






## Article

# Influence of the Morphology of Eutectoid Steels on Corrosion Resistance in NaCl Aqueous Medium with and without CO<sub>2</sub>

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**Abstract:** This study conducts a comparative electrochemical evaluation of three types of pearlitic steels used in flexible pipelines for oil transport in marine environments. The steels have been manufactured with chemical composition and geometry variations to optimize operation performance under adverse conditions. Electrochemical tests were conducted using solutions simulating marine environments with NaCl and CO<sub>2</sub>, and at high temperatures. The results indicated that spheroidized (SC) steel demonstrated the best corrosion resistance under these specific conditions. Additionally, the Raman spectroscopy characterization technique was used to analyze the layers of corrosion products formed during the tests, identifying the presence of FeCO<sub>3</sub> (siderite) and other corrosive oxides. These discoveries are valuable for selecting and improving materials in flexible pipelines used in oil production in marine waters. The study highlights the importance of the cementite morphology present in pearlite as a relevant factor in the corrosive behavior of steels, contributing to the development of more efficient and durable solutions for the offshore oil and gas industry.

**Keywords:** pearlitic steels; electrochemical evaluation; corrosion resistance; NaCl; CO<sub>2</sub>; Raman spectroscopy

## 1. Introduction

The offshore oil industry includes, as a cost of its operations, the purchase and maintenance of crude oil exploration pipelines. The pipelines, called risers, connect a floating offshore oil production structure to the subsea production system [1,2]. These risers have several layers of metallic and polymeric materials, and their main characteristics are tensile strength, collapse, internal pressure, attack by chemical products, ease of installation and reuse, and flexibility, among others [3,4].

In the region called the annular space, the tension reinforcements that support the axial loads of the pipeline and the pressure reinforcements that provide the pipeline with the capacity to withstand its internal pressures are positioned in overlapping layers. In this confined space between the layers, material failure can occur through the phenomena known as sweet corrosion and other structural damage caused by corrosion [5,6].

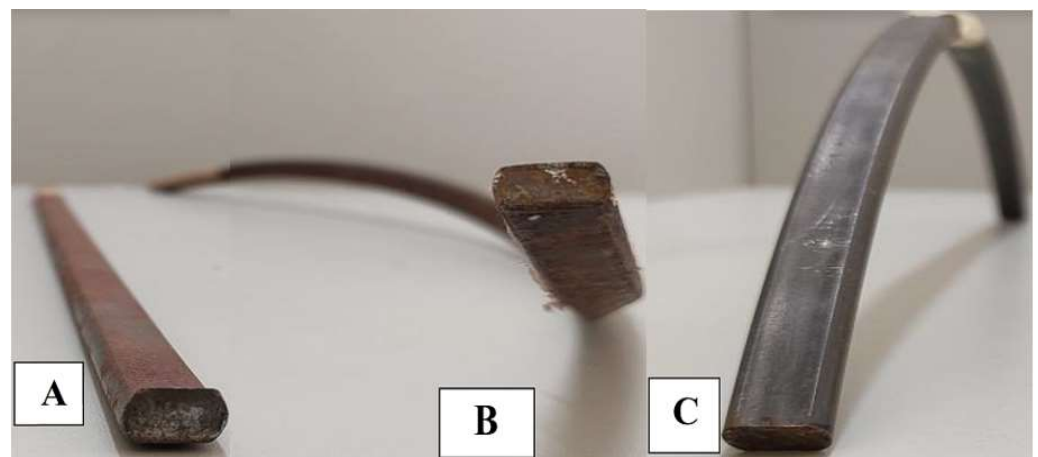
The evaluation of the tensile strength of wires subjected to severe corrosion conditions with CO<sub>2</sub> and NaCl is of interest to the oil industry, in addition to appealing to companies manufacturing risers that develop materials with different compositions and heat treatments. The application of other materials and in environmentally unfavorable conditions are relevant factors for the industries that manufacture such components since, for these reasons, the optimization and quality of the final product are essential [7,8]. CO<sub>2</sub> gas has a very damaging effect on the surface of steel, especially due to forming H<sub>2</sub>CO<sub>3</sub>, an acidic compound that can lead to the rupture of pipelines and other oil production structures [9,10].

Although this type of structure is vital for the oil industry to guarantee productivity and a safe environment, few studies have investigated them in conditions that can simulate the application of the main components of these risers when in service. Therefore, this work aims to correlate the corrosion resistance of pearlitic steels with different cementite profiles under various atmospheric conditions. NaCl solution with the presence of CO<sub>2</sub> and at different temperatures was evaluated. Raman spectroscopy was also used to verify the presence of oxides and salts resulting from corrosion products on the material's surface.

## 2. Materials and Methods

### 2.1. Materials

The materials used in this study were pearlitic steel (eutectoid) used as tensor armors, received in the form of wires, initially called wires A, B and C, as shown in Figure 1.



**Figure 1.** Wire format in its received condition: (A) flat (1.29 cm × 0.69 cm), (B) twisted (1.21 cm × 0.60 cm) and (C) bent (1.23 cm × 0.72 cm).

The chemical composition (wt%) of the wires, as shown in Table 1, was determined by using an optical emission spectrometer, model PDA 70000 (Shimadzu, Kyoto, Japan).

**Table 1.** Chemical composition (wt%) of the wires.

Wire	C	Mn	Si	S + P	Cr + Ni + V	Al	Mo + Ti	Fe
A	0.77	0.50	0.20	0.014	0.054	-	0.007	bal
B	0.76	0.56	0.18	0.014	0.041	-	0.005	bal
C	0.73	0.58	0.26	0.008	0.038	0.033	0.005	Bal

### 2.2. Microstructural Characterization of the Wires

For the acquisition of micrographs of the wires, a scanning electron microscope (SEM) Quanta 450 FEG–FEI was used. The wire samples were embedded in Bakelite, ground from 100 to 1200 mesh and polished with diamond paste from 3 to 1 μ. The samples were

sprayed with alcohol and blow dried. They were then etched with 2% Nital to reveal their microstructure. The samples were evaluated close to the external faces of the wires.

### 2.3. Electrochemical Tests

For the electrochemical tests, samples with an area of approximately 42 mm<sup>2</sup> (surfaces of the wires) were used. ASTM G1-03, Standard Practice for Preparing, Cleaning, and Evaluating Corrosion and Test Specimens, ASTM, International, West Conshohocken PA, 2011 were used. The samples were embedded in epoxy resin, ground from 100 to 600 mesh, sprayed with alcohol and blow-dried. An enamel was used to prevent crevice corrosion between the resin and the sample. For the corrosion tests, a conventional three-electrode cell was used. The working electrode was the wire samples, the counter electrode was a platinum plate and the reference electrode used was an Ag(s)/AgCl(s)/Cl-(aq)KCl electrode. All tests were performed in triplicate to ensure repeatability. The equipment used was a potentiostat model PGSTAT30 (Autolab, Methrom-Eco Chemie) connected to a computer by the software NOVA 2.1. The polarization curves were scanned from −0.1 V to 2.0 V from OCP at a rate of 1 mV/s. When the current density reached 1 mA, the tests ended. The electrolyte was an aqueous solution of 3.5% NaCl and CO<sub>2</sub> bubbling (99.99 purity) at a flow rate of 1 × 10<sup>−4</sup> m<sup>3</sup>/s. The solution was deaerated with N<sub>2</sub> before the tests. The corrosion tests were conducted at room temperature and 80 °C [11]. A different test condition was used for each wire, as shown in Table 2. From here on, wires, A, B and C are named as spheroidized (SC), lamellar (LC), and discontinuous (DC) structured, respectively. These are classified according to the specific presented microstructure in the material that a local oil company supplied.

**Table 2.** Conditions of the experiments.

Cementite Condition	Condition in 3.5% NaCl
SC, LC and DC	Saturated with CO <sub>2</sub> at 80 °C
SC, LC and DC	Saturated with CO <sub>2</sub> at room temperature
SC, LC and DC	Aerated and at room temperature

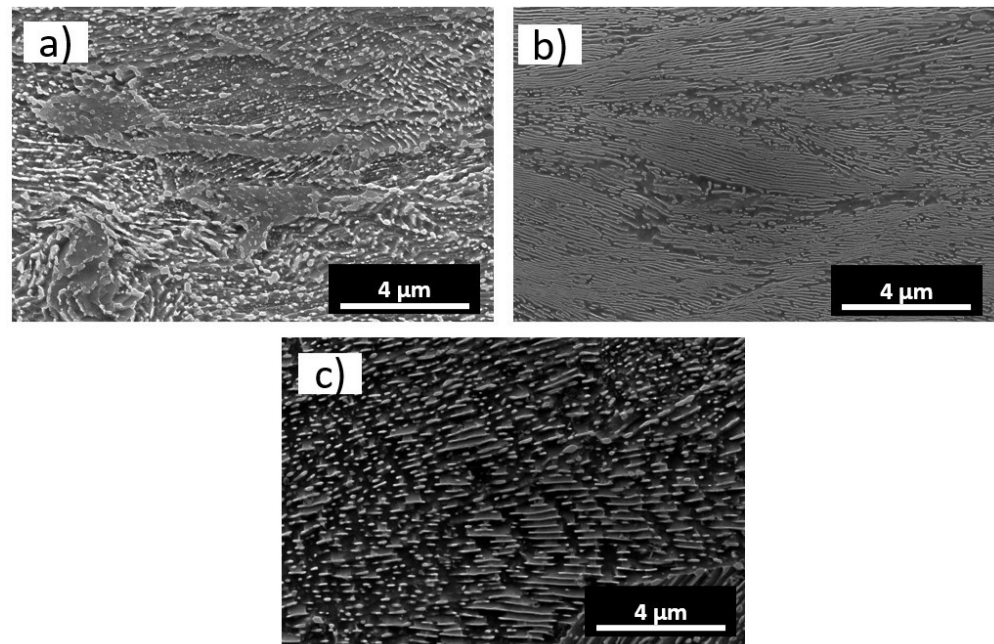
### 2.4. Raman Spectroscopy

Raman spectroscopy was performed using Witec Alpha 300 equipment, grating of 600 grooves/mm. Excitation was achieved with a 532 nm wavelength laser. The sample utilized for this characterization technique was the one that presented the most severe corrosion situation. The Raman spectral range reading was from 200 to 2000 cm<sup>−1</sup> since the main oxides found in the previously evaluated databases were in this range.

## 3. Results

### 3.1. Characterization of the Cementite of the Wires

Figure 2 shows three different types of microstructure at the cross-section of each sample. In Figure 2a, the micrograph of wire A can be seen. It is possible to visualize the presence of globular cementite and discontinuous pearlite (brittle). This wire was called spheroidized cementite (SC). A lamellar pearlite microstructure can be seen in Figure 2b. This wire was called lamellar cementite (LC). Figure 2c shows a discontinuous cementite microstructure. This wire was called discontinuous cementite (DS).



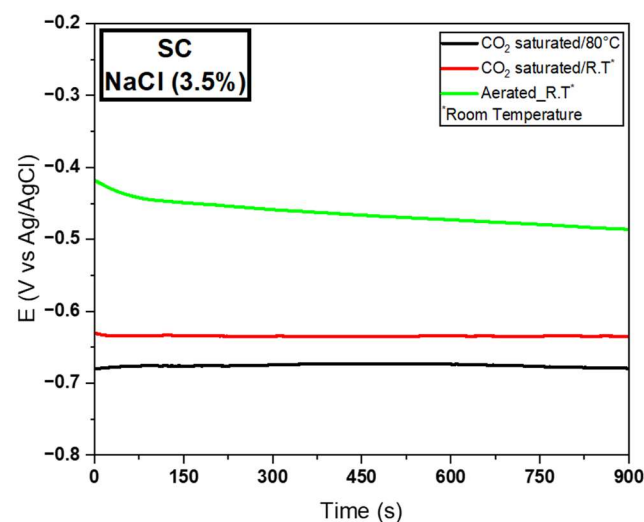
**Figure 2.** Microstructure of (a) spheroidized cementite, (b) lamellar cementite and (c) discontinuous cementite.

### 3.2. Electrochemical Tests

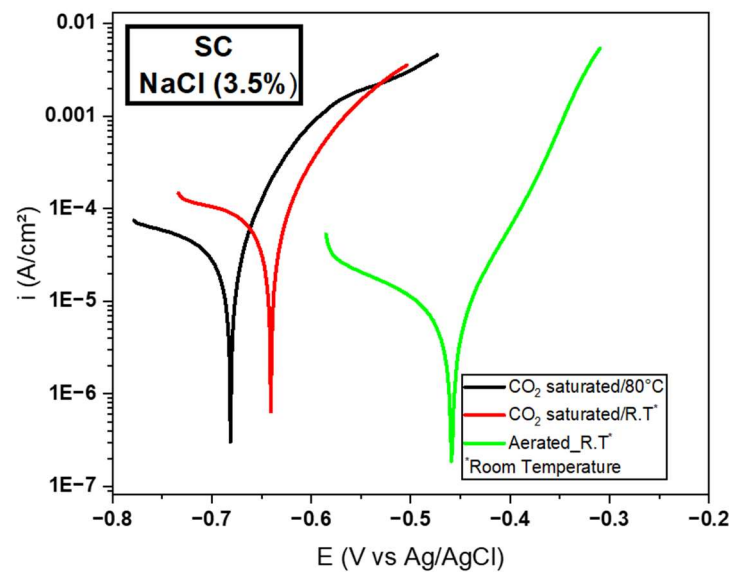
The graphs with the respective open circuit polarization (OCP) and potentiodynamic polarization curves obtained from the electrochemical tests are presented in Figures 3–8. These experiments were carried out with the three samples (SC, LP, and DP). NaCl solutions (3.5%) mixed with and without CO<sub>2</sub> were used in both specimens. Tests were performed at room temperature and at 80 °C. During the tests, the pH was recorded by means of a pH meter, obtaining the following values for the respective solutions: saturated with CO<sub>2</sub> at 80 °C (pH 3–4); saturated with CO<sub>2</sub> at room temperature (pH 4–5); aerated and at room temperature (pH 6–7).

#### 3.2.1. Electrochemical Tests for the SC Wire

The graphs of potentiodynamic polarization and OCP can be seen in Figures 3 and 4.



**Figure 3.** OCP curves for SC wire in the aqueous solution of 3.5% NaCl saturated with CO<sub>2</sub> at 80 °C, saturated CO<sub>2</sub> at room temperature (23 °C), and aerated at room temperature (23 °C).

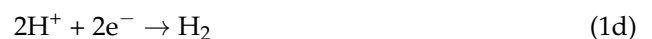
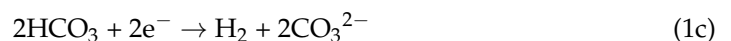
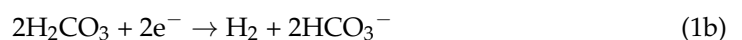


**Figure 4.** Potentiodynamic polarization curves for SC wire in the aqueous solution of 3.5% NaCl saturated with CO<sub>2</sub> at 80 °C, saturated CO<sub>2</sub> at room temperature (23 °C), and aerated at room temperature (23 °C).

Figure 3 shows that the time established for the OCP was 900 s (15 min), the same as used for the other tests, due to the stabilization of the potential applied to the material. There was great similarity between the OCP's of the SC in the condition in which there was saturated CO<sub>2</sub>. However, this same potential was not close in the absence of CO<sub>2</sub> condition. This fact demonstrates the greater sensitivity of the SC sample to carbon dioxide. On the other hand, using a different temperature from the ambient one did not show a great difference when the sample was submitted to 80 °C.

Figure 4 shows the potentiodynamic polarization curves for SC microstructure wire. In this graph, the samples in the CO<sub>2</sub> medium showed a higher anode current density (ACD), showing that the samples were more susceptible to corrosion. A difference was noticed between the corrosion resistances for each sample since, on the y-axis, it is possible to identify the current density of the specimens by the Tafel extrapolation method. These ACDs were very similar between the samples with saturated CO<sub>2</sub> at 80 °C and room temperature, while the aerated sample at room temperature showed lower ACD. Regarding the two saturated CO<sub>2</sub> specimens, this effect of low corrosion resistance is due to the acidifying power generated by CO<sub>2</sub>, forming H<sub>2</sub>CO<sub>3</sub> when in contact with water in the electrolyte [12]. This fact can be proven by the pH reduction measured in this solution, reaching values between 3 and 4.

CO<sub>2</sub> corrosion in steels can contribute to forming FeCO<sub>3</sub> (iron carbonate or siderite), significantly reducing corrosion rates. This occurs when the concentrations of Fe<sup>2+</sup> and CO<sub>3</sub><sup>2-</sup> ions exceed the solubility limit for the evaluated solution [13]. This characteristic can affect the material, depending on its morphology. This mechanism can be explained because, once CO<sub>2</sub> is dissolved in water, carbonic acid is formed, and it is harmful to carbon steel. Equations (1a)–(1e) detail these reactions [14]:





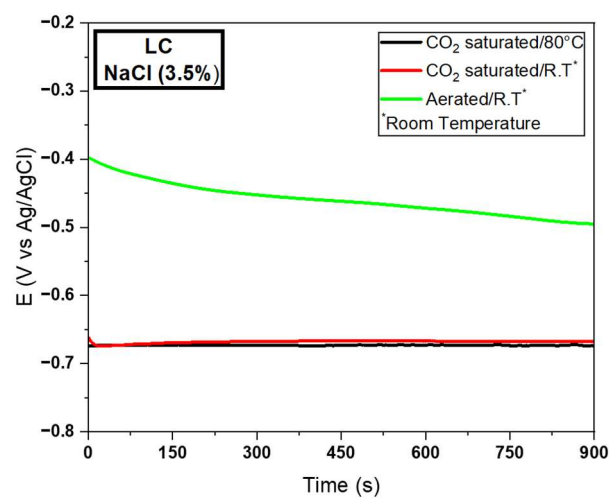
Equation (1a) (oxy-reduction) and Equations (1b)–(1d) are cathodic reactions (reduction) and Equation (1e) is anodic (oxidized). The reaction in Equation (2) helps us to understand the mechanism of iron carbonate formation, because when the  $\text{Fe}^{2+}$  and  $\text{CO}_3^{2-}$  ions exceed the solubility limit,  $\text{FeCO}_3$  is formed [15,16]:



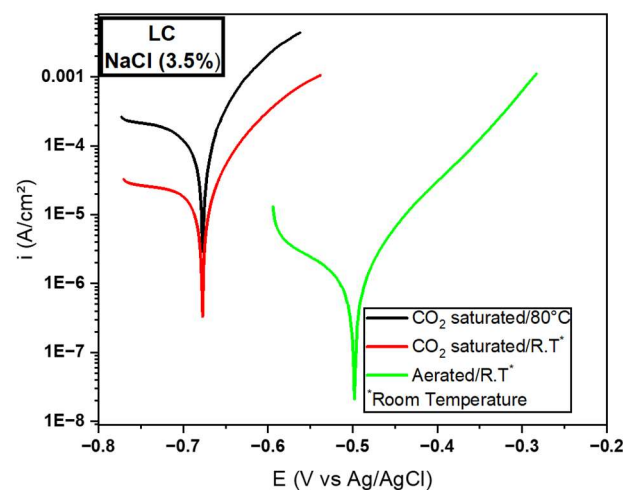
Some authors affirm the existing correlation between temperature increase and susceptibility to siderite formation, depending on the morphology [17–19].

### 3.2.2. Electrochemical Tests for the LC Wire

Figures 5 and 6 represent the OCP and potentiodynamic polarization tests for the LC wire in an aqueous medium of 3.5% NaCl saturated with  $\text{CO}_2$  at 80 °C, saturated with  $\text{CO}_2$  at room temperature, and aerated at room temperature.



**Figure 5.** OCP curves for LC wire in aqueous solution of 3.5% NaCl with saturated  $\text{CO}_2$  at 80 °C, saturated  $\text{CO}_2$  at room temperature (23 °C) and aerated at room temperature (23 °C).



**Figure 6.** Potentiodynamic polarization curves for LC wire in aqueous solution of 3.5% NaCl with saturated  $\text{CO}_2$  at 80 °C, saturated  $\text{CO}_2$  at room temperature (23 °C) and aerated at room temperature (23 °C).

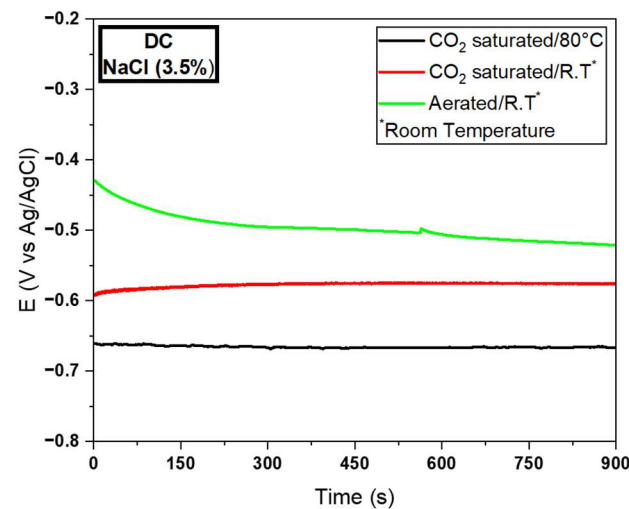
Figure 5 indicates that the LC wire OCP showed very similar results between the specimens submitted to  $\text{CO}_2$ , demonstrating once again that the influence of temperature is not very important for the established corrosion conditions when it comes to the corrosion

potential of the material. However, the OCP values were less negative when carbon dioxide was removed.

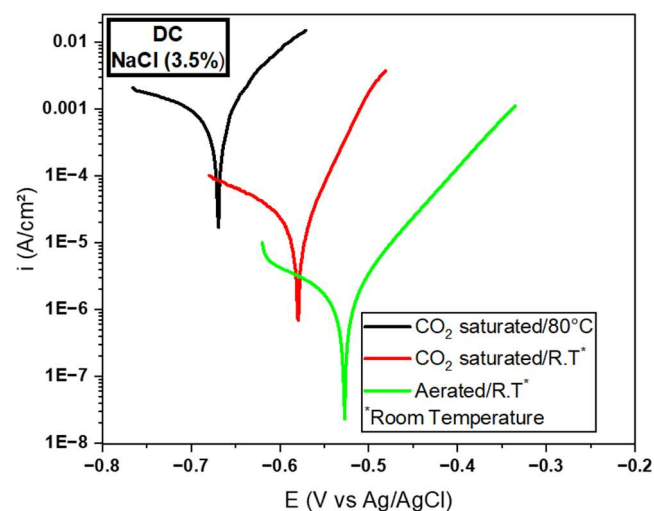
In Figure 6, which establishes the result for the polarization curves, the established difference between the current densities is denoted. Among the specimens with the presence of carbon dioxide, the sample showed greater susceptibility to corrosion. The sample tested in the presence of a temperature different from room temperature proved to be the most prone to corrosion, since its current density was higher. In turn, the aerated sample at room temperature was the one that showed the lowest susceptibility to corrosion.

### 3.2.3. Electrochemical Tests for the DC Wire

Figures 7 and 8 show the OCP and the potentiodynamic polarization curves for the DC wire.



**Figure 7.** OCP curves for the DC wire in aqueous solution of 3.5% NaCl with saturated  $\text{CO}_2$  at  $80^\circ\text{C}$ , saturated  $\text{CO}_2$  at room temperature ( $23^\circ\text{C}$ ) and aerated at room temperature ( $23^\circ\text{C}$ ).



**Figure 8.** Potentiodynamic polarization curves for the DC wire in an aqueous solution of 3.5% NaCl with saturated  $\text{CO}_2$  at  $80^\circ\text{C}$ , saturated  $\text{CO}_2$  at room temperature ( $23^\circ\text{C}$ ) and aerated at room temperature ( $23^\circ\text{C}$ ).

Unlike previous samples, in which there was considerable similarity between the OCP results for those saturated with  $\text{CO}_2$ , the DC wire (Figure 7), there was not a big difference. For this specimen, it is believed, based on the results presented, that the temperature added greater corrosion capacity, since its OCP was more negative. In a corrosion process without

the formation of a protective film on the surface, the increase in temperature increases the corrosion rate [12], which directly applies to the described case.

The current densities are indicated in Figure 8. Notably, the potential distance and current densities of the sample in saturated CO<sub>2</sub> at room temperature are quite similar to those of the other solutions.

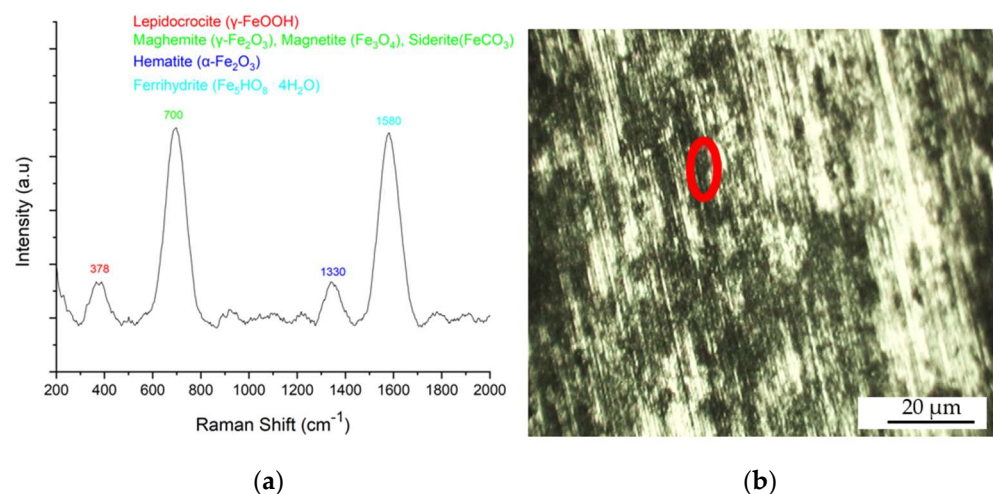
### 3.3. Characterization of the Corrosion Product by Raman Spectroscopy

The analysis of the surfaces of specimens submitted to a corrosive environment is intended to evaluate possible oxides that form on the surface of the samples. Characterization techniques such as SEM (scanning electron microscopy), EDS (energy dispersion X-ray spectroscopy) and Raman spectroscopy are the main methodologies used to obtain information on layers deposited on materials, once subjected to adverse conditions.

The technique used in this study was Raman spectroscopy. It was intended to investigate the formation of oxides on the surface of the material that can help reduce the corrosive process, simulating a passive layer. The main films that can form on the studied samples, in this study and in conditions similar to the corrosion evaluated here, are FeCO<sub>3</sub>, Fe<sub>3</sub>O<sub>4</sub>, and Fe<sub>2</sub>O<sub>3</sub> [20,21].

#### SC Sample in Aqueous NaCl Medium

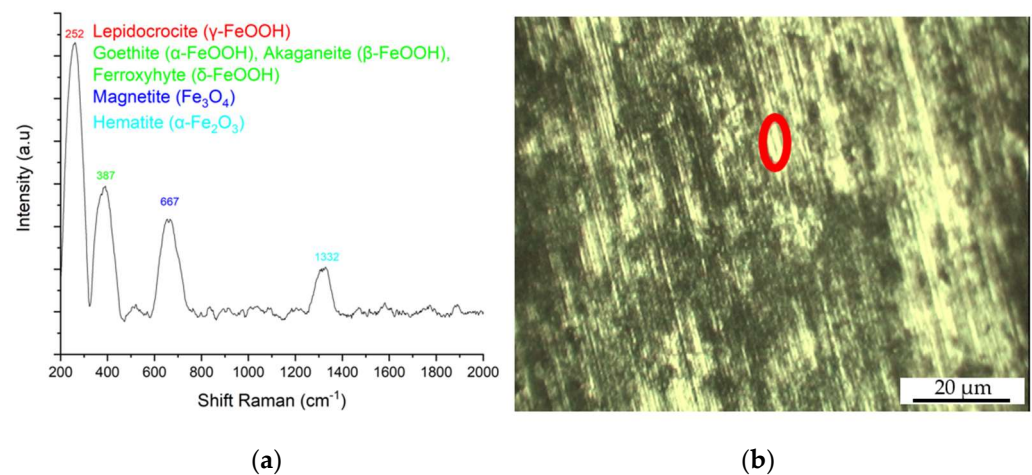
The evaluation started with the SC sample being subjected to the most severe corrosion conditions when using the NaCl solution. Figure 9b shows black regions where the material has suffered corrosion. The red circle represents the point where the Raman technique measured the corrosion. In this region, lepidocrocite, maghemite, magnetite, siderite, hematite, and ferrihydrite were found. Among these oxides, the acid salt FeCO<sub>3</sub> (siderite) stands out, obtained by the reaction of carbon dioxide with Fe<sup>+2</sup> ions [22] (Figure 9a) dissolved in the aqueous solution. Normally, those detected oxides appeared as dark spots on the sample's surface, and resulted from the reaction between Fe ions and oxygen in the water.



**Figure 9.** SC sample (immersed in 3.5% NaCl aqueous solution with saturated CO<sub>2</sub> at 80° C): (a) Raman spectroscopy characteristic peaks for generalized corrosion and (b) micrograph of the sample with emphasis on the analyzed region.

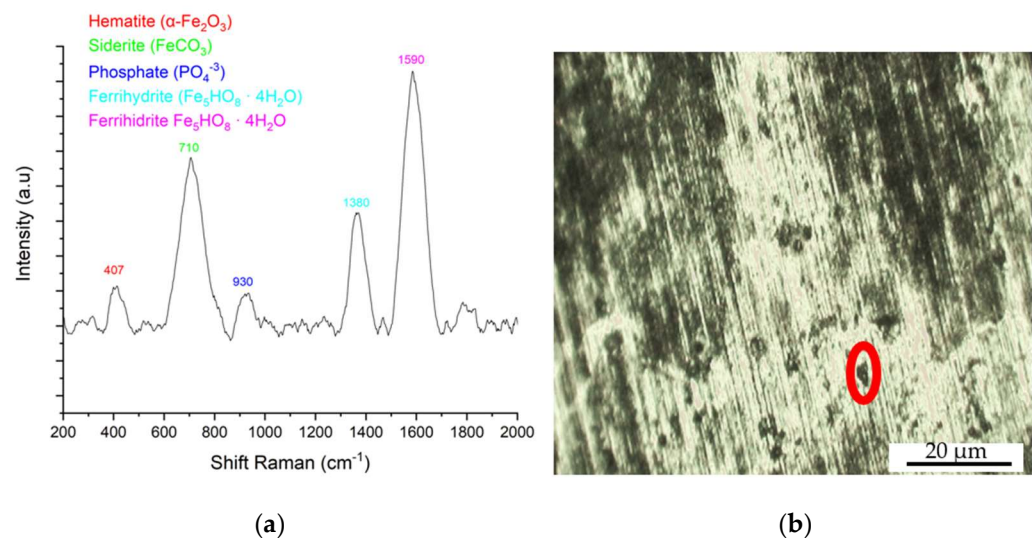
Figure 10a refers to capturing the Raman shift in a lighter sample region. This was called the area without apparent corrosion. In this region, goethite and ferroxite appeared, oxides that were not detected in Figure 9. The presence of siderite peaks was not detected, which leads to the assumption that this oxide does not necessarily appear as a protective layer from non-corroded regions, as identified by some authors [12].





**Figure 10.** SC sample (immersed in 3.5% NaCl aqueous solution with saturated  $\text{CO}_2$  at 80 °C): (a) characteristic peaks of Raman spectroscopy for the area without apparent corrosion and (b) micrograph of the sample with emphasis on the analyzed region.

The spectrum in Figure 11 was captured in a region called localized corrosion. Their characteristic spectra identified oxides and salt (siderite), as in other regions of the same sample. However, the Raman shift for phosphate ( $\text{PO}_4^{3-}$ ) is highlighted, which has not previously been identified, and is probably the result of phosphate in the chemical composition of the material. This ion was possibly formed by the presence of phosphorus in the composition of the evaluated steels. However, no published studies were found that properly discussed the ratio of Raman peaks of phosphates in steels, having only its characteristic peak [23].



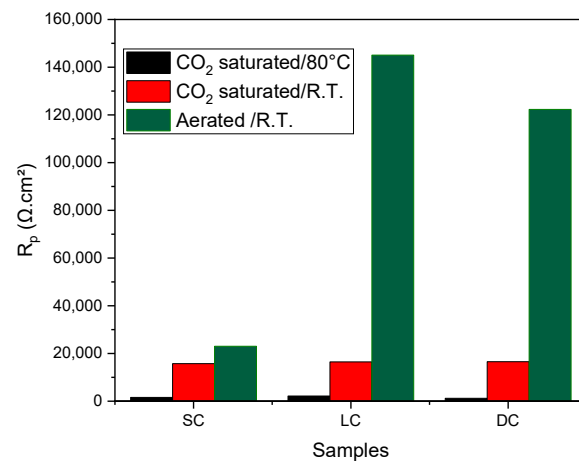
**Figure 11.** SC sample (immersed in an aqueous solution of 3.5% NaCl with saturated  $\text{CO}_2$  at 80 °C): (a) characteristic peaks of Raman spectroscopy for the area with localized corrosion and (b) micrograph of the sample with emphasis on the analyzed region.

#### 4. Discussion

Table 3 computes information about the electrochemical tests, calculating each sample's mean, standard deviation, and condition according to the studied solutions. The standard deviation indicates how uniform the test results were, consisting of measuring the level of dispersion in the measurements. The value of  $R_p$  (resistance to polarization) is essentially calculated by the ratio of OCP to current density ( $I$ ). The latter is obtained by the Tafel extrapolation method [24]. They are presented in the table and Figure 12.

**Table 3.** Statistical data from the electrochemical tests in NaCl for the three samples in the three tested solutions (MD = mean and SD = standard deviation).

SAMPLE	STATISTIC	Solution (3.5% of NaCl)								
		(1) CO <sub>2</sub> Saturated/80 °C			(2) CO <sub>2</sub> Saturated/R.T.			(3) Aerated/R.T.		
		Electrochemical Parameters								
		OCP (V vs Ag/AgCl)	I (A/cm <sup>2</sup> )	Rp (Ω.cm <sup>2</sup> )	OCP (V vs Ag/AgCl)	I (A/cm <sup>2</sup> )	Rp (Ω.cm <sup>2</sup> )	OCP (V vs Ag/AgCl)	I (A/cm <sup>2</sup> )	Rp (Ω.cm <sup>2</sup> )
SC	MD	−0.68	4.24E <sup>−4</sup>	1603.77	−0.64	4.07E <sup>−5</sup>	15,724.81	−0.44	1.91E <sup>−5</sup>	23,036.64
	SD	0.01	3.24E <sup>−4</sup>	—————	0.01	8.55E <sup>−6</sup>	—————	0.02	8.15E <sup>−6</sup>	—————
LC	MD	−0.68	3.15E <sup>−4</sup>	2158.73	−0.66	4.01E <sup>−5</sup>	16,458.85	−0.47	3.24E <sup>−6</sup>	145,061.72
	SD	0.01	2.10E <sup>−4</sup>	—————	0.01	1.01E <sup>−5</sup>	—————	0.02	7.00E <sup>−8</sup>	—————
DC	MD	−0.66	5.41E <sup>−4</sup>	1219.96	−0.56	3.38E <sup>−5</sup>	16,568.04	−0.51	4.17E <sup>−6</sup>	122,302.15
	SD	0.01	4.31E <sup>−4</sup>	—————	0.02	6.25E <sup>−6</sup>	—————	0.01	6.01E <sup>−8</sup>	—————

**Figure 12.** Comparative values of Rp (resistance to polarization) that were calculated from the ratio of OCP and the current density of each sample in different environments.

From Table 3, a comparison can be made between the samples and their different electrolyte solutions. Comparing the results of Table 3 vertically (the different samples of steels in the same solution) shows the one that had the highest potential and current density performance.

In the solution with 3.5% NaCl with saturated CO<sub>2</sub> and at a temperature of 80 °C, the three samples obtained very similar potentials, with the values varying only in the second decimal place. For the same electrolyte, a difference was observed between SC and DC wires, with the LC sample being more resistant to corrosion, since its current density obtained a lower result (greater difference than the comparison between SC and DC). Another observation to be made concerns the Rp's between these two wires (see Figure 12).

Analyzing the 3.5% NaCl solution with saturated CO<sub>2</sub> and at room temperature (23 °C), the most notable difference was shown between the LC and DC wires, dealing with the OCP of these materials under the previously mentioned conditions, obtaining a difference of almost one decimal place between the first and the second. Likewise, the current density values for these same samples proved to be more distant when comparing the LC sample with the SC one.

The medium with 3.5% NaCl aerated and at room temperature (23 °C) was generally a less aggressive medium of those evaluated. Fewer negative potentials were obtained for this medium, indicating greater resistance to corrosion by the material, emphasizing the

LC wire. However, the SC sample reached higher current densities than the other samples. The LC and DC samples reached very similar results.

Analyzing the results based on the  $R_p$  (polarization resistance), it is observed that the LC sample showed more resistance to corrosion, being less degraded, while the DC sample showed less resistance in solution 1. For solution 2, the DC obtained a high result for  $R_p$ , if compared to the SC sample, the latter being the sample that obtained the lowest resistance to corrosion. The least aggressive solution for the samples was solution 3.

Table 4 presents the results divided into levels of resistance to polarization for each sample according to its solution.

**Table 4.** Comparison of the  $R_p$ 's result of the samples in relation to the solutions of 3.5% of NaCl.

Corrosion Resistance	Solutions (3.5% de NaCl)		
	1 CO <sub>2</sub> Saturated/80 °C	2 CO <sub>2</sub> Saturated/T.A.	3 Aerated/T.A.
highest	LC	DC	LC
intermediate	SC	LC	DC
lowest	DC	SC	SC

The corrosion resistance of steels used in the oil industry was evaluated, depending on their morphologies [8]. It was possible to identify that those with globular (spheroidized) cementite obtained better results for corrosion when subjected to high temperatures (>60 °C) and with the use of CO<sub>2</sub> in a NaCl solution. This fact was also identified by another author [25], who highlighted in his work the formation of carbides and other compounds that helped to form a protective film against corrosion. The main compound formed was siderite FeCO<sub>3</sub>, which reduces the corrosion rate in steels [26–29].

Concerning Raman spectra of corrosion products, carbon dioxide, when in contact with water, forms carbonic acid; once dissociated, one of its ions is CO<sub>3</sub><sup>2-</sup> which, when reacting with Fe<sup>2+</sup>, results in carbonate of iron [30]. Such a compound could be detected by Raman spectroscopy at some wavelengths, namely, 290, 520, 734, and 1080 cm<sup>-1</sup>, due to a passivating film on the surface of metals [31].

The surface of carbon steel metals was evaluated using Raman spectroscopy. The medium used for the electrochemical tests was NaCl 3.5% and saturated CO<sub>2</sub>. The result obtained in the characterization was the formation of a magnetite (Fe<sub>3</sub>O<sub>4</sub>) and hematite (Fe<sub>2</sub>O<sub>3</sub>) film [32]. Such compounds can be explained by the decomposition of FeCO<sub>3</sub> [33], namely, by the (3a)–(3d) reactions:



The transformation of FeO takes place by reaction (3b) [34]:



In the presence of oxygen, the transformation to Fe<sub>2</sub>O<sub>3</sub> occurs.



Such characteristics help to understand the formation of oxides on the surface of metals and the possibility of finding them in characterization techniques, such as Raman.

Next, the spectra of the tested samples are cited with their respective acquisition regions. Three samples were used, according to the submitted test. It is noteworthy that the methodology used to obtain the spectra focused on submitting to the Raman spectrum those samples that obtained intermediate resistance to corrosion in a 3.5% NaCl solution saturated with CO<sub>2</sub> at 80 °C. This approach was based on the assumption that, in this way, it would be more likely to find oxides or salts that increased or decreased corrosion resistance.

Thus, these are the spectra and their respective acquisition references, according to the literature: generalized corrosion region (378 [35]; 700 [36]; 1300 [37,38]; 1580 [38]); region without apparent corrosion (252 [39]; 387 [40–44]; 667 [45]; 1332 [40]); region of localized corrosion (407 [37–39]; 710 [36]; 930 [23]; 1380 [38]; 1590 [38]).

The spectra were evaluated according to the color variation of the images obtained by the microscope of the equipment. Thus, the laser focus regions of the equipment were defined and named to obtain the spectra according to the type of corrosion found, namely, generalized corrosion region, no apparent corrosion, and localized corrosion.

## 5. Conclusions

Based on the results discussed in this research for the three wires, the following conclusions can be reached.

Under conditions of temperatures different from ambient, the pearlitic microstructure with lamellar cementite presents less susceptibility to corrosion, if compared to the other forms of cementite.

When the SC wire was subjected to an aerated environment and at room temperature, it showed lower corrosion resistance compared to other wires in the same condition (lamellar and brittle). This raises a question about CO<sub>2</sub> gas serving as a protective element for the spheroidized microstructure, since the performance of this sample in relation to corrosion was better than the other samples when in the presence of carbon dioxide.

The results of Raman spectroscopy identified the formation of siderite in the most severe corrosion condition (NaCl + CO<sub>2</sub>/80 °C), which leads to the belief that the siderite salt reduces the advance of generalized corrosion on the surface of the SC sample.

Siderite peaks were also observed in the region with localized corrosion (pitting) and not in the region with no apparent degradation, showing that this salt does not contribute to increasing the corrosion resistance of the material.

Several oxides were also found on the surface of the analyzed SC sample, especially lepidocrocite and ferrihydrite.

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