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Effect of Heating Rate on Pyrolysis Behavior and Kinetic Characteristics of Siderite

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Abstract: The pyrolysis characteristics of siderite at different heating rates under the neutral atmosphere were investigated using various tools, including comprehensive thermal analyzer, tube furnace, X-ray diffraction (XRD), scanning electron microscope (SEM), energy-dispersive spectrometry (EDS) and vibrating specimen magnetometer (VSM) measurements. The reaction of siderite pyrolysis followed the one-step reaction under the neutral atmosphere: $\text{FeCO}_3 \rightarrow \text{Fe}_3\text{O}_4 + \text{CO}_2 + \text{CO}$. As the increasing of heating rate, the start and end pyrolysis temperatures and temperature where maximum weight loss rate occurred increased, while the total mass loss were essentially the same. Increasing heating rate within a certain range was in favor of shortening the time of each reaction stage, and the maximum conversion rate could be reached with a short time. The most probable mechanism function for non-isothermal pyrolysis of siderite at different heating rates was $A_{1/2}$ reaction model (nucleation and growth reaction). With increasing heating rate, the corresponding activation energies and the pre-exponential factors increased, from 446.13 to 505.19 $\text{kJ}\cdot\text{mol}^{-1}$, and from 6.67×10^{-18} to 2.40×10^{-21} , respectively. All siderite was transformed into magnetite with a porous structure after pyrolysis, and some micro-cracks were formed into the particles. The magnetization intensity and specific susceptibility increased significantly, which created favorable conditions for the further effective concentration of iron ore.

Keywords: siderite; neutral atmosphere; heating rate; pyrolysis; mechanism

1. Introduction

China is rich in siderite resources, which are difficult to exploit by traditional beneficiation methods. The development of efficient utilization technology for siderite can increase the reserves of available iron ore, improve the tight situation of iron ore supply, and ensure the iron ore security of China. Magnetization roasting followed by magnetic separation was the most effective technology for the exploitation and utilization of siderite resources [1–4]. During the process, the siderite could be transformed to a highly magnetic mineral phase, while the magnetism of gangue minerals was barely changed, benefiting magnetic separation of the iron minerals and gangue minerals effectively [5–9]. The pyrolysis behavior and kinetic characteristics are important to better understand and further improve the magnetization roasting for siderite.

Many studies have focused on the pyrolysis mechanism, including the conversion process, pyrolysis product characteristics and pyrolysis kinetics in the process of siderite magnetization roasting. It was found that the phase transformation process in weak oxidizing atmosphere followed the steps as: $\text{FeCO}_3 \rightarrow \text{Fe}_3\text{O}_4 \rightarrow \gamma\text{-Fe}_2\text{O}_3$ at 550 °C, and $\text{FeCO}_3 \rightarrow \text{FeO} + \text{Fe}_3\text{O}_4 \rightarrow \text{Fe}_3\text{O}_4 \rightarrow \gamma\text{-Fe}_2\text{O}_3 \rightarrow \alpha\text{-Fe}_2\text{O}_3$ at 800 °C. In inert atmosphere, the decomposition pathway was: $\text{FeCO}_3 \rightarrow \text{Fe}_3\text{O}_4$ below 733 °C, and $\text{FeCO}_3 \rightarrow \text{FeO} + \text{Fe}_3\text{O}_4$ above 733 °C [10]. The final product was $\alpha\text{-Fe}_2\text{O}_3$ under the oxygen flow, and it was Fe_3O_4 in a CO_2 atmosphere, while the FeO and Fe_3O_4 were only observed in an

inert atmosphere [11–13]. Geva et al. [14] concluded the potassium markedly increased the reduction rate of wustite and changed the iron properties in subsequent reduction. Mn and Mg impurities existed in natural siderite as isomorphism state with iron and thereby influenced the decomposition process [15]. The chemical and physical characteristics of siderite before and after decomposition change tremendously. The phase transformation and microstructure were extremely complex related to atmosphere, temperature, heating rate, and its characteristics [16–19]. However, there has been little systematic and theoretical research on the influence of the heating rate on siderite decomposition.

The pyrolysis kinetics of siderite ore was studied that the mechanism of siderite pyrolysis process was applicable to each particle size interval, while the pyrolysis kinetics differed under different atmospheres [20]. The pyrolysis mechanism for magnesium-manganese siderite was investigated, it was demonstrated the mechanism function was F1, with activation energy of $192 \text{ kJ}\cdot\text{mol}^{-1}$ [21]. In air atmosphere, the Avrami–Erofe'ev reaction mechanism could explicate the pyrolysis kinetics of siderite, where the isothermal pyrolysis kinetics of siderite obeyed the nucleation and growth mechanism ($n = 1$), and the activation energy and the pre-exponential factor ($\ln A$) were $46.17 \text{ kJ}\cdot\text{mol}^{-1}$ and 23.02 min^{-1} , respectively [22]. The pyrolysis kinetics of siderite in suspended state from Wangjiatan of Shanxi Province in China was conducted by Feng et al. [23] under different isothermal conditions. The result indicated that the mechanism function was R3 (phase boundary reaction), and the pyrolysis process was accelerated with increasing temperature. Most researchers have studied the pyrolysis kinetics of siderite by an isothermal approach, not verified by a non-isothermal method. Moreover, the isothermal method is difficult to implement completely in the laboratory, especially at the beginning of the reaction.

Consequently, it is necessary to study the pyrolysis mechanism and kinetic characteristics of siderite at different heating rates to further understand siderite pyrolysis behavior, and may have implication for exploitation of Fe-bearing carbonates as a source of iron ore.

2. Experimental

2.1. Materials

The bulky iron ore rich in siderite from Hunan Province of China was used in this study. The sample was crushed through a roller crusher and ground to 85% passing 0.074 mm, then the pure siderite minerals was obtained via the gravity process using a shaking table and a low-intensity magnetic separator. The chemical composition (Table 1) indicated that the most valuable element was iron, with iron grade of 44.21% and FeO content up to 58.05%. It also contained 1.19% MgO and 2.53% MnO. Harmful elements such as sulfur, phosphorus and sodium were minor. The iron mineral analysis results (Table 2) showed that iron mainly existed in form of carbonate iron (96.23%), indicating that siderite was the major mineral in the sample. The sample belonged to high-quality pure siderite, which could satisfy the experiments well.

Table 1. Chemical composition of pure siderite mineral (in mass %).

Fe _{total}	FeO	SiO ₂	Al ₂ O ₃	CaO	MgO	Na	MnO	P	S	Loss of Ignition
44.21	58.05	0.55	0.13	0.30	1.19	0.01	2.53	undetected	0.34	36.04

Table 2. Iron mineral analysis of pure siderite mineral (in mass %).

Iron Phases	Iron Carbonate	Magnetite	Hematite	Iron Sulfide	Ferrosilite
Iron %	39.45	0.12	3.69	0.85	0.10
Mineral %	96.23	0.27	2.35	0.92	0.23

2.2. Methods

2.2.1. Pyrolysis Behavior Analysis of Siderite

The thermal decomposition was conducted by non-isothermal pyrolysis analysis using a comprehensive thermal analyzer (STA449-F3, Netzsch, Hamburg, Germany). Nitrogen (99.99%) was used as protection gas, which could restrain the diffusion effect of carbon dioxide from siderite pyrolysis. A 30 mg sample was used for each experiment at a nitrogen flow rate of 30 mL/min. The heating rates studied were 5, 10, 20 and 30 °C·min⁻¹. The temperature range was from room temperature to 900 °C, and the thermogravimetry (TG), derivative thermogravimetry (DTG), and differential scanning calorimetry (DSC) analysis results were obtained from computer processor.

2.2.2. Pyrolysis Kinetics Characteristics of Siderite at Different Heating Rates

The conversion fraction α based on the TG analysis results can be expressed as:

$$\alpha = \frac{m_0 - m_t}{m_0 - m_\infty} \quad (1)$$

where m_0 is the initial sample mass (g), m_t is the sample mass at a particular time (g), and m_∞ is the theoretical weight loss (g).

The conversion rate r can be calculated:

$$r = \frac{d\alpha}{dt} = \frac{\Delta\alpha}{\Delta t} \quad (2)$$

where $\Delta\alpha$ is the weight loss in a short span of time (g), and Δt is the short time interval (min).

Non-isothermal study of pyrolysis kinetics has great advantages compared with isothermal one. A non-isothermal thermal analysis curve often contains more information and function than isothermal analysis curves, which can make the analysis quickly and comprehensive. The differential isothermal kinetics equation of solid-state decomposition can be described by the following well-known general equation [24–29]:

$$\frac{d\alpha}{dt} = A \exp^{-E/RT} \cdot f(\alpha) \quad (3)$$

where E is the activation energy (kJ·mol⁻¹), A is the pre-exponential factor, T is the temperature (Kelvin, K), t is the time (min), R is the gas constant (8.314 J·(mol·K)⁻¹), and $f(\alpha)$ is the differential conversion function.

Some modifications are adopted in the non-isothermal kinetics compared with isothermal one. Introducing the heating rate ($\beta = dT/dt$, where β is heating rate, °C·min⁻¹), the reaction rate for a heterogeneous reaction are presented by the following well-known general equation [30]:

$$\frac{d\alpha}{dt} = \beta \frac{d\alpha}{dT} = A \exp^{-E/RT} \cdot f(\alpha) \quad (4)$$

By rearranging and integrating Equation (4), the integral form of mechanism function ($G(\alpha)$) can be obtained:

$$G(\alpha) = \int_0^T \left(\frac{A}{\beta} \right) \exp^{-E/RT} dT \quad (5)$$

Kinetic analysis is expected to obtain an adequate kinetic description of the process in terms of kinetic parameters (the reaction model, the activation energy, and the pre-exponential factor) [31]. Many methods have been established to evaluate these parameters. The Flynn–Wall–Ozawa (FWO) method based on integral isoconversional methods has been widely used to lucubrate thermodynamics [32]. By the

application of Doyle’s method [33,34], the activation energy (E) can be determined as a function of the temperatures corresponding to fixed values from experiments at different heating rates:

$$\lg(\beta) = \lg\left(\frac{AE}{RG(\alpha)}\right) - 2.315 - 0.4567\frac{E}{RT} \tag{6}$$

The most probable mechanism function can be determined by Šatava–Šesták integral method, which is given as follow [35]:

$$\lg G(\alpha) = \lg\left(\frac{AE_S}{R\beta}\right) - 2.315 - 0.4567\frac{E_S}{RT} \tag{7}$$

The most probable mechanism function can be determined by analyzing the linear relationship between $\lg(G(\alpha))$ and $1/T$. If only one $G(\alpha)$ meets the linear relationship, the $G(\alpha)$ is exactly the mechanism function. However, if several $G(\alpha)$ meet the linear relationship, the $G(\alpha)$, whose activation energy (E_s) is approximately equal to the activation energy (E) estimated by FWO method, is the most probable mechanism function [31]. The activation energy E and the pre-exponential factor A , and correlation coefficient R can be derived from the linear least-squares plot of $\lg(G(\alpha))$ versus $1/T$.

Siderite pyrolysis reaction is the typical gas-solid reaction. Based on previous studies, the most frequently-used reaction mechanism models of gas-solid reaction are listed in Table 3 [36–39].

Table 3. Gas–solid reaction models with corresponding mechanism and function [36–39].

Code	Reaction Model	Mechanism	Integral $G(\alpha)$	Differential $f(\alpha)$
A _{1/3}	Avrami–Erofeev ($n = 1/3$)	Nucleation and growth	$[-\ln(1 - \alpha)]^3$	$1/3(1 - \alpha)[- \ln(1 - \alpha)]^{-2}$
A _{1/2}	Avrami–Erofeev ($n = 1/2$)		$[-\ln(1 - \alpha)]^2$	$1/2(1 - \alpha)[- \ln(1 - \alpha)]^{-1}$
A _{2/3}	Avrami–Erofeev ($n = 2/3$)		$[-\ln(1 - \alpha)]^{3/2}$	$2/3(1 - \alpha)[- \ln(1 - \alpha)]^{-1/2}$
A ₁	Avrami–Erofeev ($n = 1$)		$-\ln(1 - \alpha)$	$1 - \alpha$
A _{4/3}	Avrami–Erofeev ($n = 4/3$)		$[-\ln(1 - \alpha)]^{3/4}$	$4/3(1 - \alpha)[- \ln(1 - \alpha)]^{1/4}$
A _{3/2}	Avrami–Erofeev ($n = 3/2$)		$[-\ln(1 - \alpha)]^{2/3}$	$3/2(1 - \alpha)[- \ln(1 - \alpha)]^{1/3}$
A ₂	Avrami–Erofeev ($n = 2$)		$[-\ln(1 - \alpha)]^{1/2}$	$2(1 - \alpha)[- \ln(1 - \alpha)]^{1/2}$
A _{5/2}	Avrami–Erofeev ($n = 5/2$)		$[-\ln(1 - \alpha)]^{2/5}$	$5/2(1 - \alpha)[- \ln(1 - \alpha)]^{3/5}$
A ₃	Avrami–Erofeev ($n = 3$)		$[-\ln(1 - \alpha)]^{1/3}$	$3(1 - \alpha)[- \ln(1 - \alpha)]^{2/3}$
A ₄	Avrami–Erofeev ($n = 4$)	$[-\ln(1 - \alpha)]^{1/4}$	$4(1 - \alpha)[- \ln(1 - \alpha)]^{3/4}$	
D ₁		One-dimensional diffusion	α^2	$1/2\alpha^{-1}$
D ₂	Valensi	Two-dimensional diffusion	$\alpha + (1 - \alpha) \ln(1 - \alpha)$	$[-\ln(1 - \alpha)]^{-1}$
D ₃	Jander	Three-dimensional diffusion	$[1 - (1 - \alpha)1/2]^2$	$(1 - \alpha)^{1/2}[1 - (1 - \alpha)1/2]^{-1}$
D ₄	Jander ($n = 2$)		$[1 - (1 - \alpha)1/3]^2$	$3/2(1 - \alpha)^{2/3}[1 - (1 - \alpha)1/3]^{-1}$
D ₅	Jander ($n = 1/2$)		$[1 - (1 - \alpha)1/3]^{1/2}$	$6(1 - \alpha)^{2/3}[1 - (1 - \alpha)1/3]^{1/2}$
D ₆	anti-Jander		$[(1 + \alpha)^{1/3} - 1]^2$	$3/2(1 + \alpha)^{2/3}[(1 + \alpha)^{1/3} - 1]^{-1}$
D ₇	Z-L-T		$[(1 - \alpha)^{-1/3} - 1]^2$	$3/2(1 - \alpha)^{4/3}[(1 - \alpha)^{-1/3} - 1]^{-1}$
D ₈	Ginstling–Brounshtein		$1 - 2\alpha/3 - (1 - \alpha)^{2/3}$	$3/2[(1 - \alpha)^{-1/3} - 1]^{-1}$
R _{1/2}	Contraction sphere ($n = 1/2$)	Phase boundary reaction	$1 - (1 - \alpha)^2$	$1/2(1 - \alpha)^{-1}$
R _{1/3}	Contraction sphere ($n = 1/3$)		$1 - (1 - \alpha)^3$	$1/3(1 - \alpha)^{-2}$
R _{1/4}	Contraction sphere ($n = 1/4$)		$1 - (1 - \alpha)^4$	$1/4(1 - \alpha)^{-3}$
R ₂	Contraction sphere ($n = 2$)		$1 - (1 - \alpha)^{1/2}$	$2(1 - \alpha)^{1/2}$
R ₃	Contraction sphere ($n = 3$)		$1 - (1 - \alpha)^{1/3}$	$3(1 - \alpha)^{2/3}$
R ₄	Contraction sphere ($n = 4$)		$1 - (1 - \alpha)^{1/4}$	$4(1 - \alpha)^{3/4}$
P _{2/3}	Mampel power law ($n = 2/3$)		$\alpha^{3/2}$	$2/3\alpha^{-1/2}$
P ₁	Mampel power law ($n = 1$)		α	1
P ₂	Mampel power law ($n = 2$)		$\alpha^{1/2}$	$2\alpha^{1/2}$
P ₃	Mampel power law ($n = 3$)		$\alpha^{1/3}$	$3\alpha^{2/3}$
P ₄	Mampel power law ($n = 4$)		$\alpha^{1/4}$	$4\alpha^{3/4}$
C ₁	Reaction order ($n = 2$)	Chemical reaction	$(1 - \alpha)^{-1} - 1$	$(1 - \alpha)^2$
C ₂	Reaction order ($n = 3/2$)		$(1 - \alpha)^{-1/2}$	$2(1 - \alpha)^{3/2}$
C ₃	Reaction order ($n = 3$)		$(1 - \alpha)^{-2}$	$1/2(1 - \alpha)^3$

2.2.3. The Characteristics of Siderite both before and after Pyrolysis

The pyrolysis products of siderite derived from comprehensive thermal analyzer were only tiny amounts, which was not available for further detection and analysis. The laboratory type tube furnace (Figure 1) was employed for the study of pyrolysis test. The tests were conducted on the temperature range from room temperature to 600 °C (based on TG/DTG/DSC analysis results), with the heating rate at 10 °C·min⁻¹ and the flow rate of nitrogen gas was 500 mL·min⁻¹. The phases, microstructures and magnetic properties analysis of siderite both before and after pyrolysis were investigated by X-ray diffraction (XRD, PW3040, PANalytical B.V., Almero, The Netherlands), scanning electron microscope (SEM, SSX-550, SHIMADZU, Kyoto, Japan), energy-dispersive spectrometry (EDS, SSX-550, SHIMADZU, Kyoto, Japan) and vibrating specimen magnetometer (VSM, JDAW-2000D, YingPu Magnetic Technology Development, Changchun, China) measurements.

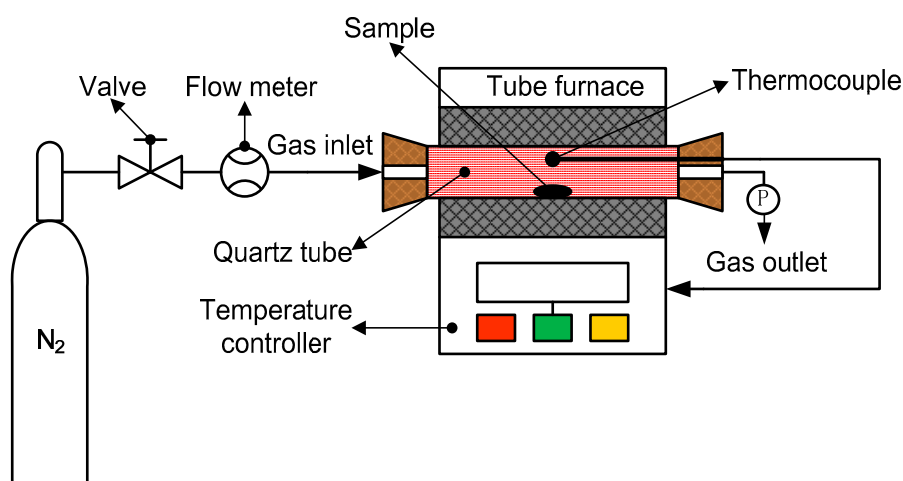
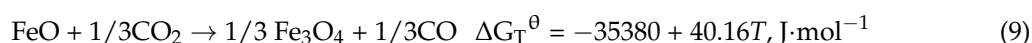
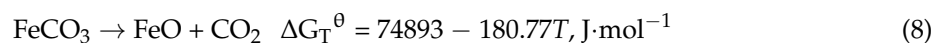


Figure 1. Schematic experimental apparatus.

3. Results and Discussion

3.1. Decomposition Behavior of Siderite

In the neutral atmosphere of pyrolysis, FeO emerged immediately during the decomposition of siderite, but the FeO, as an intermediate phase, would transform to Fe₃O₄ easily in a short period of time. The reactions can be represented as:



According to Equations (8) and (9), the pyrolysis process of siderite in the neutral atmosphere can be expressed with the following equation:



The results of the TG/DTG/DSC analyses in the neutral atmosphere with the heating rate of 10 °C·min⁻¹ are shown in Figure 2. Throughout the decomposition process of siderite, a minute mass loss (−0.83%) appeared at about 250 °C, corresponding to a tiny endothermic peak, which was most likely due to the moisture evaporation in the sample. The pyrolysis of siderite started at about 470 °C (*T*_s) with a maximum pyrolysis rate at about 500 °C (*T*_m). A significant endothermic peak was observed at 510 °C indicating that the pyrolytic reaction of siderite was an intense endothermic process, and the enthalpy change was −563.3 J·g⁻¹. The mass remained unchanged basically at about

580 °C (T_e), which suggested the completion of the pyrolytic reaction of siderite, with the total mass loss was 33.15%. The steps of mass loss on the TG curve corresponding to the peaks of the DTG/DSC curves belonged to the single and overlap-free peaks, indicating that the pyrolytic of siderite was one-step reaction (Reaction (9)).

The results of TG/DTG analysis at different heating rates are illustrated in Figure 3. It can be seen that both TG and DTG profiles shifted toward the direction of higher temperature based on the increase in thermal lag. Furthermore, the TG curves inosculated with DTG curves, and the mass loss steps of the TG curves corresponded to the peaks of the DTG curves. The starting temperature (T_s), maximum weight loss rate (T_m) and ending temperature (T_e) increased with the increasing of heating rate, and the pyrolysis temperature range became wider. However, heating rate had little effect on the total mass loss.

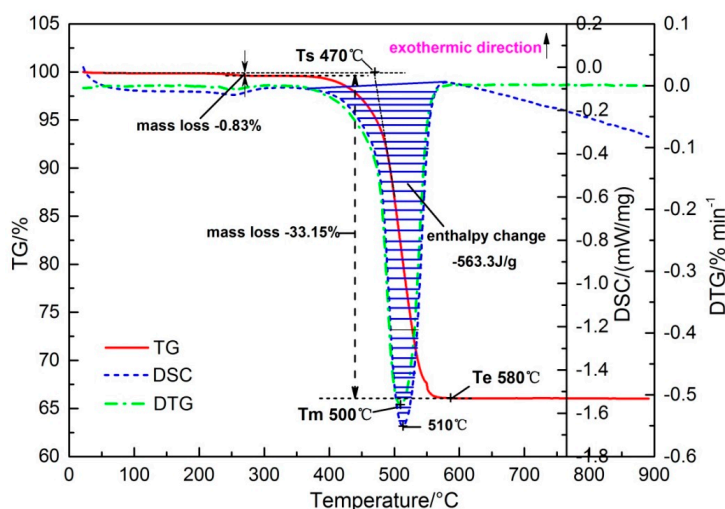


Figure 2. TG-DTG-DSC profiles of siderite sample in the neutral atmosphere.

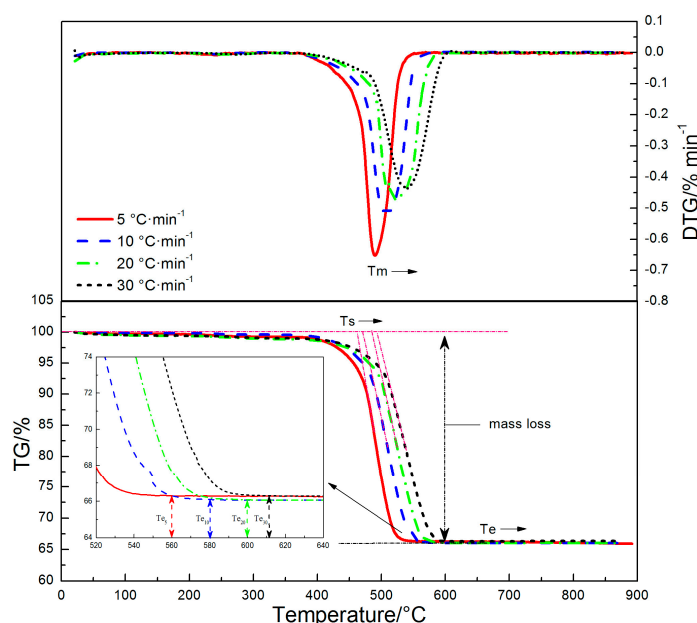


Figure 3. TG/DTG profiles of the siderite sample at different heating rates.

3.2. Pyrolysis Kinetics Characteristics of Siderite at Different Heating Rates

The conversion fraction (α) and conversion rate (r) at the different heating rates are shown in Figure 4. At the different heating rates, all conversion fractions (α) of siderite exhibited three stages, i.e., initial stage, middle stage and final stage. The curves of conversion fractions at the different heating rates did not change significantly when α ranged from 0 to 0.1, which showed the stage was the initial stage of the reaction. When α was between 0.1 and 0.9, the profiles rose linearly, indicating that the stage was the main reaction period of siderite pyrolysis with a great significance on the pyrolysis process of siderite. Thus the following study of pyrolysis kinetics parameters was determined for this stage. When α ranged from 0.9 to 1.0, the value of conversion fractions converged to 1.0, meaning that the stage was the end of the reaction (Figure 4a).

At different heating rates, all conversion rate (γ) curves reached a peak value from 0, and then dropped to 0 in the end. The peak value was increased significantly with an increasing heating rate. Consequently, the heating rate had a great impact on the conversion rate. The times to reach the peak conversion rate were 95, 50, 26 and 18 min, at the different heating rates of 5, 10, 20 and 30 °C·min⁻¹, respectively. It was found that increasing heating rate within a certain range was in favor of siderite conversion, and the peak conversion rate could be obtained in a short time (Figure 4b).

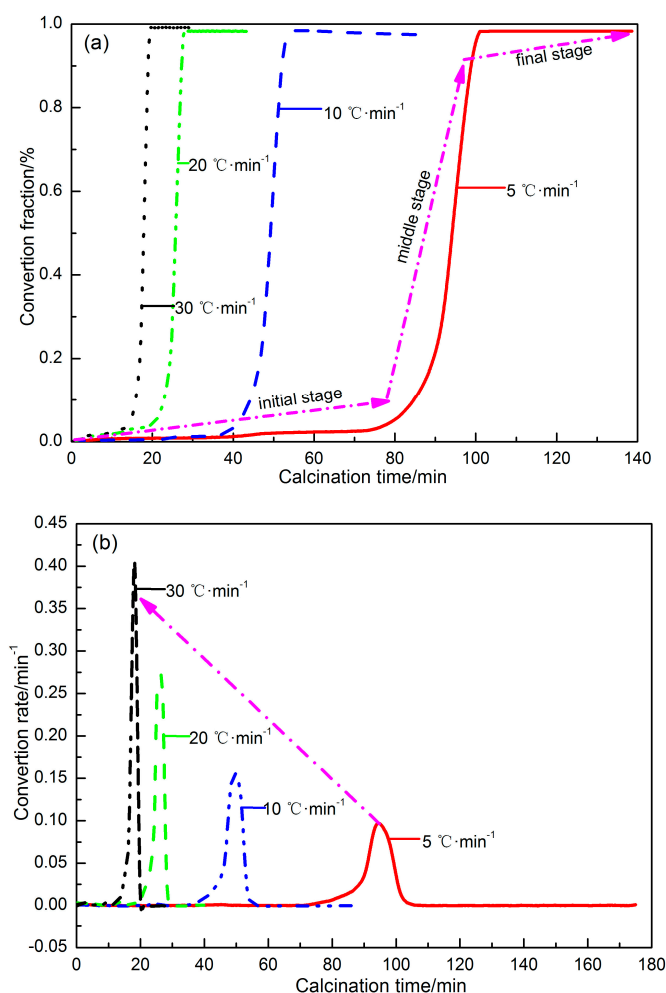


Figure 4. The conversion fraction (a) and conversion rate (b) profiles of siderite pyrolysis at different heating rates.

At each heating rate, conversion fraction (α) selected from 0.1 to 0.9 with an increment of 0.1 were used to calculate the activation energy using Equation (6). The $\lg(\beta)$ versus $1/T$ plots calculated for the

different heating rates are presented in Figure 5. From the slope of these lines in Figure 5, the E value can be determined at various conversion fraction α , as presented in Table 4.

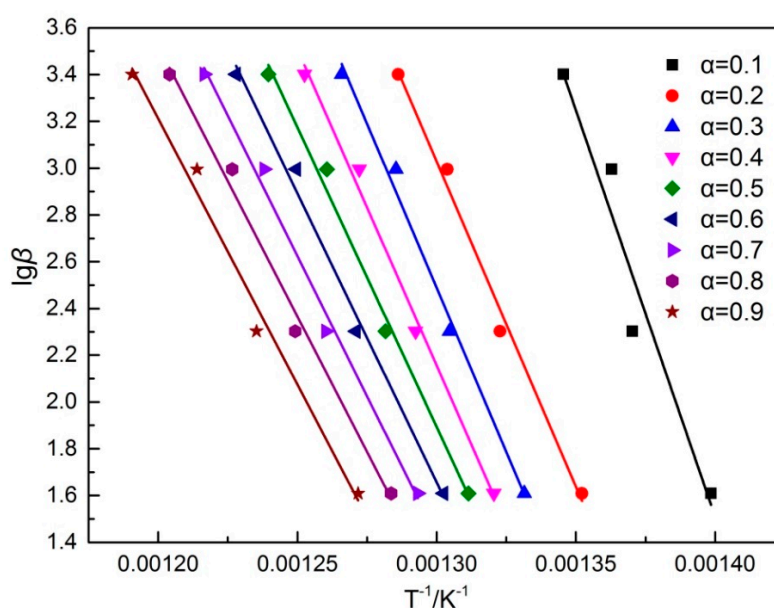


Figure 5. Regression lines for various conversion fraction based on FWO method.

Table 4. Activation energies estimated by FWO method at different conversion fraction.

Conversion Fraction/ α	$E/\text{kJ}\cdot\text{mol}^{-1}$ ^a	R ^b
0.1	532.65	0.978
0.2	516.05	0.988
0.3	510.97	0.989
0.4	492.93	0.990
0.5	466.63	0.989
0.6	450.20	0.988
0.7	436.79	0.989
0.8	421.32	0.987
0.9	413.92	0.982
Average	471.27	0.987

^a E -activation energy, calculated by Equation (6); ^b R -correlation coefficient.

As seen in Table 4, unmistakably, there was good linearity between $\lg(\beta)$ and $1/T$ at the different conversion fraction. The activation energy (E) of the siderite pyrolysis could be calculated from the corresponding slopes according to the FWO kinetic method. The average activation energy (E) was $471.27 \text{ kJ mol}^{-1}$.

The linear fitting results and relate kinetic parameters are summarized by Šatava–Šesták method. The calculated correlation coefficients (R) using mechanism function (Table 3) indicated that the linearity of $G(\alpha)$ for $A_{1/2}$ reaction model versus $1/T$ was excellent (Table 5); meanwhile, the activation energies of $A_{1/2}$ reaction model were approximately equal to those (Table 4) estimated by FWO method. Therefore it was concluded that $A_{1/2}$ reaction model, namely, $G(\alpha) = [-\ln(1 - \alpha)]^{-2}$ and $f(\alpha) = 1/2(1 - \alpha)[- \ln(1 - \alpha)]^{-1}$, was the most probable mechanism function of non-isothermal pyrolysis of siderite, and the corresponding mechanism was nucleation and growth. It was found that the activation energy (E) and pre-exponential factor (A) increased with increasing heating rate. For instance, the activation energy (E) and pre-exponential factor (A) increased from $446.13 \text{ kJ}\cdot\text{mol}^{-1}$ to 6.67×10^{-18} to $505.19 \text{ kJ}\cdot\text{mol}^{-1}$ and 2.40×10^{-21} , respectively, as heating rate increased from 5 to $30 \text{ }^\circ\text{C}\cdot\text{min}^{-1}$. More generally, the A value reflects the degree of collisions of a reaction per minute.

Therefore, the higher A value connotes the easier occurrence of siderite pyrolysis reaction, indicating that heating rate plays a positive role in the pyrolysis process of siderite.

Table 5. Kinetic parameters to most probable mechanism function at different heating rate.

Heating Rate ($^{\circ}\text{C}\cdot\text{min}^{-1}$)	Code	Intercept	Slope	R	Activation Energy E_s^a ($\text{kJ}\cdot\text{mol}^{-1}$)	Pre-Exponential Factor A^b (min^{-1})
5	$A_{1/2}$	-24506.71	20.54	0.9980	446.13	6.67×10^{-18}
10		-25768.30	20.98	0.9993	469.10	1.75×10^{-19}
20		-26800.51	22.01	0.9990	487.89	1.80×10^{-20}
30		-27750.87	23.15	0.9995	505.19	2.40×10^{-21}

^a E_s -activation energy, calculated by Equation (7); ^b A - pre-exponential factor, calculated by Equation (7).

3.3. The Characteristics of Siderite both before and after Pyrolysis

The XRD pattern of the product obtained from the tube furnace test is shown in Figure 6. All siderite was transformed into magnetite after pyrolysis, and no other mineral phases were observed.

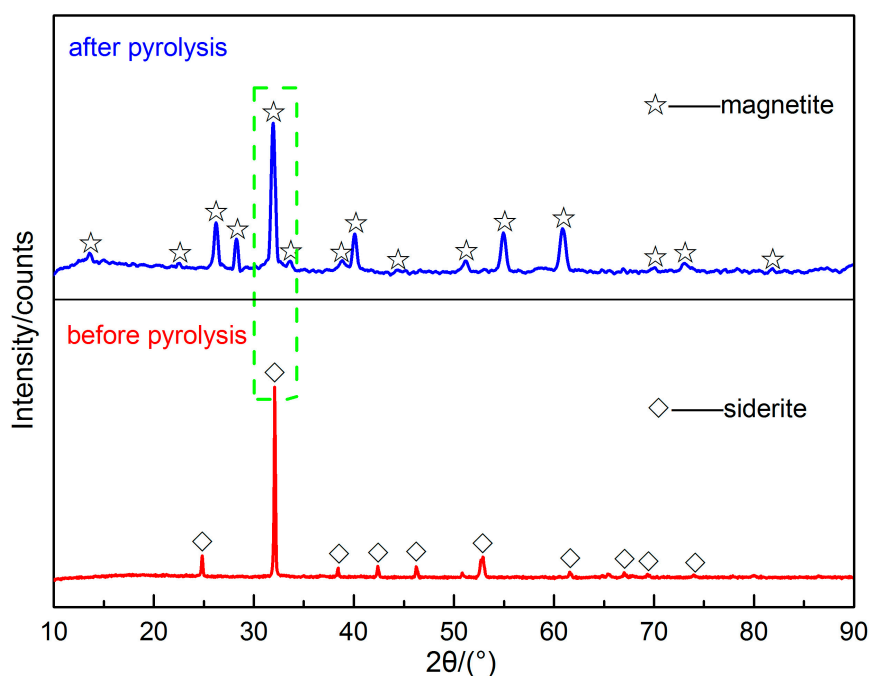


Figure 6. XRD patterns of the siderite before and after pyrolysis.

The morphologic features of the pyrolysis product from SEM and EDS analyses are shown in Figure 7. Before pyrolysis, the siderite samples were mostly chunky, with relatively even and smooth surfaces. Fe, C, O, Mn and Mg were also found on the surfaces (Figure 7a), which is in a good agreement with chemical analysis in Table 1. After pyrolysis, the surface of the newly generated magnetite appeared porous (Figure 7b), and micro-cracks were formed into the particles (Figure 7c). The carbon composition disappeared, which was most likely due to the CO_2 evaporated from the samples in the process of pyrolysis. The magnetite particles with structures of pores and micro-cracks were easy to break, which could improve the grinding efficiency drastically. On the other hand, Mn and Mg remained in samples after pyrolysis, which could lower the iron grade of the concentrate.

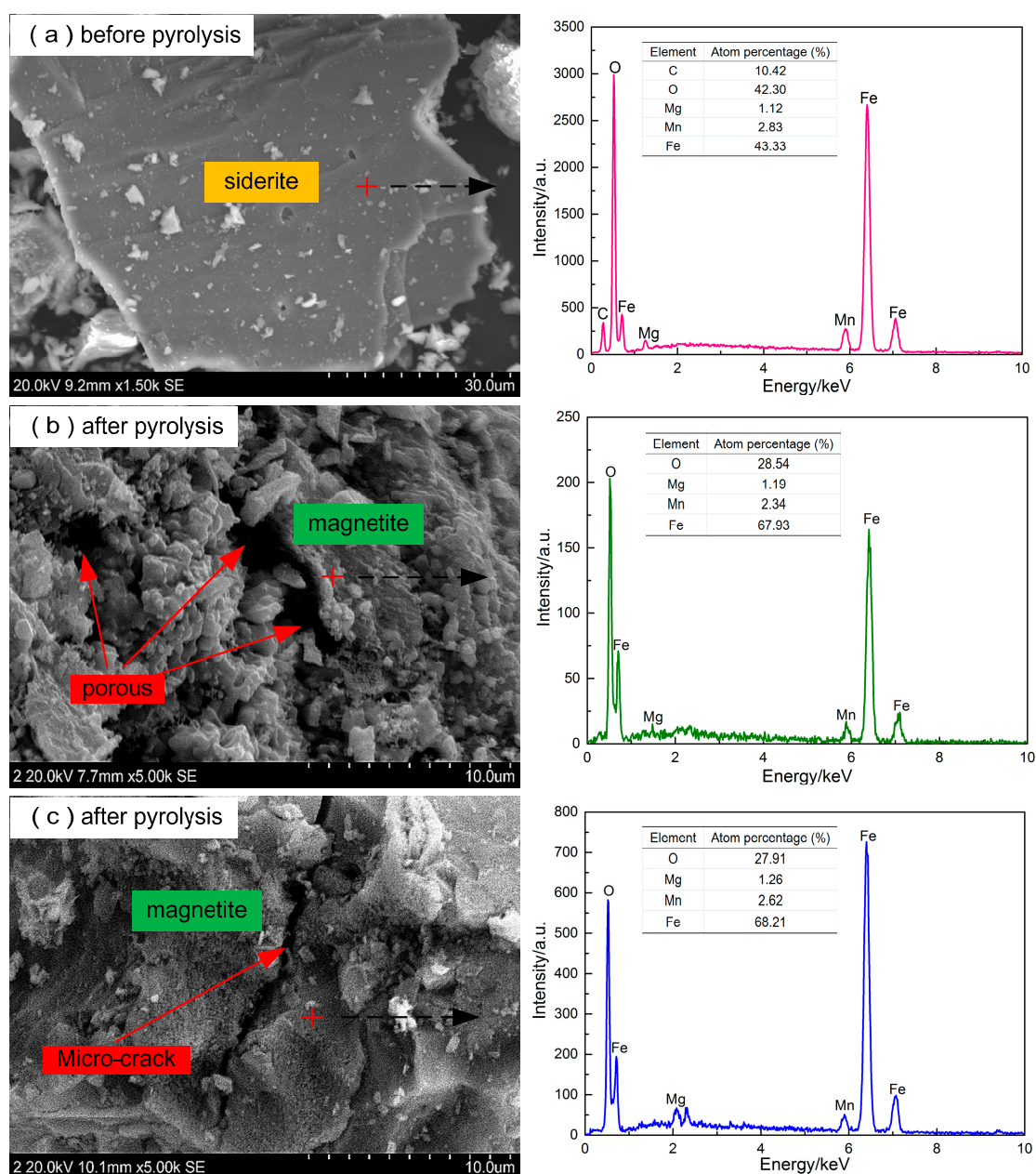


Figure 7. SEM and EDS analysis results of siderite before (a) and after (b,c) pyrolysis.

As illustrated in Figure 8, before pyrolysis, with the increase of magnetization magnetic intensity (MMI), magnetization intensity (MI) and specific susceptibility (SS) remained extremely low values; basically, it is suggested that the siderite exhibited a weakly magnetic mineral. After pyrolysis, as the MMI increased, the MI increased rapidly and then plateaued, reaching the magnetic saturation. Meanwhile, the SS briefly increased initially but decreased subsequently. A maximum of SS of $0.74 \times 10^{-3} \text{ m}^3 \cdot \text{kg}^{-1}$ was observed at the MMI of $32.87 \text{ kA} \cdot \text{m}^{-1}$. Because the MMI was lowered, magnetic domain wall in magnetite shifted rapidly with the increase of MMI, and magnetic moment turned to the magnetic field, until it reached magnetic saturation. When MMI increased further, the SS decreased gradually because the magnetic moment was essentially constraint. Therefore, the weakly magnetic siderite was transformed into strongly magnetic magnetite by pyrolysis, and the SS of sample improved significantly. In other words, the pyrolysis process expanded the difference of magnetic properties between iron mineral and gangue minerals, facilitating their effective separation in the subsequent low intensity magnetic separation.

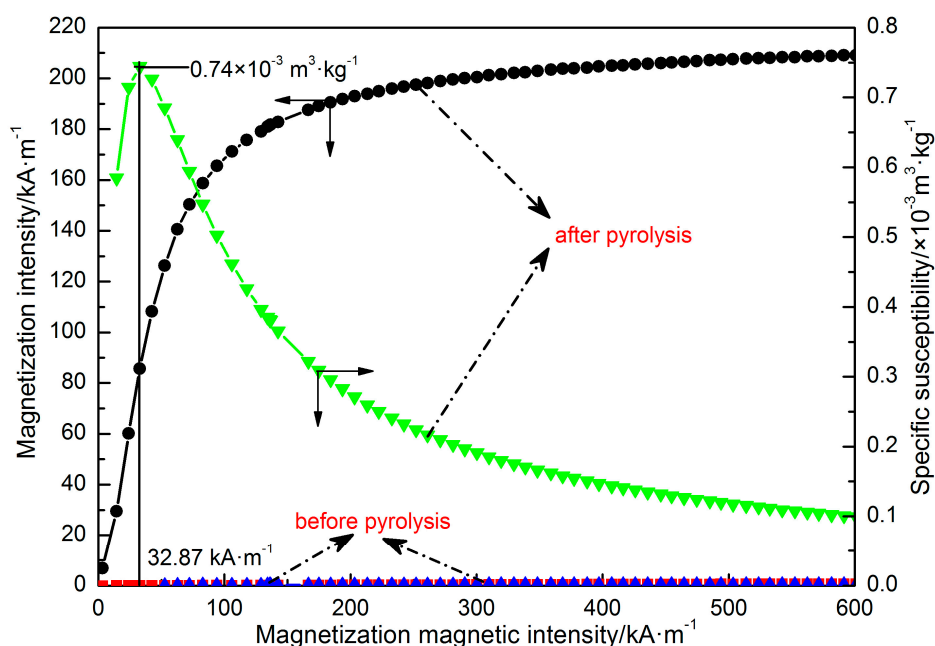


Figure 8. Magnetic analysis of the siderite before and after pyrolysis.

4. Conclusions

The pyrolysis behavior and kinetics of siderite at different heating rates were investigated using thermogravimetry technique and laboratory type tube furnace.

(1) With the increase of heating rate, the TG and DTG curves shifted to higher temperatures. The starting temperatures, maximum weight loss rate and ending temperature increased, and the pyrolysis temperature range became wider. However, heating rate had little effect on the total mass loss. The siderite pyrolysis was one-step reaction and an intense endothermic process. A higher pyrolysis rate could be achieved in a relatively short period of time by increasing of heating rate in a certain range.

(2) At heating rates of 5, 10, 20 and 30 °C·min⁻¹, the pyrolysis kinetics of siderite in neutral atmosphere was consistent with $A_{1/2}$ (nucleation and growth) mechanism function. The forms of integral and differential mechanism functions were $G(\alpha) = [-\ln(1 - \alpha)]^{-2}$ and $f(\alpha) = 1/2(1 - \alpha)[- \ln(1 - \alpha)]^{-1}$, respectively. The corresponding activation energies were 446.13, 469.10, 487.89, and 505.19 kJ·mol⁻¹ and the pre-exponential factors were 6.67×10^{-18} , 1.75×10^{-19} , 1.80×10^{-20} , and 2.40×10^{-21} min⁻¹, indicating an increasing trend with heating rate increase. The heating rate plays a vital role in controlling siderite pyrolysis.

(3) The weakly magnetic siderite was transformed into ferromagnetic magnetite after pyrolysis. The surface of the newly formed magnetite was porous with micro-cracks. Mn and Mg remained in the samples during pyrolysis. The magnetization intensity and specific susceptibility enhanced significantly, which was beneficial to further enrichment and utilization of iron minerals.

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