bi}_2\text{WO}_6$  nanoplates, which further indicate the two-dimensional microstructures of the  $\text{Bi}_2\text{WO}_6$  nanoplates.

Panel (f) of Figure 3 shows the enlarged lattice-resolved HRTEM image from a certain part of a single  $\text{Bi}_2\text{WO}_6$  nanoplate, where the spacing of the observed lattice plane is approximately 0.272 nm and 0.271 nm, consistent with the spacing for the (200) and (020) planes of orthorhombic  $\text{Bi}_2\text{WO}_6$ . This indicated that the growth orientation of two-dimensional  $\text{Bi}_2\text{WO}_6$  nanoplates is preferentially along the (001) basal plane in accordance with the XRD analysis (Figure 1) and we may speculate that the two-dimensional  $\text{Bi}_2\text{WO}_6$  nanoplates with more exposed (001) planes will show higher visible light-driven photocatalytic activities than  $\text{Bi}_2\text{WO}_6$  microstructures with other exposed crystal planes [11,17,18]. The formation mechanism of our prepared two-dimensional  $\text{Bi}_2\text{WO}_6$  nanoplates in the present reaction system could be attributed to the CTAB-assisted hydrothermal fabrication method used, as CTAB surfactant is a widely used cationic ligand in colloidal chemistry and usually plays an important role as a soft template for the shape and morphology control of semiconductor compounds [15,19,20]. Here in the as-obtained two-dimensional thin  $\text{Bi}_2\text{WO}_6$  nanoplates, CTAB makes a big difference in controlling the shape and exposed facets of  $\text{Bi}_2\text{WO}_6$  microstructures and only three-dimensional hierarchical microspheres were obtained in the absence of CTAB while keeping other fabrication conditions unchanged (Figure S2), therefore combined with the simulated  $\text{Bi}_2\text{WO}_6$  crystalline structures (numerous oxygen atoms existing in oxygen bridge bonds on the surface of (001) facets, Figure 4),  $\text{CTA}^+$  cations could selectively adsorb on the surfaces of (001) facets of  $\text{Bi}_2\text{WO}_6$  crystal reducing their surface energy. In view of the fact that the lattice planes with higher surface energy will eliminate by faster growth than lattice planes with lower surface energy [14,21,22] and the surfaces (001) facets was probably lower than that of other facets, thus unique two-dimensional  $\text{Bi}_2\text{WO}_6$  thin nanoplates with more exposed (001) facets caused by the preferential overgrowth were fabricated in this reaction solution.

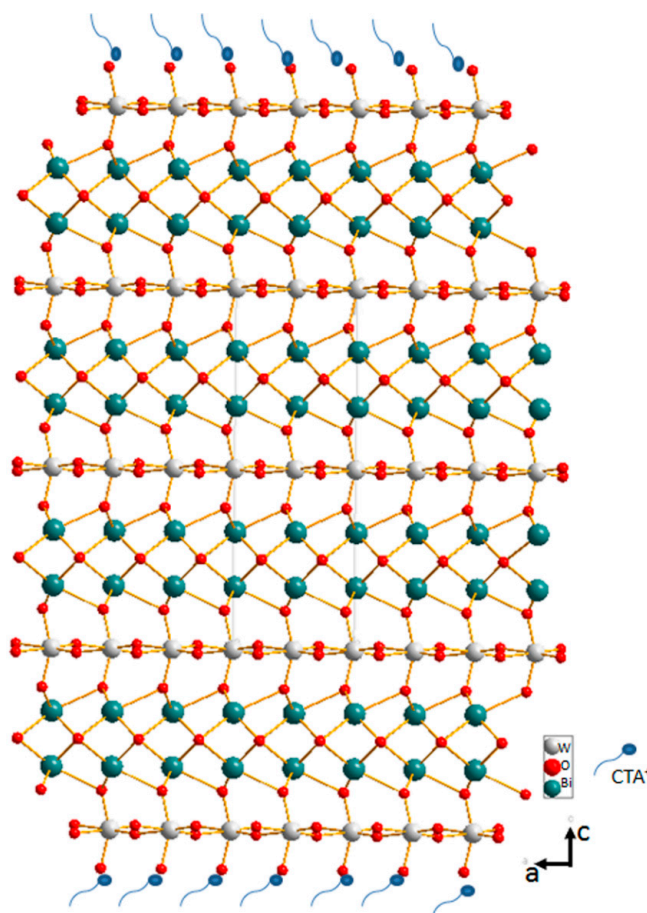
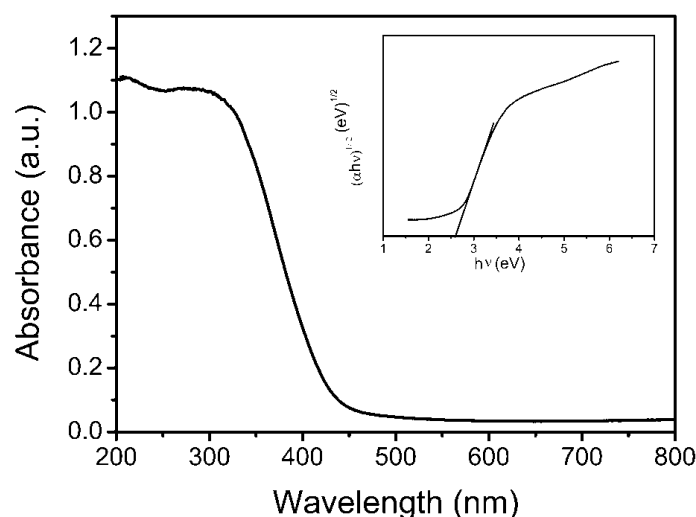


Figure 4. Structural model of the  $\text{Bi}_2\text{WO}_6$  nanoplates.

### 3.2. UV-Vis DRS Analysis

The optical absorption property of fabricated two-dimensional thin  $\text{Bi}_2\text{WO}_6$  nanoplates was investigated by the UV-visible diffuse reflectance spectroscopy technique (Figure 5), which indicated that thin two-dimensional  $\text{Bi}_2\text{WO}_6$  nanoplates exhibited strong photo-absorption in visible light region with a steep absorption edge of 451 nm. We can infer that the visible light absorption was caused by intrinsic band-gap transitions but not of transitions from the impurity level [23]. As regard to a crystalline semiconductor, the optical absorption near the band edge is in step with the equation  $\alpha h\nu = A (h\nu - E_g)^n$  for a direct band gap material, where  $\alpha$ ,  $h$ ,  $\nu$ ,  $E_g$  and  $A$  are respectively the absorption coefficient, Planck constant, light frequency, band gap and a constant, while  $n$  resolved the feature of the transition in a semiconductor and the  $n$  value of two-dimensional thin  $\text{Bi}_2\text{WO}_6$  nanoplates equals to 2 [24]. The optical band gap ( $E_g$ ) acquired by extrapolation of the plot (the inset of Figure 5) of  $(\alpha h\nu)^{1/2}$  versus  $h\nu$  is around 2.62 eV, close to that of the previously reported  $\text{Bi}_2\text{WO}_6$  microstructures in literature (2.64 eV) [25]. The color of the obtained thin  $\text{Bi}_2\text{WO}_6$  nanoplates was light-yellow, as can also be expected from the absorption spectrum. On the basis of the above UV-Vis DRS analysis results, we may infer that the as-obtained thin  $\text{Bi}_2\text{WO}_6$  nanoplates should possess the capability for visible light-driven photocatalytic degradation of organic pollutants in aqueous solution in the water treatment field.

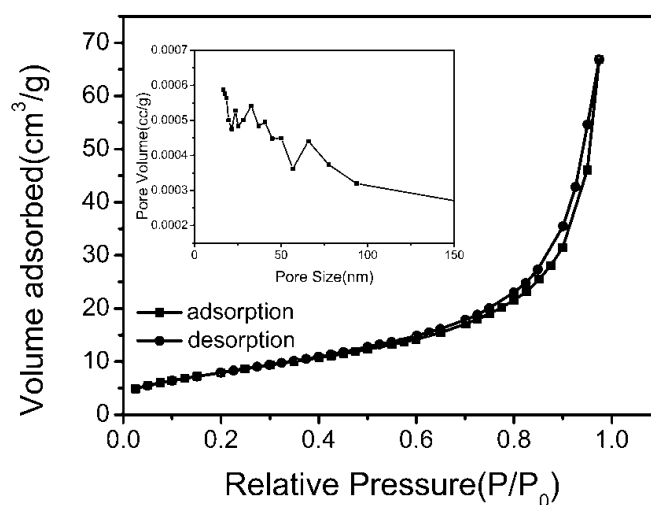


**Figure 5.** UV-vis diffuse reflectance spectra of  $\text{Bi}_2\text{WO}_6$  nanoplates. The insert is the corresponding  $(\alpha h\nu)^{1/2}$  versus photon energy plots.

### 3.3. BET Surface Area Analysis

The Brunauer-Emmett-Teller (BET) specific surface area and the pore size distribution (PSD) of the fabricated two-dimensional thin  $\text{Bi}_2\text{WO}_6$  nanoplates were studied by the  $\text{N}_2$  adsorption and desorption isotherms and the corresponding PSD curves, as shown in Figure 6. The adsorption—desorption isotherm for  $\text{Bi}_2\text{WO}_6$  nanoplates is of type IV, implying the presence of mesopores (size of 2–50 nm) [26], while the examined hysteresis loop extends to a higher relative pressure with value of  $P/P_0$  around 1 indicating the presence of macropores (size larger than 50 nm) [27]. The pore size distribution curve (inset in Figure 6) of the two-dimensional thin  $\text{Bi}_2\text{WO}_6$  nanoplates has main peaks at 23, 33 and 66 nm, so we can infer that mesoporous structures (23 nm–50 nm) can be assigned to tiny pores in individual thin  $\text{Bi}_2\text{WO}_6$  nanoplates and macropores (larger than 66 nm) are mainly caused by aggregation of  $\text{Bi}_2\text{WO}_6$  nanoplates associated with the above FESEM images shown in Figure 3. The BET surface area of two-dimensional thin  $\text{Bi}_2\text{WO}_6$  nanoplates was calculated to be  $29.76 \text{ m}^2/\text{g}$  from the  $\text{N}_2$  adsorption and desorption isotherms, higher than that of typical SSR- $\text{Bi}_2\text{WO}_6$  ( $0.6 \text{ m}^2/\text{g}$ ) [28]. Considering the

larger BET surface area may be effective for efficient visible light harvesting, the visible light-driven photocatalytic activity of  $\text{Bi}_2\text{WO}_6$  two-dimensional thin  $\text{Bi}_2\text{WO}_6$  nanoplates will be increased.

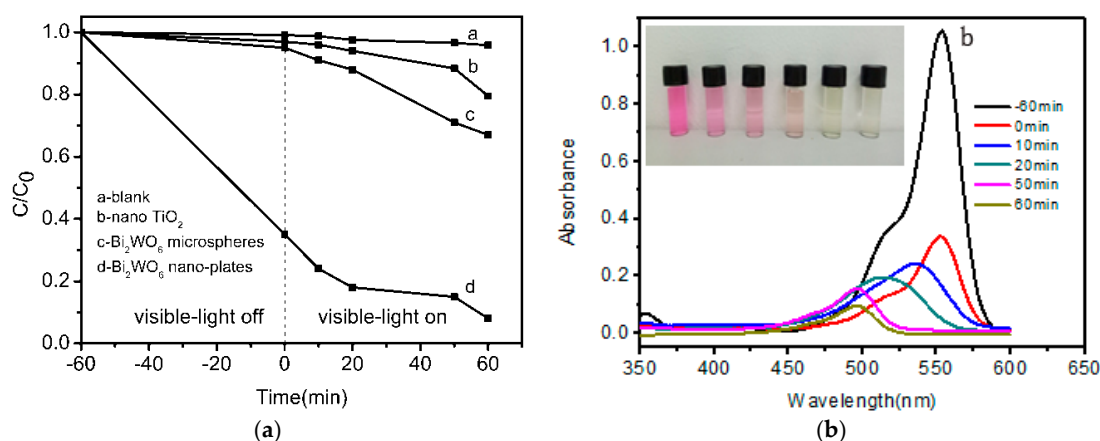


**Figure 6.**  $\text{N}_2$  adsorption and desorption isotherms and pore size distribution curve (insert) for  $\text{Bi}_2\text{WO}_6$  nanoplates.

### 3.4. Photocatalytic Activity

It was considered that a high adsorption capability of the organic pollutants on the surface of a semiconductor photocatalyst can not only strengthen the photocatalytic activity, but also expand the practical large-scale applicability for efficiently removing hazardous pollutants in the dark or under weak light irradiation, where the photocatalyst figures as a collector to gather the pollutants in aqueous solution [29,30]. Hence semiconductor photocatalysts with enhanced adsorption capability and photocatalytic activity are desirable for high removal of organic pollutants in the water treatment field.

Here the photocatalytic activities of the presently obtained two-dimensional thin  $\text{Bi}_2\text{WO}_6$  nanoplates was evaluated by using the organic pollutant RhB as target molecule in aqueous solution under visible light irradiation (Figure 7a).



**Figure 7.** (a) Comparison of photocatalytic activities on degradation of RhB from  $\text{Bi}_2\text{WO}_6$  nanoplates, microspheres and nano  $\text{TiO}_2$ ; (b) The temporal evolution of the absorption spectra of the RhB solution under visible-light irradiation in the presence of 50 mg  $\text{Bi}_2\text{WO}_6$  nanoplates, insert is the color changes of the RhB aqueous solution.

The blank test indicated that the photo-degradation efficiency was only 4% within 60 min, thus the RhB degradation is particularly slow without Bi<sub>2</sub>WO<sub>6</sub> nanoplates under visible light irradiation. Fortunately we found that the two-dimensional thin Bi<sub>2</sub>WO<sub>6</sub> nanoplates displayed much higher adsorption abilities and more enhanced visible light-driven photocatalytic performance than that of nano-TiO<sub>2</sub> and of Bi<sub>2</sub>WO<sub>6</sub> microspheres precipitated in the absence of CTAB in reaction solution, whereby 65% of the RhB organic pollutant molecules were adsorbed on the surface of the thin Bi<sub>2</sub>WO<sub>6</sub> nanoplates after 60 min absolute adsorption–desorption equilibrium without any light irradiation, and after that the RhB organic dye was almost completely degraded after 60 min of visible light illumination. An influence of the initial concentration of RhB aqueous solution on the photo-degradation performance over Bi<sub>2</sub>WO<sub>6</sub> nanoplates was observed (Figure S3), as when the initial concentration was lowered to  $2 \times 10^{-6}$  mol/L, the adsorption of RhB in the dark was more than 70% and the remaining RhB was fully degraded within only 10 min, but the degradation performance was weakened when the initial concentration of RhB was increased to  $2.5 \times 10^{-5}$  mol/L, therefore the photo-degradation performance of Bi<sub>2</sub>WO<sub>6</sub> nanoplates differed by changing the concentration of RhB in aqueous solution. When the amount of Bi<sub>2</sub>WO<sub>6</sub> nanoplate photocatalyst was decreased to 20 mg, both the adsorption and the photocatalytic performance were weaker than that of 50 mg Bi<sub>2</sub>WO<sub>6</sub> (Figure S4). The enhanced visible light-driven photocatalytic performance of two-dimensional thin Bi<sub>2</sub>WO<sub>6</sub> nanoplates can be attributed to their more exposed (001) facets [18,31], which were beneficial for promoting charge transfer and further highlighted the photocatalytic activities of Bi<sub>2</sub>WO<sub>6</sub> nanoplates under visible light irradiation. Figure 7b exhibits the temporal evolution of the absorption spectra of RhB aqueous solution degraded by Bi<sub>2</sub>WO<sub>6</sub> nanoplates under visible light irradiation. It is clearly observed that the intensity of the major absorption peaks from RhB decreased gradually with extended time, and the absorption peaks disappeared almost completely when the irradiation time reached 60 min, which illustrates a nearly complete degradation of RhB, consistent with the colour changes of the suspension from initial pink to transparent (insert of Figure 7b). It is noteworthy that two-dimensional thin Bi<sub>2</sub>WO<sub>6</sub> nanoplates with exposed (001) facets could provide more reactive sites for sunlight harvesting and the adsorption, degradation of organic dye molecules as well, thus the visible light-driven photocatalytic efficiency of Bi<sub>2</sub>WO<sub>6</sub> nanoplates was enhanced. In order to further demonstrate the visible light-driven photocatalytic activity of Bi<sub>2</sub>WO<sub>6</sub> nanoplates, the photo-degradation of colorless phenol as a substrate was performed [32], Figure S5 shows that the photo-degradation efficiency of phenol was 39% after 6 h, which further demonstrates the visible light-driven photocatalytic capability from Bi<sub>2</sub>WO<sub>6</sub> nanoplates on organic pollutant aqueous solution degradation.

#### 4. Conclusions

In summary, two-dimensional Bi<sub>2</sub>WO<sub>6</sub> thin nanoplates were prepared in the presence of CTAB using a hydrothermal route as a semiconductor photocatalyst with high adsorption and visible light-driven photocatalytic performance efficient on removal of environmental pollutants in the water treatment field. The proposed formation mechanism and visible light-driven photocatalytic activity of thin Bi<sub>2</sub>WO<sub>6</sub> nanoplates were investigated. In view of the fact more exposed (001) facets of Bi<sub>2</sub>WO<sub>6</sub> nanoplates with a typical two-dimensional nanoplate microstructure may lead to more detached photogenerated charge carriers, and the active sites on the catalyst surface will grow in number, therefore the visible light-driven photocatalytic activity of Bi<sub>2</sub>WO<sub>6</sub> thin nanoplates on degradation of organic pollutants will be greatly enhanced, indicating possible application of the as-prepared two-dimensional Bi<sub>2</sub>WO<sub>6</sub> thin nanoplates in the water treatment field.

**Supplementary Materials:** The following are available online: Figure S1. EDS pattern of Bi<sub>2</sub>WO<sub>6</sub> thin nanoplates prepared at 180°C for 20h; Figure S2. SEM images of Bi<sub>2</sub>WO<sub>6</sub> microspheres, 180°C, 20h; Figure S3. Comparison of photocatalytic activities on degradation of different concentration of RhB from Bi<sub>2</sub>WO<sub>6</sub> nanoplates (a)  $2.5 \times 10^{-5}$ , (b)  $1 \times 10^{-5}$ , (c)  $2 \times 10^{-6}$ ; Figure S4. Comparison of photocatalytic activities on degradation of RhB ( $1 \times 10^{-5}$ ) from different amounts of Bi<sub>2</sub>WO<sub>6</sub> nanoplates (a) 20mg, (b) 50mg; Figure S5. Photocatalytic activity on degradation of color-less phenol (15 mg/L) of Bi<sub>2</sub>WO<sub>6</sub> nanoplates.



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**Author Contributions:** Yuxue Zhou designed the experiments and wrote the paper; Pengfei Lv, Yanping Tang and Pingping Huang performed the experiments; Yuxue Zhou and Xiangdong Meng analyzed the data; Xiaobing Chen, Xiaoshuang Shen and Xianghua Zeng contributed materials and analysis tools.

**Conflicts of Interest:** The authors declare no conflict of interest.

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**Sample Availability:** Not available.



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