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# Structure and Physicochemical Properties of Water Treated under Methane with Low-Temperature Glow Plasma of Low Frequency

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**Abstract:** Our former studies delivered a strong evidence that water indirectly treated with low-temperature, low-pressure glow plasma of low frequency (GP) changed its structure depending on the atmosphere in which such treatment was performed (air, ammonia, and nitrogen) and on the time of the treatment (0 to 120 min). In every case, water of different physicochemical characteristics and interesting biological functions was produced. Therefore, the relevant studies were extended to treating deionized water with GP under methane. The resulting samples were characterized by means of ultraviolet/visible (UV/VIS), Fourier transformation infrared-attenuated total reflectance (FTIR-ATR), electron spin resonance (ESR) and Raman spectroscopies, differential scanning calorimetry (DSC), thermogravimetry, pH, conductivity, and refractive index. The generated samples of water had entirely different physicochemical properties from those recorded for water treated with GP in the air and under both ammonia and nitrogen. The treatment of water with GP under methane did not produce clathrates hosting methane molecules. Thermogravimetry delivered an evidence that the treatment with GP increased the aqueous solubility of methane. That solubility non-linearly changed against the treatment time.

**Keywords:** methane clathrates; synthetic naftusya water; water macrostructure reconstruction

## 1. Introduction

In our recent paper, the structure and physicochemical properties of water treated with low-temperature, low-pressure glow plasma of low frequency (GP) in the air [1] was presented. GP was developed in the proximity of the treated water. Under fixed parameters of the GP generation, the structure and physicochemical properties of the resulting water depended on the treatment time. In a series of subsequent papers, the benefits resulting from the use of that water were shown. Thus, that water stimulated growth and pathogenicity of entomopathogenic fungi used as biopesticides [2]. Murawski et al. demonstrated the benefits of cryopreservation of ram [3] and boar [4] semen in

such water. The barley malt quality could be improved when GP-treated water was applied [5]. Watering peppermint with GP-treated water had no essential effect on the yield of the essential oil and a plant crop yield, but the composition of the essential oil dramatically changed. The menthol content in it was distinctly lower and, simultaneously, the bactericidal properties of that oil increased. Several new components of that oil could be characterized. Bioaccumulation of cations and anions in leaves, stems, and roots also changed [6]. Several interesting applications of water treated with GP in contact with the air are presented in a newly published monograph [7]. Applications presented therein were based on unpublished research reports dealing with plant cultivation and animal breeding.

These findings prompted us to study the treatment of water with GP under ammonia. The treatment of hard water with GP under ammonia protected precipitation of a scale in kettles, heat exchangers, and pipe-lines. On heating of so treated hard water, a separated fine precipitate did not agglomerate into a hard scale [8]. Recently, a specific kind of water was presented [9] when water was treated with GP under oxygen-free nitrogen [9]. Treating water with GP in contact with the air resulted in the formation of aqueous clathrates including excited singlet oxygen molecules. Water treated with GP under nitrogen contained aqueous clathrates with various free radical species of molecular nitrogen. Their nature depended on the treatment time. The effect of such water upon the growth and composition of essential oil of basil (*Ocimum basilicum* L.) [10] and another study upon the growth of cress [11] prompted us to study a treatment of water with GP under methane, following identical conditions of the GP treatment, as applied in former studies.

Under normal atmospheric pressure, methane poorly solubilizes in water and its amount in g/kg distilled water ranges from almost 0.04 to 0.0235 at 0 °C and 20 °C, respectively [12]. As the pressure elevates, that solubility increases. That solubility depends also on the water salinity [13]. In the Solar System as well as in the floor of oceans aqueous methane clathrates  $\text{CH}_4 \cdot 5.75\text{H}_2\text{O}$  or  $4\text{CH}_4 \cdot 23\text{H}_2\text{O}$  are formed, respectively, as a result of low temperature and high pressure [14–17].

Since GP deteriorates the macrostructure of water, the treated water solubilizes various compounds better and its smaller clusters more readily permeate biological membranes. It makes the GP-treated water a good vector for its solutes [1–7]. Hence, one might anticipate that the treatment water with GP would increase the aqueous solubility of methane.

Spas in Iwonicz Zdrój, Rymanów Zdrój (Poland), and Truskawiec (Ukraine) are known for mineral water, so-called “naftusia” (naphtusya). Apart from some minerals (about 0.82 mg/L), it also contains hydrocarbons penetrating the water from underground deposits of petroleum located nearby [18–20]. Naftusia is recommended for curing several health disorders [21,22]. Thus, among the potential applications of water treated with GP under methane, its therapeutic value could be taken into account.

## 2. Materials and Methods

### 2.1. Materials

#### 2.1.1. Water

Commercially available deionized water of  $\text{pH} 7.02 \pm 0.04$ , electromotive force (EMF) =  $351.8 \pm 0.3$  mV (glass versus calomel electrode at 25 °C) and conductivity  $\gamma = 0.444 \pm 0.004$  mS/cm was used.

#### 2.1.2. Methane

Methane from the municipal gas supply system of Częstochowa (Poland) was used. It contained 95.5% of methane, 3.816% ethane, 0.216% propane, 0.031% *i*-butane, 0.030% *n*-butane, 0.06% *i*-pentane, 0.04% *n*-pentane, 0.06% C<sub>6</sub>+ hydrocarbons, 0.605% N<sub>2</sub>, and 0.030% CO<sub>2</sub>.

## 2.2. Methods

### 2.2.1. Treatment of Water with GP

Through the deionized water a stream of methane was bubbled for 15 min. Its flow rate depended on the volume of the water sample. In case of the 200 mL sample, the rate of 10 mL/min was maintained. That water (200 mL) was placed in 250-mL glass bottles and free space over the liquid was additionally filled with methane. The whole system was placed in the metallic chamber at the distance of 5 cm from the plasma generator (Scheme shown in a previous paper [9]) and exposed to GP for 5, 15, 30, 60, and 90 min. Plasma of 38 °C was generated at  $5 \times 10^{-3}$  mbar, 800 V, 50 mA, and 10 KHz frequency. The produced water was stored at ambient temperature in 100-mL closed teflon containers [23].

### 2.2.2. pH

The measurements at 25 °C for all samples (both control samples and those treated with GP for 5 to 90 min) were performed in triplicates with a laboratory multifunction meter CP-501 (Elmetron, Zabrze, Poland) equipped with a HYDROMET, Type ERH-11 combined glass-calomel electrode.

### 2.2.3. Conductivity

Estimations were performed at 25 °C with an ELMETRON CPC-505 instrument (Elmetron, Zabrze, Poland) equipped in an ELMETRON EC-60 sensor. The estimations were taken in triplicates for both control samples and these were treated with GP for 5 to 90 min).

### 2.2.4. Refractive Index

Lab refractometer RL from PZO Warsaw, Poland was used. Measurements performed at 22 °C were triplicated.

### 2.2.5. Differential Scanning Calorimetry (DSC)

Differential scanning calorimeter DSC 2500 (TA Instruments, New Castle, DE, USA) was used. Multipoint calibration (Hg, In, Sn, Bi, Zn, and CsCl) was involved. Analyzed samples (12 mg) were hermetically sealed in alumina capsules. The capsules were then punctured and cooled down to −20 °C with the rate of 10 °C/min followed with heating to 140 °C with the same rate at 50 mL/min nitrogen flow rate. Characteristic temperatures and enthalpy of transitions ( $T_{onset}$ ,  $T_{mid}$ ,  $T_{end}$ , and  $\Delta H$ ) were determined with Proteus Analysis (Netzsch, Selb, Germany) software. Experiments were duplicated.

### 2.2.6. Thermogravimetry

DSC 2500, instrument manufactured by TA Instruments, New Castle, DE, USA was used. The sample was heated from +20 to 120 °C with the rate of 10 °C/min at the 50 mL/min nitrogen flow.

### 2.2.7. Fourier Transformation Infrared-Attenuated Total Reflectance (FTIR-ATR) Spectra

The FTIR-ATR spectra of the film were recorded following the method previously described [1]. Thus, the measurements were performed in the range of 4000–700  $\text{cm}^{-1}$  at resolution of 4  $\text{cm}^{-1}$  using a Mattson 3000 FT-IR (Madison, WI, USA) spectrophotometer. That instrument was equipped with a 30SPEC 30° reflectance adapter fitted with the MIRacle ATR accessory from PIKE Technologies Inc., Madison, WI, USA.

### 2.2.8. Electron Spin Resonance (ESR) Spectra

The spectra as 1st derivative were recorded at room temperature employing an instrument constructed at Wrocław Technical University. The range of the X-band ( $\nu = 9.5$  GHz,  $\lambda = 3.158$  cm) was employed at the 20–25 dB attenuation and 2048 s swiping time. An EPR controller software designed for taking the EPR spectra was applied. For samples saturated with nitrogen,  $g \approx 2.16$ .

### 2.2.9. Ultraviolet/Visible (UV/VIS) Absorption Spectra

The spectra were recorded with a Thermo Scientific Evolution 220 (Thermo Fischer Scientific, USA) spectrophotometer in the wavelength range of 190–1100 nm, in a quartz gas tight cell of 10-mm path length. The method was described previously [1]. The instrument was set on automatic measuring mode at medium scan speed and 1.0 nm slit width.

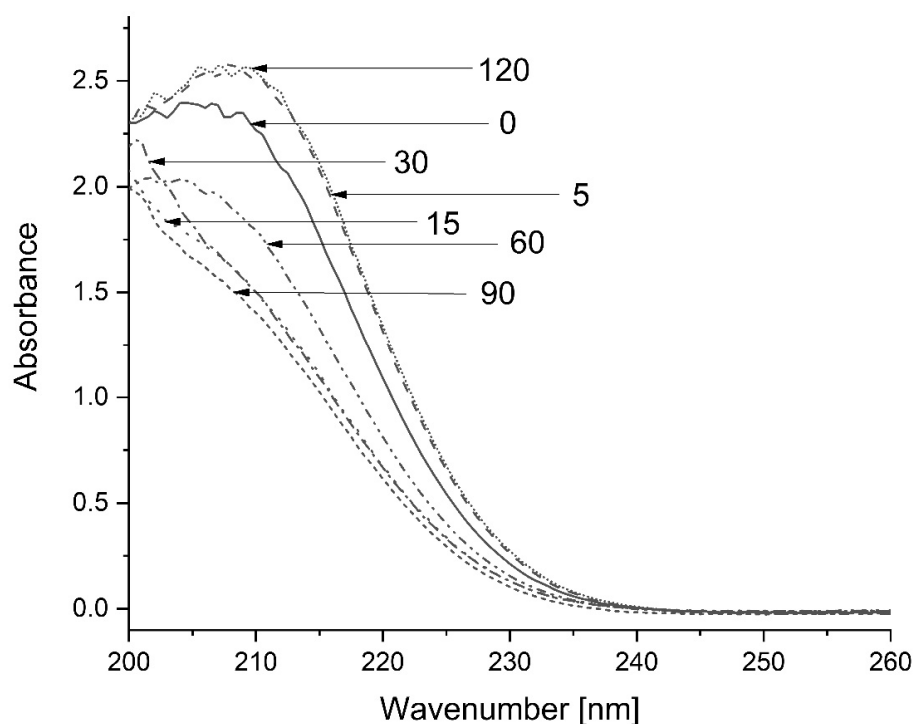
### 2.2.10. Raman Spectra

The spectra were taken with a Perkin-Elmer MPF44A Fluorescence Spectrophotometer (Waltham, MA, USA) equipped with a xenon lamp (excitation at  $330 \pm 1$  nm and slit 2.0 nm) and a 4-mL quartz gas tight cell, following the method described in previous reports [1,5,9]. The spectra were recorded at 22 °C for control distilled water stored in contact with the air, control distilled water saturated with methane, and distilled water saturated with methane exposed to GP for 5, 15, 30, 60, 90, and 120 min.

## 3. Results and Discussion

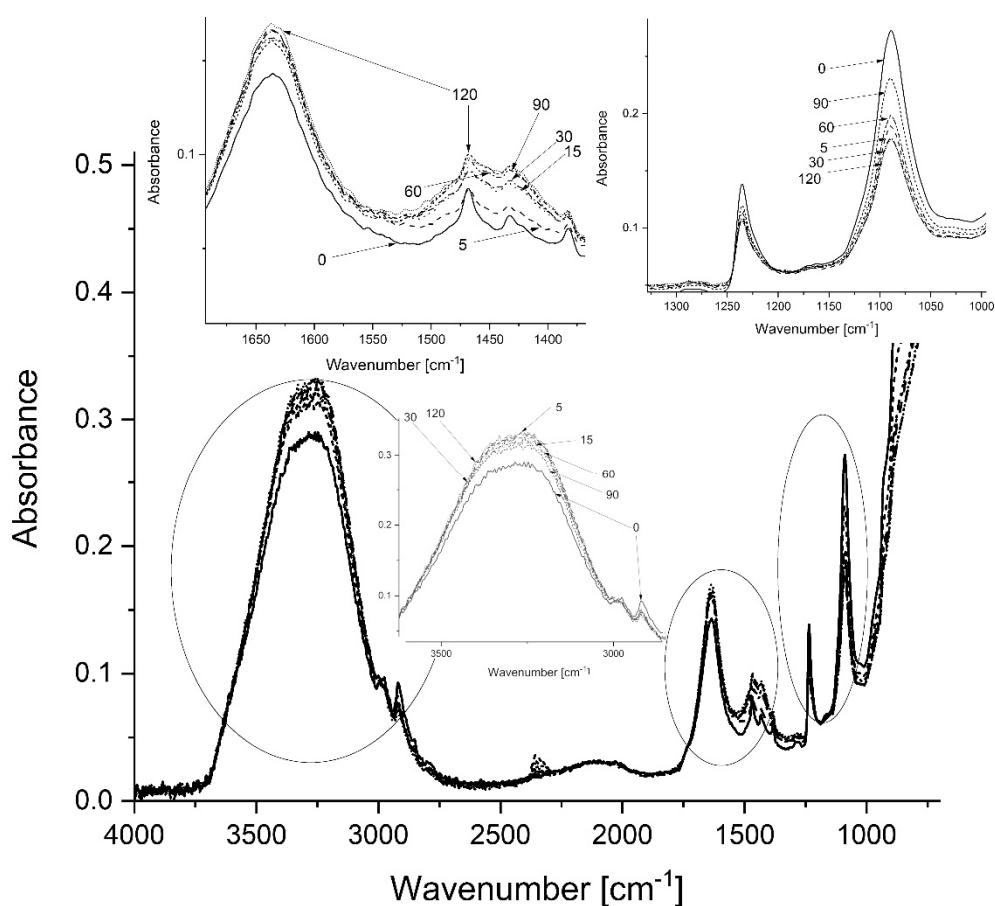
The effect of the treatment of water with GP upon its macrostructure was well documented by spectral measurements.

The UV absorption spectra invariably demonstrated the band at 220 nm which resulted from the  $^1A_1 \rightarrow ^1B_1$  transition [24] (Figure 1). In the spectrum of water saturated with methane prior to the treatment with GP, that band of absorbance of about 2.4 a.u. showed a vibrational structure resulting from simultaneous low transmissions in both light beams. The GP treatment for 5 min elevated the maximum of that band up to about 2.6 a.u., accompanied by a subtle 10-nm upshift. The spectrum of the water treated for 120 min perfectly fitted the same pattern of that band. Prolonged treatment led to a decrease in the intensity of that band to an extent providing a less intensive shoulder of the Reyleigh scattering band with its maximum below 200 nm. The intensity of that shoulder depended on the treatment time. It decreased in the order: treated for 60 min > for 30 min > for 90 min. That is along the increasing ordering and rigidity of the macrostructure.



**Figure 1.** Ultraviolet/visible (UV/VIS) absorption spectra of water treated with glow plasma of low frequency (GP) for 0 to 120 min.

The FTIR-ATR absorption spectra (Figure 2) demonstrated an increase in the intensity of the  $\nu_{\text{OH}}$  band centered around  $3250\text{ cm}^{-1}$ . That increase reflected some deterioration of the macrostructure releasing an increasing number of the hydroxyl groups not engaged in the formation of intermolecular hydrogen bonds. Differences in the absorbance of that band in the spectra of the GP-treated water were fairly subtle, but an insight in these spectra showed that it decreased as a consequence of the treatment time in the order: 120 min > 5 min > 30 min > 60 min > 15 min > 90 min. Absorbance of the  $\nu_{\text{OH}}$  around  $1650\text{ cm}^{-1}$  changed in a similar order. Absorbance of the bands which could be assigned to the C–H vibrations and bending in the molecules of methane, that is, the bands at  $\sim 2900$ ,  $1400$ ,  $1200$ , and  $1100\text{ cm}^{-1}$  decreased with the treatment time. Again, the decrease was not linear against the treatment time, which meant that methane was somehow incorporated into the water macrostructure.



**Figure 2.** Fourier transformation infrared-attenuated total reflectance (FTIR-ATR) spectra of deionized water GP treated under methane for 0 to 120 min.

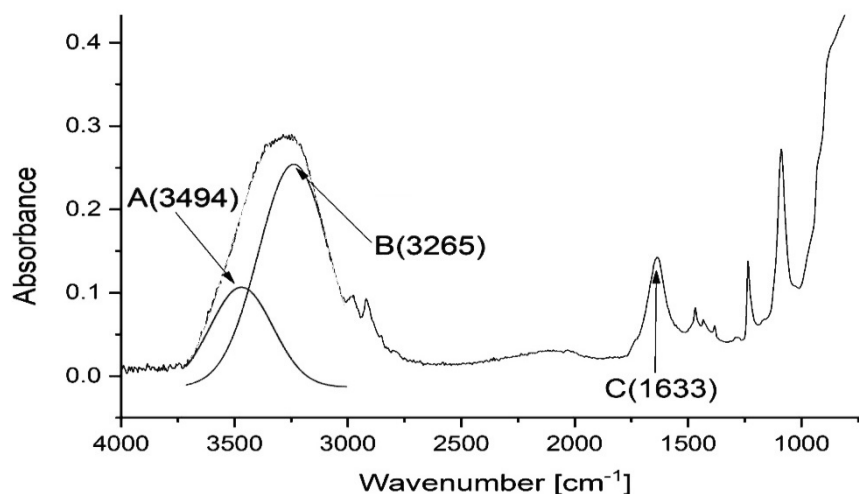
Table 1 demonstrates that removal of dissolved air by blowing it with methane increased the number of the water molecules vibrating asymmetrically. The GP treatment of water saturated with methane for the first 5 min increased their number, then gradually fairly regularly decreased, against the treatment time, their number. Solely, there was practically no effect of the treatment time upon the number of asymmetrically vibrating water molecules treated for 5 and 15 min.

**Table 1.** Results of the Gaussian distribution of the  $\nu_{\text{OH}}$  in FTIR spectra of deionized water GP-treated under methane for 0 to 120 min.

Sample <sup>a</sup>	A(3494)	B(3265)	C(1633)	A/B	A/C	B/C
W	0.0642	0.3266	0.1633	0.197	0.393	2.000
W0	0.1062	0.2542	0.1424	0.418	0.746	1.785
W05	0.1776	0.2748	0.1662	0.646	1.069	1.653
W515	0.1657	0.2666	0.1617	0.622	1.025	1.649
W1530	0.1658	0.2808	0.1662	0.590	0.998	1.690
W3060	0.1539	0.2818	0.1623	0.546	0.948	1.736
W6090	0.1517	0.2753	0.1607	0.551	0.944	1.713
W90120	0.1380	0.3004	0.1696	0.459	0.814	1.771

<sup>a</sup> Deionized control water.

Figure 3 presents an approach to the Gaussian distribution of the FTIR-ATR spectra presented in Figure 2.

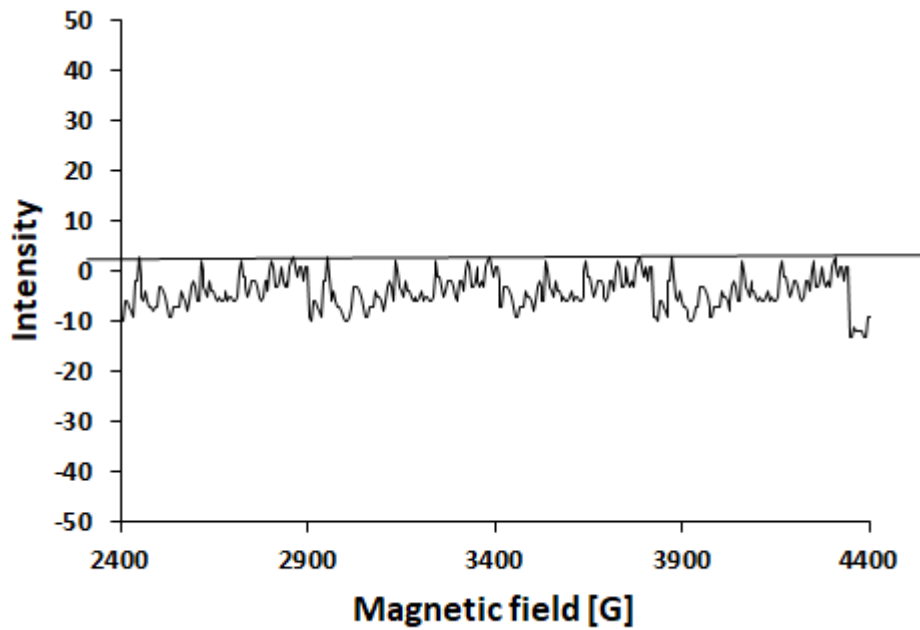
**Figure 3.** Notation of the bands taken under consideration in the Gaussian distribution of the  $\nu_{\text{OH}}$  band.

Regardless of the treatment time of water with GP, its ESR spectra demonstrated no free radical signals (Figure 4).

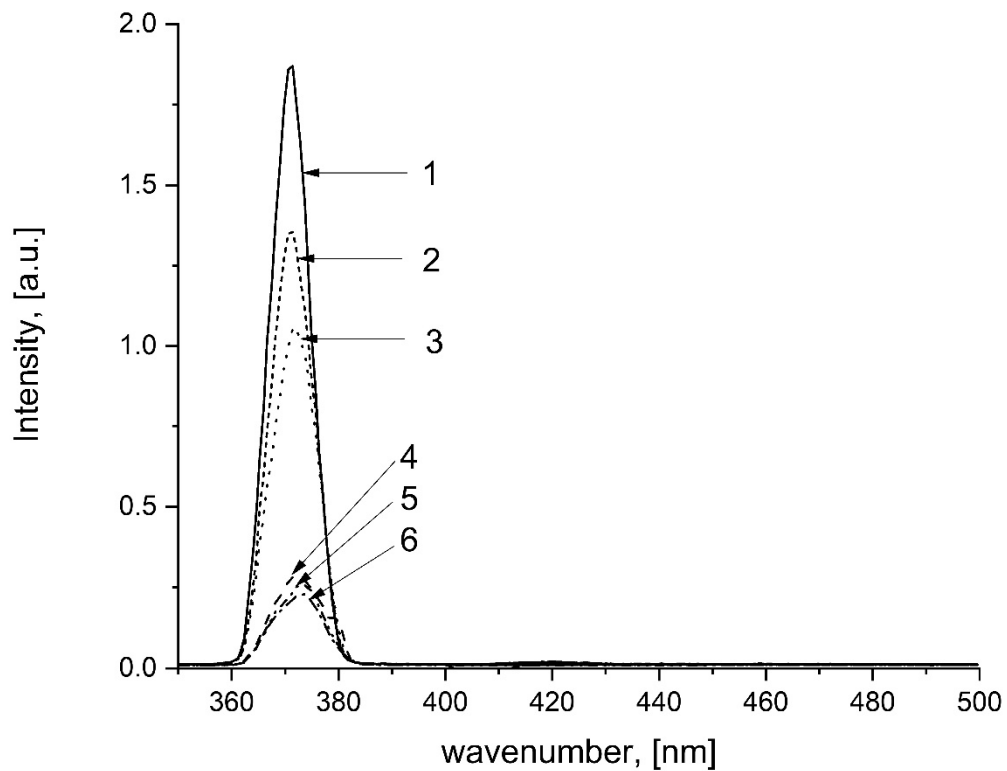
In the case of water treated with GP in the air [1], under ammonia [8], and nitrogen [9], based on the Raman spectra, the formation of aqueous clathrates of the corresponding gases was postulated. The appearance of the bands centered around 400 nm ( $5303 \text{ cm}^{-1}$  Raman shift) following the intensive band at around 370 nm ( $3276 \text{ cm}^{-1}$  Raman shift) was the criterion for the clathrate formation.

Except for the water treated for 15 min, the Raman band of the treated water located at 373 nm (Figure 5) was not followed by any longer wavelength band. In that exceptional spectrum, a longer wavelength shoulder could be observed. The corresponding  $C_p$  value was another sign of a certain specificity of so treated water.

The intensity of the 373 nm band of the scattering spectrum (the corresponding Raman shift of  $3494 \text{ cm}^{-1}$ ) was sensitive to the GP treatment time and that dependence on the treatment time was non-linear. The intensity of that band decreased in the order:  $W0 > W120 > W30 > W15 > W90 > W60$ .



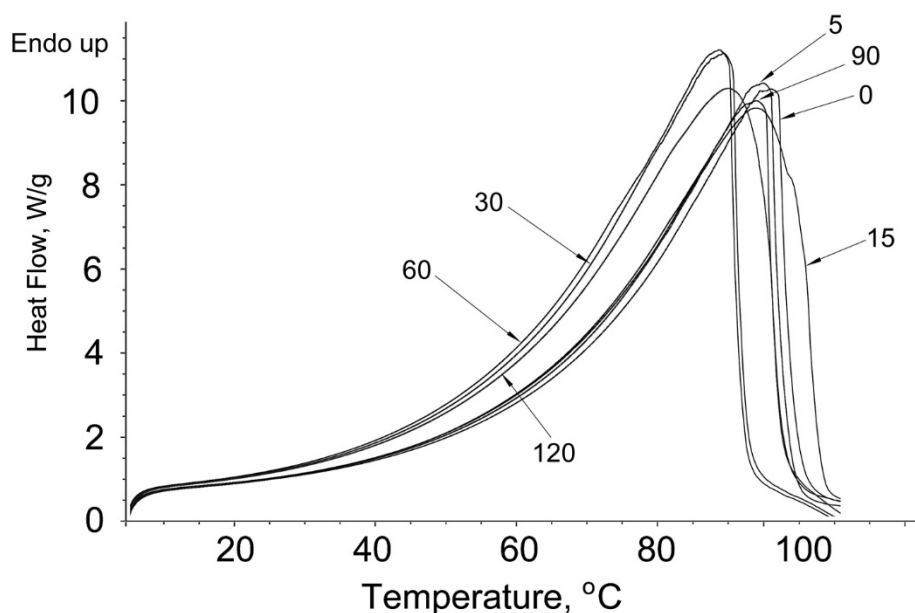
**Figure 4.** Electron Spin Resonance (ESR) spectra of the samples of water saturated with methane. The spectra of such water after the GP treatment regardless the GP treatment time are identical.



**Figure 5.** The scattering spectrum included Raman band of water treated with GP under methane for 0 to 120 min. Excitation by  $330 \pm 1$  nm/ 2 nm slit at 22 °C. Notation: 1, non-treated water saturated with methane (control); 2–6, water saturated with methane treated for 120, 30, 15, 90, and 60 min, respectively.

Saturation with methane and GP treatment did not influence the original pH and conductivity of water. Similarly, the refractive index ( $RI = 1.3314$ ) remained unchanged. DSC measurements revealed that GP had a minor effect on temperature of the principal endothermal effect of water saturated with methane (Figure 6 and Table 2). The temperature measured at maximum was always lower than that

for untreated water. It could be interpreted as a result of a transformation of primary macrostructure into another one containing objects of smaller size more readily leaving the liquid phase. Similarly as in the spectral measurements presented above, DSC measurements also pointed to a specific structure of water treated with GP for 15 min. The corresponding DSC graph (Figure 6) demonstrated an additional shoulder located at about 103 °C.



**Figure 6.** Differential scanning calorimetry (DSC) diagrams of the water saturated with methane treated with GP for 0 to 120 min. Speed: 10 °C/min.

**Table 2.** Characteristics of the differential scanning calorimetry (DSC) curves.

Sample <sup>a</sup>	Temperature at Peak (°C)	Heat Flow Intensity (W/g)	Heat of Vaporization (J/g)
Standard <sup>b</sup>	100.00 <sup>c</sup>	n.d. <sup>d</sup>	2257.0 <sup>c</sup>
WM0	96.11	−10.26	2057.1
WM5	94.81	−10.48	2050.3
WM15	94.05	−9.82	2049.9
WM30	89.37	−11.20	2132.2
WM60	88.86	−11.23	2154.1
WM90	94.11	−10.00	2048.4
WM120	90.11	−10.26	2106.3

<sup>a</sup> Numbers following W (water saturated with methane) correspond to the time (min) of the GP treatment;

<sup>b</sup> Deionized, methane-free, non-treated water; <sup>c</sup> Data for pure water at 1013 mbar; <sup>d</sup> Not determined.

The increase in the endothermal effect did not parallel the temperature changes. The sample treated for 60 min exhibited the most endothermal effect. Depending on the treatment time it declined in the order:

$$\text{standard} > \text{non-treated water} > 90 > 5 > 15 > 120 > 30 > 60 \text{ min}$$

Non-linear changes of both parameters against the treatment of the samples with GP reflected building different, treatment time-dependent macrostructures. Such effects were also observed in our studies on water treated with GP in the air [1], ammonia [8], and nitrogen [9], as well as computations performed for corresponding aqueous clathrates [25].

The thermodynamic data for GP-treated deionized water saturated with methane (Table 3) also non-linearly changed against the treatment time. That showed that the treatment for 15 and 60 min required the lowest and the highest specific heat,  $C_p$ , of the thermal transition, respectively. The saturation of the water with methane significantly increased the enthalpy of the thermal transition already prior to the treatment with GP. The treatment influenced also the enthalpy in the manner



irregular against the time. The lowest enthalpy was noted for the water treated for 30 min. The resulting preparation (WM30) disposed with the highest entropy ( $\Delta S$ ).

**Table 3.** Thermodynamic data for deionized water saturated with methane (WM) treated with GP for 0 to 120 min.

Sample <sup>a</sup> (Immediately)	$C_p$ (J/g.K)	$\Delta H$ (J/g)	$\Delta S$ (J/g.K)
W	34.377	−1770	4.7690
WM0	35.837	−1346	3.6266
WM5	32.980	−1451	3.9095
WM15	26.107	−1487	4.0065
WM30	33.935	−1845	4.9710
WM60	36.378	−1765	4.7555
WM90	31.559	−1788	4.8175
WM120	32.847	−1742	4.6935

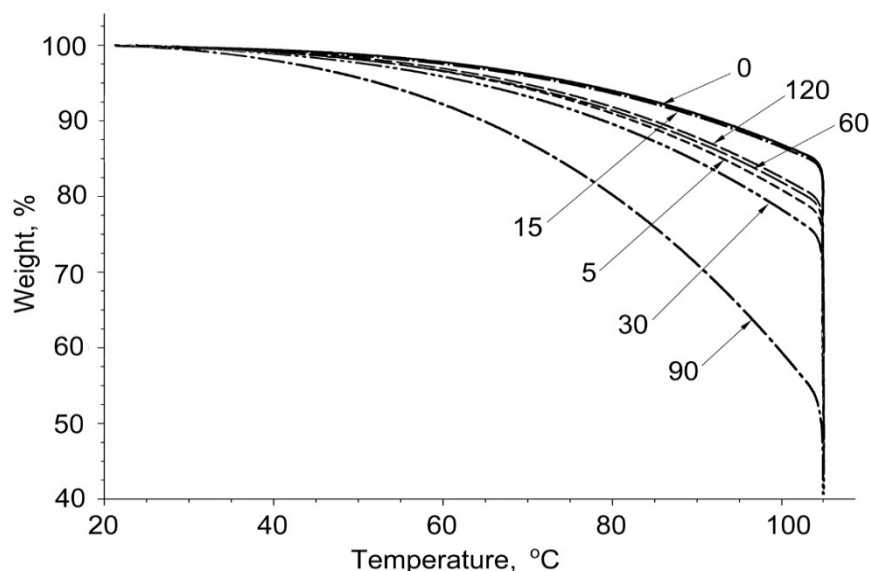
<sup>a</sup> W—control pure deionized water.

Treatment time-dependent enthalpy values ( $\Delta H$ ) declined in the same order as did entropy values, that is:

$$30 > 90 > \text{methane-free water} > 60 > 120 > 15 > 5 > 0 \text{ min}$$

Thermogravimetric measurements threw a light on the solubility of methane in water depending on the treatment time (Figure 7). The thermogravimetric curves did not provide any signal related to the temperature-dependent evolution of methane from particular samples. Instead, clear differences in the sudden loss of the weight above 100 °C could be seen and the magnitude of that loss for depending on their GP treatment time increased in the order:

$$90 > 30 > 5 > 60 > 120 > 15 > 0 \text{ min}$$



**Figure 7.** Thermograms of water treated with GP for 0 to 120 min measured after 10 days. Speed 10 °C/min.

It meant that until the boiling point of those samples was reached the samples lost consecutively 46.0, 24.7, 22.5, 20.6, 19.5, 16.0, and 15.5% of their weight.

Certainly, the observed weight loss could not be directly related to the loss of dissolved methane. The order shift

$$15 > 120 > 5 > \text{methane-free water} > 90 > 0 \text{ (non-treated water)} > 60 > 30 \text{ min}$$

reflected, first of all, the degree of deterioration of the water macrostructure.

However, this order did not fall in line with the order of the thermal effects observed in the DSC measurements (Figure 5 and Table 1). The orders of the treatment time-dependent thermal effects

$$15 < 0 < 5 < 120 < 90 < 30 < 60$$

and their intensity

$$30 < 60 < 0 < 90 < 5 < 120 < 15$$

differed from one another. Such incompatibility of those orders supported the assumption that the aqueous solubility of methane changed depending on the GP treatment time. These orders could speak in favor of arresting methane molecules in some niches of the water macrostructures.

#### 4. Conclusions

Removal of the air dissolved in water by blowing with methane increased the number of the water molecules vibrating asymmetrically. The GP treatment of water for over 5 min gradually fairly regularly against the treatment time decreased their number against the treatment time. The formed macrostructure of water presented a dynamic system which developed smaller size units more readily leaving the liquid phase. The process involved much higher energy for defeating forces binding these newly formed structures with the surrounding macrostructure. The most remarkable changes of the macrostructure were observed in water treated with GP for 15 min. The pattern of the Raman spectrum suggested the formation of specific aqueous clathrates of methane.

**Author Contributions:** A.C. run pH, conductivity measurements, W.C. run ESR spectra, K.K. run UV-VIS and Infrared Spectra, H.K. run Raman spectra, refractive index, D.K. Thermogravimetry measurements, Z.O. equipped a research team in nanowater, P.T. invented the project, coordinated study and designed the text of this report, J.A.S. run substantial remarks and discussion of the subject. All authors jointly participated in interpretation of all data and in writing report. All authors have read and agreed to the published version of the manuscript.

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