

A Multifunctional Coating on Sulfur-containing Carbon-based Anode for High-performance Sodium-ion Batteries

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Table S1. Summary of some different carbon-based materials for SIBs

Number	Material	Reversible capacity (mAh g ⁻¹ , at A g ⁻¹)	ICE (%)	Reference
1	G-NCs	329 (0.05 A g ⁻¹)	32	Ref. [51]
2	Natural graphite	127 (0.1 A g ⁻¹)	-	Ref.[52]
3	A-CG	280 (0.04 A g ⁻¹)	38	Ref. [53]
4	HPCCNFs-1	215 (0.1 A g ⁻¹)	25.3	Ref.[15]
5	P-CNSs	237.5 (0.05 A g ⁻¹)	42.65	Ref.[37]
6	HCS-1900	295 (0.05 A g ⁻¹)	68.8	Ref. [54]
7	SNMHCSs	240 (0.05 A g ⁻¹)	29	Ref. [55]
8	PC-800	263.2 (0.05 A g ⁻¹)	79.6	Ref. [56]

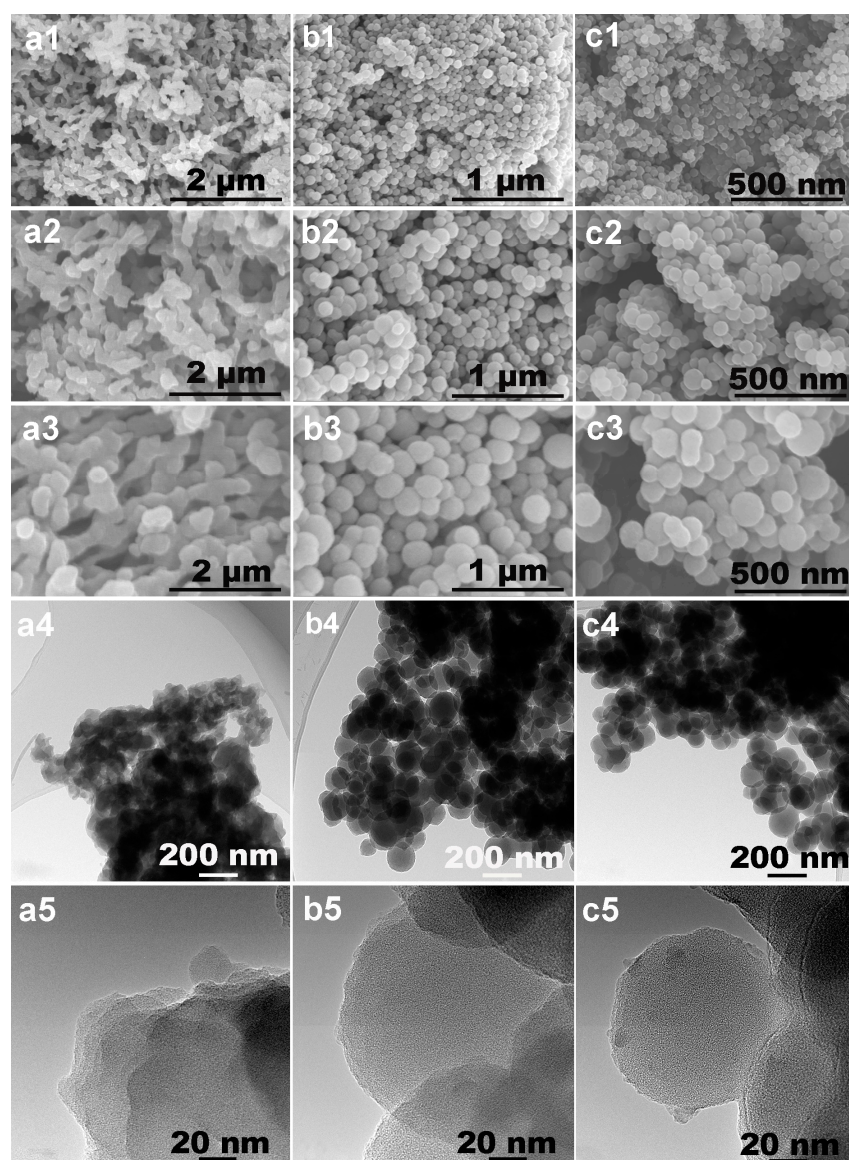


Figure S1. The SEM images of **a1, a2, a3**) NSC, **b1, b2, b3**) SGCS and **c1, c2, c3**) GCS at different magnifications. TEM and HRTEM images of **a4, a5**) NSC, **b4, b5**) SGCS and **c4, c5**) GCS.

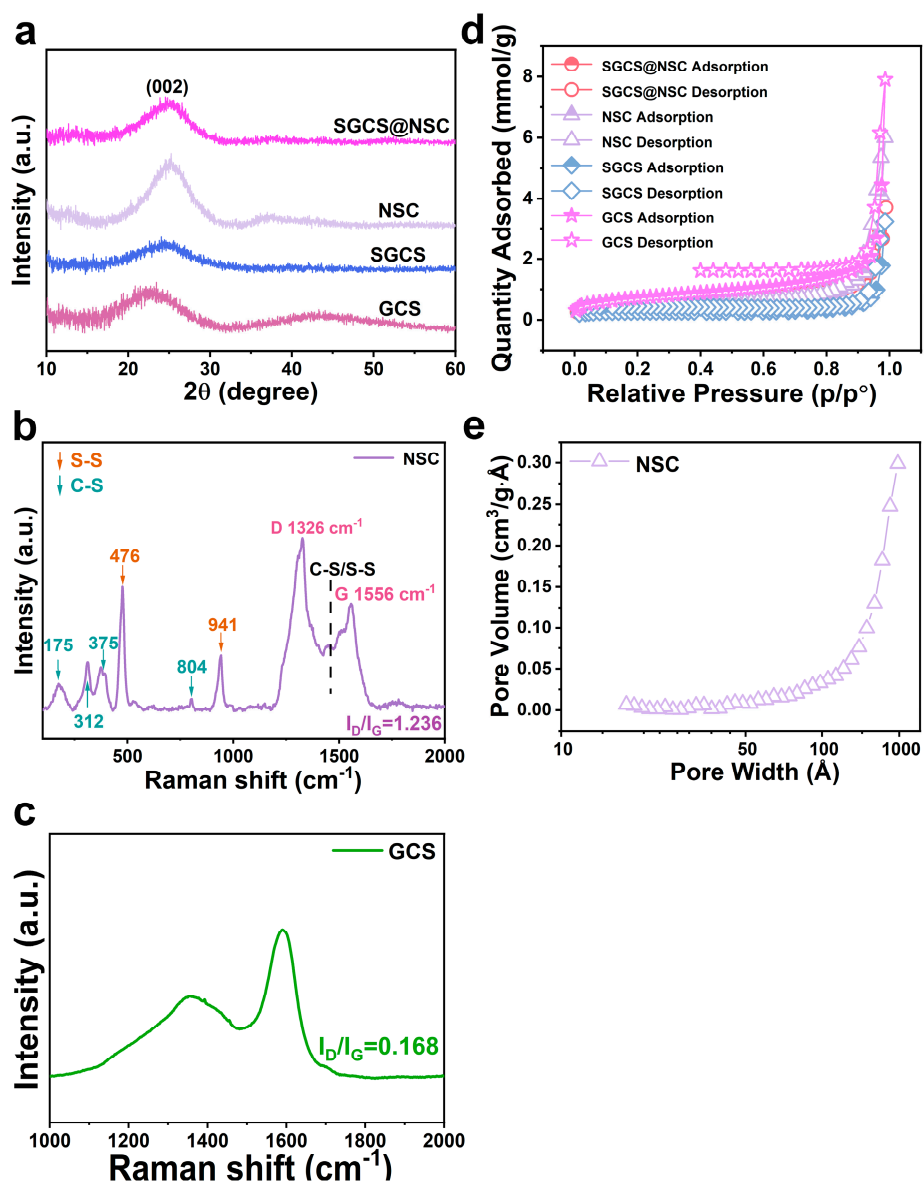


Figure S2. (a) X-ray diffraction patterns of SGCS@NSC, NSC, SGCS, and GCS. (b, c) Raman scattering spectra of NSC and GCS. (d) N_2 adsorption-desorption isotherms of SGCS@NSC, NSC, SGCS, and GCS. (e) Pore distribution curves of NSC.

Table S2. Analysis of sulfur bonding area ratio of SGCS@NSC and SGCS.

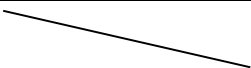
Sulfur bonding	SGCS@NSC	SGCS
HS _x C-	8.1	
C-S	70.3	
S-S	11.3	
Sulfate	10.3	

Table S3. Analysis of carbon bonding area ratio of SGCS@NSC and SGCS.

Carbon bonding	SGCS@NSC	SGCS
C-C	35.9	36.3
C-N/C-S	18.0	12.5
C=O/C=N	46.0	51.2

Table S4. Specific surface areas and pore volume analysis of SGCS@NSC, SGCS, GCS, and NSC.

Samples	$S_{\text{BET}}(\text{m}^2/\text{g})$	$V_{\text{pore}}(\text{cm}^3/\text{g})$	$R_{\text{pore}}(\text{nm})$
SGCS@NSC	34.956	0.128	3.683
SGCS	18.670	0.113	6.052
GCS	56.147	0.274	4.880
NSC	31.168	0.208	8.984

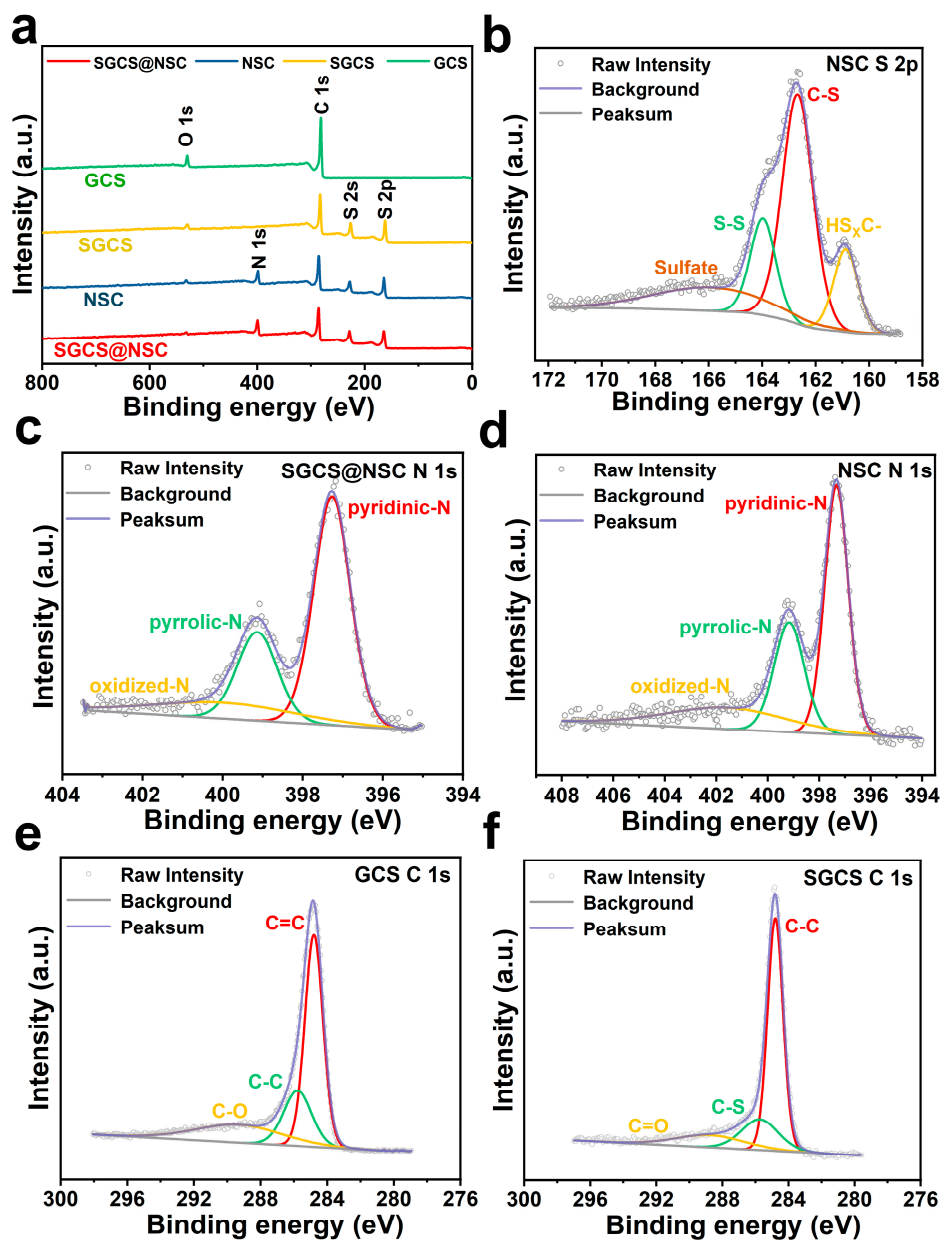


Figure S3. (a) Typical XPS survey spectra. (b, d, f) The high-resolution XPS of S 2p, N 1s, and C 1s for NSC, respectively. (c) The high-resolution XPS of N 1s for SGCS@NSC. (e) The high-resolution XPS of C 1s for GCS

Table S5. Element composition and ratio of the surface of four groups of samples calculated from XPS data (wt.%).

Samples	C (at%)	N (at%)	S (at%)	O (at%)
SGCS@NSC	63.019	12.637	19.075	2.471
NSC	61.019	16.542	19.860	1.813
SGCS	71.471		22.083	5.446
GCS	92.808			7.192

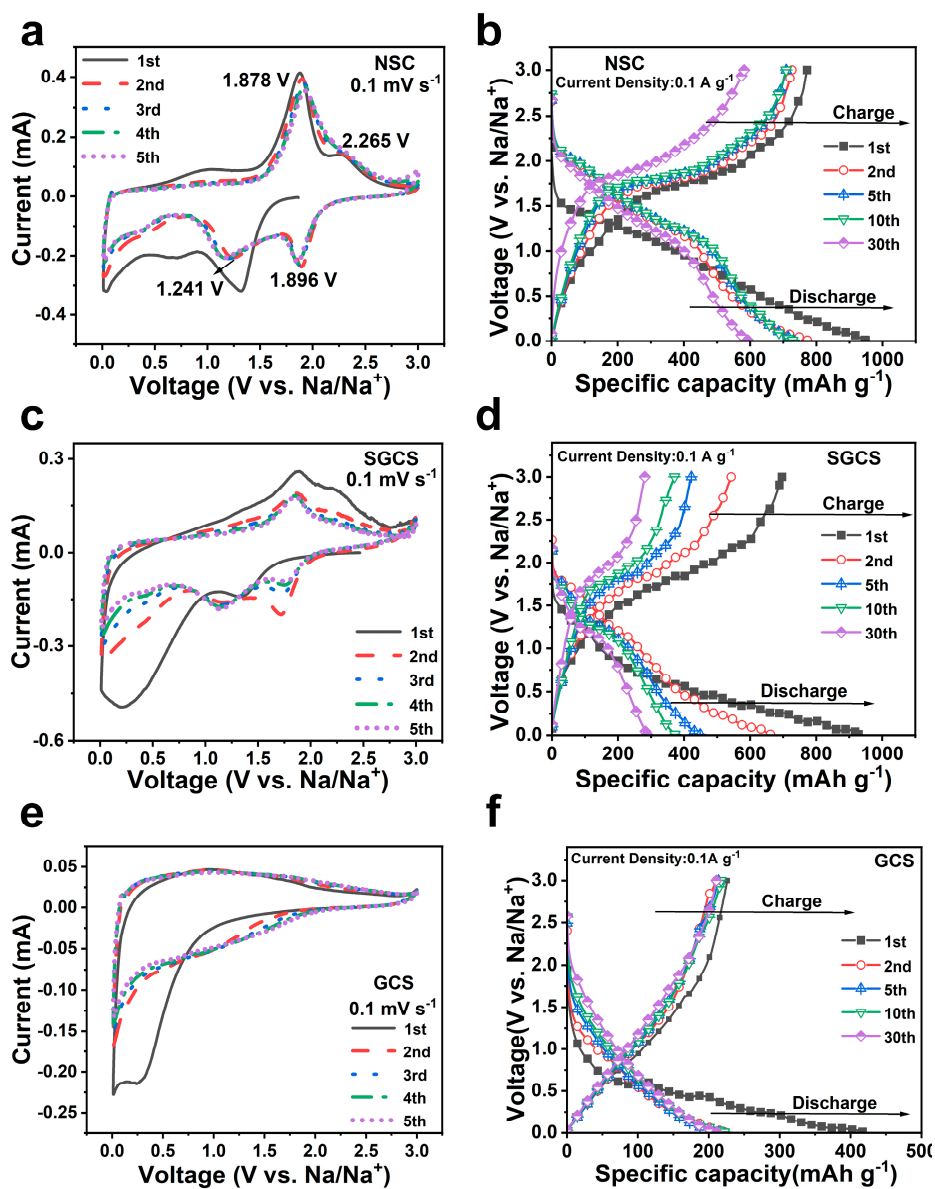


Figure S4. The first five successive CV curves of a) NSC, c) SGCS, and e) GCS at the scan rate of 0.1 mV s^{-1} . Galvanostatic discharge-charge profiles (100 mA g^{-1}) at the 1st, 2nd, 5th, 10th, and 30th cycles of b) NSC, d) SGCS, and f) GCS.

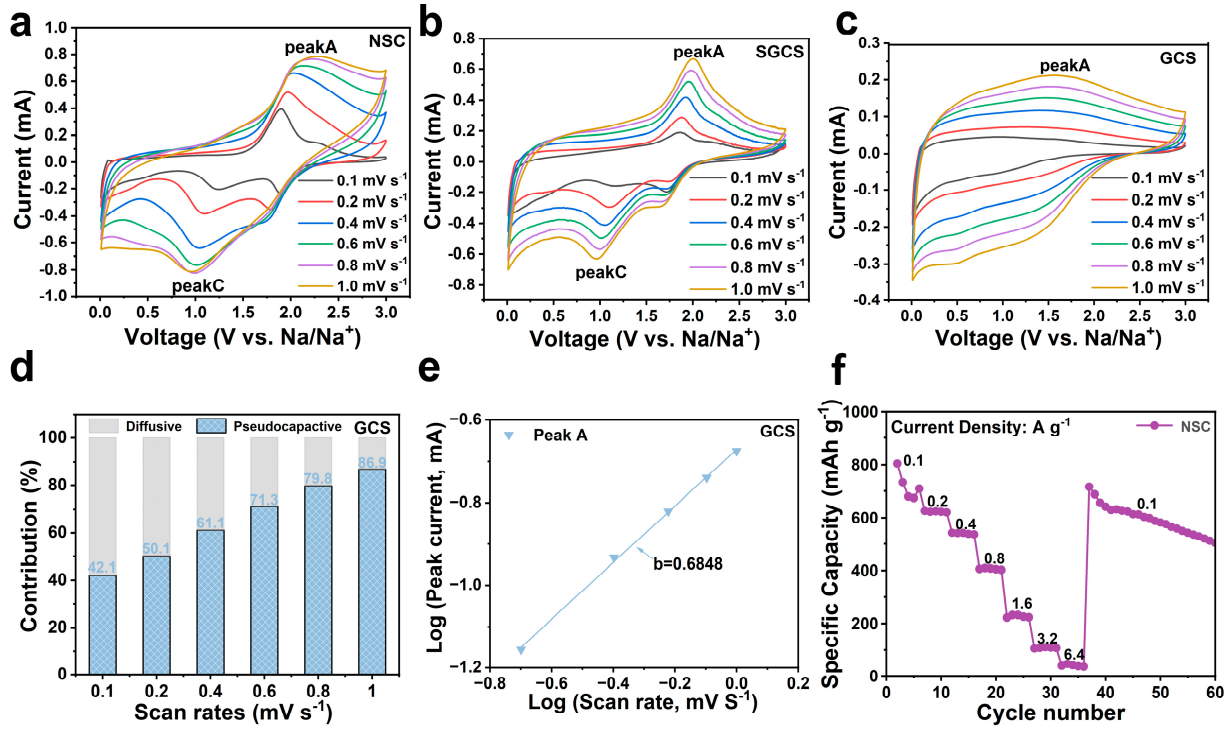


Figure S5. CV curves at different scan rates of **a)** NSC, **b)** SGCS, and **c)** GCS from 0.1 to 1.0 mV s⁻¹. **(d)** Normalized pseudocapacitive contribution ratios of GCS at scan rates from 0.1 to 1.0 mV s⁻¹. **(e)** Log (*i*) versus log (*v*) plots of GCS at selected peak currents. **(f)** Rate capability of NSC electrodes at increasing current densities from 0.1 to 6.4 A g⁻¹.

The galvanostatic intermittent titration technique is used to calculate the Na⁺ diffusion coefficient according to the following formula (S1):

$$D = \left(\frac{\Delta E_s}{\Delta E_t} \right) \frac{4L^2}{\pi\tau} \quad (\text{S1})$$

In the formula, ΔE_s represents the change in steady-state potential of a current pulse (V), ΔE_t is the elimination of constant current pulse potential (V) changes after iR drops, *t* is the duration of the current pulse, τ is the relaxation time, and *L* is the diffusion length of Na⁺[42,57].

Table S6. The EIS fitting parameters of SGCS@NSC, NSC, SGCS, and GCS.

Samples	$R_s(\Omega)$	$R_{ct}(\Omega)$
SGCS@NSC	4.9	637.2
NSC	7.832	975.1
SGCS	11	1875
GCS	13.1	981