

# Supporting Information

## Carboxymethyl $\beta$ -Cyclodextrin Assistance for the 4-Nitrophenol Reduction Using Cobalt-Based Layered Double Hydroxides

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## I) Determination of the CM $\beta$ CD amount inside LDHs

The amount of CM $\beta$ CD can be estimated thanks to the corresponding TGA profile but also with different hypothesis that have to be considered.

**Hypothesis 1:** the anionic cyclodextrin (CM $\beta$ CD with a substitution degree of 3.2 is linked twice with the 2 cationic layers (like the carbonate anion). There is 1.2 carboxylate group left which is finally linked with 1.2 sodium cation.

The molecular weight can be calculated as follows:

$$M(\text{CM}\beta\text{CD in LDH}) = (\beta\text{CD}) - 3.2 \cdot M(\text{H}) + 3.2 M(\text{CH}_2\text{COO}) + 1.2 \cdot M(\text{Na})$$

$$M(\text{CM}\beta\text{CD in LDH}) = 1135 - 3.2 \cdot 1 + 3.2 \cdot (14 + 12 + 32) + 1.2 \cdot 23.$$

$$M = 1345 \text{ g/mol}$$

**Hypothesis 2:** the hybrid LDH can be composed with CM $\beta$ CD but also carbonate ions.

Taking into account the general molecular formula  $[\text{M(II)}_a \text{M'(III)}_b (\text{OH})_{2a+2b}] (\text{A}^{n-b/n})_4 \cdot 4 \text{H}_2\text{O}$  and the b value of 2 (hypothesis 1) :  $[\text{Co}_6\text{Al}_2(\text{OH})_{16}](\text{CO}_3^{2-} \text{ and/or CM}\beta\text{CD}) \cdot 4 \text{H}_2\text{O}$

$M(\text{hybrid LDH}) = 6 \cdot M(\text{Co}) + 2 \cdot M(\text{Al}) + 16 \cdot M(\text{OH}) + X \cdot M(\text{CM}\beta\text{CD in LDH}) + (1-X) \cdot M(\text{CO}_3)$  with X the molar ratio of intercalated CM $\beta$ CD

$$M(\text{hybrid LDH}) = 6 \cdot 59 + 2 \cdot 27 + 16 \cdot 17 + X \cdot 1345 + (1-X) \cdot 60 + 4 \cdot 18$$

$$M(\text{hybrid LDH}) = 812 + 1285X$$

**Hypothesis 3:** the calcinated product obtained during the TGA analysis under air is the same for control and hybrid LDHs, i.e.  $\text{Co}_6\text{Al}_2\text{O}_{11}$ , so 584 g/mol.

- For the hybrid material CoAl\_CM $\beta$ CD[0.06], the final weight is 51%, so the initial weight is  $M_i =$

$$\frac{584}{0.51} = 1168 \text{ g/mol. With the previous equation, } 812 + 1285 X = 1168, \text{ so } X = 0.28$$

- For the hybrid material CoAl\_CM $\beta$ CD[0.03], the final weight is 52%, so the initial weight is  $M_i =$

$$\frac{584}{0.52} = 1123 \text{ g/mol. With the previous equation, } 812 + 1285 X = 1123, \text{ so } X = 0.24$$

**CONCLUSION: the estimated amount of CM $\beta$ CD**

**in CoAl\_CM $\beta$ CD[0.06] is 28%**

**in CoAl\_CM $\beta$ CD[0.03] is 24%**

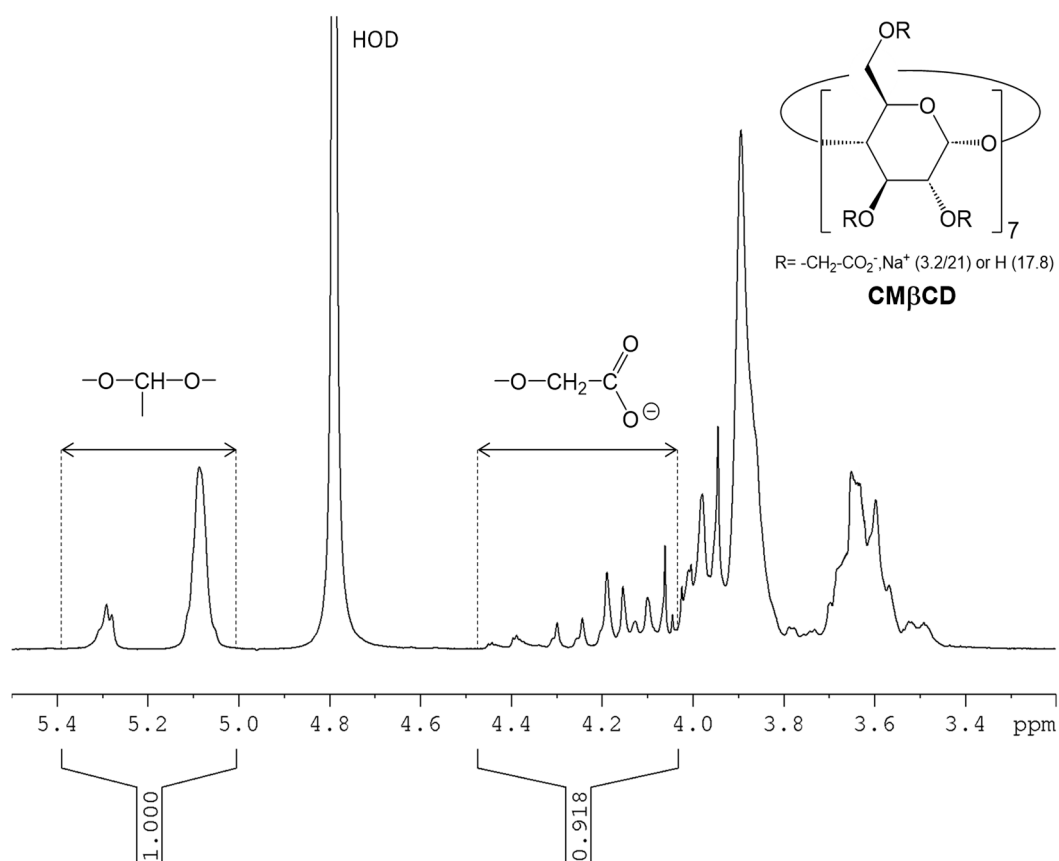
## II) Synthesis of cobalt-based hydroxides

### *Without CM $\beta$ CD*

3.13 g of  $\text{Co}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$  were dissolved in 20 mL of deionized water and the resulting metallic cation solution was poured in an addition funnel and dropwise poured into 30 mL of an alkaline solution during 10 minutes. The pH of the mixture was kept at a value of 10,5 by using a  $2 \text{ mol.L}^{-1}$  sodium hydroxide solution. When the addition was complete, the resulting suspension was stirred 24h at  $60^\circ\text{C}$ , filtrated, washed with deionized water until filtrate pH corresponded to the water pH. The solid was dried during 64h at  $60^\circ\text{C}$ . A dark grey solid (0.97 g) was obtained and crushed in a mortar.

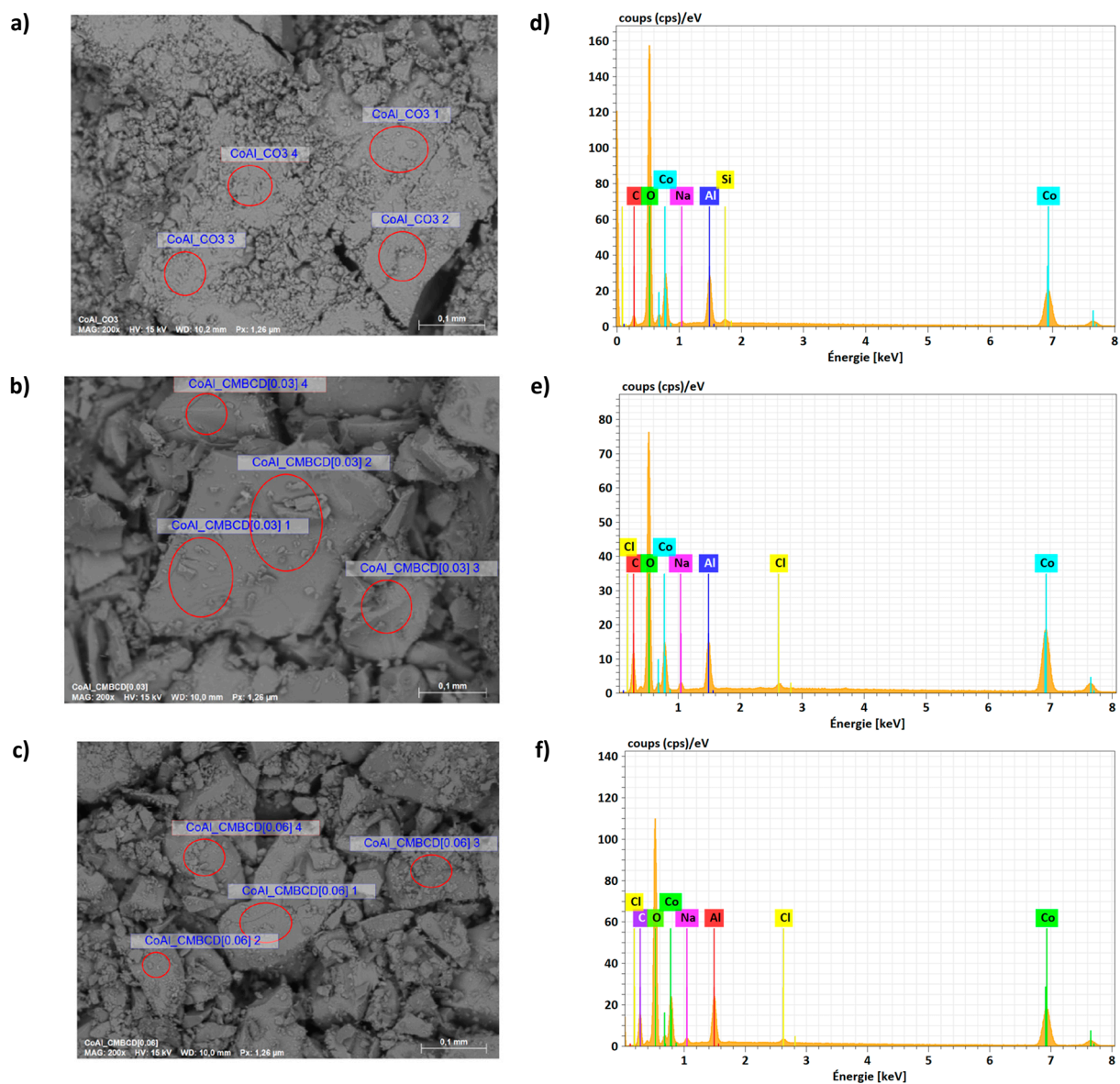
### *With CM $\beta$ CD at an initial concentration of 0.06M*

3.13 g of  $\text{Co}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$  were dissolved in 20 mL of deionized water and the resulting metallic cation solution was poured in an addition funnel (solution A). 1.30 g of CM $\beta$ CD, dissolved in 30 mL of an alkaline solution, were put in a becher with a fixed pH of 10,5 (solution B). Solution A was dropwise poured into solution B during 10 minutes. The pH of the mixture was kept at a value of 10,5 by using a  $2 \text{ mol.L}^{-1}$  sodium hydroxide solution. When the addition was complete, the resulting suspension was stirred 24h at  $60^\circ\text{C}$ , filtrated, washed with deionized water until filtrate pH corresponded to the water pH. The solid was dried during 64h at  $60^\circ\text{C}$ . An olive green solid (1,16 g) was obtained and crushed in a mortar.

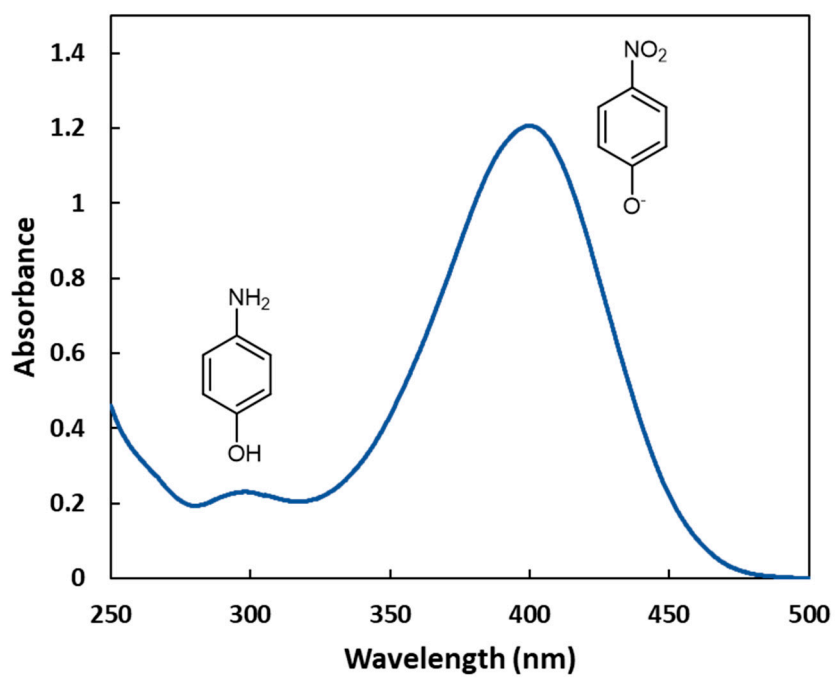


**Figure S1.  $^1\text{H}$  NMR spectrum of CMβCD ( $\text{D}_2\text{O}$ , 25 °C, 300 MHz)**

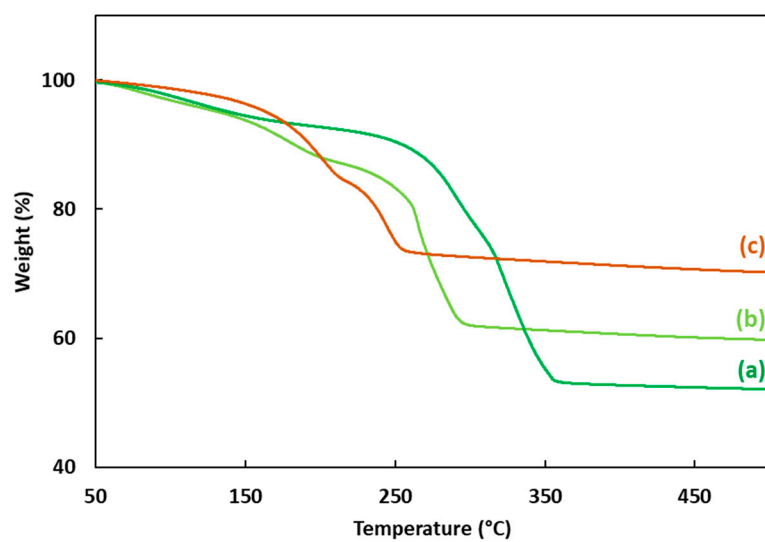
Using the  $^1\text{H}$  NMR spectrum of CMβCD it is possible to determine the average degree of substitution (DS) of this cyclodextrin. In the case where the integration of the signals of the  $\text{H}_1$  protons of the CD is calibrated to 1 (1  $\text{H}_1$  proton per glycosidic unit), the integration of the  $\text{CH}_2$  signals of the carboxymethyl groups gives the average number of protons present on this modified cyclodextrin. The DS was found by dividing this integration by 2 and multiplying by 7 ( the number of glycosidic units). Using this calculation, we obtain an average degree of substitution equal to 3.2.



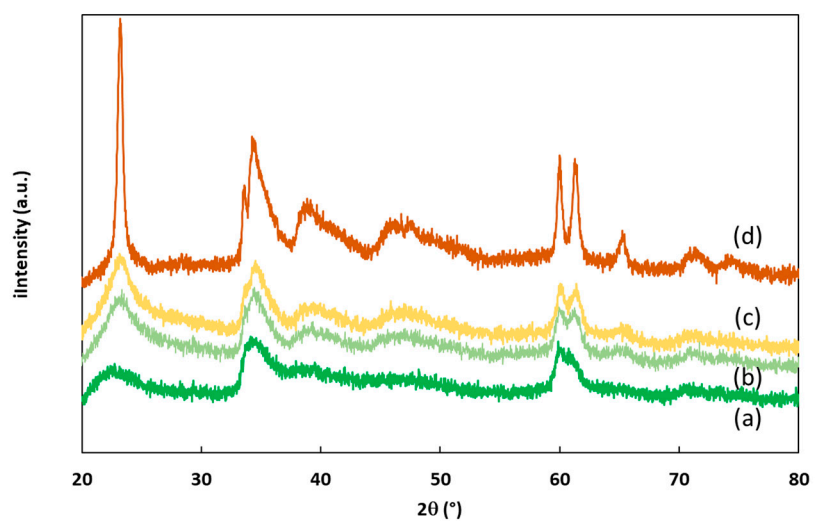
**Figure S2.** Scanning electron micrographs and the corresponding average elemental analysis of the four arbitrary domains through EDX of  $\text{CoAl\_CO}_3$  (a,d),  $\text{CoAl\_CMBD}[0.03]$  (b,e) and  $\text{CoAl\_CMBD}[0.06]$  (c,f).



**Figure S3.** UV-Vis spectrum of an equimolar mixture of 4-nitrophenol and 4-aminophenol in a sodium borohydride aqueous solution

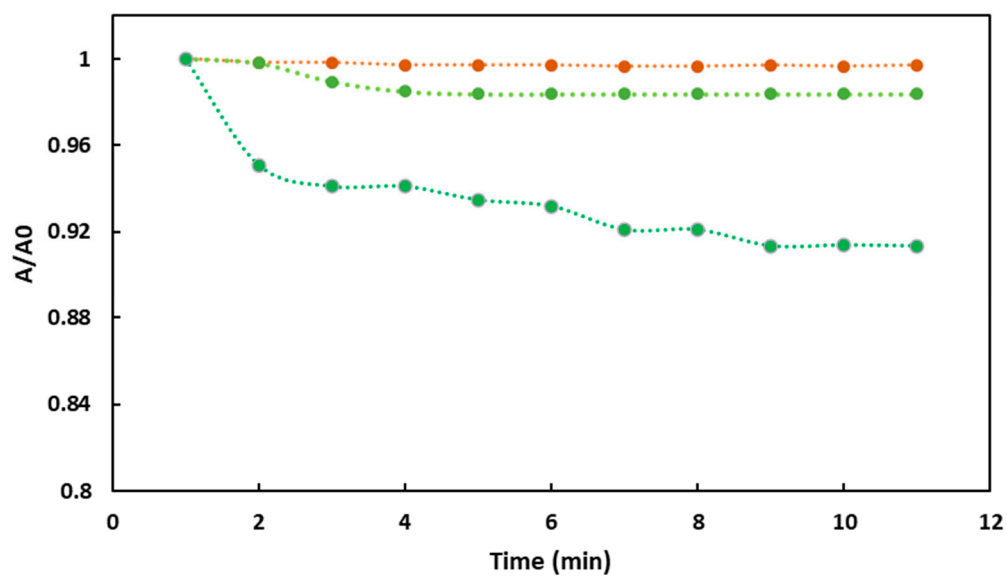


**Figure S4.** Thermograms of CoAl\_CMβCD[0.06] (fresh (a); after the 3<sup>rd</sup> run (b), CoAl\_CO<sub>3</sub> (c))



**Figure S5.** Diffractograms of CoAl\_CM $\beta$ CD[0.06] during recycling study (fresh (a); after the 1<sup>st</sup> run (b), after the 3<sup>rd</sup> run (c), and CoAl\_CO<sub>3</sub> diffractogram (d))

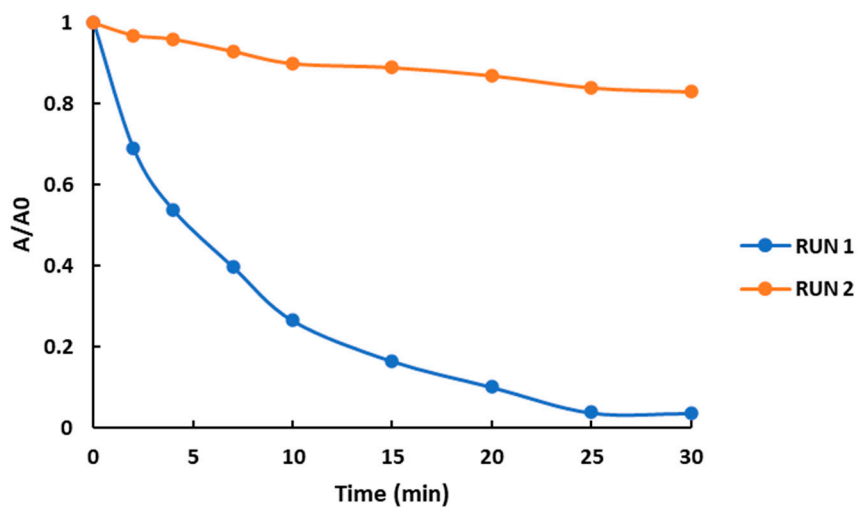




**Figure S6.** 4-NPhO<sup>-</sup> adsorption on cobalt-based materials (CoAl<sub>2</sub>CO<sub>3</sub> (brown line); CoAl<sub>2</sub>CO<sub>3</sub> + CMβCD (light green line); CoAl\_CMβCD[0.06] (dark green line))

### III) Material Recycling procedure

After 30 minutes of reaction (RUN 1), the suspension is centrifuged and the supernatant was removed. This supernatant was mixed with a fresh 4-NPhOH solution containing sodium borohydride (RUN 2). The kinetic of the reaction was monitored by UV-Vis spectrophotometry.



**Figure S7.** CoAl\_CMβCD[0.06] recycling for the 4-NPhOH reduction monitored by UV-Vis spectrophotometry