

Activation of Peroxymonosulfate by Co-Ni-Mo Sulfides/CNT for Organic Pollutant
Degradation

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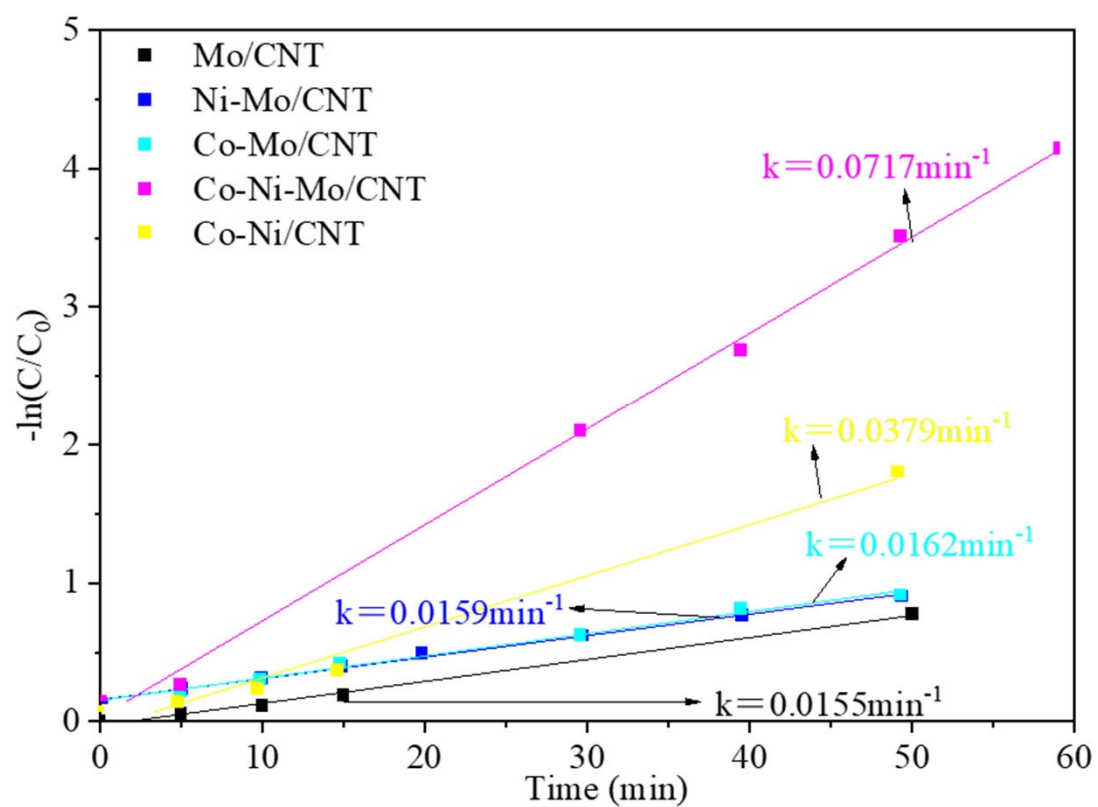


Figure S1. Degradation rate constants of Co-Ni-Mo/CNT, Co-Mo/CNT, Co-Ni/CNT, Ni-Mo/CNT, and Mo/CNT.

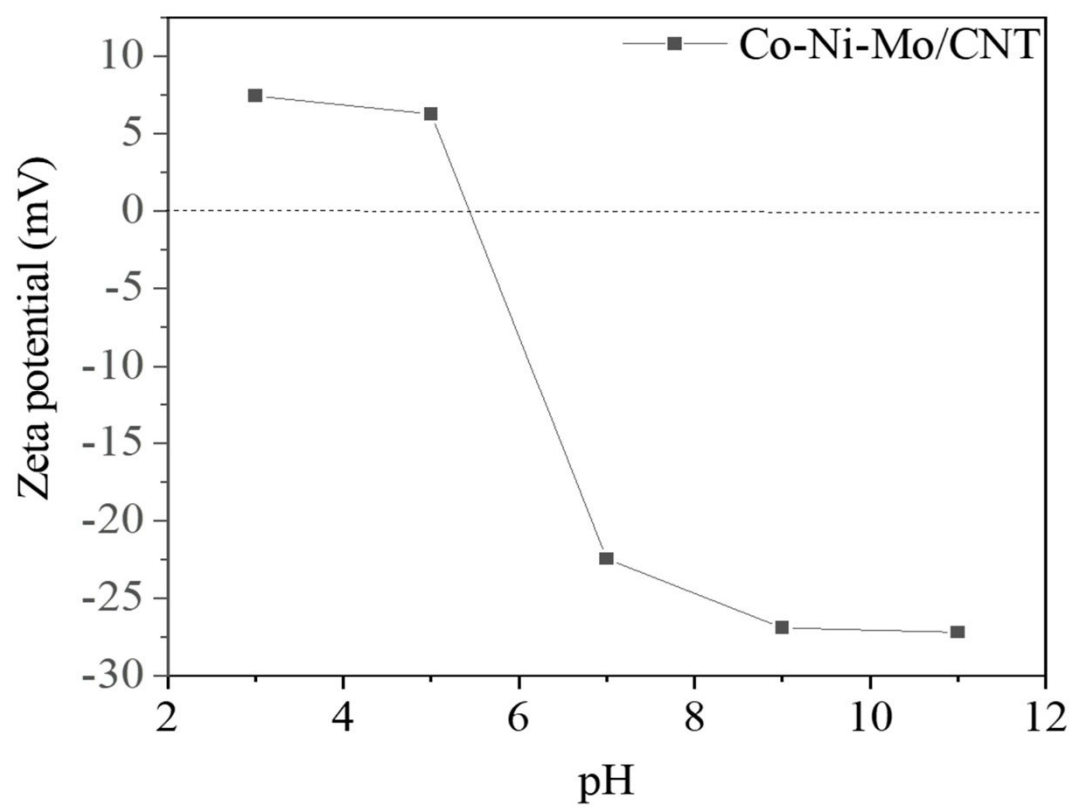


Figure S2. The change of zeta potential of Co-Ni-Mo/CNT with pH.

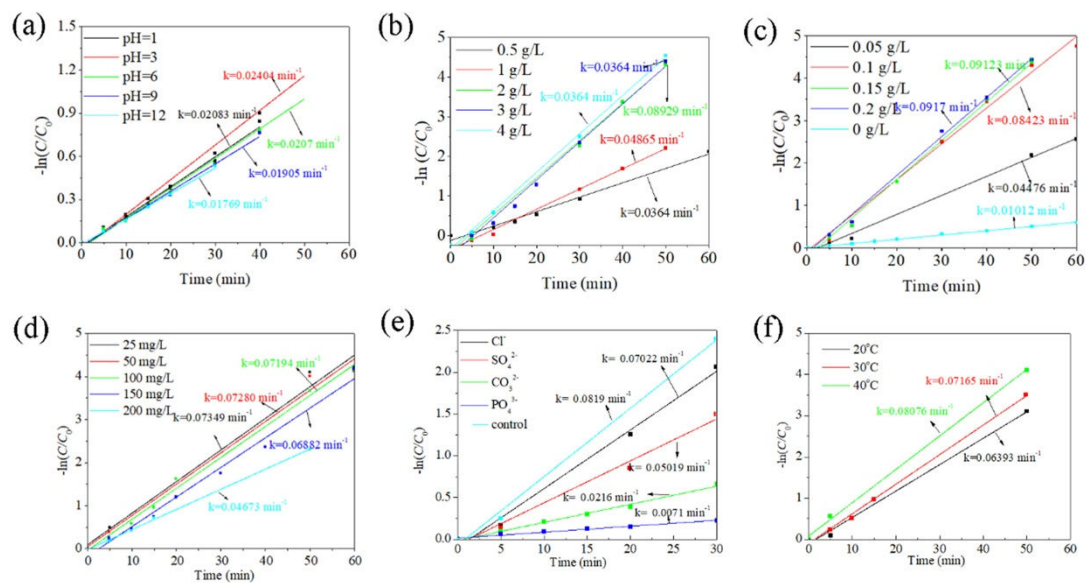


Figure S3. Degradation rate constants on different pH(a), PMS concentration (b), catalyst dosage (c), Rhodamine B concentration(d), different ion species(e), and different temperatures (f).

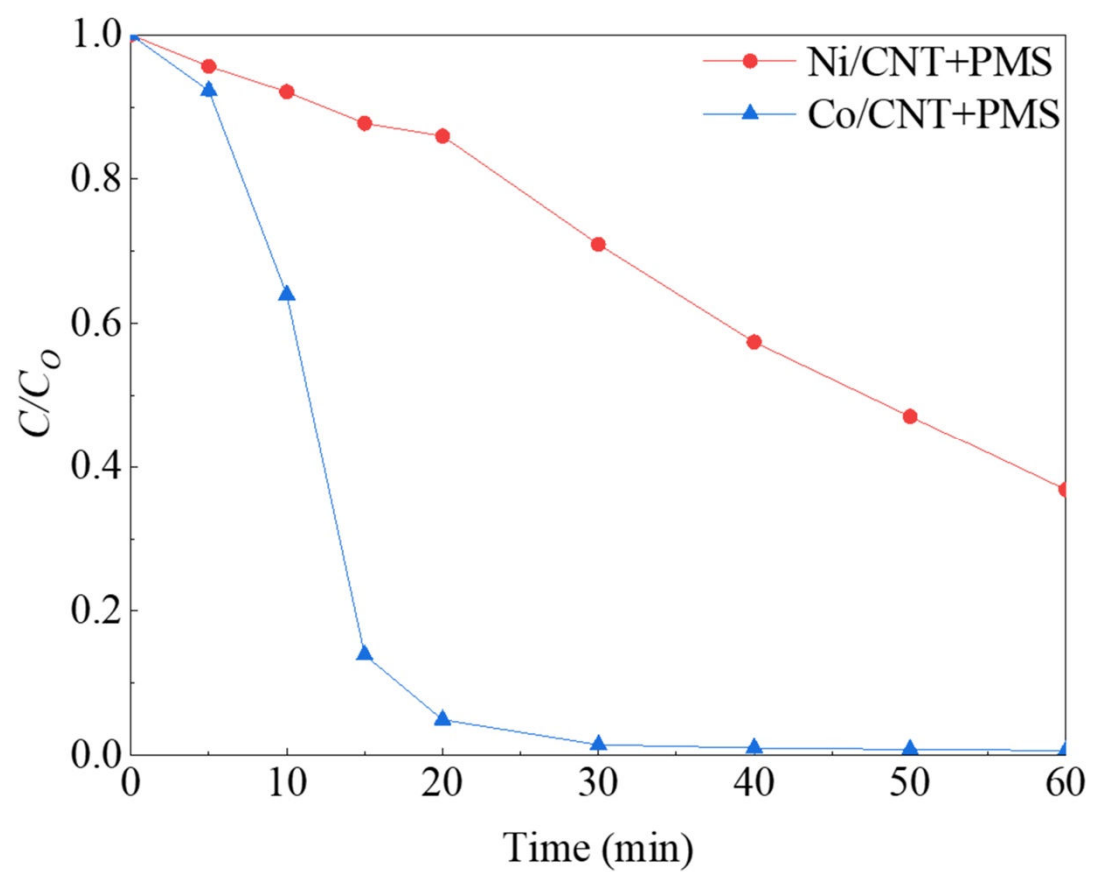


Figure S4. Degradation of Rhodamine B catalyzed by Ni/CNT or Co/CNT.

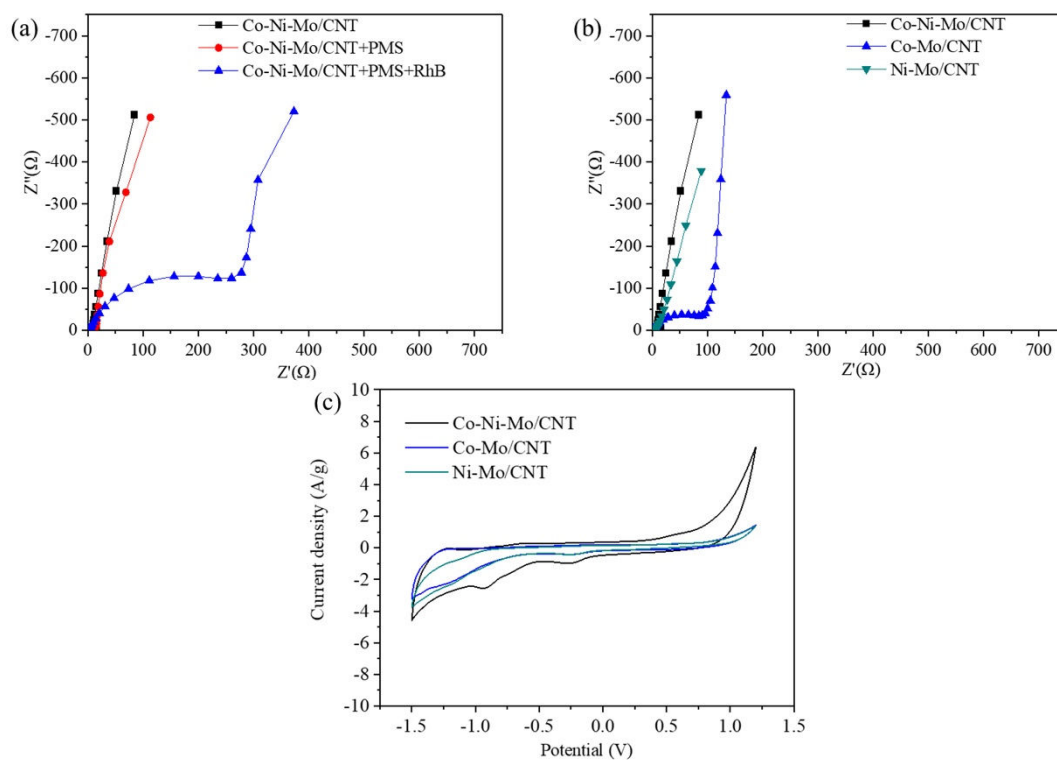


Figure S5. (a) EIS of Co-Ni-Mo/CNT in different systems. (b) EIS and (c) CV curves of Co-Ni-Mo/CNT, Co-Mo/CNT and Ni-Mo/CNT in the presence of PMS.

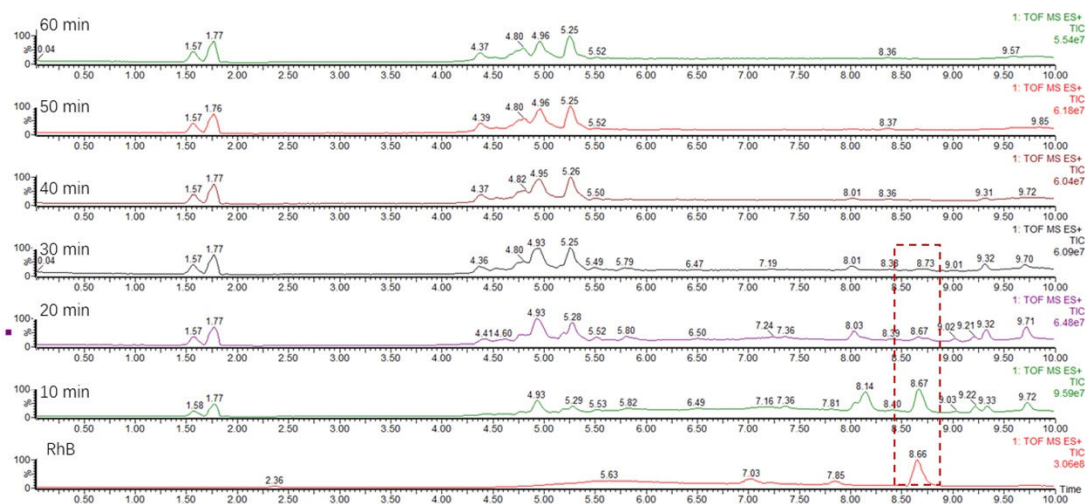


Figure S6. UPLC spectra of Rhodamine B solution at different times in the presence of Co-Ni-Mo/CNT + PMS.

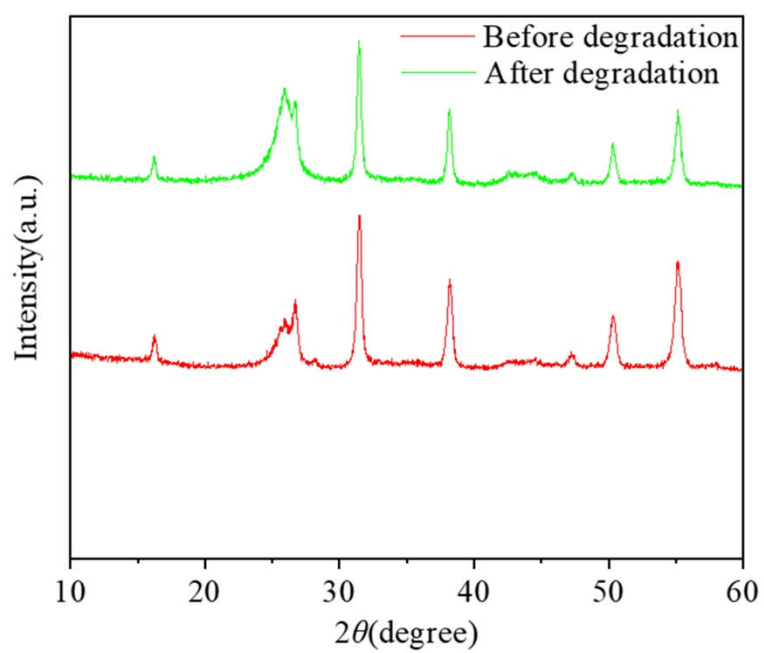
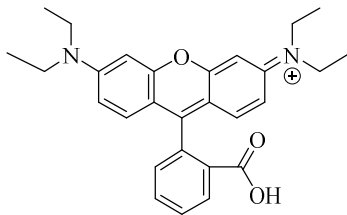
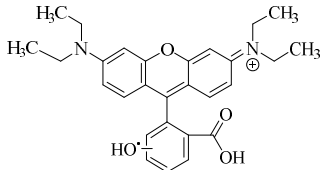
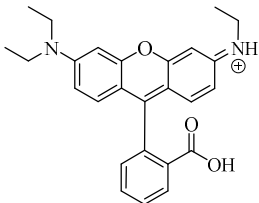
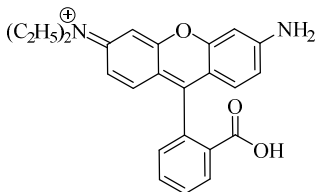
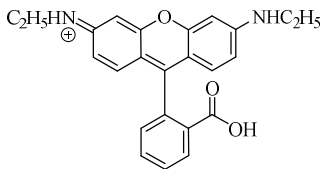
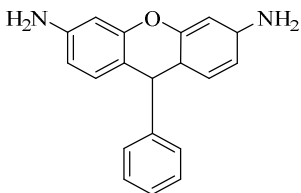
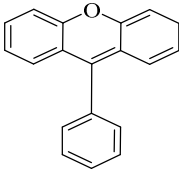
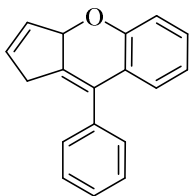
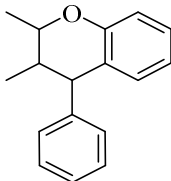
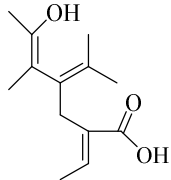
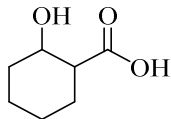
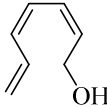
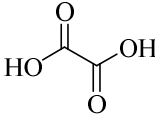
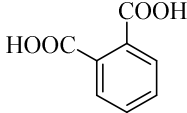
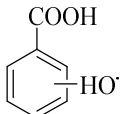
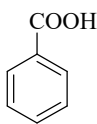


Figure S7. XRD patterns of the recycled catalyst.

Table S1, Intermediates of Rhodamine B degradation in presence of Co-Ni-Mo/CNT + PMS

No.	m/z	Formula	Structure	Intermediates
1	443.2	$C_{28}H_{31}N_2O_3^+$		Rhodamine B
2	459.2	$C_{28}H_{31}N_2O_4^+$		/
3	415.2	$C_{26}H_{27}N_2O_3^+$		N,N-diethyl-N-ethylrhodamine
4	387.2	$C_{24}H_{23}N_2O_3^+$		N,N-diethylrhodamine
5	387.2	$C_{24}H_{23}N_2O_3^+$		N-ethyl-N-ethylrhodamine
6	290.1	$C_{19}H_{18}N_2O$		Me-rhodamine
7	258.1	$C_{19}H_{14}O$		9-phenyl-3H-xanthene
8	246.1	$C_{18}H_{14}O$		9-phenyl-1,3a-dihydrocyclopenta[b]chromene

9	238.1	$C_{17}H_{18}O$		2,3-dimethyl-4-phenyl chromane
10	224.1	$C_{13}H_{20}O_3$		(2E,5Z)-2-ethylidene- 6-hydroxy-5-methyl-4 -(propan-2-ylidene)he pt-5-enoic acid
11	144.1	$C_7H_{12}O_3$		2-hydroxycyclohexan e-1-carboxylic acid
12	110.1	$C_7H_{10}O$		(2Z,4Z)-hepta-2,4,6-tr ien-1-ol
13	90.0	$C_2H_4O_2$		Oxalic acid
14	166.0	$C_8H_6O_4$		Phthalic acid
15	138.0	$C_7H_6O_3$		/
16	122.0	$C_7H_6O_2$		Benzoic acid
