

Supplementary Materials:

Crystallization, Luminescence and Cytocompatibility of Hexagonal Calcium Doped Terbium Phosphate Hydrate Nanoparticles

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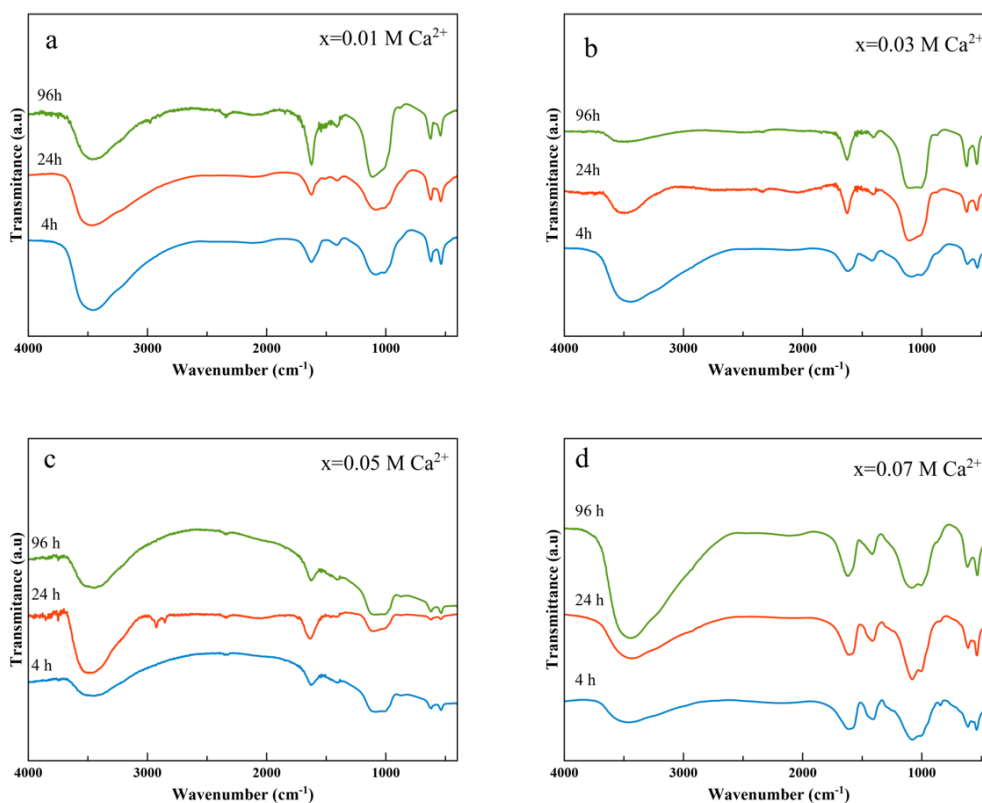


Figure S1. Evolution of FTIR spectra with time of samples prepared with different Ca²⁺ doping concentrations $x = 0.01$ to 0.07 M.

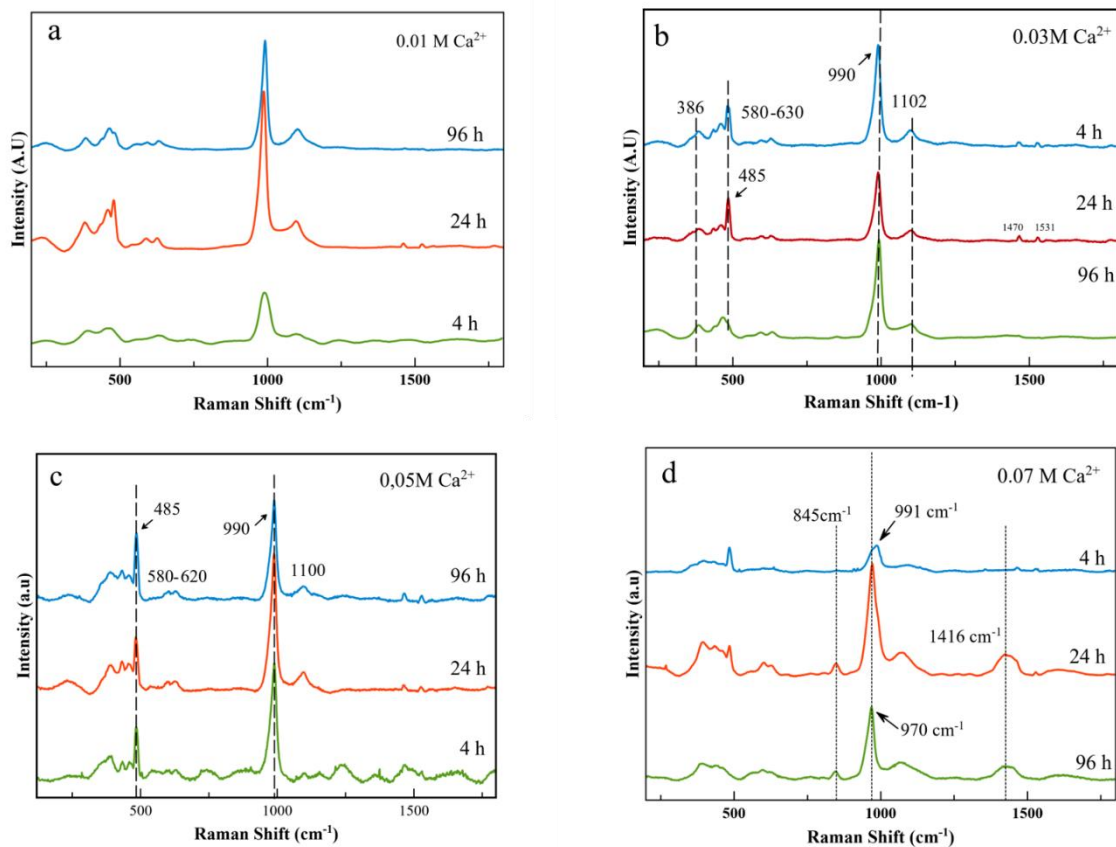


Figure S2. Evolution of Raman spectra with time of samples prepared with different Ca^{2+} doping concentrations $x = 0.01$ to 0.07 M.

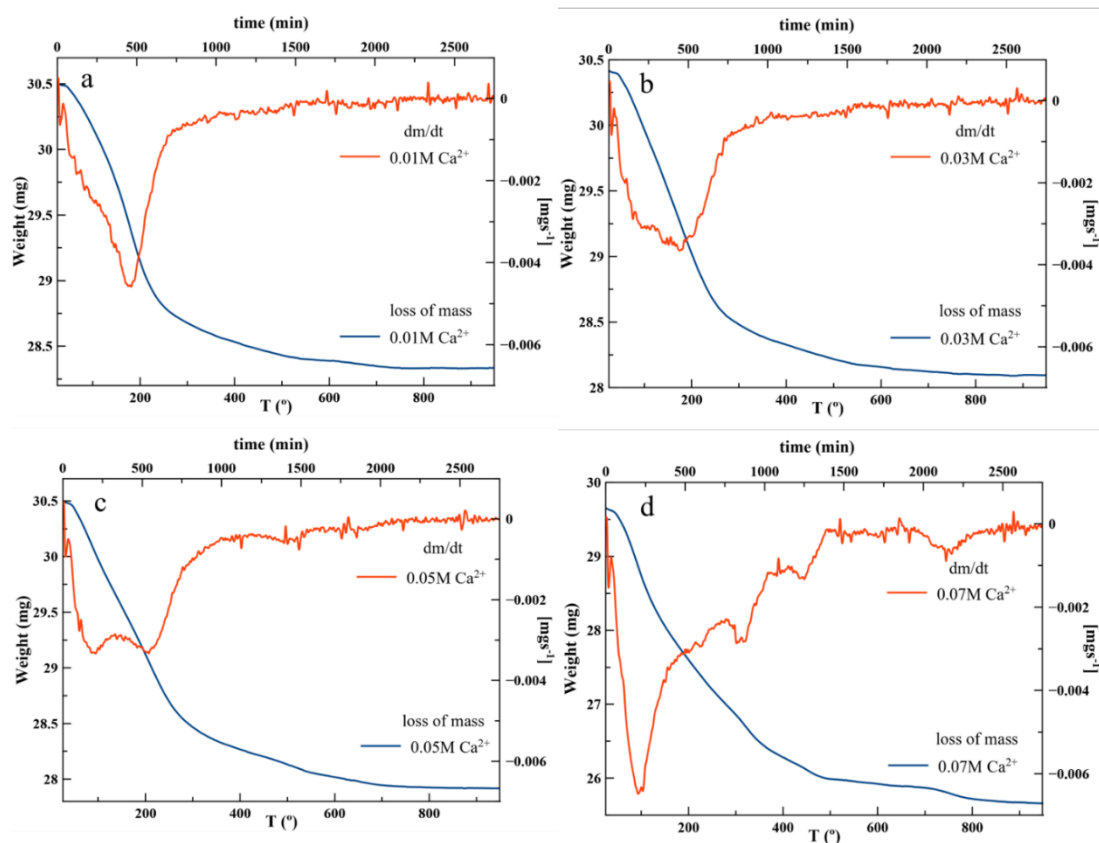


Figure S3. TGA analyses of precipitates obtained at 96 hours from solutions with Ca^{2+} doping concentrations ranging from a) $x = 0.01$ to d) $x = 0.07$ M.

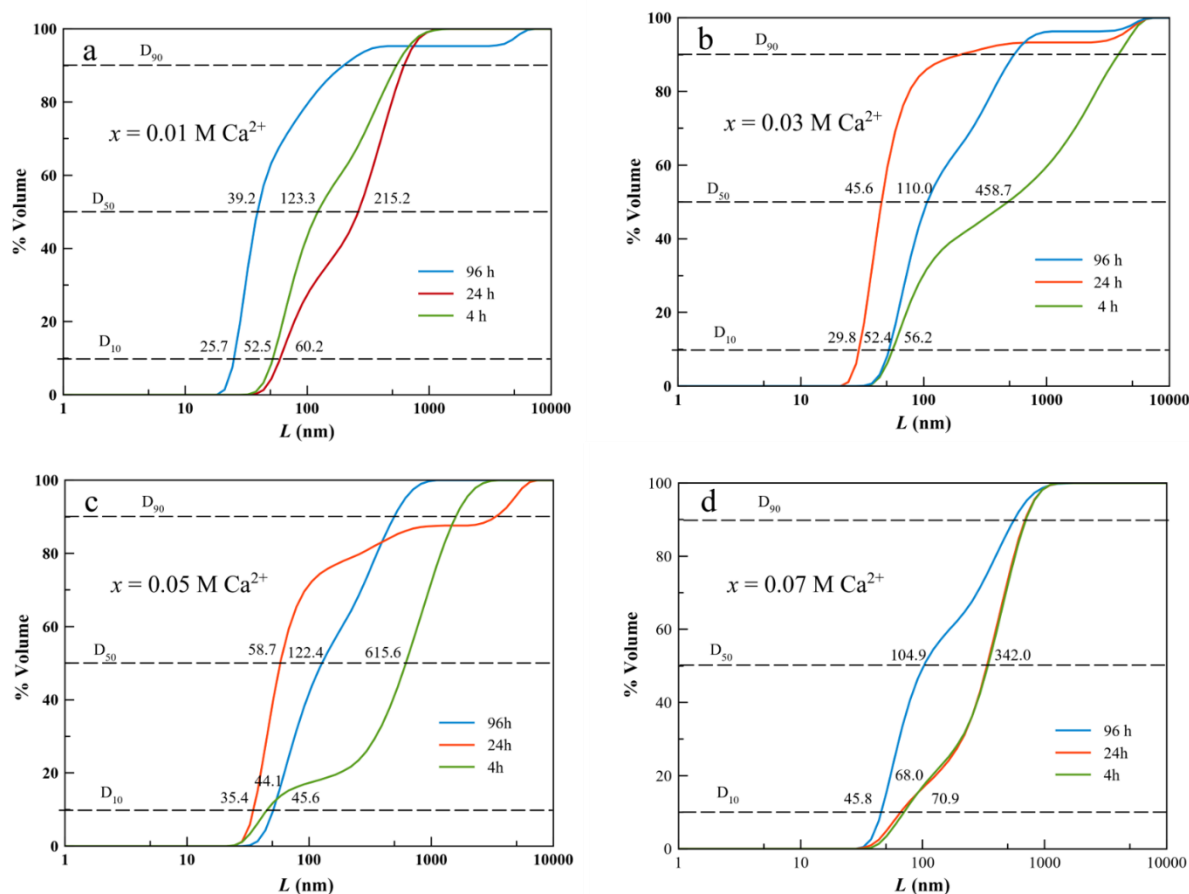


Figure S4. Cumulative volume oversize distribution of the cit-Ca²⁺:TbPO₄.nH₂O nanocrystals prepared with Ca²⁺ doping concentrations $x = 0.01, 0.03, 0.05$ and 0.07 M at 4, 24 and 96 hours.

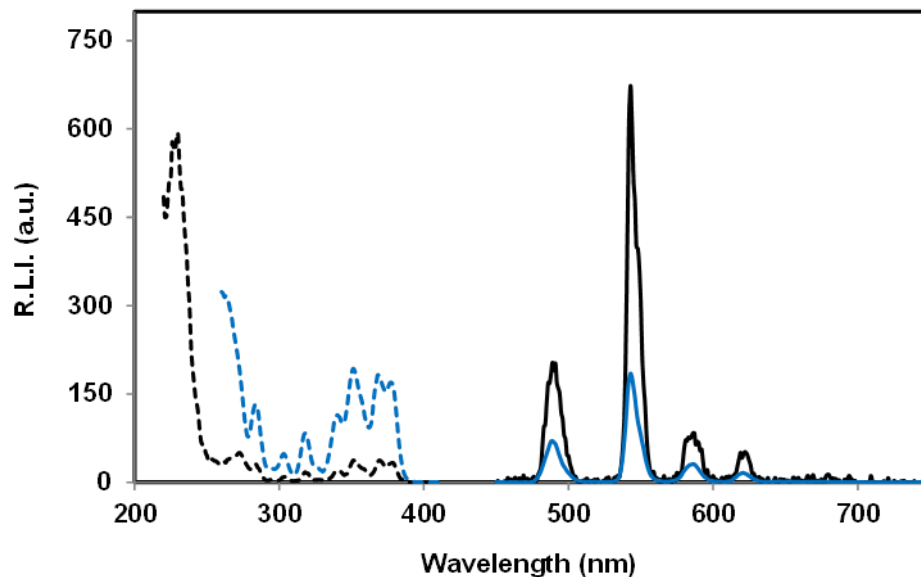


Figure S5. Excitation (dashed lines) and emission (solid lines) uncorrected spectra of solid cit-Ca²⁺:TbPO₄.nH₂O samples prepared with $x = 0.01$ M Ca²⁺ at maturation times of 96 h using $t_d = 120 \mu\text{s}$, $t_d = 5\text{ms}$ and a) $\lambda_{\text{exc/em}} = 230/545$ nm, slit width $\text{exc/em} = 2.5/2.5$ nm, detector voltage 545 V for the black line; b) $\lambda_{\text{exc/em}} = 375/545$ nm, slit width $\text{exc/em} = 5/5$ nm, detector voltage 470 V for the blue line.

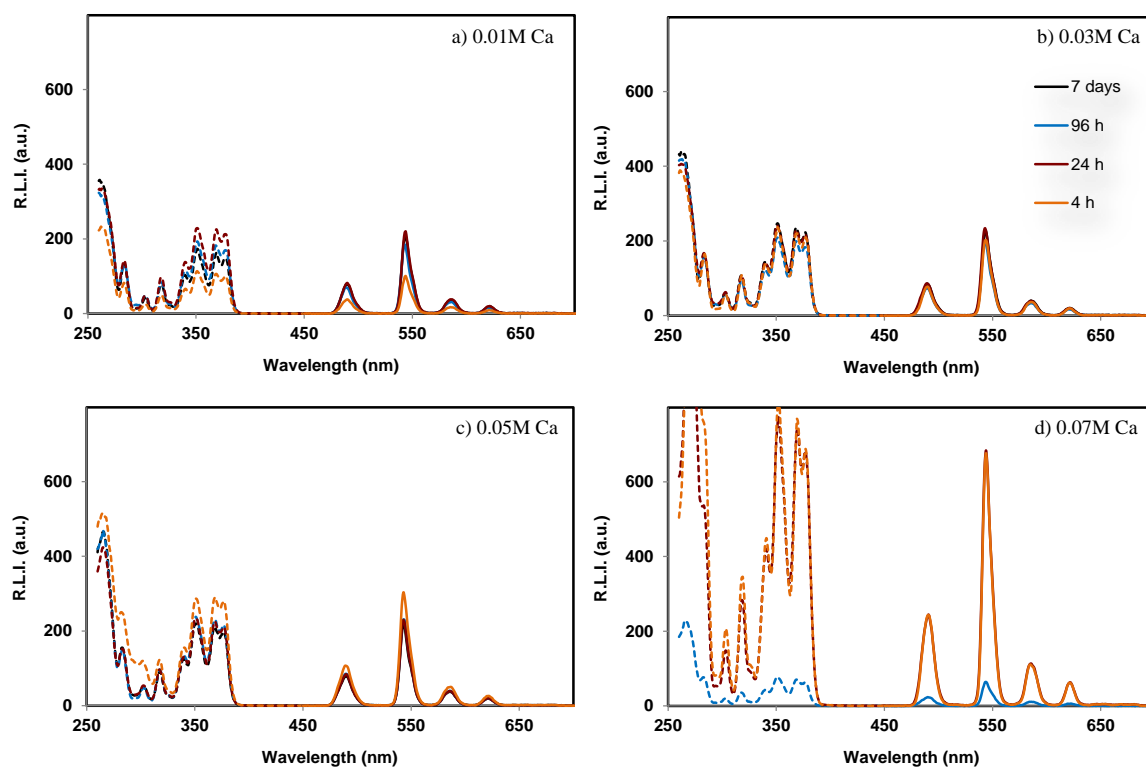


Figure S6. Excitatin (dashed lines) and emission (solid lines) uncorrected spectra of solid cit- Ca^{2+} : $\text{TbPO}_4 \cdot n\text{H}_2\text{O}$ samples prepared with different Ca^{2+} doped concentration at maturation times of 4 h, 24 h, 96 h and 7 days.

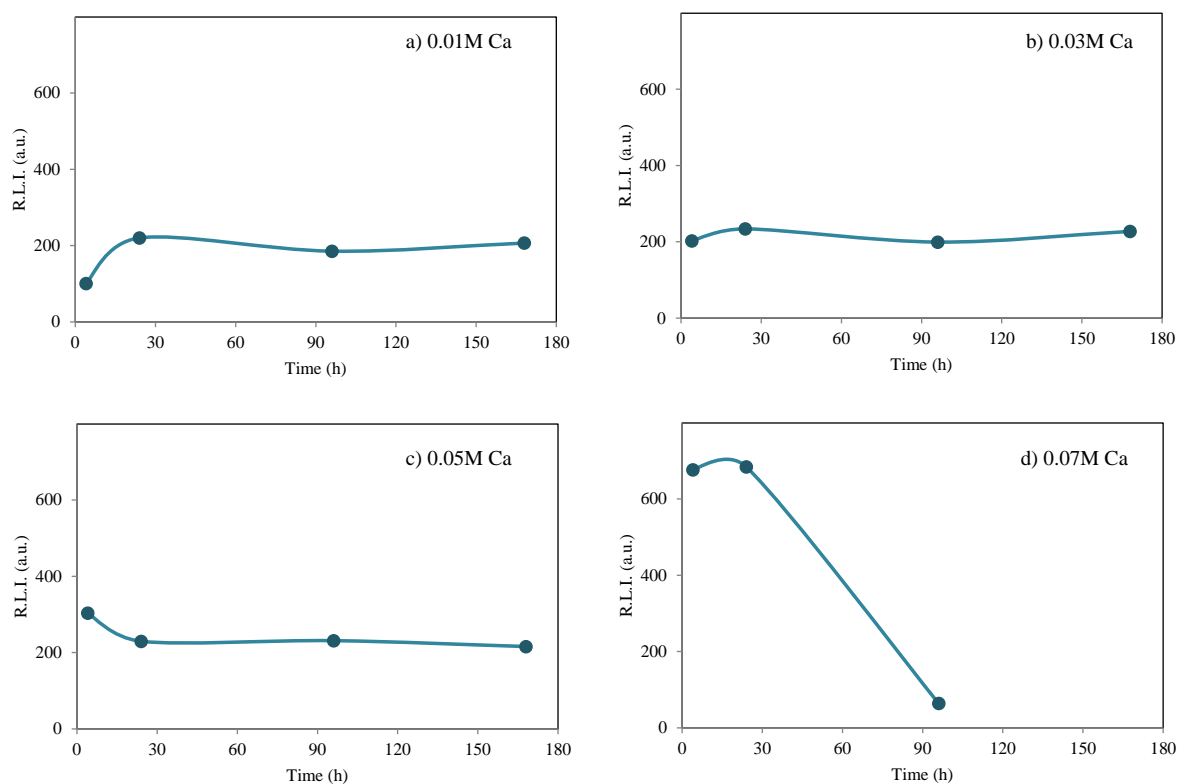


Figure S7. Variation of the R.L.I. of the solid cit- Ca^{2+} : $\text{TbPO}_4 \cdot n\text{H}_2\text{O}$ samples at the maximum excitation and emission wavelengths at several Ca^{2+} concentrations when the maturation time is changed.

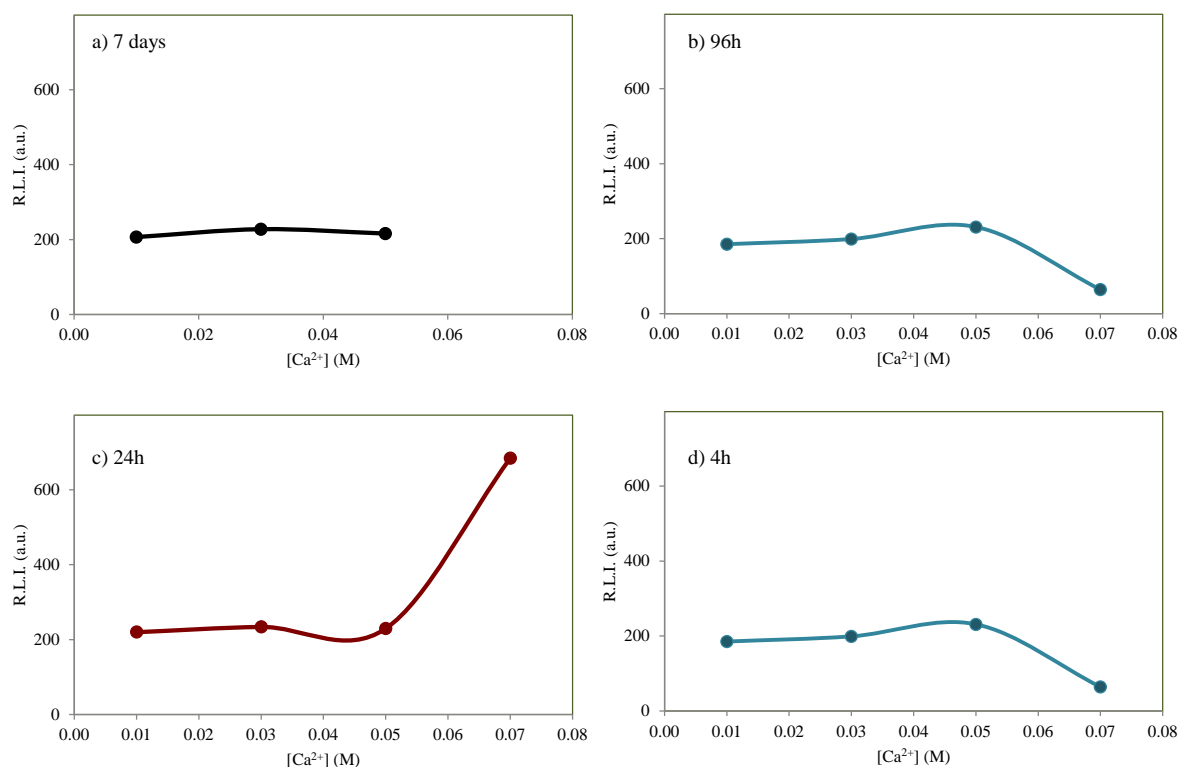


Figure S8. Variation of the R.L.I. of the solid cit-Ca²⁺:TbPO₄·nH₂O samples at the maximum excitation and emission wavelengths at several maturation time when the Ca²⁺ concentration is changed.

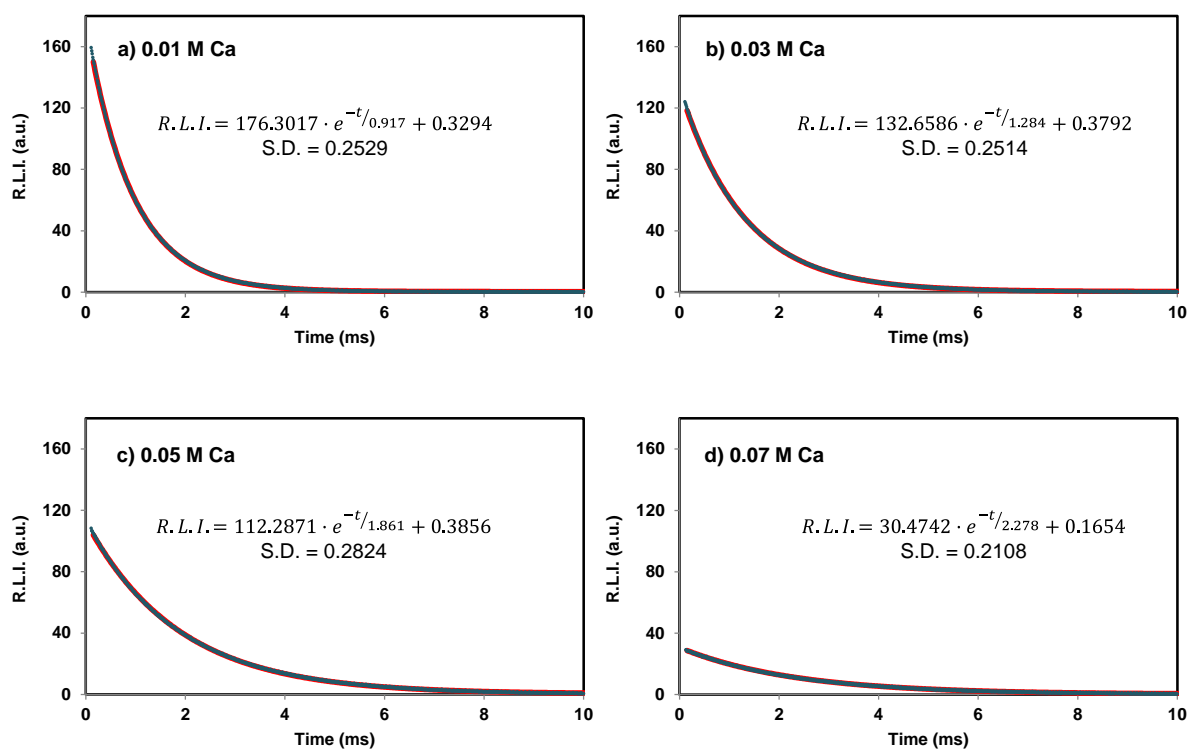


Figure S9. Luminescence decay curve of different solid cit-Ca²⁺:TbPO₄·nH₂O samples at maturation times of 96h, $t_d = 100 \mu\text{s}$, $t_g = 0.01 \text{ ms}$, $\lambda_{\text{exc/em}} = 375/545 \text{ nm}$, slit-width_{exc/em} = 10/10 nm, and detector voltage = 600 V. Circles correspond to experimental data (100 cycles) and lines to the fitting equation.

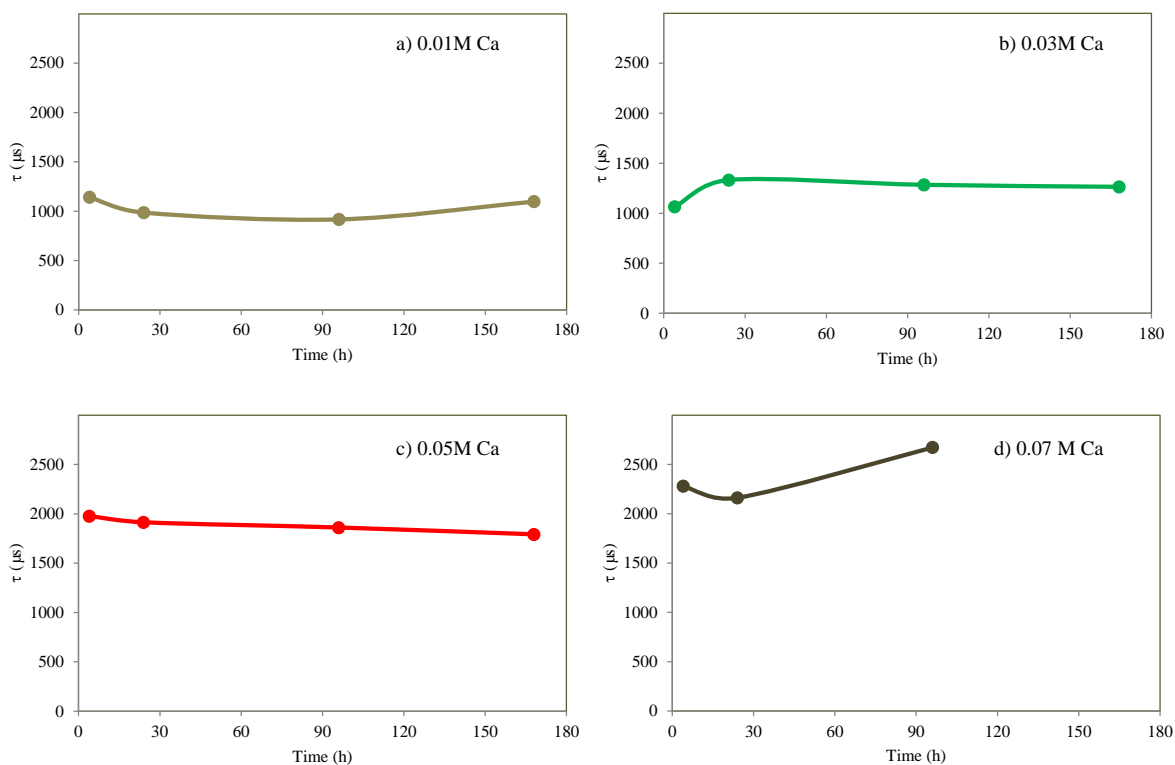


Figure S10. Variation of the luminescence lifetime of the solid $\text{cit-Ca}^{2+}:\text{TbPO}_4 \cdot n\text{H}_2\text{O}$ nanoparticles prepared at several Ca^{2+} concentrations when the maturation time is changed.

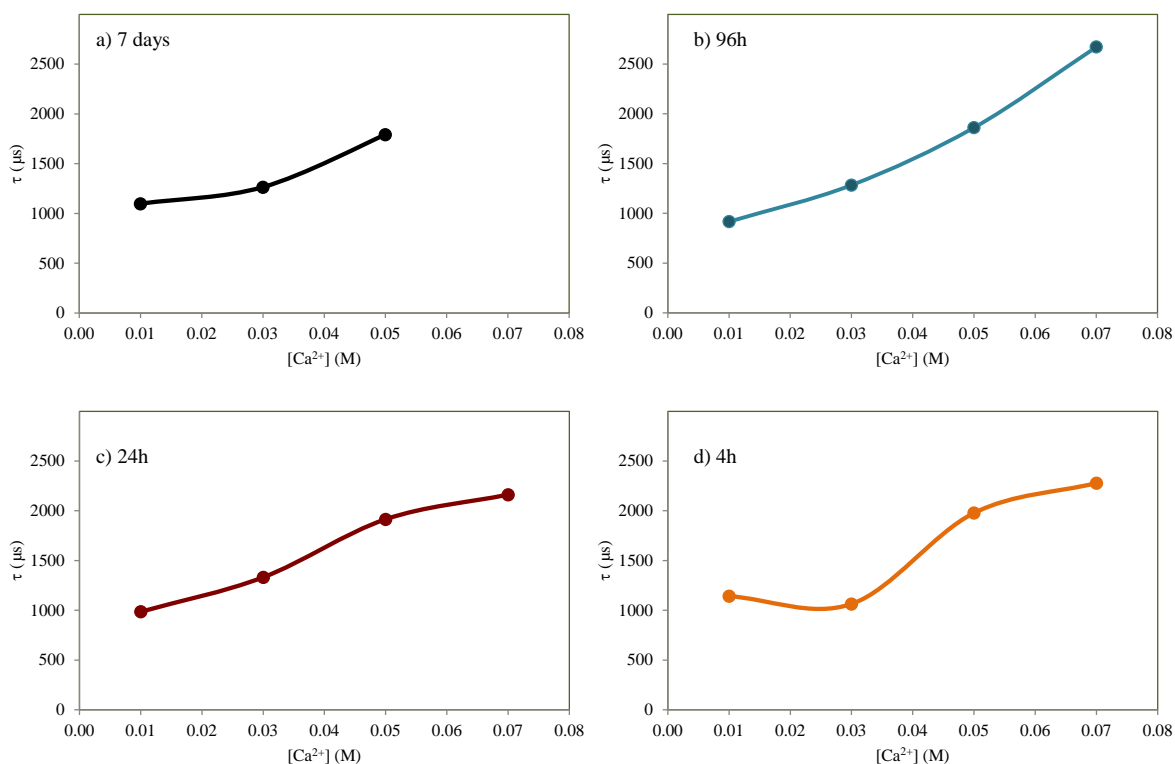


Figure S11. Variation of the luminescence lifetime of the solid $\text{cit-Ca}^{2+}:\text{TbPO}_4 \cdot n\text{H}_2\text{O}$ samples at several maturation time when the Ca^{2+} concentration is changed.

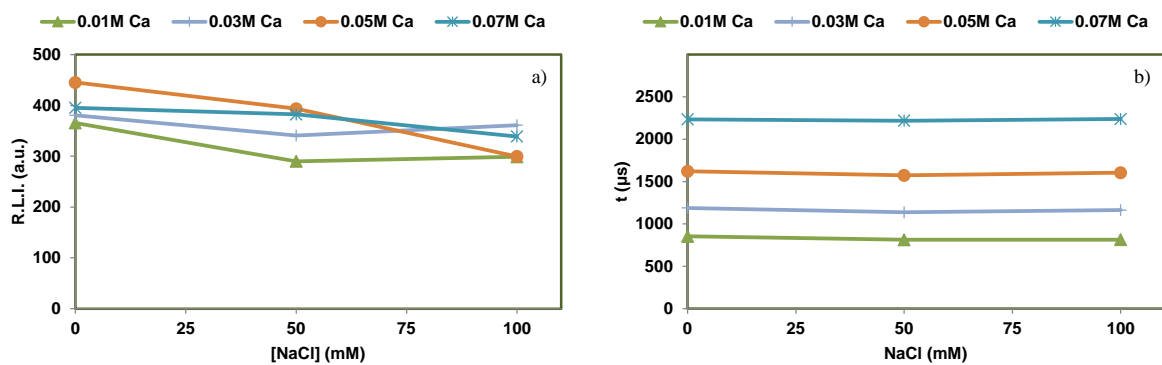


Figure S12. Effect of the ionic strength over the a) R.L.I. and b) luminescence lifetime of the cit-Ca²⁺:TbPO₄·nH₂O samples at 96 h maturation time dispersed in aqueous media at several Ca²⁺ concentration.

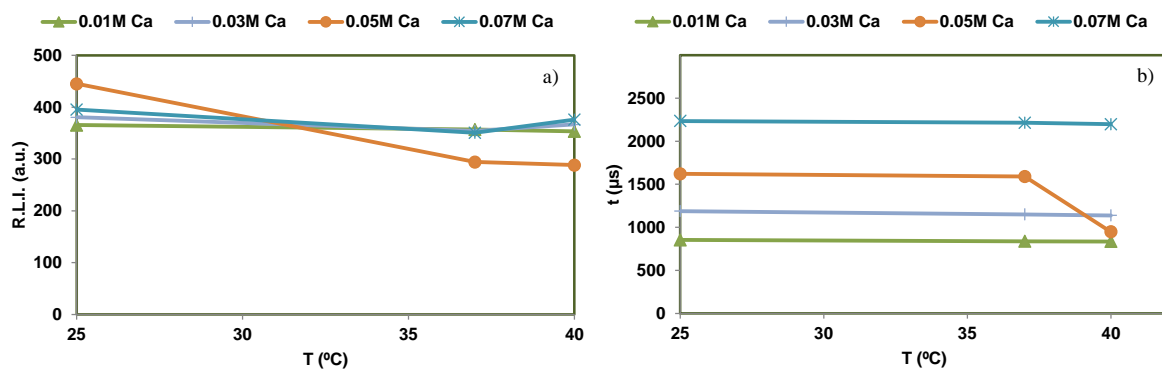


Figure F13. Effect of the temperature over the a) R.L.I. and b) luminescence lifetime of the cit-Ca²⁺:TbPO₄·nH₂O samples at 96 h maturation time dispersed in aqueous media at several Ca²⁺ concentration is changed.