

# One-Pot Synthesis of $\text{TiO}_2$ -rGO Photocatalysts for the Degradation of Groundwater Pollutants

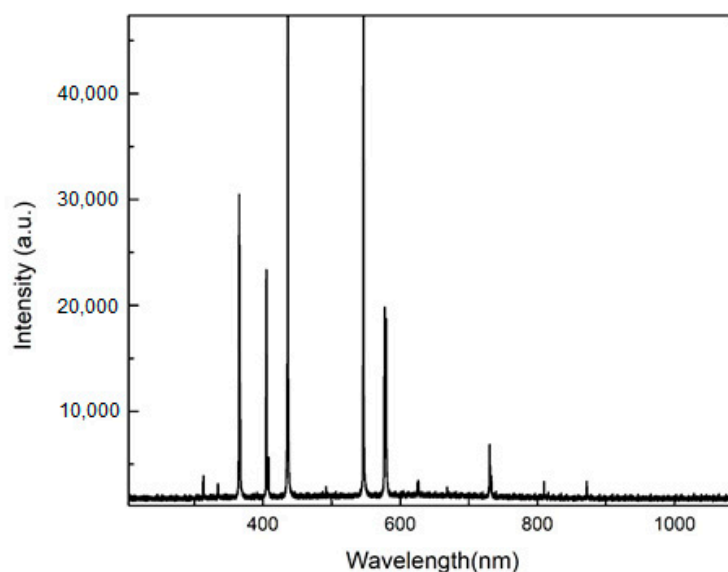
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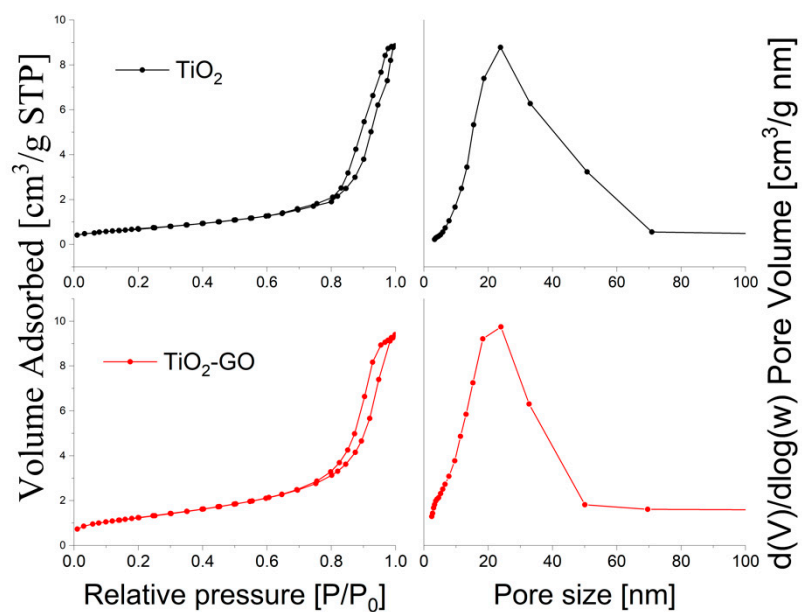
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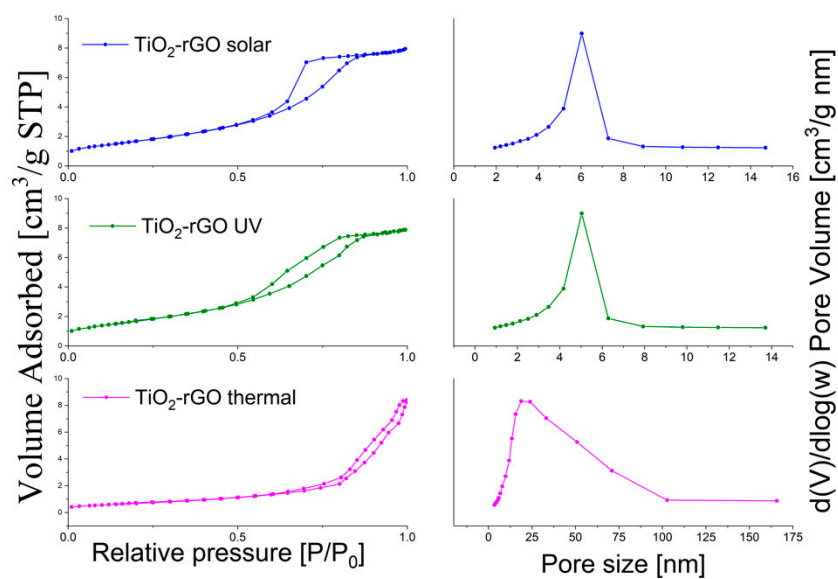
## Supplementary Materials



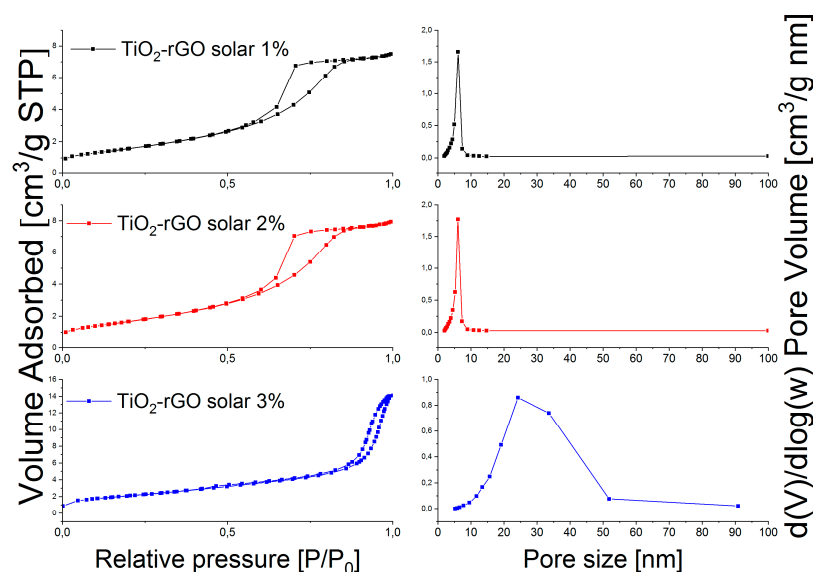
**Figure S1.** Measured emission spectrum of the Osram Ultra Vitalux lamp.



**Figure S2.** N<sub>2</sub> adsorption-desorption isotherms and BJH pore size distribution for the examined samples.

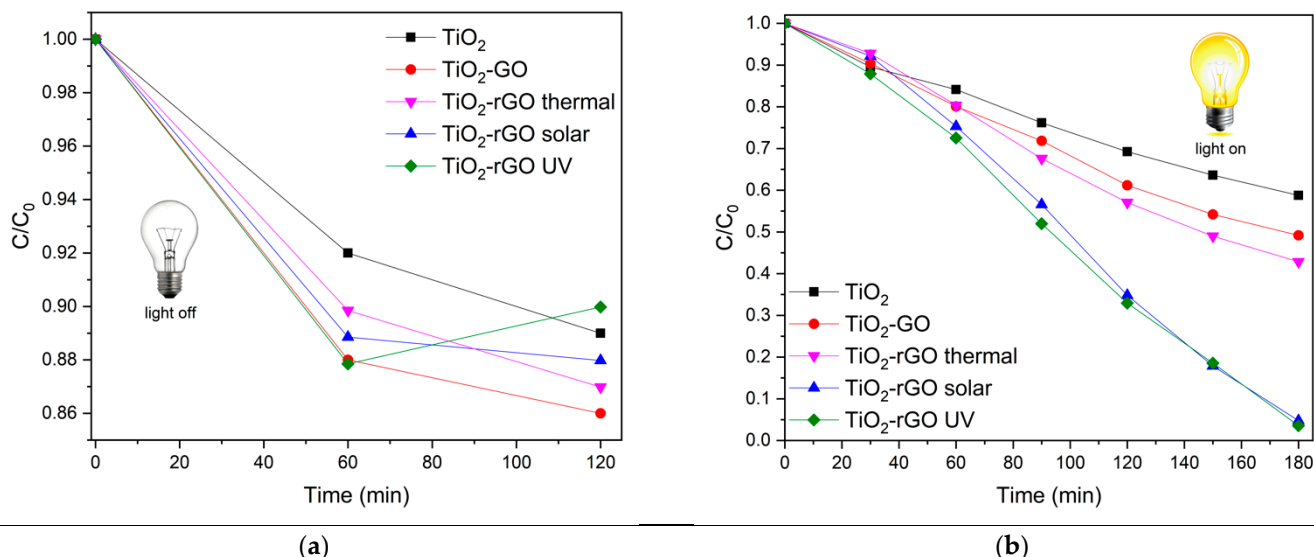


**Figure S3.** N<sub>2</sub> adsorption-desorption isotherms and BJH pore size distribution for the examined samples.

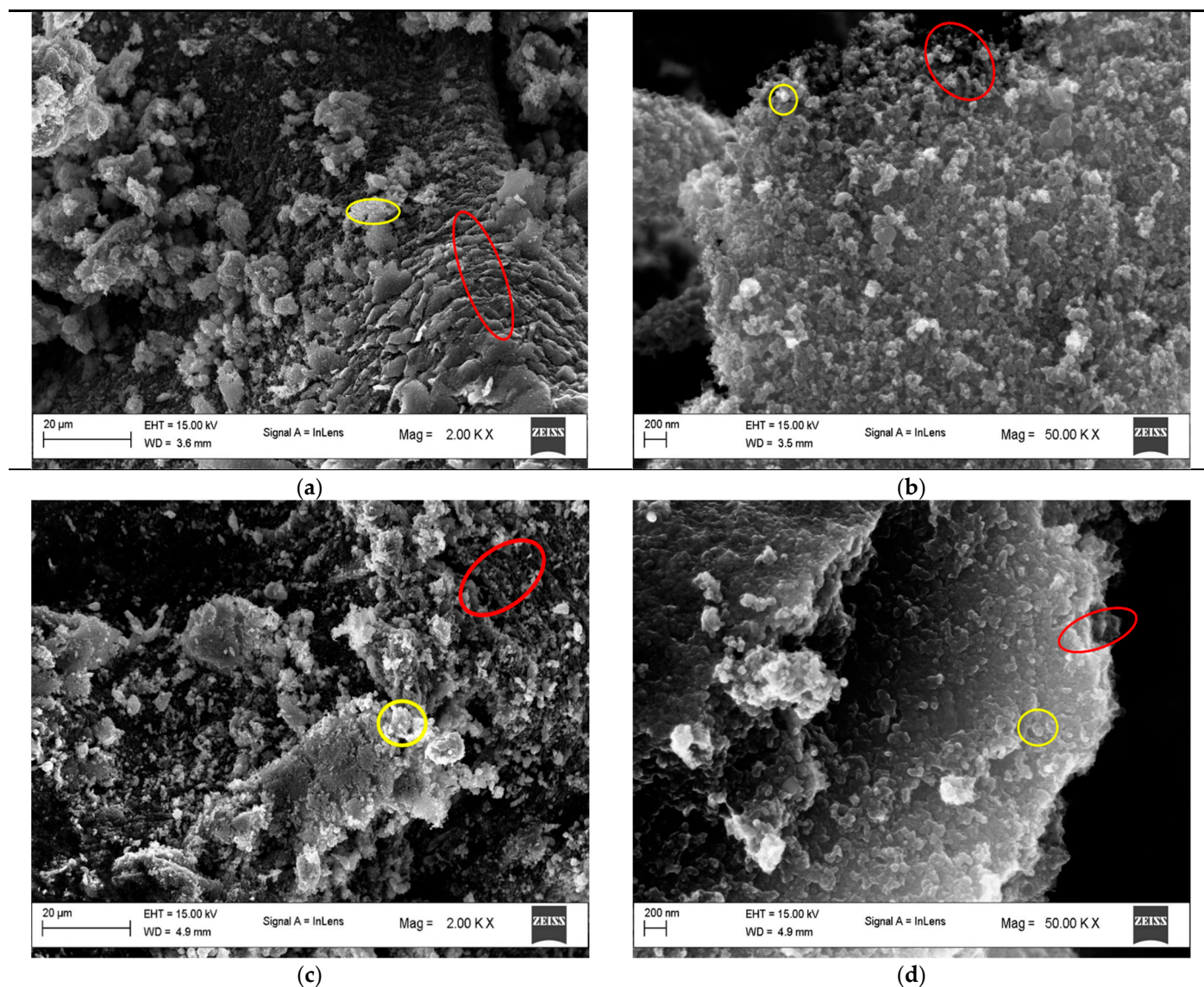


**Figure S4.** N<sub>2</sub> adsorption-desorption isotherms and BJH pore size distribution for the solar samples with different percentage content of rGO.

From the data shown in the Figure S4, it is possible to note how a small amount of rGO (1%, 2% wt) did not change the textural properties of the TiO<sub>2</sub>-rGO composites. However, with the 3 wt.% of rGO, despite the surface area remaining similar, ( $S_{\text{BET}} = 51 \text{ m}^2\cdot\text{g}^{-1}$ ,  $53 \text{ m}^2\cdot\text{g}^{-1}$  and  $50 \text{ m}^2\cdot\text{g}^{-1}$  for TiO<sub>2</sub>-1%rGO solar, TiO<sub>2</sub>-2%rGO solar, and TiO<sub>2</sub>-3%rGO solar, respectively) the pore shapes and dimensions were different. The hysteresis changes from H2 to H3, related to slit-shaped pores, while the mean pores diameter increases from 6 nm of 2% sample to 16 nm of TiO<sub>2</sub>-3%rGO solar, pointing to as the increase of the amount of rGO deposited on the surface of TiO<sub>2</sub> caused the pores' modification.

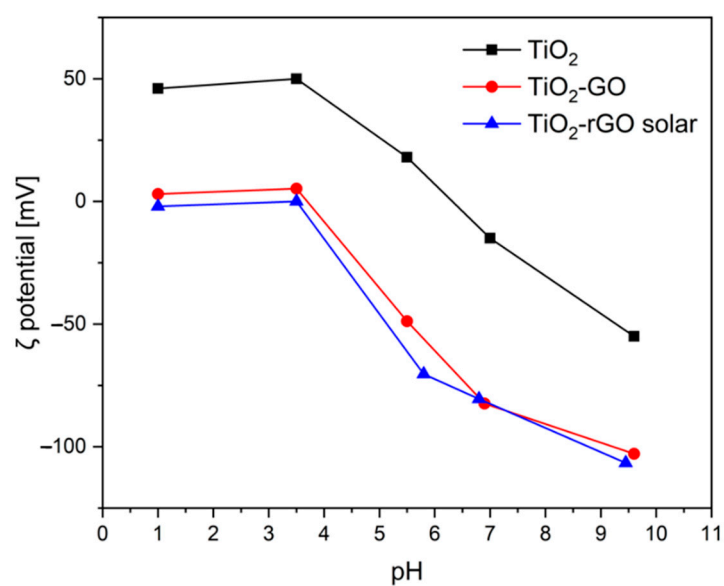


**Figure S5.** (a) Adsorption-desorption equilibrium in the first two hours. (b) Photocatalytic degradation after the reached equilibrium.



**Figure S6.** SEM images of TiO<sub>2</sub>-GO (a) and (b) at 2K and 50K magnification respectively; the same for TiO<sub>2</sub>-rGO solar in (c) and (d) respectively.

From the SEM images of TiO<sub>2</sub>-GO and TiO<sub>2</sub>-rGO solar, it is possible to distinguish the carbonaceous structures marked by the red circles, and the TiO<sub>2</sub> agglomerates by the yellow ones.



**Figure S7.** Zeta potential of the principal sample in function of pH.

**Table S1.** Kinetic constants measured at different wt.% of rGO.

wt.% of rGO	Total degradation (%)	Kinetic constant·10 <sup>3</sup> (s <sup>-1</sup> )
TiO <sub>2</sub>	48	0.06 ± 0.01
1%	80	0.13 ± 0.02
2 %	97	0.17 ± 0.01
3 %	90	0.15 ± 0.01

Consequently, the 2 wt.% was chosen as optimal content of rGO. Indeed, a lower amount reaches only the 80% of degradation, whereas a higher amount decreases the activity to 90% of degradation. The same considerations are similar considering the kinetic constants.