

Article

Study on Phase Transfer Catalyst Used in the Synthesis of Sodium Isobutyl Xanthate

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Abstract: Xanthates are by far the most widely used collectors in the froth flotation beneficiation of sulfide ores. However, the xanthate production process suffers from low yield, low productivity, long reaction time and environmental pollution. To address these issues, an effective method was developed for the synthesis of xanthates using phase transfer catalyst. Sodium isobutyl xanthate was synthesized from isobutyl alcohol ((CH₃)₂CHCH₂OH), sodium hydroxide (NaOH) and carbon disulfide (CS₂) with dichloromethane (CH₂Cl₂) as solvent and cetyltrimethylammonium bromide (CTAB), cetyltrimethylammonium chloride (CTAC), tetrabutylammonium bromide (TBAB) and tetrabutylammonium chloride (TBAC) as phase transfer catalyst. The compound was characterized by elemental analysis, infrared spectrum, ¹H NMR and ¹³C NMR. The influencing factors on the content and yield of sodium isobutyl xanthate including phase transfer catalyst type, phase transfer catalyst dosage and reaction time were studied by single-factor experiments. The influencing factors on the product purity and yield including reaction temperature, solvent volume, material ratio and rotating speed were studied by orthogonal experiments. The results showed that when the amount of TBAC was 3.0% *m*_{alcohol}, the reaction temperature was 35 °C, the solvent volume was 3.5 *V*_{alcohol}, the rotating speed was 180 rpm, the reaction time was 4 h and the material ratio was n[(CH₃)₂CHCH₂OH]:n(NaOH):n(CS₂) = 1:1:1.10, the product yield could be up to 86.66% and the product purity reached 82.56%.

Keywords: phase transfer catalyst; sodium isobutyl xanthate; solvent method; orthogonal experiment; weight matrix



Citation: Ma, Z.; Cheng, L.; Weng, X.; Gao, Y.; Huang, J. Study on Phase Transfer Catalyst Used in the Synthesis of Sodium Isobutyl Xanthate. *Minerals* **2021**, *11*, 850. <https://doi.org/10.3390/min11080850>

Academic Editors: Saeed Chehreh Chelgani and Reza Dehghan

Received: 29 June 2021
Accepted: 4 August 2021
Published: 6 August 2021

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1. Introduction

Sulfide ores are main mineral resource of non-ferrous metals and are widely used to extract elements such as copper, lead, zinc, molybdenum and nickel [1–6]. Demand for non-ferrous metals is growing in response to continuous advances in science and technology [7,8]. Currently, most non-ferrous metals are extracted from sulfide mineral, which is recovered from sulfide ores by flotation [9–11].

Typically, xanthates are used as sulfide mineral collectors [12–15]. Research into the synthesis technology of xanthates is important for improving the quality and reducing the cost [16]. There are various synthesis methods of xanthates, the difference lies in the proportion and sequence of the addition, the type of medium or solvent, the reaction equipment and the stirring method. The synthesis methods mainly include: direct synthesis method, crystallization method, solvent method and wet alkali method. Solvent synthesis of xanthates does not require an alkali-making system, the sodium hydroxide added in the synthesis process is easy to disperse, the solvent can remove the reaction heat in time and the conversion rate of alcohol is high [17]. However, there have been few studies that examine the involvement of phase transfer catalysts in the synthesis process.

Compared to other available collectors such as dithiocarbamate, xanthates are the most widely used as flotation reagent, the maximum consumption of collectors and the

most economical sulfide ores collectors [18,19]; a medium-sized producer in China can sell 30,000 tons per year domestically and internationally [20]. Xanthates (R-O-CSSMe) have been shown to have weaker selectivity and better collectivity for sulfide minerals than dithiophosphate [21–24]. The xanthate molecule is composed of a non-polar group and a polar group. The non-polar group plays a hydrophobic role, and the polar group plays a role of immobilization [25,26]. This research has led to reduce the reaction time and improve the quality of xanthates such as ethyl xanthate, butyl xanthate and pentyl xanthate as collectors for flotation recovery of sulfide minerals.

Synthesis of xanthates is a complex physical and chemical process in which the phase transfer catalyst plays an essential role [27–30]. Adding the phase transfer catalyst is a useful approach for the longer reaction time of xanthate. Other synthesis methods have been proven to be inadequate, they take longer reaction time to obtain the same product quality [31]. There have been a number of investigations of synthesis processes for xanthates by adding the phase transfer catalysis [32]. In each case conditions, including a suitable phase transfer catalyst, reaction temperature, material ratio, solvent volume and reaction time, have been established for the synthesis of the desired product and depression of unwanted product.

In our study, we used CTAB, CTAC, TBAB and TBAC as phase transfer catalyst. We investigated the influence of the type and dosage of phase transfer catalyst and reaction time on the synthesis process. Further, we studied the internal connection and influence of reaction temperature, solvent volume, rotation speed and material ratio in the synthesis process of sodium isobutyl xanthate by the orthogonal experiment method and matrix analysis method.

2. Materials and Methods

2.1. Materials

2.1.1. Reagents

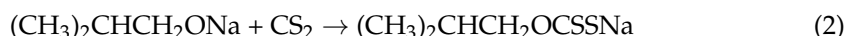
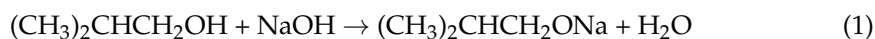
Analytical-grade isobutyl alcohol ((CH₃)₂CHCH₂OH), carbon disulfide (CS₂) and sodium hydroxide (NaOH) were used to synthesize sodium isobutyl xanthate. Analytical-grade CTAB, CTAC, TBAB and TBAC are used as phase transfer catalyst.

2.1.2. Instruments

The four-necked flask, water bath, spiral condenser, rotary evaporator, thermometer and electronic balance were used as synthetic instruments. The Bruker Advance III 400 (Karlsruhe, Germany) nuclear magnetic resonance spectrometer (TMS was internal standard, CDCl₃ was solvent), Perkin Elmer Frontier (Waltham, MA, USA) transform infrared spectrometer (KBr tablet) and Elementar Vario Unicube (Hanau, Germany) elemental analyzer were used as characterization instruments.

2.1.3. Synthesis Mechanism

The (CH₃)₂CHCH₂OH, NaOH and CS₂ were used as raw materials to synthesize sodium isobutyl xanthate, the reaction could be written as follows:



The reaction, which was exothermic, for the synthesis of sodium isobutyl xanthate was very fast. During the synthesis process, if the temperature was too high, some side reactions would occur, as shown in the following formulas:



The orange-red by-product generated by the side reaction affected the quality of sodium isobutyl xanthate.

In the synthesis process of sodium isobutyl xanthate, firstly $(\text{CH}_3)_2\text{CHCH}_2\text{OH}$ reacted with NaOH to form $(\text{CH}_3)_2\text{CHCH}_2\text{ONa}$ on the surface of granular sodium hydroxide; $(\text{CH}_3)_2\text{CHCH}_2\text{ONa}$ continued to collide with CS_2 at the solid-liquid phase interface to react, reacting to produce $(\text{CH}_3)_2\text{CHCH}_2\text{OCSSNa}$. $(\text{CH}_3)_2\text{CHCH}_2\text{ONa}$ had a low solubility in isobutyl alcohol and was insoluble in dichloromethane. The speed of peeling $(\text{CH}_3)_2\text{CHCH}_2\text{ONa}$ from the surface of the granular sodium hydroxide was slow, resulting in a longer reaction time. The kinetic study results showed that the process of ion pairs entering the organic phase from the solid surface was the key to the reaction speed in the solid-liquid phase reaction. After the phase transfer catalyst was added, ion exchange was carried out at the interface between the solid phase and the liquid phase. $(\text{CH}_3)_2\text{CHCH}_2\text{O}^-$ was brought into the liquid phase, which would greatly increase the chance of collision between $(\text{CH}_3)_2\text{CHCH}_2\text{O}^-$ and CS_2 , and increase the nucleophilic reaction rate, increasing the chemical reaction speed of the entire reaction process.

2.2. Methods

2.2.1. Synthesis Method

Into a 500 mL four-neck flask equipped with an electric stirrer, a spiral condenser and a thermometer, a certain amount of dichloromethane was added, the isobutyl alcohol and carbon disulfide according to a certain molar ratio was added, and the phase transfer catalyst according to a certain mass ratio was added. The four-necked flask was placed in a water bath at a set temperature, starting the stirrer to stir, adding a certain amount of granular sodium hydroxide in batches. After reacting for a certain time at a certain temperature, a yellow slurry could be obtained. After the completion of the reaction, the obtained slurry was rotary evaporated under reduced pressure with a rotary evaporator for 40 min at room temperature to obtain a dry product of sodium isobutyl xanthate.

2.2.2. Experiment Method

Firstly, the appropriate phase transfer catalyst type and amount were selected in the catalyst experiment; then the orthogonal experiments were performed on the amount of solvent, material ratio, reaction temperature and rotating speed to optimize the process parameters (the factors and levels investigated by the orthogonal experiments were shown in Table 1, the $L_9(3^4)$ orthogonal table was designed according to the parameters in Table 1 and 9 groups of orthogonal experiments were carried out, each group of orthogonal experiments was repeated 3 times and the results were averaged); finally, a single factor experiment of reaction time was carried out to systematically study the changes of xanthate content, yield and free alkali content in the presence of phase transfer catalyst.

Table 1. Synthesizing orthogonal factor level.

Level	Solvent Volume	Material Ratio	Reaction Temperature/ $^{\circ}\text{C}$	Rotating Speed/rpm
1	3.0 V_{alcohol}	1:1:1.05	25	120
2	3.5 V_{alcohol}	1:1:1.10	30	180
3	4.0 V_{alcohol}	1:1:1.10	35	240

2.2.3. Content Analysis of Sodium Isobutyl Xanthate

The method of titrating xanthate using sodium rosate as indicator and lead acetate standard solution was to calculate the content of xanthate as follows:

$$w_1 = \frac{2 \times c_1 \times V_1 \times M_1 / 1000}{5 \times m_1 / 500} \times 100\% \quad (6)$$

w_1 was the content of sodium isobutyl xanthate, %; c_1 was the concentration of lead acetate standard solution, mol/L; V_1 was the volume of lead acetate standard solution consumed

during titration, mL; M_1 was the molecular weight of sodium isobutyl xanthate, g/mol; m_1 was the sample weight, g.

2.2.4. Content Analysis of Free Alkali

The method of titrating sodium hydroxide using phenolphthalein as indicator and acetic acid standard solution was to calculate the content of free alkali in the product as follows:

$$w_2 = \frac{c_2 \times V_2 \times M_2 / 1000}{m_2 \times 1 / 500} \times 100\% \quad (7)$$

w_2 was the content of free alkali in the product, %; c_2 was the concentration of acetic acid standard solution, mol/L; V_2 was the volume of acetic acid standard solution consumed during titration, mL; M_2 was the molecular weight of sodium hydroxide, g/mol; m_2 was the sample weight, g.

2.2.5. Multi-Target Weight Analysis

The commonly used method of multi-index orthogonal experiment is to take the average value of each index weight and adopt the average weighting method, use the test results to convert the multi-index test problem into a single-index test problem, and then use the single-index analysis method for comprehensive selection of schemes. This method ignores the differences and importance of various indicators, and the analysis results are inevitably remote. In order to solve the problem including huge calculated amount and confirming the weights unreasonably which exist in multi-target orthogonal test method, multi-target orthogonal test design was optimized by matrix analysis.

An orthogonal experiment was designed and a three-layer structure model was established based on its data structure. As shown in Table 2, the first layer was the investigation index layer, the second layer was the factor layer and the third layer was the horizontal layer. The index layer matrix M , the factor layer matrix T , the horizontal layer matrix S and the weight matrix w were established that affected the experimental index value [33,34].

$$M = \begin{bmatrix} K_{11} & 0 & 0 & 0 \\ K_{12} & 0 & 0 & 0 \\ K_{13} & 0 & 0 & 0 \\ 0 & K_{21} & 0 & 0 \\ 0 & K_{22} & 0 & 0 \\ 0 & K_{23} & 0 & 0 \\ 0 & 0 & K_{31} & 0 \\ 0 & 0 & K_{32} & 0 \\ 0 & 0 & K_{33} & 0 \\ 0 & 0 & 0 & K_{41} \\ 0 & 0 & 0 & K_{42} \\ 0 & 0 & 0 & K_{43} \end{bmatrix} \quad (8)$$

$$T = \begin{bmatrix} T_1 & 0 & 0 & 0 \\ 0 & T_2 & 0 & 0 \\ 0 & 0 & T_3 & 0 \\ 0 & 0 & 0 & T_4 \end{bmatrix} \quad (9)$$

$$S = \begin{bmatrix} S_1 \\ S_2 \\ S_3 \\ S_4 \end{bmatrix} \quad (10)$$

$$w = MTS \quad (11)$$

Table 2. Data structural of multi-target orthogonal test.

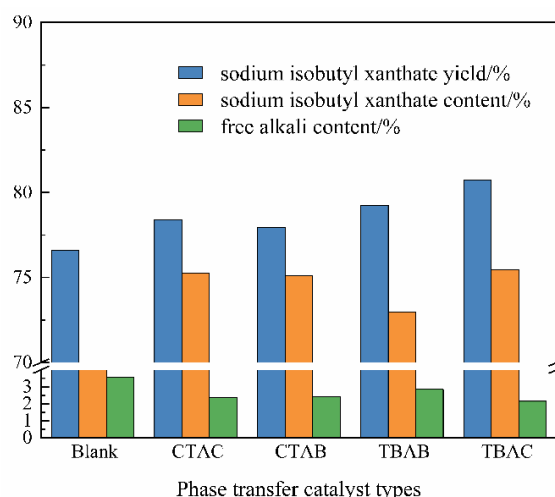
First Layer	Investigation Index											
Second layer	Factor A			Factor B			Factor C			Factor D		
Third layer	A1	A2	A3	B1	B2	B3	C1	C2	C3	D1	D2	D3

In the above matrix, $w_1 = K_{11}T_1S_1$, $K_{11}T_1$ was $K_{11} / \sum_{j=1}^m K_{ij}$, which was the ratio of the index value of the first level of factor A to the total index value of all levels of factor A; S_i was $s_i / \sum_{i=1}^l s_i$, which was the ratio of the range of factor A to the total range of all factors. The value of the product of the two could not only reflect the degree of influence of the first level of factor A on the index value, it could also reflect the magnitude of the extreme difference of factor A, as well as other factors and levels. Through calculation, the weight of the influence of each factor and level on the test results could be obtained. According to the weight, the optimal plan and the priority order of the influencing factors could be obtained.

3. Results and Discussion

3.1. The Influence of the Type of Phase Transfer Catalyst on the Synthesis Process

The amount of phase transfer catalyst was 3% m_{alcohol} , the reaction temperature was 25 °C, the solvent volume was 3.5 V_{alcohol} , the rotating speed was 120 rpm, the material ratio was $n[(\text{CH}_3)_2\text{CHCH}_2\text{OH}]:n(\text{NaOH}):n(\text{CS}_2) = 1:1:1.05$, the reaction time was 4 h, and the rotating speed was 120 rpm. Figure 1 shows that the content and yield of sodium isobutyl xanthate in the product are relatively low, and the content of free alkali is relatively high without the addition of catalyst. After adding the catalysts CTAC, CTAB, TBAB and TBAC, the content and yield of sodium isobutyl xanthate in the product are increased, and the content of free alkali is reduced. Among the four catalysts, TBAC has the greatest impact on product quality, the content and yield of sodium xanthate are relatively highest, and the content of free alkali is relatively lowest. TBAC has excellent properties of molecular symmetry and anion hydrophilicity, the ion exchange effect is relatively good, which greatly increases the chance of collision between $(\text{CH}_3)_2\text{CHCH}_2\text{O}^-$ and CS_2 , and the product purity is relatively high. The results show that the presence of the catalyst optimizes the effect of the reaction under certain conditions.

**Figure 1.** The effect of catalyst types on the content and yield of sodium isobutyl xanthate.

3.2. The Influence of the Dosage of Phase Transfer Catalyst on the Synthesis Process

The type of phase transfer catalyst was 3% m_{alcohol} , the reaction temperature was 25 °C, the solvent volume was 3.5 V_{alcohol} , the rotating speed was 120 rpm, the material ratio was $n[(\text{CH}_3)_2\text{CHCH}_2\text{OH}]:n(\text{NaOH}):n(\text{CS}_2) = 1:1:1.05$, and the reaction time was 4 h, and the rotating speed was 120 rpm. Figure 2 shows that the content and yield of sodium

isobutyl xanthate in the product are increasing, and the content of free alkali is decreasing with the increase in the amount of catalyst. When the amount of TBAC reaches 3.0% m_{alcohol} , the content and yield of sodium isobutyl xanthate in the product are relatively minimal. As the amount of TBAC increases, the chance of $(\text{CH}_3)_2\text{CHCH}_2\text{O}^-$ colliding with CS_2 gradually increases. When the amount of TBAC exceeds 3%, the ion exchange reaches equilibrium and the product purity changes smoothly. The results show that the appropriate amount of catalyst optimizes the effect of the reaction under certain conditions.

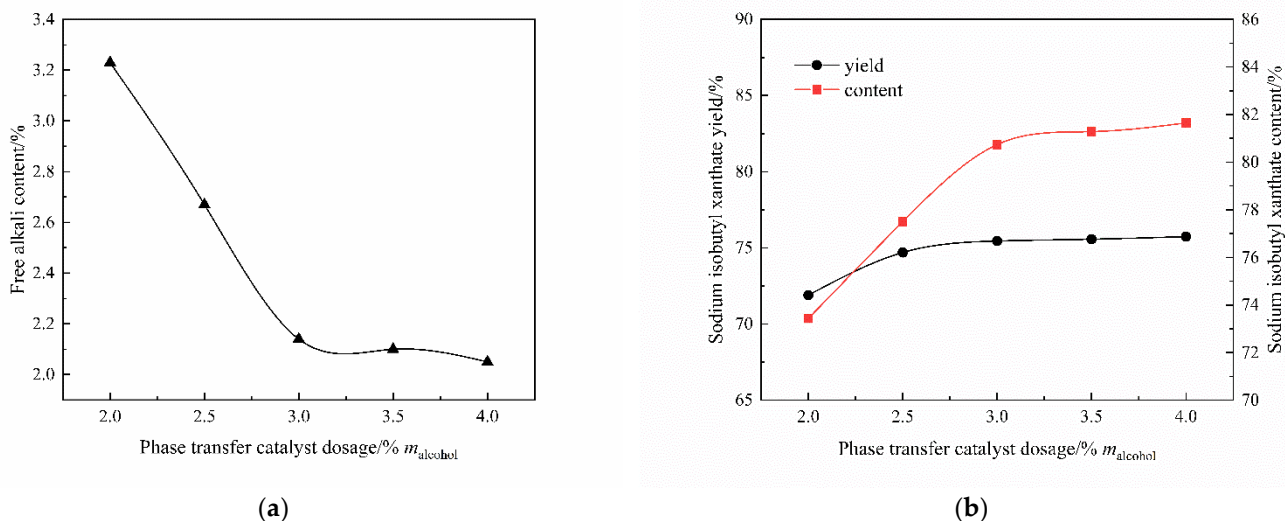


Figure 2. The effect of catalyst dosage: (a) The content of free alkali; (b) The content and yield of sodium isobutyl xanthate.

3.3. Orthogonal Test

Orthogonal experiments mainly investigated the influence of solvent dosage, material ratio, reaction temperature, rotating speed and other factors on the synthesis process of sodium isobutyl xanthate. There were 3 levels of each influencing factor in the experiment, and the factor level table was shown in Table 1. The required quality indicators were the content of sodium isobutyl xanthate (the higher the better), the yield of sodium isobutyl xanthate (the higher the better) and the content of free alkali (the lower the better). The test was a four-factor three-level experiment, so the $L_9(3^4)$ orthogonal table was selected.

Table 3, which is the intuitive analysis of the experimental results, shows that the optimal solution is A2B2C2D3 for the content of sodium isobutyl xanthate; the optimal solution is A1B3C3D2 for the yield of sodium isobutyl xanthate; the optimal solution is A2B2C3D2 for the content of free alkali. Which one of these three options to choose, usually needs to be solved by the comprehensive balance method, comprehensive scoring method and drawing method. If the matrix analysis method is used to calculate the weight matrices of the three inspection indicators that affect the test results, the optimal solution can be quickly obtained.

Table 3. Orthogonal experimental results.

No.	Solvent Volume	Material Ratio	Reaction Temperature/°C	Rotating Speed/rpm	Xanthate Content/%	Xanthate Yield/%	Free Alkali Content/%
1	3.0 V_{alcohol}	1:1:1.05	25	120	73.39	72.69	4.15
2	3.0 V_{alcohol}	1:1:1.10	30	180	76.86	81.39	0.28
3	3.0 V_{alcohol}	1:1:1.15	35	240	76.71	81.54	0.27
4	3.5 V_{alcohol}	1:1:1.05	30	240	78.32	80.14	0.56
5	3.5 V_{alcohol}	1:1:1.10	35	120	76.70	78.14	0.07
6	3.5 V_{alcohol}	1:1:1.15	25	180	76.63	73.75	2.73
7	4.0 V_{alcohol}	1:1:1.05	35	180	75.99	78.38	0.09
8	4.0 V_{alcohol}	1:1:1.10	25	240	74.61	71.76	4.00
9	4.0 V_{alcohol}	1:1:1.15	30	120	74.61	76.20	1.79
K1	226.96	227.70	224.63	224.70	The content of sodium isobutyl xanthate Visual analysis		
K2	231.65	228.17	229.79	229.48			
K3	225.21	227.95	229.40	229.64			
Range	6.44	0.47	5.16	4.94			
Level	A2	B2	C2	D3			
K1	235.62	231.21	218.20	227.03	The yield of sodium isobutyl xanthate Visual analysis		
K2	232.03	231.29	237.73	233.52			
K3	226.34	231.49	238.06	233.44			
Range	9.28	0.28	19.86	6.49			
Level	A1	B3	C3	D2			
K1	4.70	4.80	10.88	6.01	The content of free alkali Visual analysis		
K2	3.36	4.35	2.63	3.10			
K3	5.88	4.79	0.43	4.83			
Range	2.52	0.45	10.45	2.91			
Level	A2	B2	C3	D2			

The first inspection index was the content of isobutyl xanthate sodium. By the matrix analysis method, the calculation process of the weight matrix w_1 as follows (the other two indicators were similar):

$$M_1 = \begin{bmatrix} 226.96 & 0 & 0 & 0 \\ 231.65 & 0 & 0 & 0 \\ 225.21 & 0 & 0 & 0 \\ 0 & 227.70 & 0 & 0 \\ 0 & 228.17 & 0 & 0 \\ 0 & 227.95 & 0 & 0 \\ 0 & 0 & 224.63 & 0 \\ 0 & 0 & 229.79 & 0 \\ 0 & 0 & 229.40 & 0 \\ 0 & 0 & 0 & 224.70 \\ 0 & 0 & 0 & 229.48 \\ 0 & 0 & 0 & 229.64 \end{bmatrix} \quad T_1 = \begin{bmatrix} \frac{1}{68382} & 0 & 0 & 0 \\ 0 & \frac{1}{68382} & 0 & 0 \\ 0 & 0 & \frac{1}{68382} & 0 \\ 0 & 0 & 0 & \frac{1}{68382} \end{bmatrix} \quad S_1 = \begin{bmatrix} \frac{6.44}{17.01} \\ \frac{0.47}{17.01} \\ \frac{5.16}{17.01} \\ \frac{4.94}{17.01} \end{bmatrix}$$

$$w_2 = M_2 T_2 S_2 = \begin{bmatrix} 0.08774 \\ 0.08640 \\ 0.08428 \\ 0.00260 \\ 0.00260 \\ 0.00260 \\ 0.17389 \\ 0.18945 \\ 0.18971 \\ 0.05912 \\ 0.06081 \\ 0.06079 \end{bmatrix} \quad w_3 = M_3 T_3 S_3 = \begin{bmatrix} 0.04825 \\ 0.06750 \\ 0.03857 \\ 0.00887 \\ 0.00979 \\ 0.00889 \\ 0.02102 \\ 0.08697 \\ 0.53193 \\ 0.04260 \\ 0.08259 \\ 0.05301 \end{bmatrix}$$

The total weight matrix of the indicators examined in this orthogonal experiment was the average of the weight matrices of the three indicator values. The calculation process was as follows:

$$w = \frac{w_1 + w_2 + w_3}{3} = \frac{1}{3} \times \begin{bmatrix} 0.12566 \\ 0.12825 \\ 0.12469 \\ 0.00920 \\ 0.00922 \\ 0.00921 \\ 0.09965 \\ 0.10194 \\ 0.10176 \\ 0.09543 \\ 0.09746 \\ 0.09753 \end{bmatrix} + \begin{bmatrix} 0.08774 \\ 0.08640 \\ 0.08428 \\ 0.00260 \\ 0.00260 \\ 0.00260 \\ 0.17389 \\ 0.18945 \\ 0.18971 \\ 0.05912 \\ 0.06081 \\ 0.06079 \end{bmatrix} + \begin{bmatrix} 0.04825 \\ 0.06750 \\ 0.03857 \\ 0.00887 \\ 0.00979 \\ 0.00889 \\ 0.02102 \\ 0.08697 \\ 0.53193 \\ 0.04260 \\ 0.08259 \\ 0.05301 \end{bmatrix} = \begin{bmatrix} 0.08722 \\ 0.09405 \\ 0.08251 \\ 0.00689 \\ 0.00720 \\ 0.00690 \\ 0.09819 \\ 0.12612 \\ 0.27447 \\ 0.06572 \\ 0.08029 \\ 0.07044 \end{bmatrix} = \begin{bmatrix} A_1 \\ A_2 \\ A_3 \\ B_1 \\ B_2 \\ B_3 \\ C_1 \\ C_2 \\ C_3 \\ D_1 \\ D_2 \\ D_3 \end{bmatrix}$$

Figure 3 shows that the three indicators of sodium isobutyl xanthate content, sodium isobutyl xanthate yield and free alkali content affect the synthesis process as follows: C > A > D > B; factors A2, B2, C3 and D2 have the largest weights. The optimal solution of the orthogonal experiment is A2B2C3D2, that is, the amount of TBAC is 3% $m_{alcohol}$, the volume of methylene chloride is 3.5 $V_{alcohol}$, the material ratio is $n[(CH_3)_2CHCH_2OH]:n[NaOH]:n[CS_2] = 1:1:1.10$, the reaction temperature is 35 °C, and the rotating speed is 180 rpm, the best synthesis effect can be obtained.

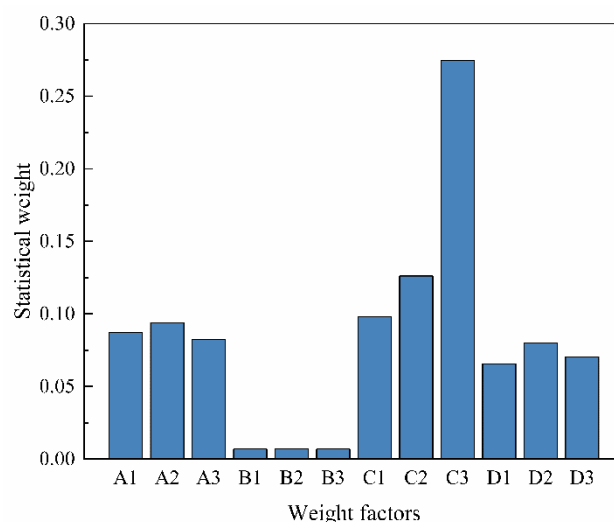


Figure 3. Comprehensive statistical weight of each weight factor.

3.4. The Influence of the Time on the Synthesis Process

The amount of TBAC was 3.0% m_{alcohol} , the reaction temperature was 35 °C, the solvent volume was 3.5 V_{alcohol} , the rotating speed was 180 rpm, reaction time was 4 h and the material ratio was $n[(\text{CH}_3)_2\text{CHCH}_2\text{OH}]:n(\text{NaOH}):n(\text{CS}_2) = 1:1:1.10$. Figure 4 shows that the content and yield of sodium isobutyl xanthate in the product present an overall upward trend with the reaction time increases, and the content of free alkali presents an overall downward trend; after the reaction time reaches 4 h, the content and yield of sodium isobutyl xanthate in the product present an overall gentle downward trend, and the overall change of free alkali content present a stable trend. In the heterogeneous catalytic synthesis system of sodium isobutyl xanthate, when the reaction time exceeds 4 h, the reaction is close to equilibrium, side reactions gradually intensify and the product purity slightly decreases. The results show that the optimal reaction time optimizes the effect of the synthesis under certain conditions.

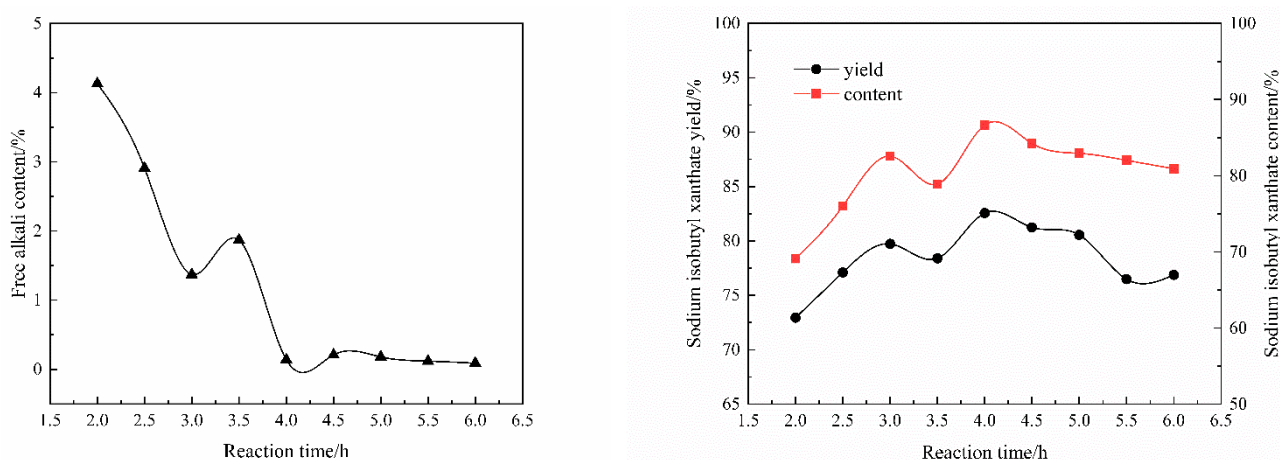


Figure 4. The effect of reaction time: (a) The content of free alkali; (b) The content and yield of sodium isobutyl xanthate.

3.5. Repetitive Experiment

Under suitable synthetic process conditions, the amount of TBAC was 3% m_{alcohol} , the solvent volume was 3.5 V_{alcohol} , the material ratio was $n[(\text{CH}_3)_2\text{CHCH}_2\text{OH}]:n(\text{NaOH}):n(\text{CS}_2) = 1:1:1.10$, the reaction temperature was 35 °C, the rotating speed was 180 rpm, the reaction time was 4 h and the reproducibility experiment was performed. Table 4 shows that the results of the 4 verification experiments are similar, indicating that the phase transfer catalyst used in the synthesis of sodium isobutyl xanthate has good reproducibility and strong operability.

Table 4. Reproducible experimental results.

No.	Xanthate Content/%	Xanthate Yield/%	Free Alkali Content/%
1	82.44	86.21	0.15
2	82.47	86.08	0.16
3	82.56	86.66	0.14
4	82.65	86.32	0.14
Average	82.53	86.32	0.15

3.6. Product Characterization Analysis

The measured value of the elemental analysis of the product (%), calculated value of $\text{C}_2\text{H}_9\text{OS}_2\text{Na}$: C 32.81 (34.89), H 5.04 (5.27), S 35.17 (37.24).

^1H NMR (400 MHz, CDCl_3), δ : 4.23 (d, 2H, CH_2), 2.18–2.09 (m, 1H, CH), 0.88 (s, 6H, CH_3). ^{13}C NMR (400 MHz, CDCl_3), δ : 229.8 (s, O-C(=S)-S), 79.6 (s, O- CH_2), 26.0 (s, CH) and 19.2 (s, CH_3).

The infrared spectrum of the product was measured with KBr tablet in the range of 400–4000 cm^{-1} . Figure 5 shows that the weak peak at 630.57 cm^{-1} is the vibration peak of C-S; the peak at 1061.83 cm^{-1} is the absorption peak of C=S; the peak at 1111.95 cm^{-1} is the stretching peak of C-O-C; the peak at 1371.94 cm^{-1} is the absorption peak of CH_3 ; the peak at 1464.25 cm^{-1} is the absorption peak of CH_2 ; the peak at 2873.00 cm^{-1} is the stretching vibration peak of CH_3 ; the peak at 2960.26 cm^{-1} is the stretching vibration peak of CH_3 .

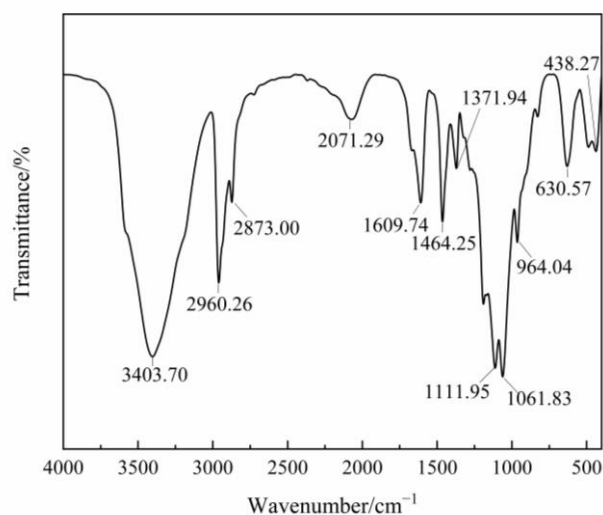


Figure 5. The infrared spectrum of the product.

The above detections and structural characterizations confirmed that the obtained compound was the target product.

4. Conclusions

The experimental results of the phase transfer catalyst showed that the addition of the phase transfer catalyst accelerated the reaction speed and shortened the reaction time during the synthesis of sodium isobutyl xanthate under certain conditions. When the amount of TBAC was 3% m_{alcohol} , the catalytic effect was the best, the xanthate content and yield were the highest and the free alkali content was the least.

Orthogonal experimental matrix analysis results showed that the order of the degree of influence of each factor was: reaction temperature > solvent volume > rotating speed > material ratio. The appropriate process conditions for TBAC to participate in the synthesis of sodium isobutyl xanthate were: the volume of the dichloromethane was 3.5 V_{alcohol} , the material ratio was $n[(\text{CH}_3)_2\text{CHCH}_2\text{OH}] : n(\text{NaOH}) : n(\text{CS}_2) = 1 : 1 : 1.10$, the synthesis time was 4 h, the synthesis temperature was 35 °C and the reaction speed was 180 rpm.

The results of repeated experiments showed that the content of sodium isobutyl xanthate in the product was up to 82.53%, the yield reached 86.32% and the content of free alkali was less than 0.15%. It was a qualified product. This synthesis process of sodium isobutyl xanthate did not require a refrigeration system, only used cooling water below 15 °C, so the synthesis process had low power consumption; and the addition of a phase transfer catalyst shortened the reaction time, saved energy and reduced the production cost of sodium isobutyl xanthate.

Author Contributions: Validation, Z.M. and L.C.; resources, X.W. and Y.G.; writing—original draft preparation, L.C. and J.H. All authors have read and agreed to the published version of the manuscript.

Funding: This research received no external funding.

Institutional Review Board Statement: Not applicable.

Informed Consent Statement: Not applicable.

Data Availability Statement: The data presented in this study are available in this article.

Conflicts of Interest: The authors declare no conflict of interest.

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