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# Preparation of *n*-Alkane/Polycaprolactone Phase-Change Microcapsules via Single Nozzle Electro-Spraying: Characterization on Their Formation, Structures and Properties

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**Abstract:** The phase change microcapsule (mPCM) is one of the primary candidates in the fields of energy storage and thermal regulation. In this study, electro-spraying, as a green, high-efficiency electrohydrodynamic atomization technology, is applied to the microencapsulation of two phase change materials (PCM) (*n*-hexadecane and *n*-eicosane) with three loading contents (30%, 50%, and 70% by weight) in a polycaprolactone matrix. Ethyl acetate (EA) and chloroform (Chl) were chosen as solvents to prepare the working solutions. The objective of this study is to clarify the microencapsulation process during electro-spraying and to optimize the structure and properties of the electro-sprayed mPCM. The structures, morphologies, and thermal properties of the mPCM were characterized by optical microscopy (OM), scanning electron microscopy (SEM), differential scanning calorimeter (DSC), thermogravimetric analysis (TGA), and fourier transform infrared spectroscopy (FT-IR). Electro-sprayed spherical and non-porous mPCM have been successfully prepared. The mean diameter and the particle size distribution depend mainly on the choice of the *n*-alkane, as well as the solvent used to prepare the working solutions. Meanwhile, the structure formation of electro-sprayed mPCM and the loading content of PCM were mainly influenced by the evaporation of the solvent and the phase separation between PCM and poly(caprolactone) (PCL) matrix. During the shell formation or PCL solidification, the control of the PCM leaching out of the matrix allows improving the loading content. Finally, based on a high latent heat and simple formation process, the electro-spraying route of PCM is a green, non-toxic, and high-efficiency direction for energy storage and heat regulation.

**Keywords:** phase-change microcapsules; electro-spraying; poly(caprolactone); paraffin wax; formation mechanism; structure and morphology; thermal properties

## 1. Introduction

Phase change materials (PCM) are inorganic or organic compounds, which can store or release a large amount of latent heat during their phase change transition induced with the surrounding temperature variations [1–5]. Thus, they are mainly used in space applications, smart textile structures, battery thermal management, industrial waste heat recovery, building energy conservation, and solar energy, in order to manage the thermal management due to their high energy storage density [6–9]. The use of organic phase change materials has some limitations due to their low thermal conductivity, low density, high volatilization, and flammability [10,11]. Since they are sometimes in their liquid phase, they cannot be applied on a substrate without being stabilized, which overcomes their drawbacks [12–16]. The PCM shape stabilization technique is one of the well-known ways to achieve the best benefits in energy storage in practical applications [17]. The PCM shape stabilizing materials can be obtained from

various routes. Their microencapsulation improves heat transfer by increasing the surface-to-volume ratio [18]. Their incorporation into a polymer matrix is a low-cost technology. The use of nanomaterials in composite design improves their stability [19], while their impregnation into porous materials prevents their leakage, as well as the formulation of solid-solid phase change materials.

The encapsulation step leads, not only to protect the active substance and to prevent leaching, but also to an increase in the specific surface contact area, and therefore, improve the thermal transfer management. Microencapsulation of PCM is mainly carried out via in-situ polymerization [18], interfacial polycondensation [20], sol-gel method [21,22], simple or complex coacervation [23,24], phase separation or suspension process [25,26], and spray drying [27–29]. These methods have some drawbacks in the preparation of phase-change microcapsules, such as the complex shell formation process, the use of toxic substances or organic solvent, chemical additives and surfactants, and the difficulty in the control of structure, as well as morphology (i.e., size, shape, and porosity). Electro spraying method, as a one-step electrohydrodynamic atomization technology, has great potential and advantages to fabricate polymeric microcapsules or nanocapsules in a green and simple way [30–33]. Furthermore, the control of the process and formulation parameters avoids particle aggregation and ensures a narrow and mono-dispersed particle size distribution under stable cone-jet mode. It also ensures the the shape, the surface state and porosity, and phase structure of microcapsules to be adjusted [34,35].

Among the electrohydrodynamic atomization technologies, the electrospinning method is the most used with the phase change materials to prepare core-sheath structure nanofibers from a coaxial system [36–38]. The use of electro spraying for PCM microencapsulation, even if some authors recognize it as a suitable method, is seldom described in the scientific literature [39]. In their works, Moghaddam et al. prepared *n*-nonadecane/sodium alginate mPCM, using melting coaxial electro spraying [39–41]. The formation of the shell was carried out via the cross-linking solidification reaction between alginate solution (outer liquid) and calcium chloride in collecting bath. The investigation of the alginate solution concentration and the optimization of the process parameters, controlling the mean diameter of the particles, allowed them to reach an encapsulation efficiency of about 59%, a phase change enthalpy of their capsules about 88 J/g, and a mean diameter between 200 to 700  $\mu\text{m}$ . *N*-octadecane was successfully entrapped in a polyamic acid shell via a coaxial electro spraying process [42]. The authors observed that the core feed rate affected the surface morphology as well as the mean diameter of the capsules. Furthermore, microcapsules with a mean diameter between 1 to 2  $\mu\text{m}$ , with narrow size distribution and a loading content of 60%, were obtained.

In single nozzle electro spraying, a solvent able to solubilize both polymer shell and PCM is used to form the working solution. Then, this solution is pumped out from the electro-sprayed nozzle with an applied voltage to prepare mPCM. During the evaporation of the solvent, the solvent's richest phase diffuses to form a shell, and poor solvent phase stays in the inner part of the microcapsules to form the core, due to the solubility differences between matrix and PCM in the working solvent. The cone-jet mode is challenging to reach when compounds with low electrical conductivity, such as PCM, are used. Therefore, the use of a suitable solvent was required to be solubilized and improve its electro-sprayability. Thus, the formation and the final properties of electro-sprayed mPCM particles are not only determined by the feed ratio between the polymeric shelling the core, but also by the breakup of electro-sprayed droplets, the evaporation of solvent, the diffusion of PCM and the phase separation, to lead the formation of the shell around the PCM during the electro spraying process. Until now, there is no publication to investigate the effects of these processes mentioned later on the structure characteristics (mean diameter, morphology, and loading content of PCM) and thermal properties of the electro sprayed mPCM.

In this work, the design of PCM microencapsulation from a single nozzle electro spraying process, based on the entrapment of PCM-loaded microparticles in a polyester matrix or shell was proposed, and its impact on the resulting thermal properties was assessed. *N*-hexadecane and *n*-eicosane have been selected as liquid and solid PCMs at room temperature, ethyl acetate (EA), and chloroform (Chl)

as solvents for the working solutions, and polycaprolactone (PCL) as shell material. The loading content, as well as the encapsulation efficiency, were also investigated according to the core/shell ratio. The structures, morphologies, and properties of electro-sprayed MPCM were characterized by optical and scanning electron microscopies (OM and SEM), the chemical structure of the obtained microparticles were assessed by Fourier transform infrared spectroscopy (FT-IR), and the thermal properties by differential scanning calorimetry (DSC) and thermogravimetric analysis (TGA).

## 2. Materials and Methods

### 2.1. Materials

Polycaprolactone (PCL), used as matrix or shell polymer (Capa™ 6400, mean molecular weight  $37,000 \text{ g}\cdot\text{mol}^{-1}$ ) was purchased from Perstorp (Malmö, Skåne County, Sweden). Ethyl acetate (EA) and chloroform (Chl) were purchased from Sigma Aldrich (Dorset, UK). *N*-hexadecane (C<sub>16</sub>-parafol 16-97) and *n*-eicosane (C<sub>20</sub>-parafol 20z) were obtained from Sasol Germany GmbH (Hamburg, Germany). All reagents were used as received without any further treatment.

### 2.2. Preparation of PCM/PCL Electro spraying Solutions

A series of PCM/PCL electro spraying solutions with different weight ratios, PCM to PCL, were prepared from the solubilization of PCL in ethyl acetate or chloroform (10 wt.%) at 40 °C under vigorous magnetic stirring during 1 h. The required weight of PCM was subsequently added to achieve final PCM/PCL weight ratios of 0/100, 30/70, 50/50, and 70/30, under continuous agitation during 2 h.

### 2.3. Production of Microencapsulation Structures by Electro spraying

The previous formulations were processed using a CAT000002 Electro spraying Starter Kits purchased from Spraybase® AVECATS (Kildare, Ireland). The solutions were pumped with a digitally controlled syringe pump at a flow-rate of 1 mL/h through a single stainless steel nozzle (26 gauge, single nozzle, outer diameter: 460 µm, inner diameter: 260 µm). The applied voltage between nozzle and collector was set 4.0 kV. Samples were collected on a grounded copper plate covered with aluminum foil for SEM analyses, with glass slides for OM analyses and metal dish for the other characterizations, with a collecting time of 10, 3, and 300 min, respectively. All the containers were placed at 17 cm from the tip of the nozzle, and during the various experiments, temperature and relative humidity were kept constant at 25 °C, and 45%, respectively.

### 2.4. Morphological Characterization of the Particles

Scanning electron microscopy was conducted on an Inspect F50 (FEI Company, Hillsboro, OR, USA) at an accelerating voltage of 20 kV. The microparticles samples were coated with gold before analysis. Electro-sprayed phase-change microparticles were observed via optical microscopy (Motic BA410, Barcelona, Spain) coupled to a Moticam 5 Digital Camera (Barcelona, Spain), controlled by the software Motic image plus 3.0 which was used to analyze the microparticles. Particles diameters were measured from the OM and SEM micrographs using the Image J software. Size distributions were obtained from a minimum of 200 random measurements. The obtained data were statistically analyzed in a spreadsheet (Microsoft Excel 2011, Microsoft Company, Redmond, WA, USA) to calculate the mean diameter as well as the particle size distribution.

### 2.5. Fourier Transform Infrared (FT-IR) Analysis of the Samples

The chemical structure of electro-sprayed mPCM was analyzed by Fourier transform infrared spectroscopy (Nexus-560 spectrometer, Nicolet, Madison, WI, USA). The powdery electro-sprayed materials were dispersed in spectroscopic grade potassium bromide (KBr), and analyzed in absorbance mode. All spectra were acquired in the wavenumber range of 400–4000  $\text{cm}^{-1}$ , by averaging 256 scans at 2  $\text{cm}^{-1}$  resolution.

## 2.6. Thermal Analyses of the Samples

The thermal behavior of *n*-hexadecane, *n*-eicosane, PCL microparticles and mPCM was recorded using a DSC 3+ (Mettler Toledo, Columbus, OH, USA) piloted on PC with STARe software. The sample space was purged with nitrogen flow during the experiments. The onset temperatures of solid-liquid and liquid-solid transitions ( $T_{onset}$ ), end temperatures of phase transitions ( $T_{end}$ ), and maximum temperatures of melting ( $T_m$ ) and crystallization ( $T_c$ ), as well as melting ( $\Delta H_m$ ) and crystallization ( $\Delta H_C$ ) enthalpies of each transitions were determined. 3–8 mg of microcapsules (or *n*-alkanes) were placed in an aluminum pan and placed in DSC apparatus, heated from  $-20$  °C to  $80$  °C at  $5$  °C·min $^{-1}$ , held constant at  $80$  °C for 10 min, and then cooled to  $-20$  °C at the same scanning rate, and held for 10 min. Each test was conducted twice and carried out with a least three independent experiments. Values were taken on every second cycle. On the base of phase transitions enthalpies, the loading content (LC) as well as the encapsulation efficiency (EE) were calculated from the Equations (1) and (2) [43],

$$LC = \frac{\Delta H_{mPCM}}{\Delta H_{PCM}} \quad (1)$$

$$EE = \frac{LC}{LC_{th}} \quad (2)$$

where,  $\Delta H_{mPCM}$  and  $\Delta H_{PCM}$  are the melting enthalpies of microparticles, and *n*-alkanes (*n*-hexadecane  $\sim 199$  J·g $^{-1}$ , and *n*-eicosane  $\sim 205$  J·g $^{-1}$ ), respectively.  $LC_{th}$  is the theoretical loading content.

The relative crystallinity of the PCL matrix in the samples ( $X_c$ ) was calculated according to Equations (3) and (4),

$$X_{c(m)} = \frac{\Delta H_m (PCL \text{ phase in } mPCM)}{\Delta H_{m(PCL)}^0 (1 - w_m (m))} \times 100 \quad (3)$$

$$X_{c(th)} = \frac{\Delta H_m (PCL \text{ phase in } mPCM)}{\Delta H_{m(PCL)}^0 (1 - w_m (th))} \times 100 \quad (4)$$

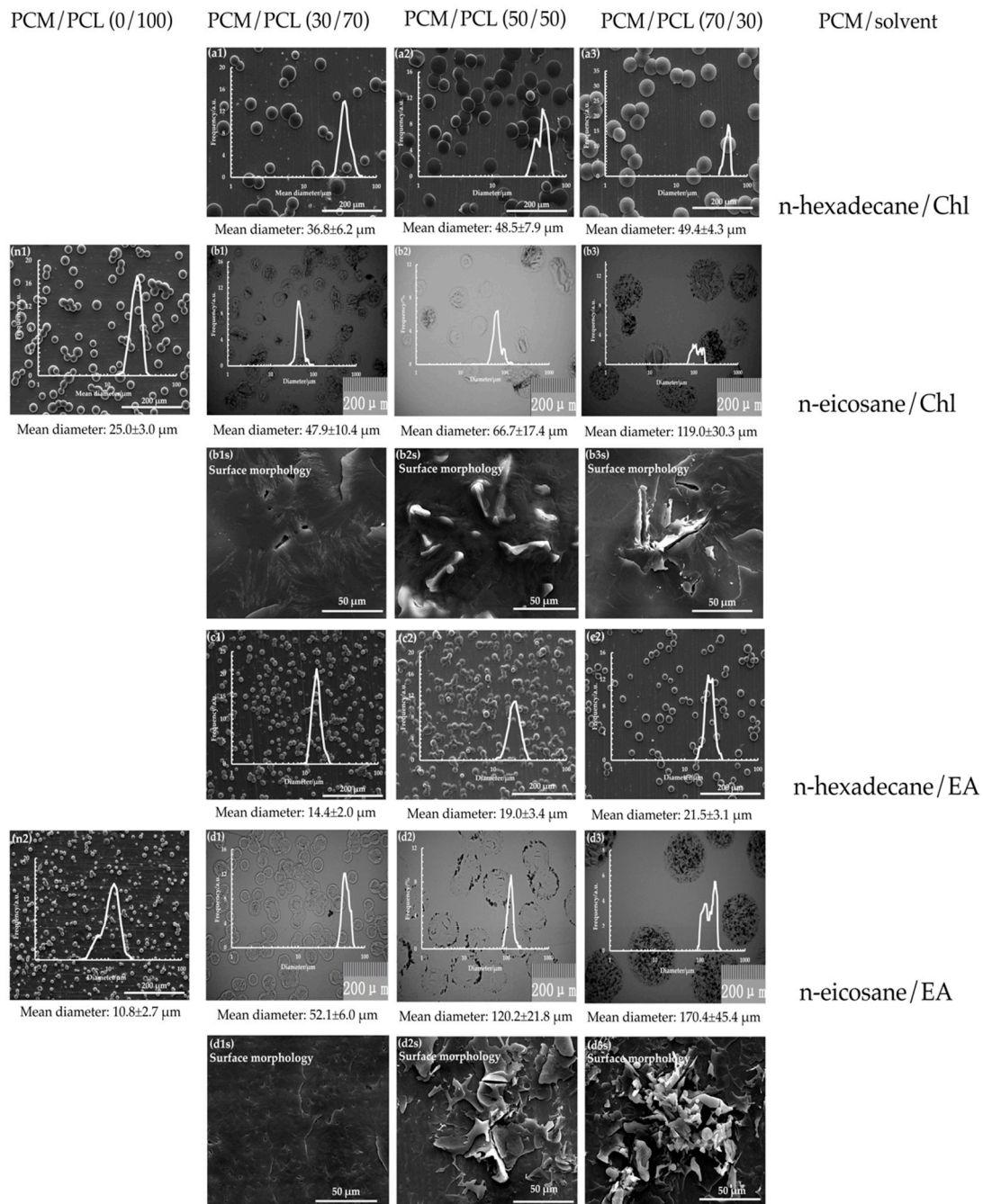
where,  $\Delta H_m$  is the specific melting heat,  $\Delta H_{m}^0$  is the theoretical specific heat of 100% crystalline PCL, which was taken as  $139.5$  J·g $^{-1}$  [44], and  $w_m$  was the weight fraction of PCM in the samples.

The thermal stability of *n*-alkanes, PCL microparticles and mPCM were assessed by thermogravimetric analysis, carried out on a TGA 3+ (Mettler Toledo, Columbus, OH, USA) under inert atmosphere at a purge rate of  $50$  mL·min $^{-1}$ . A heating rate of  $10$  °C·min $^{-1}$  was applied, and the temperature raised from  $20$  to  $700$  °C. For each experiment,  $3$  to  $8$  mg was used in the TGA test. In addition to the percentage weight (TG) curves, the derivative weight loss percentage (DTG) was calculated for each sample.

## 3. Results and Discussion

### 3.1. Size and Morphology of Electro-Sprayed mPCM

The size, the surface state, and the morphology of electro-sprayed mPCM were characterized by SEM and OM observations (Figure 1). They are mainly influenced by the PCM content, the phase change temperature of the PCMs used during the process, and also the choice of the solvents. Blank PCL particles present a spherical shape, whatever the solvent selected to solubilize PCL, i.e., chloroform or ethyl acetate (Figure 1(n1,n2)). The use of chloroform allows the particles to have a smooth surface state. In contrast, the solubilization in ethyl acetate solution leads to the formation of a single-pore morphology as previously described by Zhang et al. [45]. Adding PCMs in the working solution does not have a substantial effect on the obtained morphology, since the obtained particles are also spherical. Nevertheless, the increase of *n*-alkanes in the solution leads to an increase in the mean diameter and the particle size distribution. Furthermore, in the case of the use of EA as the solvent, single-pore morphology turns to a non-porous morphology for PCM/PCL blends.



**Figure 1.** Scanning electron microscopy (SEM) and optical microscopy (OM) graphs of neat poly(caprolactone) (PCL) microspheres and mPCM obtained from different conditions (n1) neat PCL microspheres obtained from chloroform (n2) neat PCL microspheres obtained from ethyl acetate (a1–a3) *n*-hexadecane/PCL/chloroform microcapsules (b1–b3) *n*-eicosane/PCL/chloroform microcapsules (b1s–b3s: Their surface morphologies) (c1–c3) *n*-hexadecane/PCL/ethyl acetate microcapsules (d1–d3) *n*-eicosane/PCL/ethyl acetate microcapsules (d1s–d3s: their surface morphologies) (scale bar: 200 μm) (b1s–b3s and d1s–d3s) the magnification SEM images of related sample surface morphology (scale bar: 50 μm).

The use of *n*-hexadecane and *n*-eicosane does not lead to the same granulometric properties, even if their molecular chain length differs by four carbon atoms. *N*-hexadecane is in its liquid state at room temperature, while *n*-eicosane is in its solid state. Under the action of Coulombic repulsions, the break-up process of *n*-eicosane/PCL charged droplets is different from that of *n*-hexadecane/PCL



one. During the flying process, from the nozzle to the collector, droplets containing *n*-eicosane start their solidification at the earliest stages due to the evaporation of the solvent, which induces an enhancement of *n*-eicosane chain interactions and, therefore, their crystallization. Therefore, it leads to a higher specific surface to volume ratio, which prevents the further break-up process of the electro-sprayed droplets, unlike *n*-hexadecane, which is in a liquid state throughout the process. Furthermore, the crystallization being an exothermic phenomenon increases the local temperature in the inner core of the droplets, which leads to a growth of the solvent evaporation rate. Therefore, the use of *n*-eicosane leads to the formation of particles having a higher mean diameter and particle size distribution than with *n*-hexadecane one. Since solid particles are obtained in the earliest stages of the flying process, in the last steps, the solvent release results in the obtaining of the rough and wrinkled surface state of the particles, which is more pronounced with the increase of *n*-eicosane content. Thus, the black color observed in OM micrographs of *n*-eicosane/PCL particles belonged to *n*-eicosane laying on the surface of the PCL matrix (Figure 1(b3,d3)). At low content in *n*-eicosane, the core compound is completely entrapped by the PCL shell, but when it increased due to the thermodynamical phenomena and solvent evaporation, it diffuses on the surface of the particles.

The granulometric properties of the electro-sprayed particles, as well as their morphology, are also affected by the solvent choice, that is used to solubilize the polymer and the *n*-alkane. Thus, in the presence of ethyl acetate (Figure 1(c1–c3)), the obtained particles with *n*-hexadecane have a lower mean diameter and a narrower size distribution than with chloroform (Figure 1(a1–a3)). This is related to the higher vapor pressure solution and viscosity of the chloroform based solution [45]. During the earliest stages of the flying process, chloroform evaporates quicker in the surrounding medium compared to ethyl acetate solution, and in the absence or at low solvent content, the process break-up stops before to reach the collector. Furthermore, high viscosity also depresses the break-up process of the charged droplets, resulting in a higher size of the electro-sprayed particles [46].

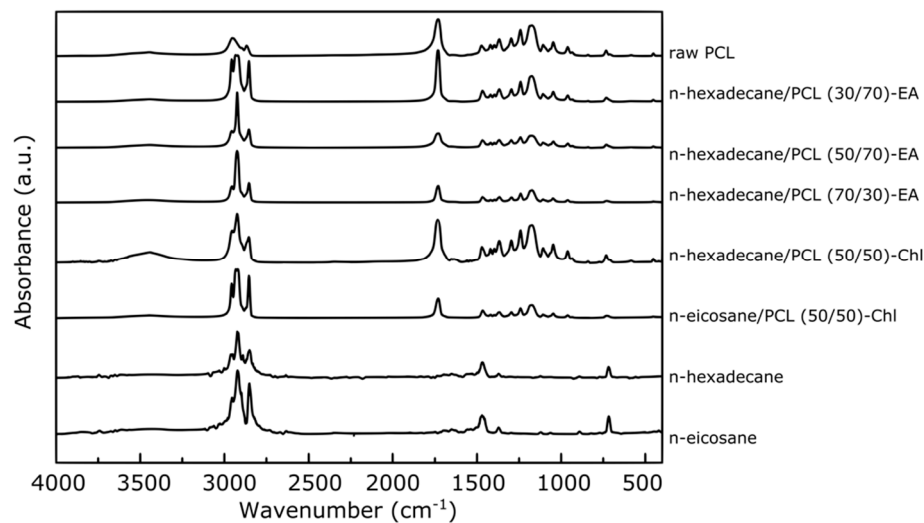
On the other hand, the mean diameter and the particle size distribution of the electro-sprayed *n*-eicosane particles are determined in the earliest stage of the process due to core compound crystallization, which allows solid particles to be obtained. Then, in the presence of chloroform (Figure 1(b1–b3)), the process break-up coupled to the solvent evaporation allows the formation of tinier particles than with ethyl acetate (Figure 1(d1–d3)) since more solvent is released.

Also, the crystallization of *n*-eicosane plays a primary role in the obtained morphology. Thus, *n*-eicosane/PCL microcapsules, obtained from ethyl acetate, are rougher and more wrinkled than those obtained from chloroform (Figure 1(b1s–b3s) and Figure 1(d1s–d3s)). During the solidification of the *n*-alkane, ethyl acetate gradually is released from these particles until reaching the collector, which allows the diffusion of the *n*-alkanes chains to the surface of the particles. Whereas, in the presence of chloroform, due to its quick evaporation, *n*-eicosane is located in the inner core of the particles.

### 3.2. Structure of Electro-Sprayed *m*PCM

FT-IR spectra of PCL microparticles, *n*-hexadecane/PCL microparticles in EA (30/70, 50/50; 70/30), *n*-hexadecane/PCL microparticles in Chl, and *n*-eicosane/PCL microparticles in Chl are depicted in Figure 2. PCL microparticles spectrum shows a strong absorption band at  $1750\text{ cm}^{-1}$  corresponding to the stretching vibration of carbonyl group, and characteristic bands at  $2915$  to  $2860\text{ cm}^{-1}$  associated with the aliphatic C–H stretching vibrations. The characteristic bands at  $1475\text{ cm}^{-1}$ ,  $1397\text{ cm}^{-1}$ , and  $1361\text{ cm}^{-1}$  are attributed to the bending vibration of the alkyl group, and those located at  $1230\text{ cm}^{-1}$ ,  $1107\text{ cm}^{-1}$ , and  $1042\text{ cm}^{-1}$  corresponding to the stretching vibration of –CO–C– group. Stretching vibrations of –C–C– and –CO– are found at  $1290$ , and  $1160\text{ cm}^{-1}$ , respectively [47]. PCM spectra exhibit a strong absorption band at  $2920$ – $2852\text{ cm}^{-1}$  associated with the aliphatic C–H stretching vibrations of the *n*-hexadecane or *n*-eicosane. The in-plane rocking vibration of the CH<sub>2</sub> groups is observed at  $717\text{ cm}^{-1}$ , and C–H bending vibration in CH<sub>2</sub> is found at  $1470$  and  $1370\text{ cm}^{-1}$ . All of the characteristic peaks for *n*-hexadecane, or *n*-eicosane, and PCL, are clearly distinguished in the spectra of the FT-IR spectra of

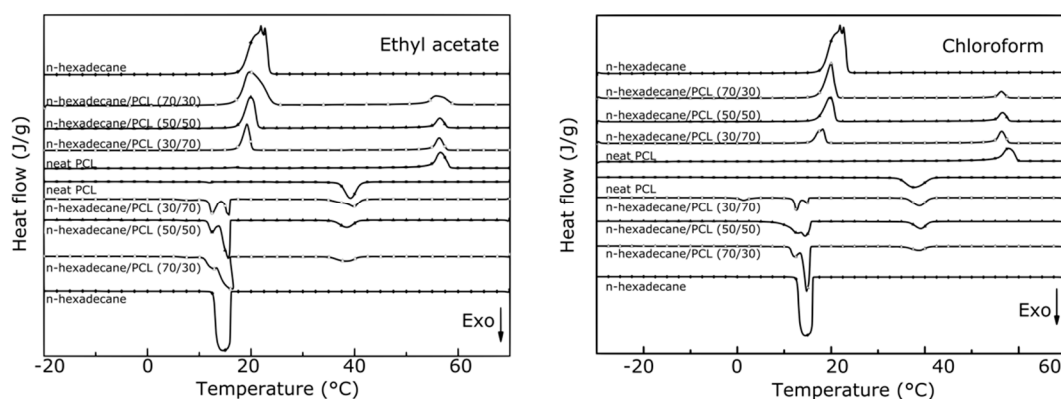
the electro-sprayed micro-particle samples, which verifies that *n*-alkane are successfully entrapped by PCL.



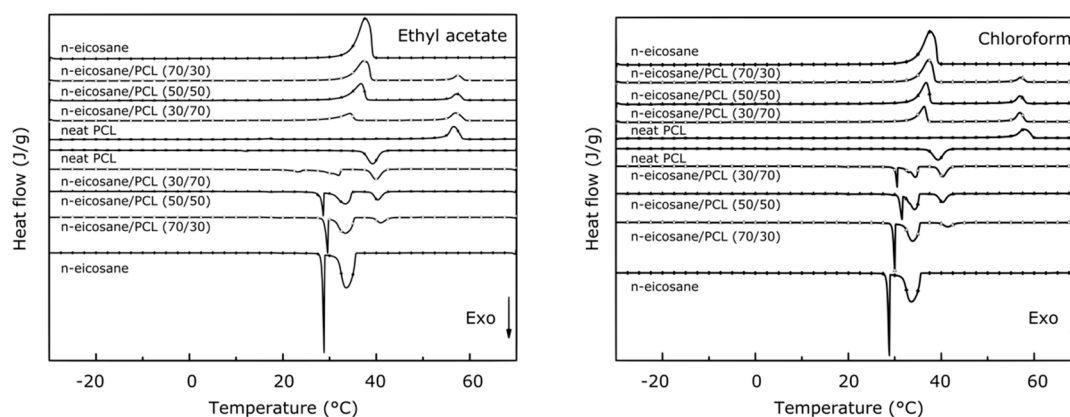
**Figure 2.** The fourier transform infrared spectroscopy (FT-IR) spectra of *n*-eicosane, *n*-hexadecane, neat PCL microsphere and a series of mPCM obtained from different conditions.

### 3.3. Phase Change Properties of Electro-Sprayed mPCM

The phase change temperatures and the latent heats of melting and crystallization of raw *n*-alkanes, neat PCL microspheres, and mPCM were obtained using differential scanning calorimeter. They are presented in Figures 3 and 4, and gathered in Tables 1 and 2. On the one hand, the latent heats of fusion of raw *n*-hexadecane and *n*-eicosane are 199, and 205 J·g<sup>-1</sup>, respectively. Their melting and crystallization temperatures are found at 17.9 °C, 16.2 °C, 34.3 °C, and 35.7 °C. On the other hand, the onset of melting and crystallization temperatures of PCL microspheres are determined at 54.9 and 41.2 °C, when they have been prepared from ethyl acetate solution, and 55.6 and 41.5 °C from chloroform. The solvent used modifies to a limited extent the crystallinity since it is about 40.8% and 42.7% for the samples prepared from EA, and Chl, respectively. DSC thermograms of microencapsulated PCM have both the thermal transitions of the *n*-alkanes used and PCL.



**Figure 3.** DSC curves of *n*-hexadecane, neat PCL microspheres and a series of *n*-hexadecane/PCL microcapsules obtained from different solvents with different mass ratios between *n*-hexadecane and PCL.



**Figure 4.** Differential scanning calorimeter (DSC) curves of *n*-eicosane, neat PCL microspheres and a series of *n*-eicosane/PCL microcapsules obtained from different solvents with different mass ratios between *n*-eicosane and PCL.

The increase of *n*-hexadecane to PCL from 30 to 70 wt.% in the EA working solution results in the increase of the melting (and crystallization) enthalpy from 52 to 113 J·g<sup>-1</sup>, corresponding to an increase of the loading content from 26 to 56%, which is a quite lower than the theoretical content. On the other hand, the encapsulation efficiency reaches a maximum of 89% for 50 wt.%, then decreases to 81% for 70 wt.%. The loading content does not have a significant effect on the phase change transition temperatures, which are close to the raw *n*-hexadecane one. The mPCM obtained from chloroform solution have lower melting enthalpy than those with ethyl acetate. The encapsulation efficiency is slightly lower too. Furthermore, the melting and crystallization temperatures for low loading content are found relatively lower compared to the other content. The loss of *n*-hexadecane during electro-spray microencapsulation results from the evaporation of the solvent in flight, which promotes phase separation between *n*-hexadecane and PCL, and sometimes the presence of pure tiny droplet of *n*-hexadecane. The rigidification of the PCL hinders this phenomenon. *n*-hexadecane is highly soluble in chloroform than in ethyl acetate, and therefore, has a better compatibility [48,49]. Thus, the use of the chloroform working solution induces a higher loss in *n*-hexadecane during its evaporation. The use of ethyl acetate may prevent this phenomenon, and therefore, higher loading content in PCM in the microcapsules, are reached.

Apart from *n*-hexadecane/PCL microcapsules, the thermal properties of *n*-eicosane-based mPCM, obtained from different solvents and various weight ratios between the *n*-alkane and PCL, are also presented in Figure 4 and Table 2. On the one hand, the presence of the PCL and the choice of the solvent used for the working solution does not affect the phase temperatures, which are quite similar to the raw *n*-hexadecane one. On the other hand, the use of chloroform allows entrapping higher content of *n*-eicosane (more than 10%) compared to the use of ethyl acetate. These results are the opposite of those obtained in the case of *n*-hexadecane, highlighting the influence or importance of the phase change that takes place during the process. Thus, the solvent evaporation, which occurs in the early stages of the electro-spraying process, induces a fast phase separation and further the PCM crystallization, which also promotes the stiffening of the PCL. In this context, the more volatile the solvent of the working solution, the more PCM is entrapped in the microparticles. Furthermore, the calculated loading content of *n*-eicosane when working with chloroform is quite the same as those obtained with *n*-hexadecane and ethyl acetate. Higher encapsulation efficiencies are obtained.



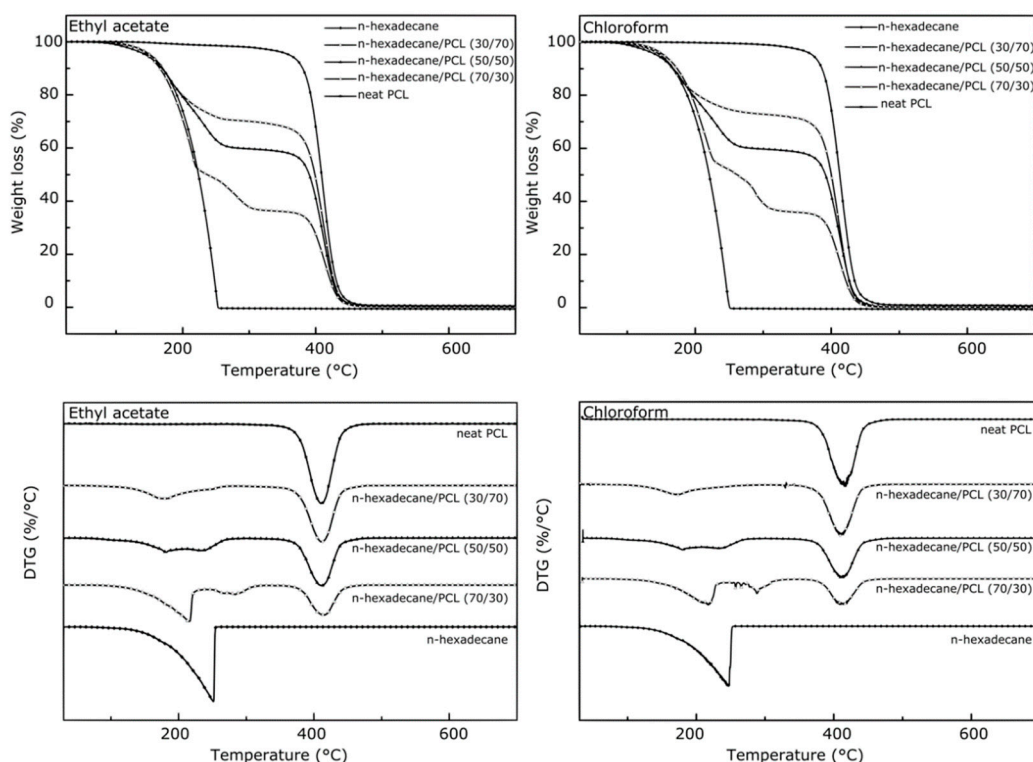
**Table 1.** Thermal properties of *n*-hexadecane, neat poly(caprolactone) (PCL) microspheres and a series of *n*-hexadecane/PCL microcapsules obtained from ethyl acetate and chloroform solutions at three weight ratios between *n*-hexadecane and PCL.

Sample Label		Latent Heat (J·g <sup>-1</sup> )		<i>T</i> <sub>onset</sub> (°C)	<i>T</i> <sub>c</sub> or <i>T</i> <sub>m</sub> (°C)		<i>T</i> <sub>end</sub> (°C)		<i>LC</i> (%)	<i>EE</i> (%)	<i>X</i> <sub>c(th)</sub> (%)	<i>X</i> <sub>c(m)</sub> (%)	
<i>n</i> -hexadecane	heating	199.4 ± 2.2		17.9	20.7		23.7						
	cooling	195.9 ± 4.7		16.2	16.0		13.0						
PCL-EA	heating	57.0 ± 1.1		54.9	56.4		58.5						
	cooling	54.7 ± 1.3		41.2	39.5		37.2						
<i>n</i> -hexadecane/PCL(30/70)-EA	heating	52.1 ± 7.3	32.5 ± 1.5	17.7	55.0	19.9	56.2	20.2	57.7	26.1 ± 3.7	87.1 ± 12.2	33.3 ± 1.5	31.5 ± 1.5
	cooling	51.7 ± 6.6	33.6 ± 3.9	16.0	41.5	15.9	39.9	14.7	37.1				
<i>n</i> -hexadecane/PCL(50/50)-EA	heating	88.7 ± 4.1	25.0 ± 0.6	17.6	55.0	19.2	56.2	21.2	57.8	44.5 ± 2.0	89.0 ± 4.1	35.8 ± 0.8	32.0 ± 0.7
	cooling	88.1 ± 1.8	25.7 ± 0.5	16.1	40.9	16.4	38.7	14.0	36.0				
<i>n</i> -hexadecane/PCL(70/30)-EA	heating	113.2 ± 7.2	24.1 ± 1.1	17.5	55.0	19.5	56.2	22.7	58.4	56.8 ± 3.7	81.1 ± 5.2	57.7 ± 2.8	40.0 ± 1.9
	cooling	111.7 ± 5.1	25.9 ± 1.8	16.1	40.8	16.4	38.8	13.6	35.8				
PCL-Chl	heating	59.6 ± 2.0		55.6	57.6		59.8						
	cooling	57.2 ± 1.9		41.5	37.9		34.4						
<i>n</i> -hexadecane/PCL(30/70)-Chl	heating	35.6 ± 0.7	31.6 ± 1.6	15.9	54.9	18.1	56.2	19.1	57.8	17.8 ± 0.4	59.4 ± 1.2	32.4 ± 1.6	27.6 ± 1.4
	cooling	35.3 ± 1.0	33.2 ± 1.6	13.6	41.2	12.8	39.0	11.9	36.1				
<i>n</i> -hexadecane/PCL(50/50)-Chl	heating	72.4 ± 0.6	25.0 ± 0.3	17.2	55.0	19.2	56.3	21.0	58.0	36.3 ± 0.3	72.6 ± 0.6	35.8 ± 0.4	28.1 ± 0.1
	cooling	71.1 ± 0.4	26.0 ± 0.5	15.8	41.1	14.8	39.4	12.4	36.6				
<i>n</i> -hexadecane/PCL(70/30)-Chl	heating	99.4 ± 5.7	17.0 ± 1.5	17.5	55.0	19.5	56.2	21.0	57.7	49.9 ± 2.8	71.2 ± 4.1	40.5 ± 2.7	24.2 ± 2.2
	cooling	96.9 ± 4.7	15.8 ± 0.4	15.7	41.0	15.5	38.8	13.8	35.6				

**Table 2.** Thermal properties of *n*-eicosane, neat PCL microspheres and a series of *n*-eicosane/PCL microcapsules obtained from ethyl acetate and chloroform solutions at three weight ratios between *n*-eicosane and PCL.

Sample Label		Latent Heat (J·g <sup>-1</sup> )		<i>T</i> <sub>onset</sub> (°C)	<i>T</i> <sub>m</sub> or <i>T</i> <sub>c</sub> (°C)		<i>T</i> <sub>end</sub> (°C)		<i>LC</i> (%)	<i>EE</i> (%)	<i>X</i> <sub>c(th)</sub> (%)	<i>X</i> <sub>c(m)</sub> (%)	
<i>n</i> -eicosane	heating	205.3 ± 1.7		34.3	36.9		39.2						
	cooling	204.2 ± 1.9		35.7	30.7		28.7						
PCL-EA	heating	57.0 ± 1.1		54.9	56.4		58.5						
	cooling	54.7 ± 1.3		41.2	39.5		37.2						
<i>n</i> -eicosane /PCL(30/70)-EA	heating	42.3 ± 2.0	32.1 ± 1.1	32.8	55.3	35.5	57.0	36.6	59.0	20.6 ± 1.0	68.6 ± 3.2	32.9 ± 1.1	29.0 ± 1.0
	cooling	41.2 ± 3.1	31.3 ± 2.7	34.2	41.4	28.1	40.3	27.3	38.5				
<i>n</i> -eicosane /PCL(50/50)-EA	heating	72.4 ± 7.3	21.2 ± 1.2	33.2	55.5	36.3	57.0	37.8	58.7	35.3 ± 3.4	70.6 ± 7.2	30.5 ± 1.7	23.5 ± 1.3
	cooling	72.1 ± 7.1	21.6 ± 1.6	34.8	42.0	29.4	40.6	28.2	39.0				
<i>n</i> -eicosane /PCL(70/30)-EA	heating	107.2 ± 2.4	17.8 ± 0.5	33.9	55.8	36.4	57.1	38.4	58.6	52.3 ± 1.1	74.7 ± 1.6	42.5 ± 1.3	26.7 ± 0.9
	cooling	106.9 ± 2.6	17.7 ± 0.7	35.4	42.9	31.5	41.2	30.0	39.6				
PCL-Chl	heating	59.6 ± 2.0		55.6	57.6		59.8						
	cooling	57.2 ± 1.9		41.5	37.9		34.4						
<i>n</i> -eicosane /PCL(30/70)-Chl	heating	60.2 ± 0.2	37.3 ± 2.4	34.0	55.4	36.0	56.7	37.1	58.1	29.3 ± 0.1	97.8 ± 0.4	38.2 ± 2.4	37.9 ± 2.4
	cooling	59.9 ± 1.3	38.0 ± 5.5	35.2	41.6	30.9	40.6	30.2	39.0				
<i>n</i> -eicosane /PCL(50/50)-Chl	heating	94.3 ± 0.8	28.1 ± 2.5	34.2	55.4	36.3	56.7	37.7	58.6	45.9 ± 0.4	91.9 ± 0.8	40.3 ± 3.5	37.2 ± 5.4
	cooling	93.1 ± 3.3	29.8 ± 4.1	34.5	41.9	32.1	40.6	30.9	39.0				
<i>n</i> -eicosane /PCL(70/30)-Chl	heating	121.1 ± 1.9	17.4 ± 0.6	34.3	55.6	36.7	57.0	39.6	58.3	59.0 ± 0.9	84.3 ± 0.7	41.6 ± 1.3	30.4 ± 1.0
	cooling	117.4 ± 1.2	17.2 ± 0.5	35.5	43.4	31.2	41.5	29.7	39.9				

Differential Scanning Calorimetry was also used to determine the effect of solvent and core compounds on the thermal properties of PCL microparticles. All data were determined from the second run, and the first run was used to clear the thermal history. The DSC thermograms (Figures 4 and 5) show single melting peaks with maximum melting start and end temperatures of 54.9 and 56.4 °C for ethyl acetate solution, which increase slightly to 55.6 and 57.6 °C for chloroform. The latent heat of fusion associated with these transitions is approximately 57.0 and 59.5 J·g<sup>-1</sup>, which corresponds to a crystallinity degree of 40.8% and 42.7%, respectively. The presence of *n*-alkanes in the microparticles does not affect the melting temperatures of the PCL, but at high *n*-hexadecane content, there is a slight decrease in crystallization temperatures, while *n*-eicosane acts as a nucleating agent of the PCL, and therefore the crystallization temperatures increase. However, *n*-alkanes play a significant role in the formation of the crystalline phase of PCL. The increase in the *n*-alkane content leads to a decrease in the crystalline phase. This trend is quite similar, regardless of the solvent and *n*-alkane used.



**Figure 5.** TG and DTG curves of *n*-hexadecane, neat PCL and the microencapsulated *n*-hexadecane.

In addition, the calculated crystallinity index of PCL in the mPCM is influenced by the solvent/*n*-alkane pair used, which is correlated with the crystallization rate during the electrospaying process. When the measured crystallinity index values ( $X_{PCL\ m}$ ) are compared with the calculated (theoretical) values ( $X_{PCL\ th}$ ), it can be seen that they are lower, except low weight ratio, where they are close. Thus, the presence of *n*-alkanes in the microparticles, and undoubtedly in the PCL matrix, reduces the PCL chain mobility, which gives rise to a lower crystallinity index in both the selected solvent and *n*-alkane systems.

Thus, the presence of *n*-alkanes leads to a decrease in the crystallinity index. In the early stages of the electrospay process, the droplets contain solvent, PCL macromolecular chains and PCM molecules. Depending on the solubility of PCM and PCL in each solvent system, the PCM molecules are dispersed together to form a mixed solution in which the macromolecular PCL chains have an intermolecular interaction or are entangled in the PCM chains when the solvent evaporates. The presence of PCM molecules reduces the mobility of PCL macromolecular chains and thus limits their availability for crystallization. This results in a more amorphous matrix [50–53]. In addition, the solvent used also

affects the crystallinity index, PCL and *n*-alkanes are more soluble in Chl than in EA, which leads to the preparation of the most homogeneous working solutions.

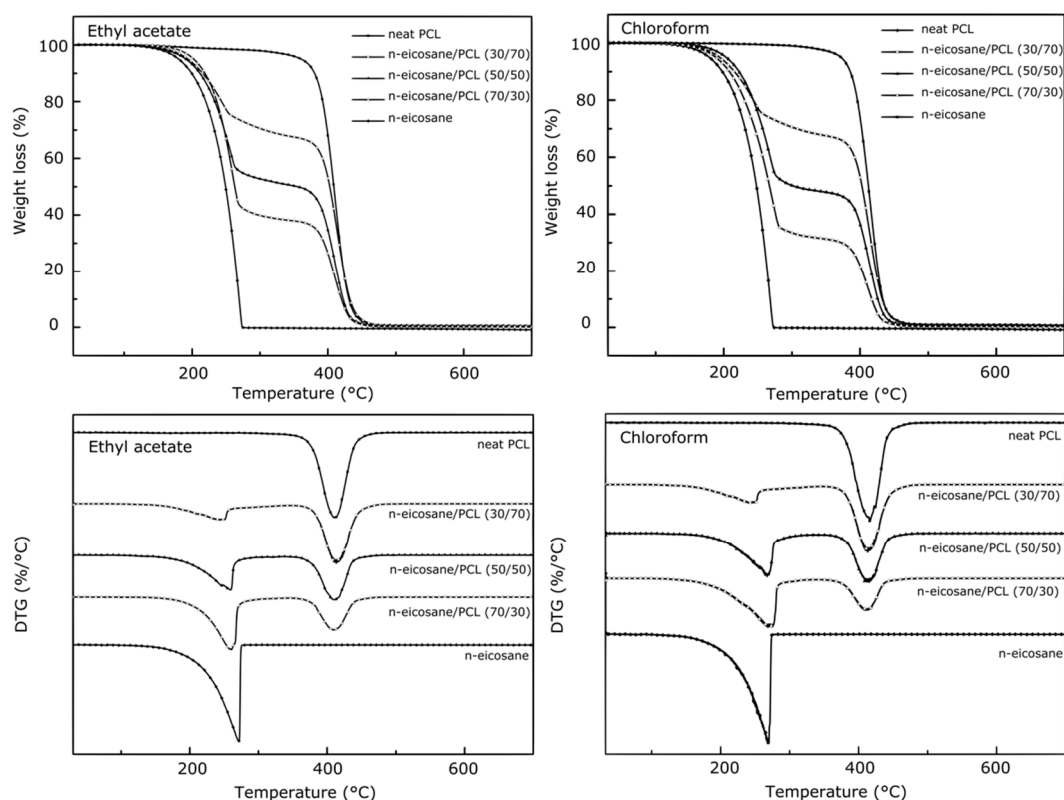
The rapid evaporation rate of Chl, during the early stages of the process, induces a higher phase separation between PCL and PCM, and therefore, a higher crystallinity index in the case of *n*-eicosane compared to the use of EA. The low evaporation rate of EA during the process leads to a higher degree of entanglement of the PCL macromolecular chains and PCM molecules, and therefore to a more amorphous matrix. Considering that *n*-hexadecane is liquid throughout the electro-spraying process, the molecules diffuse into the inner nucleus of the droplets during the slow evaporation of EA, allowing the PCL macromolecular chains to organize themselves to build the shell. On the other hand, the crystallinity index of PCL was found to decrease with increasing PCM content when Chl was used as a solvent. This observation suggests that the organization of the macromolecular chains is limited.

However, the difference between the theoretical content and the determined content of *n*-alkanes in the microparticles also suggests the presence of a PCM chain dispersed in the PCL matrix. Thus, the microparticles contain in the inner nucleus the *n*-alkanes, surrounded by a mixture of PCL/PCM chains, entrapped in a PCL shell.

#### 3.4. Thermal Stability of Electro-Sprayed mPCM

The thermal stability of raw PCM, neat PCL electro-sprayed microparticles and mPCM were analyzed by TGA. TG and DTG curves of PCL microparticles, *n*-hexadecane (or *n*-eicosane) and *n*-hexadecane (or *n*-eicosane) based microparticles from chloroform and ethyl acetate are shown in Figure 5 (or Figure 6). The thermal behavior of *n*-hexadecane and *n*-eicosane displays single step degradation under N<sub>2</sub> atmosphere. The decomposition of *n*-alkanes, characterized by the onset temperature at 5% weight loss, starts at 151.6 and 178.8 °C for *n*-hexadecane, and *n*-eicosane, respectively. The main decomposition peaks of *n*-hexadecane and *n*-eicosane corresponding to the *n*-alkanes backbone decomposition take place at 246 and 269, and the maximum rates of decomposition are about 2.2, and 2.6% °C<sup>-1</sup>, respectively. Furthermore, there is no residue for these compounds at over 280 °C. Under an inert atmosphere, PCL electro-sprayed microparticles display a single step degradation between 350 to 500 °C, corresponding to polymer pyrolysis with a maximum of degradation temperature at 411 or 413 °C, according to the solvent used in the electro-spraying process. The initial degradation temperatures ( $T_{onset}$ ) were found to occur at 359 and 369 °C for samples prepared from ethyl acetate and chloroform, respectively. Furthermore, it seems that the solvent used has a few influences on the temperatures and maximum rate degradation (Table 3).

*N*-alkanes based microparticles exhibit a two-step degradation. The first one, located from about 100 to 350 °C is ascribed to the PCM degradation, whereas the second one to the PCL matrix between 350 to 500 °C. The presence of PCL does not have an effect on the initial degradation temperature in the case of *n*-hexadecane, but a shift from 10 to 18 °C is observed for samples containing *n*-eicosane (Table 4). In the first degradation step, two types of phenomena occur, the first one corresponds to the degradation of *n*-alkane at the surface of the sample until 250 °C. When the temperature reaches 60–62 °C, PCL melts, and then all the microparticles samples are in a liquid state. Due to its lower density during the increase of temperature, a part of *n*-alkane chains can diffuse at the crucible surface and starts its degradation. The second phenomenon, between 250 to 300 °C, is related to the entrapped *n*-alkane chains in the PCL matrix and their entanglements, which required more thermal energy to be degraded since PCL acts as a barrier to prevent the degradation. In addition, with the increasing of *n*-alkane content, the maximum degradation rate in the first stage increases gradually; on the other hand, the second step is similar to that of the PCL.



**Figure 6.** TG and DTG curves of *n*-eicosane, neat PCL and the microencapsulated *n*-eicosane.

On the one hand, even if the initial degradation temperatures are close to the *n*-hexadecane, the maximum degradation temperatures in the first stage are significantly lower than the raw compound, increase with the increase in the loading content, and are not very dependent on the solvent used. On the other hand, when *n*-eicosane was used as a core compound, the initial thermal degradation temperatures are 10 to 18 °C higher than that of the raw compound one and decrease with the increase in the loading content. The maximum degradation temperatures in the first stage are 30 °C lower than *n*-eicosane one for low loading content and increase up to the *n*-alkane one at higher loading content. Furthermore, chloroform based microparticles have better thermal stability than the ethyl acetate-based one. This stability is related to the affinity between PCL and *n*-alkanes chains in the solvent solutions, which can favorize the entanglement during the electro-spraying process.



**Table 3.** Thermogravimetric data for of *n*-hexadecane, neat PCL microsphere and *n*-hexadecane microcapsules in inert atmosphere at a heating rate of 10 °C·min<sup>-1</sup>.

Sample	Initial Degradation Temperature (°C)	First Step			Second Step		
		Weight Loss (100–350 °C) (%)	Maximum Degradation Temperature (°C)	Maximum Degradation Rate (%/°C)	Weight Loss (350–500 °C) (%)	Maximum Degradation Temperature (°C)	Maximum Degradation Rate (%·°C <sup>-1</sup> )
PCL-EA	358.9 ± 3.1	4.2 ± 0.9	-	-	95.0 ± 1.1	411.4 ± 0.4	2.2 ± 0.1
PCL-Chl	369.2 ± 3.6	3.0 ± 0.4	-	-	96.1 ± 0.4	413.4 ± 1.6	2.3 ± 0.02
Raw <i>n</i> -hexadecane	151.6 ± 1.1	99.4 ± 0.1	246.6 ± 0.7	2.2 ± 0.1	-	-	-
<i>n</i> -hexadecane/PCL (30/70)-EA	150.8 ± 1.9	29.6 ± 1.7	174.3 ± 9.1	0.3 ± 0.1	69.6 ± 1.9	412.9 ± 3.4	2.0 ± 0.9
<i>n</i> -hexadecane/PCL (50/50)-EA	145.8 ± 2.3	39.5 ± 5.7	183.5 ± 10.3	0.5 ± 0.2	59.3 ± 6.3	411.9 ± 2.4	1.4 ± 0.2
<i>n</i> -hexadecane/PCL (70/30)-EA	149.5 ± 5.8	59.2 ± 4.3	214.9 ± 5.9	0.9 ± 0.2	39.9 ± 8.5	413.3 ± 1.4	1.0 ± 0.2
<i>n</i> -hexadecane/PCL (30/70)-Chl	150.1 ± 3.1	27.3 ± 0.9	173.5 ± 2.7	0.3 ± 0.03	72.0 ± 1.0	413.7 ± 0.3	1.8 ± 0.02
<i>n</i> -hexadecane/PCL (50/50)-Chl	147.3 ± 4.6	37.4 ± 2.4	177.5 ± 1.6	0.4 ± 0.04	61.8 ± 2.8	412.5 ± 1.5	1.5 ± 0.1
<i>n</i> -hexadecane/PCL (70/30)-Chl	151.8 ± 6.1	60.4 ± 4.0	217.0 ± 1.0	0.9 ± 0.1	38.9 ± 7.3	414.7 ± 2.4	0.9 ± 0.1

**Table 4.** Thermogravimetric data for of *n*-eicosane, neat PCL microsphere and *n*-eicosane microcapsules in inert atmosphere at a heating rate of 10 °C·min<sup>-1</sup>.

Sample	Initial Degradation Temperature (°C)	First Step			Second Step		
		Weight Loss (100–350 °C) (%)	Maximum Degradation Temperature (°C)	Maximum Degradation Rate (%/°C)	Weight Loss (350–500 °C) (%)	Maximum Degradation Temperature (°C)	Maximum Degradation Rate (%/°C)
PCL-EA	358.9 ± 3.1	4.2 ± 0.9	-	-	95.0 ± 1.1	411.4 ± 0.4	2.2 ± 0.1
PCL-Chl	369.2 ± 3.6	3.0 ± 0.4	-	-	96.1 ± 0.4	413.4 ± 1.6	2.3 ± 0.02
raw <i>n</i> -eicosane	178.8 ± 1.3	99.9 ± 0.1	269.1 ± 0.01	2.6 ± 0.1	-	-	-
<i>n</i> -eicosane/PCL (30/70)-EA	196.4 ± 9.7	26.9 ± 3.0	236.1 ± 6.1	0.4 ± 0.1	72.4 ± 5.4	414.9 ± 0.6	1.7 ± 0.1
<i>n</i> -eicosane/PCL (50/50)-EA	194.0 ± 13.6	44.1 ± 4.9	254.8 ± 3.5	0.9 ± 0.04	55.3 ± 5.4	413.6 ± 2.3	1.3 ± 0.2
<i>n</i> -eicosane/PCL (70/30)-EA	188.8 ± 14.0	58.1 ± 6.8	248.2 ± 12.9	1.2 ± 0.2	41.6 ± 6.8	407.0 ± 5.1	1.0 ± 0.3
<i>n</i> -eicosane/PCL (30/70)-Chl	194.3 ± 11.8	30.6 ± 4.4	237.3 ± 4.9	0.4 ± 0.1	68.8 ± 4.3	414.4 ± 1.8	1.6 ± 0.1
<i>n</i> -eicosane/PCL (50/50)-Chl	195.9 ± 7.8	49.5 ± 4.0	268.1 ± 1.0	1.0 ± 0.1	49.8 ± 3.7	414.1 ± 2.2	1.2 ± 0.1
<i>n</i> -eicosane/PCL (70/30)-Chl	186.8 ± 10.5	69.3 ± 1.1	269.5 ± 5.6	1.3 ± 0.2	30.2 ± 0.8	413.2 ± 3.5	0.7 ± 0.04

### 3.5. Structural Stability of the Electro-Sprayed mPCM

The shape stabilization of the electro-sprayed microparticles shape was investigated by SEM analyses after the heating and cooling process from  $-30$  to  $50$  °C and  $50$  to  $-30$  °C (Figure 7). After the thermal treatment, the shape and the particle morphology present any changes. The observed particles are spherical shape, smooth, and compact surface morphology. The absence of *n*-hexadecane on the SEM micrographs indicated that the PCL shell protects the PCM core from the leakage. Therefore the entrapment of PCM in the PCL matrix with this electro-spraying process allows obtaining suitable stability.

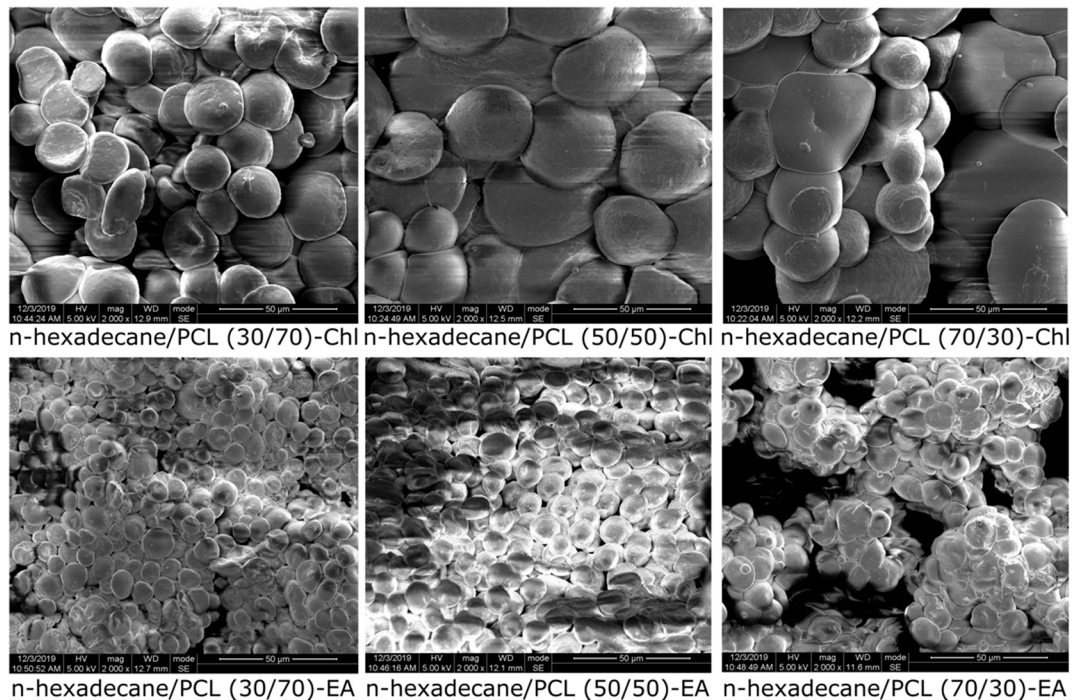


Figure 7. SEM micrographs of *n*-hexadecane/PCL microparticles after thermal treatment.

## 4. Conclusions

The microencapsulated phase change materials in the PCL matrix were successfully manufactured by an electro-spray process, using two types of solvents, ethyl acetate and chloroform, and three loading contents (30%, 50%, and 70% by weight). The process not only produces particles with a mono-dispersed particle size distribution, an average micrometric diameter, a non-porous morphology, but also a high encapsulation efficiency. The microparticles have a latent heat correlated to the theoretical load and good thermal stability. The break up process in the early stages of the electroplating method mainly determines the particle size of the mPCM. Thus, depending on the choice of *n*-alkane, *n*-hexadecane, or *n*-eicosane, two phenomena occur. On the one hand, in the case of *n*-hexadecane, which is in a liquid state at room temperature, the break-up of the charged droplets takes place throughout the flight process, from the nozzle to the collector, which allows the solvent to evaporate entirely and thus obtain a low mean diameter. On the other hand, the droplets loaded with *n*-eicosane break-up only in the earliest stages of the flight process, as the evaporation of the solvent leads to the *n*-alkane crystallization, which stops the process, resulting in higher mean diameters and broad size distribution. The surface morphology of the electro-sprayed mPCM is either smooth, and dense or rough and wrinkled when *n*-hexadecane, and *n*-eicosane are used, respectively. Besides, *n*-eicosane tends to be on the surface of particles as its loading content increases. Also, the choice of solvent affects the particle size and morphology of the particles, which may be related to changes in the physico-chemical properties of solutions, such as surface tension, vapor pressure, and viscosity. The loading content depends

mainly on the diffusion of *n*-alkane chains during the process, which is also influenced by solvent evaporation or phase separation, the interactions between PCL macromolecular chains, and *n*-alkane chains, and the evolution of the physical state of PCM. Thus, the encapsulation efficiency and charge content of an *n*-alkane liquid at room temperature are respectively higher and lower than those of a solid *n*-alkane in ethyl acetate, but in chloroform, it is the opposite.

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