





High-Efficiency Catalysis of Peroxymonosulfate by MgO for the Degradation of Organic Pollutants

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Figure S1. X-ray diffraction patterns of three kinds of MgO.



Figure S2. X-ray diffraction patterns of MgO before and after reaction.



Figure S3. Catalytic degradation of various organic dyes in PMS/MgO system.



Figure S4. Effects of the radical scavengers on MB degradation (**a**) tert- butanol, (**b**) methanol (TBA), (**c**) SOD and (**d**) furfuryl alcohol. Control conditions: (MB) = 100 mg/L, (PMS) = 0.8 g/L, (MgO) = 0.4 g/L, pH = 7, T = 25 °C.



Figure S5. Effect of dissolved oxygen on the degradation of MB by PMS/MgO system. Control condition: (MB) = 100 mg/L, (PMS) = 0.8 g/L, (MgO) = 0.4 g/L, T = 25 °C, pH = 7.



Figure S6. The effect of pH on Mg^{2+} concentration in solution.



Figure S7. Differences in catalytic effects of MgO with or without calcination in argon. Control conditions: (MB) = 100 mg/L, (PMS) = 0.8 g/L, (MgO) = 0.4 g/L, T = 25 °C, pH = 7.



Figure S8. The effect of the calcined MgO with different sizes on MB degradation.



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