

Figure S1. The double-reciprocal plot of copper hexacyanoferrate nanoparticles. Error bars represent standard deviations ($n = 3$).

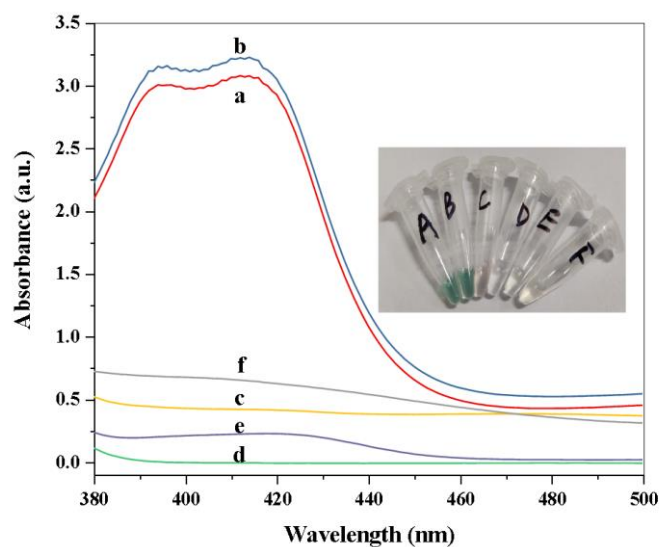


Figure S2. UV-vis absorption spectra of (a) AAO+AA+K₃[Fe(CN)₆]+CuCl₂·2H₂O +ABTS mixture solution, (b) AAO+K₃[Fe(CN)₆]+CuCl₂·2H₂O+ABTS mixture solution, (c) AA+K₃[Fe(CN)₆]+CuCl₂·2H₂O+ABTS mixture solution, (d) AAO+AA+CuCl₂·2H₂O+ABTS mixture solution, (e) AAO+AA+K₃[Fe(CN)₆] +ABTS mixture solution, (f) AAO+AA+K₃[Fe(CN)₆]+CuCl₂·2H₂O mixture solution (Insets: the corresponding images).

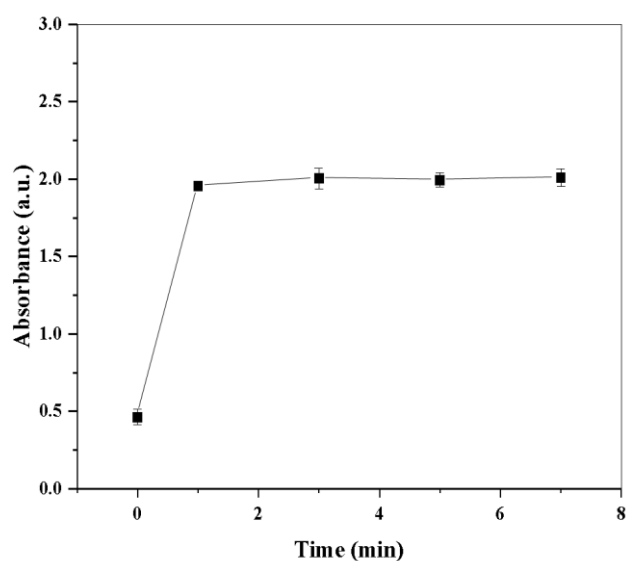


Figure S3. The effect of reaction time between copper hexacyanoferrate nanoparticles and ABTS on the absorbance value at 412 nm. Error bars represent standard deviations ($n = 3$).

Table S1. Comparison of apparent Michaelis–Menten Constant (K_m) of copper hexacyanoferrate nanoparticles for the oxidation of ABTS with other reported oxidase-like nanozymes.

Oxidase-like nanozymes	K_m (μM)	Ref.
Copper hexacyanoferrate NPs ^a	22.30	This study
MoO ₃ NPs	1676.9	[29]
UiO-66-NH ₂ @Ce	720	[30]
AuNCs ^b -GMP/PtNCs-GMP/Au–PtNCs-GMP	95.3/1440/132.1	[31]
Co ₃ O ₄ NPs	37	[32]
Ag NPs	7.2	[33]

^aNP: nanoparticles; ^bNCs: nanocluster.