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Data-Based Modeling, Multi-Objective Optimization, and Multi-Criterion Decision-Making to Maximize the Electro-Oxidation of Metoprolol over Boron-Doped Diamond Electrodes in a Flow-By Reactor

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Abstract: Metoprolol is a cardioselective beta-blocker drug often used to treat hypertension, but it is considered as a hazardous organic persistent contaminant in wastewater. In this study, a 2.5 L solution of metoprolol (50 mg/L) underwent electro-oxidation in a flow-by reactor using boron-doped diamond electrodes in the batch recirculation mode. The study used multi-objective optimization and multi-criterion decision-making to determine the optimal operating parameters. The response surface methodology and a central composite rotatable design were used with three factors (pH₀: 5–8, I: 2.5–4 A, and Q: 0.8–1.7 L/min) to model the chemical oxygen demand's (COD's) removal efficiency and the total organic carbon's (TOC's) removal efficiency. The experimental responses were modeled by reduced third- and second-order polynomials with determination coefficients (R²) of 0.9816 and 0.9430. The optimal operating parameters were found to be pH₀ 5, an I value of 3.84 A, and a Q value of 0.8 L/min with an electrolysis time of 7.5 h, resulting in a maximum COD removal efficiency of 60.8% and a TOC removal efficiency of 90.1%. The specific energy consumption was calculated as 9.61 kWh/mg of TOC, with a total operating cost of 0.77 USD/L. In conclusion, this study showed that the electrochemical process is efficient and reliable for treating wastewater containing metoprolol.

Keywords: BDD electrodes; data-based modeling; electro-oxidation; flow reactor; metoprolol; optimization



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1. Introduction

Because of the scarcity of water on Earth, caused by global warming, and increased contamination of water sources because of global industrialization in recent years, the United Nations' member states have declared the decontamination and sanitation of wastewater as priority tasks through Agenda 2030, which is the basis for objective 6 (clean water and sanitation) of the sustainable development goals [1]. Regarding contamination by anthropogenic activities, the emerging contaminant (EC) molecules in oceans, rivers, lakes, and soils from sources such as farming, cattle raising, and highways can provoke human health and environmental harm [2]. ECs include dyes, personal care products, endocrine disruptors, and drugs (e.g., antibiotics and beta-blockers among others), which are usually discarded in water bodies [3–5]. These pollutants are persistent, transported over long-range distances, bioaccumulate, and are highly toxic because they produce allergic reactions, immune system damage, endocrine disruption, and liver and kidney damage;

increase cancer risk; cause neuronal damage; possess mutagenicity; and impact children's development [6]. Metoprolol (MET) is the beta-blocker most used worldwide as a treatment for hypertension, tachycardia, cardiac insufficiency, heart failure, and myocardial infarction [7]. Approximately 1.28 billion people are suffering from arterial hypertension [8] worldwide, and in Mexico, around 30 million Mexicans are affected by arterial hypertension [9]. Between 3 and 10% of the medication dosage is excreted without undergoing any transformation and is released through urine [10]. Also, MET has been detected in treated wastewater from conventional treatment plants in Asia, America, and Europe, at concentrations of 0.268, 0.212, and 5.762 $\mu\text{g/L}$, respectively [11,12]. Therefore, it is crucial to remove this medication because of its health risks for both humans and aquatic organisms.

Numerous technologies have been developed to eliminate ECs (e.g., metoprolol) from wastewater, including physical, chemical, and biological. However, physical treatments are unable to completely degrade ECs; chemical methods produce secondary pollutants and are costly, and biological methods struggle to treat high-concentration pollutants, require long treatment times, and are sensitive to the environment [13]. In this sense, electrolysis, an advanced oxidation process, has emerged as an efficient and green technology for removing ECs, such as MET. It offers high efficiency, environmental compatibility, flexibility for automation, easy maintenance, and the possibility of the mineralization of recalcitrant pollutants and does not need additional chemicals to conduct the process [6,14–16].

In the process of electrolysis, electrodes play a crucial role. Various types of electrodes (e.g., Pt, PbO_2 , SnO_2 , $\text{Ta}_2\text{O}_5\text{-IrO}_2$, and $\text{Nb}_2\text{O}_5\text{-IrO}_2$) have been developed, including carbon-based materials (e.g., graphite, graphene, carbon nanotubes, and boron-doped diamond (BDD) electrodes). BDD electrodes are especially preferred for electrolysis because of their chemical, electrochemical, and mechanical stabilities [17]. They are also known for their ability to produce a large amount of hydroxyl radicals ($\cdot\text{OH}$, with an oxidation potential of 2.8 V) [18–20] and other oxidized species, depending on the electrolyte medium used. For example, in Na_2SO_4 media, persulfate, peroxydisulfate, and sulfate radicals are formed [21–26].

The removal of MET has been studied primarily using ozonation [27,28], UV/ H_2O_2 [29], photo-oxidation [30], Fenton [31], photo-Fenton [7], heterogeneous photocatalyst [32], and photo-Fenton-like [7] methods, with a batch reactor being the most used configuration. Only a few studies on electro-oxidation for MET removal were found [33–36], with a batch lab reactor also being the commonly used configuration in these studies. The flow-by reactor is the best reactor configuration for electro-oxidizing recalcitrant pollutants because of its high mass transport coefficients and homogeneous current distribution on the electrodes [37]. It is important to note that no studies have reported optimizing the removal of MET (see Table 1).

Regarding the optimization of the electro-oxidation process for ECs carried out in a flow-by reactor setup, the most common method for performing the optimization is the response-surface-methodology-driven DoE [38–40]. Also, the central composite rotatable design (CCRD) is widely utilized [41] because it offers sufficient data for testing the lack of the fit.

Table 1 summarizes the literature results related to the removal of MET using different treatment technologies. It can be observed that the related literature reveals the high removal efficiency of MET; however, it is worth pointing out that the treated volume used in many cases is very low, on the order of fractions of liters.

The focus of this study was to maximize the degradation of MET in terms of the COD and TOC removal efficiencies. This was achieved through electro-oxidation in a flow-by reactor using two Nb/BDD electrodes operating in the batch recirculation mode. The multi-objective optimization was carried out using RSM driven by CCRD, with three operating parameters, namely, the current intensity (I), initial pH (pH_0), and volumetric flow rate (Q). In addition to the optimization process, this study also involved model validation, determination of the total operating cost, kinetic rate models for the degradation of MET, COD removal, TOC removal, and phytotoxicity testing.

Table 1. Removal of MET using different non-optimized technologies.

Process Type	Reaction Conditions	Results		Ref.
		MET Removal (%)	TOC Removal (%)	
UV	[MET] ₀ = 50 mg/L; H ₂ O ₂ = Stoich.; pH ₀ 5.6; λ = 254 nm; V _{treated} = 0.1 L; Catalyst: Cu-TiO ₂ ; t = 100 min	92	25	[7]
Ozonation (O ₃)	[MET] ₀ = 200 µg/L; [O ₃] = 380 mg/L h; V _{treated} = 0.8 L; pH ₀ 11; t = 120 min	63.4	---	[27]
Adsorption and Photocatalysis	[MET] ₀ = 50 mg/L; pH ₀ 9; λ = 250–400 nm; V _{treated} = 0.025 L; 100 pm; Catalyst: TiO ₂ ; t = 300 min	100	63	[42]
Thermal plasma and UV/H ₂ O ₂	[MET] ₀ = 5 µg/L; Plasma 150 W; 1 MHz; t = 120 min; V _{treated} = 0.1 L; T = 35 °C	64.7	---	[29]
Fenton	[MET] ₀ = 148.5 mg/L; V _{treated} = 0.1 L; T = 35 °C; pH ₀ 3; [Fe ³⁺] ₀ = 0.5 mM; t = 360 min	---	80	[43]

2. Materials and Methods

2.1. Reagents and Aqueous MET Solution

Sulfuric acid (H₂SO₄), sodium hydroxide (NaOH), and metoprolol tartrate salt (MET), with purities of 95, 97, and 99.9%, respectively, were purchased from the Sigma-Aldrich Company (Ciudad de México, Mexico). Additionally, sodium sulfate (Na₂SO₄), with a purity of 99%, was supplied by the Karal Group (Ciudad de México, Mexico, a Mexican Company).

Before each experiment, a synthetic aqueous solution containing 50 mg/L of MET in 2.5 L with 0.1 M Na₂SO₄ was freshly prepared. This MET solution was added to the reservoir tank of the electrochemical plant (as shown in Figure 1) and homogenized by recirculating for 15 min. Distillate water was used for preparing all the aqueous solutions containing 50 mg/L of MET. Table 2 shows the physicochemical properties of the MET compound.

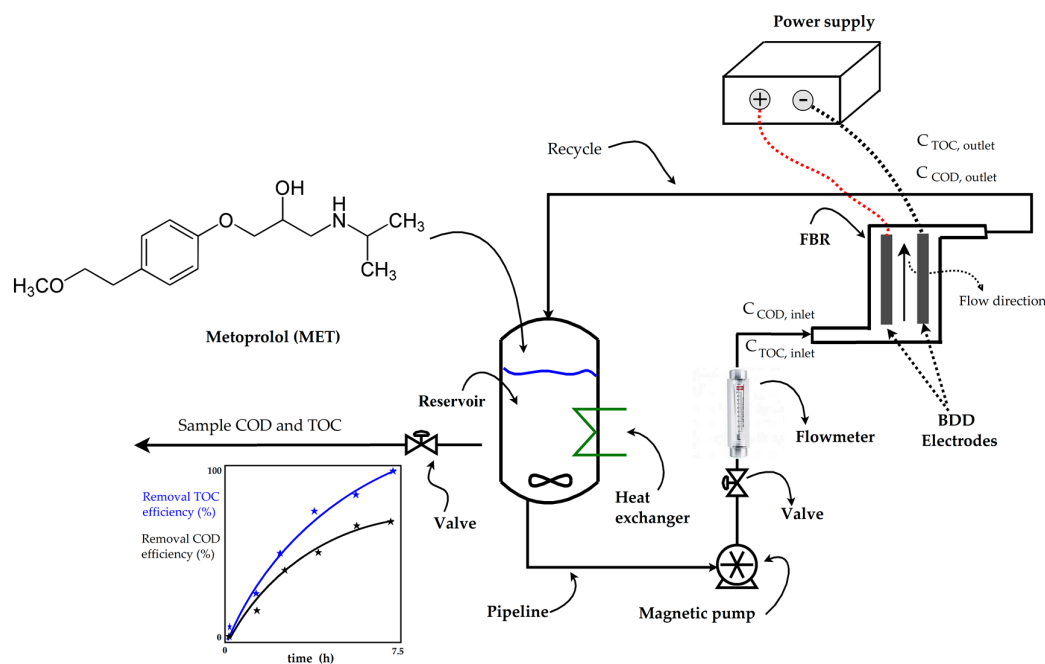
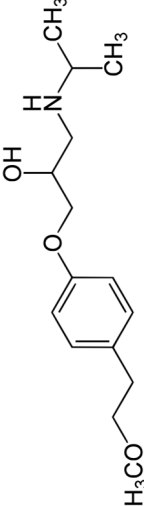
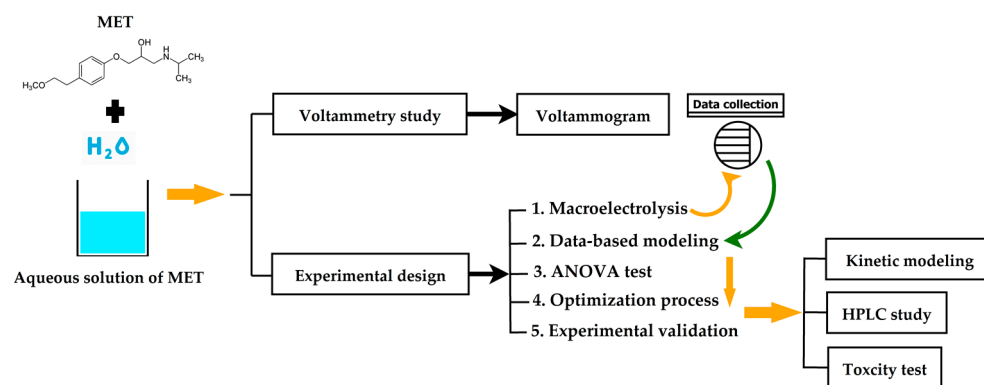
**Figure 1.** Flowchart of the electrochemical plant.

Table 2. Aqueous solution characteristics and physicochemical properties.

Chemical Structure of MET	Aqueous Solution Characteristics						
	C_{MET} (mg/L)	pH	Conductivity ($\mu\text{S}/\text{cm}$)	TOC (mg/L)		COD (mg/L)	
				Theo.	Exp.	Theo.	Exp.
	50	7	14,180	29.81	31.53 ^a	98.13	96.00 ^a
	Physicochemical properties						
	Property			Value			
	Molar weight (g/mol)			684.81			
	Melting point ($^{\circ}\text{C}$)			120 [44]			
	Boiling point ($^{\circ}\text{C}$)			695.37 [44]			
	Water solubility (mg/mL) at 25 $^{\circ}\text{C}$			>1000 [44]			
	pKa			9.7 [45]			
	λ (nm)			223 [42]			
	Chemical formula: ($\text{C}_{15}\text{H}_{25}\text{NO}_3$) ₂ • $\text{C}_4\text{H}_6\text{O}_6$			CAS No.: 56392-17-7			

^a The average value of all the trials; Theo.: theoretical value; Exp.: experimental value.

Scheme 1 depicts the flowchart diagram of the methodology employed in this research. Also, this scheme shows the stages followed to conduct the data-based modeling, multi-objective optimization, and multi-criterion decision-making to maximize the electro-oxidation of the metoprolol over boron-doped diamond electrodes in a flow-by reactor. Additionally, each step is thoroughly described in detail in the following subsections.

**Scheme 1.** Flowchart diagram of the methodology.

2.2. Voltametric Experiment

A micro-electrolysis test was performed using an Autolab potentiostat/galvanostat[®] 30 controlled using the GPE software V.4.3. The test used a three-electrode cell (100 mL) with Ag/AgCl, a graphite bar, and a BDD electrode (with an electroactive area of 1 cm²) serving as reference, auxiliary, and working electrodes, respectively. Cyclic voltammetry was conducted in a voltage range from 0 to 2 V, at a scan rate of 100 mV/s, at atmospheric pressure, and at a temperature of 25 $^{\circ}\text{C}$. Also, the aqueous solution containing 50 mg/L of MET was stirred for 1 min to attain a homogeneous mixture.

2.3. Electrochemical Plant

The electro-oxidation of metoprolol took place at an electrochemical plant (see Figure 1) consisting of an electrochemical flow-by reactor (FBR) with a width of 4.5 cm, a channel thickness of 0.5 cm, an electroactive length of 20 cm; two BDD electrodes (20 cm) that function as an anode and a cathode; a 3.5 L polycarbonate reservoir; a heat exchanger to maintain a constant room temperature; a magnetic pump operating at 300 rpm; a glass flowmeter for the liquid flow rate control; a power supply for electrode energization; and PVC pipelines, connectors, and valves. Additional information about the electrochemical FBR can be found in the authors' previous work [46].

2.4. Experimental Design

A 2.5 L aqueous solution containing 50 mg/L of MET was added to the electrochemical plant reservoir (refer to Figure 1). The studied operating parameters included the initial pH (pH_0), current intensity (I), and volumetric flow rate (Q), as outlined in Table 3.

Table 3. Levels and studied operating parameters.

K Design Parameters	Design Parameter Levels				
	$-\alpha$	-1	0	$+1$	$+\alpha$
x_A : Initial pH	3.98	5.00	6.50	8.00	9.02
x_B : Current intensity (A)	1.99	2.50	3.25	4.00	4.51
x_C : Volumetric flow rate (L/min)	0.49	0.80	1.25	1.70	2.00

In Table 3, the lower, central, highest, and axial ($\pm\alpha$) levels of each operating parameter are encoded as $-\alpha$, -1 , 0 , $+1$, and $+\alpha$, following the Equation (1):

$$\begin{aligned} \ell &= \frac{x_{\ell,i} - x_{\ell,0}}{\Delta x_{\ell}} \\ \ell &= A, B, \text{ and } C \\ i &: -\alpha, -1, 0, +1, +\alpha \end{aligned} \quad (1)$$

A central composite rotatable design (CCRD) with three operating parameters was adopted as an experimental design comprising 2^K factorial trials, $2K$ axial trials, and K central trials, giving 17 trials (see Section 3). The experimental design, data-based modeling, ANOVA test, and optimization were conducted using Design Expert[®] V.10 software (license number 9839-2917-5744-6919 from State-Ease, Inc., Minneapolis, MN, USA). Additionally, the Q range was chosen according to the authors' previous work, where ibuprofen was electrochemically oxidized in the same FBR [40].

2.5. Chemical Analysis

Samples of the MET solution for each trial were taken directly from the reservoir of the electrochemical plant to determine the chemical oxygen demand (COD) and total organic carbon (TOC) of MET solutions before and after the electro-oxidation treatment. The COD determination was gauged using the standard method (5220B) given in [47]. The TOC determination was measured using a 6001 TOC analyzer (ShimadFzu, Kyoto, Japan). The pH was measured with an HI2210 Hanna potentiometer.

2.6. Calculations

The COD removal efficiency of the MET (η_{COD}) was computed based on the change in the COD (mg/L) quantity between the initial sample and the treated sample, according to Equation (2):

$$\eta_{COD} = \frac{[COD]_0 - [COD]_t}{[COD]_0} \times 100 \quad (2)$$

where $[COD]_0$ and $[COD]_t$ are the COD of the MET solution before and after the electro-oxidation treatment.

The TOC removal efficiency of the MET (η_{TOC}) was computed based on the change in the TOC (mg/L) quantity between the initial sample and the treated sample, according to Equation (3):

$$\eta_{TOC} = \frac{[TOC]_0 - [TOC]_t}{[TOC]_0} \times 100 \quad (3)$$

where $[TOC]_0$ and $[TOC]_t$ are the TOC of the MET solution before and after the electro-oxidation treatment.

The total electricity energy consumed (E (kWh)), including the energy consumed by the electrodes ($E_{electrodes}$), energy consumed by the recirculation pump ($E_{pump, flow}$), and energy consumed by the heat exchanger pump ($E_{pump, heat exchanger}$), was calculated using the set of Equations (4)–(7):

$$E_{electrode} = \frac{\Psi \times I \times t}{1,000} \quad (4)$$

$$E_{pump, flow} = P_n \times t \quad (5)$$

$$E_{pump, heat exchanger} = P_m \times t \quad (6)$$

$$E = E_{electrodes} + E_{pump, flow} + E_{pump, heat exchanger} \quad (7)$$

where E is the total electricity consumed (kWh), Ψ is the average cell potential (23.75 V), I is the applied current intensity (A), t is the total electro-oxidation time (7.5 h), P_n is the supplier catalog's nominal power (0.198 kW) for the recirculating pump, and P_m is the supplied catalog's nominal power (0.123 kW) for the heat exchanger pump.

The total operating cost was determined as a function of the electricity and electrolyte prices according to Equation (8):

$$cost = p_{electricity} E + p_{electrolyte} m_{electrolyte} \quad (8)$$

where $p_{electricity}$ is the electricity price (0.058 USD/kWh, based on 1 USD/MXN = 18.36) for commercial use in Mexico, $p_{electrolyte}$ is the price of the electrolyte (0.81 USD/kg), and $m_{electrolyte}$ is the consumed electrolyte's mass (0.36 kg).

2.7. Data-Based Modeling and Optimization Procedure

Experimental data responses (η_{COD} and η_{TOC}) were modeled using surrogate models, according to Equation (9), in terms of encoded variables:

$$\begin{aligned} \eta_j = & \beta_0 + \beta_1 A + \beta_2 B + \beta_3 C + \beta_4 AB + \beta_5 AC + \beta_6 BC + \beta_7 A^2 + \beta_8 B^2 + \beta_9 C^2 + \dots \\ & \beta_{10} ABC + \beta_{11} A^2 B + \beta_{12} A^2 C + \beta_{13} AB^2 + \beta_{14} AC^2 + \beta_{15} B^2 C + \beta_{16} BC^2 + \dots \\ & \beta_{17} A^3 + \beta_{18} B^3 + \beta_{19} C^3 \\ j = & \text{COD and TOC} \end{aligned} \quad (9)$$

where η_j is the surrogate model for the responses of the COD and TOC removal efficiencies; the β -values are the coefficients of the third-order polynomial function (see Equation (9)); and A , B , and C are the encoded operating parameters according to Equation (1). The β -values were calculated using the stepwise regression method in the Design Expert® V.10 software package.

It is essential to prioritize accuracy when building models. In this study, the accuracy of the adjusted surrogate models was evaluated using the determination coefficient (R^2), mean absolute error (MAE), and root-mean-square error (RMSE). The performance indexes, R^2 ,

MAE, and RMSE, are represented by Equations (10) [48], (11) [49], and (12) [50], respectively. Also, a high-quality model has a high R^2 value and low values of MAE and RMSE.

$$R^2 = \frac{\sum_q^n (\eta_{Exp, q} - \bar{\eta}_{Exp, q}) \sum_q^n (\eta_{Model, q} - \bar{\eta}_{Model, q})}{\sqrt{\sum_q^n (\eta_{Exp, q} - \bar{\eta}_{Exp, q})^2 \sum_q^n (\eta_{Model, q} - \bar{\eta}_{Model, q})^2}} \quad (10)$$

$$MAE = \frac{1}{n} \sum_q^n |\eta_{Exp, q} - \eta_{Model, q}| \quad (11)$$

$$RMSE = \sqrt{\frac{1}{n} \sum_q^n (\eta_{Exp, q} - \eta_{Model, q})^2} \quad (12)$$

where n is the total of number of trials, η_{Exp} is the experimental response, η_{Model} is the modeled response, and $\bar{\eta}_{Exp}$ and $\bar{\eta}_{Model}$ are the average values of the experimental and modeled values, respectively.

The multi-objective optimization and multi-criterion decision-making for the determination of the optimal operating parameters were assessed using Design Expert® V.10 software to attain the maximum COD and TOC removal efficiencies of the MET solution by applying a surrogate model of the responses (η_{COD} and η_{TOC}) represented by Equation (9). For this purpose, the Design Expert software package performs multi-objective optimization and multi-criterion decision-making using the steepest ascent method. It is necessary to input additional information into the software to achieve the multi-objective optimization in this study, as shown in Table 4. The two objectives for the selected responses (η_{COD} and η_{TOC}) need to be maximized. Also, the importances of each objective and the operating parameters (A , B , and C) were set as +++ (the default importance). Also, the desirability function was used to indicate the degree to which all the objectives were to be reached. Values close to 1 indicate that all the objectives were met [51].

Table 4. Multi-objective optimization and multi-criterion decision-making of the electrochemical process.

Response	Objective	Limits		Unit	Importance
		Min.	Max.		
pH ₀	Is in range	5	8	Dimensionless	+++
I	Is in range	2.5	4.0	A	+++
Q	Is in range	0.80	1.7	L/min	+++
η_{COD}	Maximize	37.5	62.5	%	+++
η_{TOC}	Maximize	54.58	91.24	%	+++

2.8. Phytotoxicity Test

After determining the best operating conditions for the electro-oxidation of a MET aqueous solution, a phytotoxicity test was conducted before and after the electrochemical wastewater treatment containing MET, using mung bean seeds (*Vigna radiata*) because their fast-duration crop is suitable for various soil types [52]. For this purpose, 30 mung bean seeds were planted in four 250 mL glasses. Over ten days, each set of seeds was irrigated with 10 mL of distilled water, tap water, and MET aqueous solution before and after the electrochemical treatment. After ten days, the germination percentage and lengths of the shoot and root were measured. Also, the germination percentage was computed utilizing Equation (13) [53]:

$$GP(\%) = \frac{NSG}{NSS} \times 100 \quad (13)$$

where GP is the germination percentage, NSG is the number of seeds germinated, and NSS is the number of seeds sown.

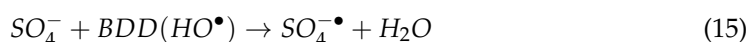
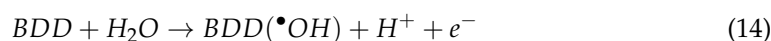
2.9. HPLC and Mass Spectrometry

The MET electro-degradation progress was conducted using reverse-phase high-performance liquid chromatography (RP-HPLC) equipped with a photodiode detector (Agilent 6410) in the range 200–600 nm. The column used was a Hypersil GOLD, 5 μm , 150 \times 4.6 mm, and the mobile phase was methanol (grade HPLC) and an aqueous solution of 0.1% formic acid. Isocratic separation was performed at 90/10 (*v/v*) at a flow rate of 0.5 mL/min. The sample was concentrated for a Sep-Pak[®] C18 cartridge (water-acetonitrile), the injection volume was 15 μL , and a temperature of 30 $^{\circ}\text{C}$ was used. To identify the MET electro-degradation species, triple Quad LC/MS and an electrospray ionization (ESI) source were used. Mass spectrometry (MS) was carried out in the positive ion mode at a capillary voltage of 3000 V and a nitrogen gas flow rate of 10 L/min, and the gas temperature was 380 $^{\circ}\text{C}$.

3. Results and Discussion

3.1. Voltammetry Study

The electrochemical behavior of the aqueous solution containing 50 mg/L of MET on the Nb/BDD anode was studied using cyclic voltammetry measurements. Figure 2 illustrates the cyclic voltammetry of the Na_2SO_4 (black line) and Na_2SO_4 + MET (blue line) with the Nb/BDD anode at a potential scan rate of 100 mV/s. This figure exhibits the typical hysteresis behavior of a non-reversible system when wastewater contains MET, with a well-delineated peak at around +1.3 V, indicating that MET is directly oxidized on the Nb/BDD anode's surface. Also, when the potential is increasing, a smoothed current signal is seen after the peak at +1.3 V, indicating a signal of the oxidation byproducts from MET. Furthermore, these intermediates could be directly oxidized by the free heterogeneous radicals ($\bullet\text{OH}$) on the Nb/BDD anode's surface [54,55] and other oxidants (e.g., sulfate radicals and persulfate ions) present in the aqueous medium because Na_2SO_4 was utilized as the electrolyte support [56–58], which could contribute to the fast degradation/mineralization of organic pollutants, such as MET, in synthetic wastewater. The formation of sulfate radicals ($\text{SO}_4^{\bullet-}$) and the persulfate ions ($\text{S}_2\text{O}_8^{2-}$) formed on the Nb/BDD anode follow the reaction mechanisms given by Equations (14)–(17) [23,26,59,60]:



3.2. Data-Based Model Fitting

Table 5 records the 17 trials in the experimental design for the electro-oxidation of MET, as proposed in Section 2.4. The reported experimental values of the COD and TOC removal efficiencies of the MET are in the ranges 37.50–62.50% and 54.50–91.24%, respectively. Given the experimental responses of η_{COD} and η_{TOC} , reduced cubic and quadratic polynomials were fitted for η_{COD} and η_{TOC} , respectively. Table 6 shows the coefficients (β values) of the data-based polynomials fitted.

The data-based models' fit statistics, shown in Table 6, indicate that the fitted models are in excellent agreement with the experimental values because the determination coefficients (R^2) for responses η_{COD} and η_{TOC} (0.9816 and 0.9430, respectively) are close to 1 [61], as confirmed by the parity graphics' shape in Figure 3. Also, the Modeled_ R^2 values (0.7826 and 0.7036) for η_{COD} and η_{TOC} are in reasonable agreement with the Adjusted_ R^2 values (0.9632 and 0.8698) for η_{COD} and η_{TOC} because the differences are less than 0.2. Addition-

ally, the adequate precision measures (24.28 and 14.36 for η_{COD} and η_{TOC} , respectively) are greater than 4, indicating that the fitted data-based models can be used to navigate the design space [62]. In addition, the obtained values of the coefficient of variance (C.V., 2.69 and 4.556% for η_{COD} and η_{TOC} , respectively) showed the great accuracy of the performed test because the C.V.s were under 10%, indicating that the adjusted data-based models (η_{COD} and η_{TOC}) have a high degree of reproducibility [63]. Furthermore, the performance indexes, MAE (0.79 and 1.80 for η_{COD} and η_{TOC} , respectively) and RMSE (0.93 and 2.21 for η_{COD} and η_{TOC} , respectively), for η_{COD} and η_{TOC} indicate that the reduced cubic and quadratic polynomial models are appropriate for modeling the responses (η_{COD} and η_{TOC}) because the values of MAE and RMSE are less than the standard deviations (1.36 and 3.45 for η_{COD} and η_{TOC} , respectively), and the ratios of $RSME/(\text{Std. Desv.})$ are 0.68 and 0.64 for η_{COD} and η_{TOC} , respectively, indicating satisfactory data-based model simulations for both responses (η_{COD} and η_{TOC}) because the $RSME/(\text{Std. Desv.})$ values are less than or equal to 0.70 [64].

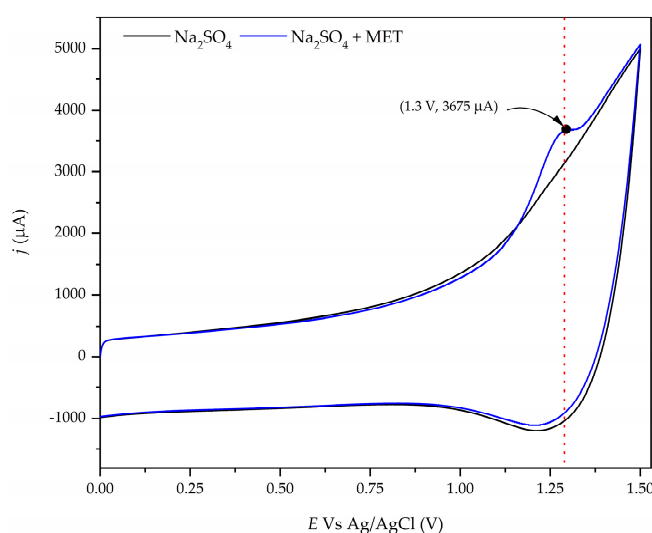


Figure 2. Cyclic voltammogram for a solution containing 50 mg/L of MET in a 0.1 M Na_2SO_4 solution at the Nb/BDD anode. The potential scan was at 100 mV/s.

Table 5. Design matrix (CCRD) and experimental results of the electro-oxidation of MET.

Std.	Run	Design Space	Actual Parameters			Encoded Parameters			Responses	
			pH ₀	I (A)	Q (L/min)	A	B	C	η_{COD}	η_{TOC}
1	1	Factorial	5.00	2.50	0.80	-1	-1	-1	45.83 ± 0.67	65.41 ± 0.83
9	2	Axial	3.98	3.25	1.25	-α	0	0	62.50 ± 0.70	80.15 ± 0.57
10	3	Axial	9.02	3.25	1.25	+α	0	0	41.67 ± 0.57	83.46 ± 0.68
12	4	Axial	6.50	4.51	1.25	0	+α	0	56.33 ± 0.35	78.72 ± 0.55
3	5	Factorial	5.00	4.00	0.80	-1	+1	-1	58.33 ± 0.67	91.24 ± 0.64
16	6	Central	6.50	3.25	1.25	0	0	0	52.17 ± 0.76	70.91 ± 0.56
8	7	Factorial	8.00	4.00	1.70	+1	+1	+1	49.83 ± 0.64	81.87 ± 0.85
4	8	Factorial	8.00	4.00	0.80	+1	+1	-1	45.83 ± 0.72	84.99 ± 0.80
13	9	Axial	6.50	3.25	0.49	0	0	-α	38.58 ± 0.50	79.15 ± 0.75
5	10	Factorial	5.00	2.50	1.70	-1	-1	+1	37.50 ± 0.53	59.64 ± 0.74
7	11	Factoria	5.00	4.00	1.70	-1	+1	+1	50.00 ± 0.80	82.86 ± 0.25
11	12	Axial	6.50	1.99	1.25	0	-α	0	49.99 ± 0.50	54.58 ± 0.55
2	13	Factorial	8.00	2.50	0.80	+1	-1	-1	51.00 ± 0.62	67.29 ± 0.61
6	14	Factorial	8.00	2.50	1.70	+1	-1	+1	58.30 ± 0.95	77.02 ± 0.87
14	15	Axial	6.50	3.25	2.00	0	0	+α	58.33 ± 0.46	71.08 ± 0.35
17	16	Central	6.50	3.25	1.25	0	0	0	52.33 ± 0.59	76.77 ± 0.97
15	17	Central	6.50	3.25	1.25	0	0	0	50.67 ± 0.74	79.82 ± 0.53

Table 6. Beta values of the data-based models in encoded terms.

Response/ β -Values	Terms of Data-Based Polynomial Models											
	Intercept	A	B	C	AB	AC	BC	A ²	B ²	C ²	A ² C	AB ²
η_{COD}/β_i	51.75	-6.19	1.61	5.87	-4.83	3.50	---	---	---	-1.50	-6.54	7.86
	$R^2 = 0.9816$; $Adj_R^2 = 0.9632$; $Model_R^2 = 0.7826$; C.V. = 2.69%; Std. Dev. = 1.36; Adeq_Precision = 24.80											
η_{TOC}/β_i	75.71	1.29	8.22	-1.55	-3.31	2.59	-1.93	2.52	-2.84	0.15	---	---
	$R^2 = 0.9430$; $Adj_R^2 = 0.8698$; $Model_R^2 = 0.7036$; C.V. = 4.56%; Std. Dev. = 3.45; Adeq_Precision = 14.36											

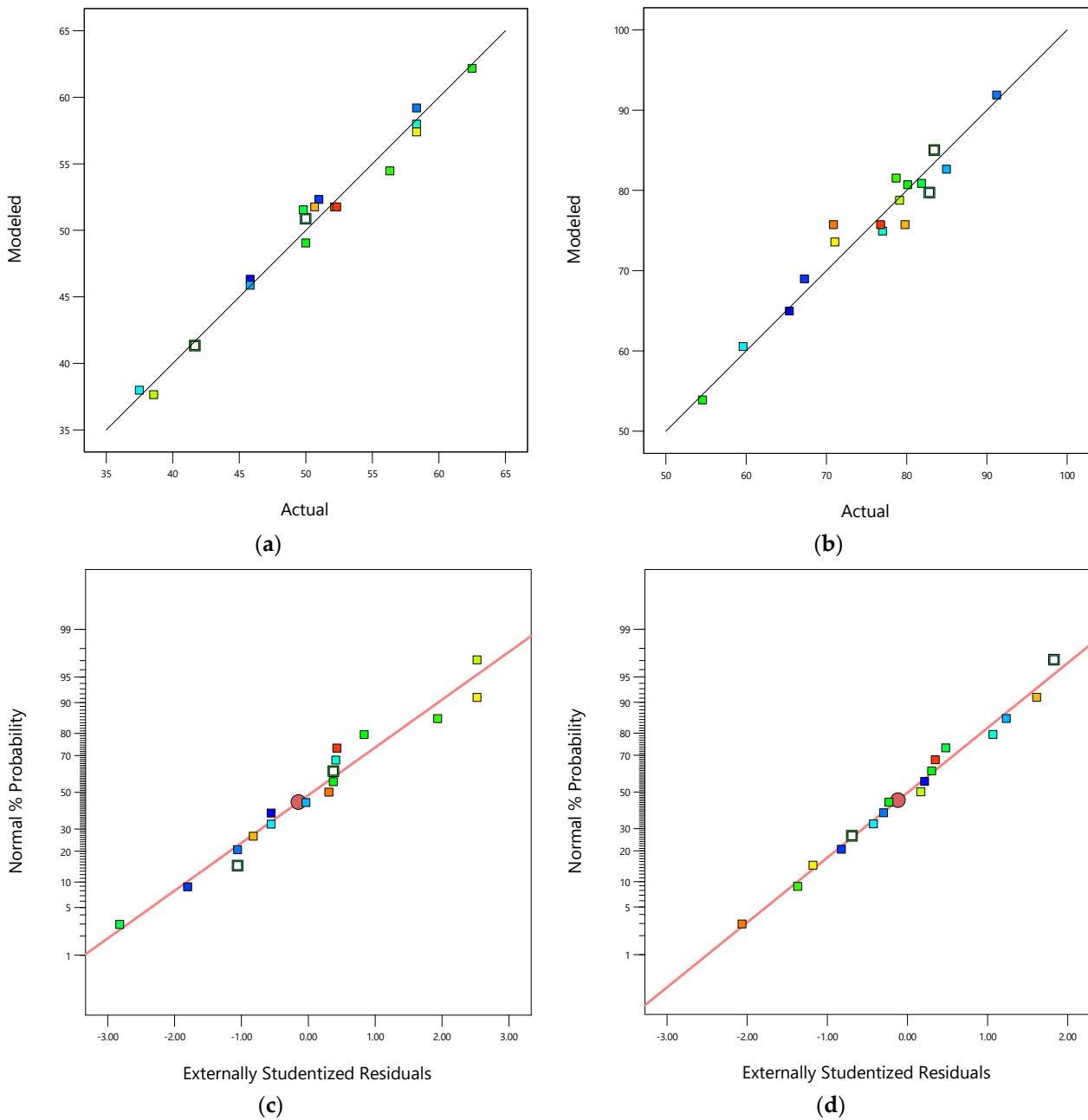


Figure 3. (a) Parity graphic for COD removal efficiency; (b) parity graphic for TOC removal efficiency; (c) normal percentage probability residual versus externally studentized residual for COD removal efficiency; (d) normal percentage probability residual versus externally studentized residual for TOC removal efficiency.

The β_0 values (51.75 and 75.71 for η_{COD} and η_{TOC} , respectively) represent the averages of all the runs for both responses (η_{COD} and η_{TOC}). An analysis of the variance inflation factors (VIFs) for the reduced cubic data-based model (η_{COD}) reveals that the VIF values of the terms B , AB , AC , and C^2 are 1, indicating that these terms are orthogonal. Also, the VIF values of the terms A , C , A^2C , and AB^2 are 2.42, suggesting that these terms are multicollinear because all the VIF values are greater than 1. A similar analysis of the VIF values for the quadratic data-based model (η_{TOC}) reveals that the VIF values of the terms A , B , C , AB , AC , and BC are 1, suggesting that these terms are orthogonal. Additionally, the VIF values of the terms A^2 , B^2 , and C^2 are 1.16, indicating that these terms are multicollinear because all the VIF values are greater than 1. Moreover, a higher value than 1 indicates a hard correlation factor. As a general rule of thumb, VIF values lower than 10 are acceptable [65].

The parity graphic (see Figure 3a,b) and the normal percentage probability residual versus the studentized residual plot (see Figure 3c,d) demonstrate the model's adequacy to predict both responses (η_{COD} and η_{TOC}). Figure 3a,b reveal high-quality correlations between the experimental values of both responses (η_{COD} and η_{TOC}) and both modeled responses (η_{COD} and η_{TOC}). Figure 3c,d show normal distributions because the plotted points align closely to a straight line at a 45-degree angle. This suggests that the experimental responses follow a normal distribution. Additionally, the outlier values fall within the ranges ± 1.05 and ± 0.99 for η_{COD} and η_{TOC} , respectively, indicating that the data-based models used to fit the responses were reasonably accurate, with minimal data register errors and high degrees of reproducibility.

3.3. ANOVA Test

Table 7 displays the results of the ANOVA test for the data-based models fitted. The test results indicate that the adjusted model for the COD removal efficiency is statistically suitable for modeling the electro-oxidation of MET within the studied operational parameter range. The model's F -value of 54.38 for η_{COD} suggests that the adjusted polynomial data-based model is significant because the F -value is higher than the p -value (0.0001). Additionally, the lack of fit is not significant for the pure error, as the F -value of 2.6 is higher than the p -value (0.3034). Furthermore, the terms A , B , C , AB , AC , C^2 , A^2C , and AB^2 are significant because their p -values (0.0001, 0.0024, 0.0001, 0.0001, 0.0001, 0.0041, 0.0001, and 0.0001) are less than 0.05. Similarly, the model's F -value of 1379.37 for η_{TOC} suggests that the adjusted polynomial data-based model is significant because the F -value is higher than the p -value (0.0014). Moreover, the lack of fit is not significant for the pure error, as the F -value of 42.31 is higher than the p -value (0.8162). Moreover, the terms B , AB , A^2 , and B^2 are significant because their p -values (0.0001, 0.0300, 0.0439, and 0.0281) are less than 0.05.

Table 7. ANOVA test results for the data-based models (η_{COD} and η_{TOC}).

Source	Sum of Squares	Degrees of Freedom	Mean Squared	F-Value	p-Value	Remark
<i>COD removal efficiency (η_{COD})</i>						
Model	78.15	8	98.39	53.28	<0.0001	Significant
A	216.94	1	216.94	117.48	<0.0001	
B	35.36	1	35.36	19.15	0.0024	
C	195.03	1	195.03	105.62	<0.0001	
AB	186.92	1	186.92	101.23	<0.0001	
AC	97.93	1	97.93	53.03	<0.0001	
C²	29.20	1	29.20	15.81	0.0041	
A²C	141.64	1	141.64	76.71	<0.0001	
AB²	204.67	1	204.67	110.84	<0.0001	
Residual	14.77	8	1.85			
Lack of fit	13.10	6	2.18	2.60	0.3034	Not significant
Pure error	1.68	2	0.8385			
Cor total	801.92	16				

Table 7. Cont.

Source	Sum of Squares	Degrees of Freedom	Mean Squared	F-Value	p-Value	Remark
<i>TOC removal efficiency (η_{TOC})</i>						
Model	1379.37	9	153.26	12.88	0.0014	Significant
A	22.65	1	22.65	1.90	0.2102	
B	921.77	1	921.77	77.44	<0.0001	
C	32.64	1	32.64	2.74	0.1417	
AB	87.78	1	87.78	7.37	0.0300	
AC	53.87	1	53.87	4.53	0.0709	
BC	29.88	1	29.88	2.51	0.1571	
A²	71.76	1	71.76	6.02	0.0439	
B²	90.73	1	90.73	7.62	0.0281	
C²	0.2739	1	0.2739	0.030	0.8837	
Residual	83.32	7	11.90			
Lack of fit	42.31	5	8.46	0.4127	0.8162	Not significant
Pure error	41.01	2	20.51			
Cor total	1462.69	16				

3.4. Influences of the Operational Parameters

From the previous analysis, the models are summarized in Table 6. The negative signs of the β -values of the terms A , AB , C^2 , and A^2C display negative effects on η_{COD} compared to the positive signs of the β -values of the terms B , C , AC , and AB^2 . Similarly, the negative signs of the β -values of the terms C , AB , BC , and B^2 exhibit negative effects on η_{TOC} compared to the positive signs of the β -values of the terms A , B , AC , A^2 , and C^2 [66].

The perturbation graphics (Figure 4a,b) reveal that all the operating parameters (A , B , and C ; i.e., the initial pH, current intensity, and volumetric flow rate, respectively) are important because of their distinguished slopes. However, the importances of the operating parameters for the response of the η_{COD} follow the order of the initial pH (A), current intensity (B), and volumetric flow rate (C) because of their slopes. Additionally, the importances of the operating parameters for the response of the η_{TOC} follow the order of the initial pH (A), volumetric flow rate (C), and current intensity (B). Additionally, the Spearman rank values are shown in Figure 4c. Figure 4c reveals that the current intensity (B) is the most significant variable for η_{TOC} because the Spearman rank value is 0.773. Meanwhile, the volumetric flow rate (C) is the most significant variable for η_{COD} because the Spearman rank value is 0.20. Also, pH_0 exerts a negative effect on η_{COD} because of its negative Spearman rank value (-0.211), and Q displays a similar effect on η_{TOC} because of its negative Spearman rank value (-0.125).

The Pareto test was performed utilizing Equation (18) to assess the impacts of the terms on the responses (η_{COD} and η_{TOC}).

$$P_i = \left(\frac{\beta_i}{\sum \beta_i} \right) \times 100 \quad (18)$$

Figure 5 shows the Pareto graphic analysis. This figure suggests that the interaction terms A^2C and AB^2 have large effects (17.26 and 20.74%) on η_{COD} , respectively. Meanwhile, A , C , and AB have moderate effects (16.33, 15.49, and 12.74%) on η_{COD} , respectively. Also, the rest of the terms (B , AC , and C^2) have the least effects (4.25, 9.23, and 3.96%) on η_{COD} , respectively. Similarly, the term B has a higher effect (33.69%) on η_{TOC} . Also, the terms AB , AC , A^2 , and B^2 have moderate effects (13.57, 10.61, 10.33, and 11.64%) on η_{TOC} , respectively. In addition, the rest of the terms (A , C , BC , and C^2) have the least effects (5.29, 6.35, 9.9, and 0.61%) on η_{TOC} .

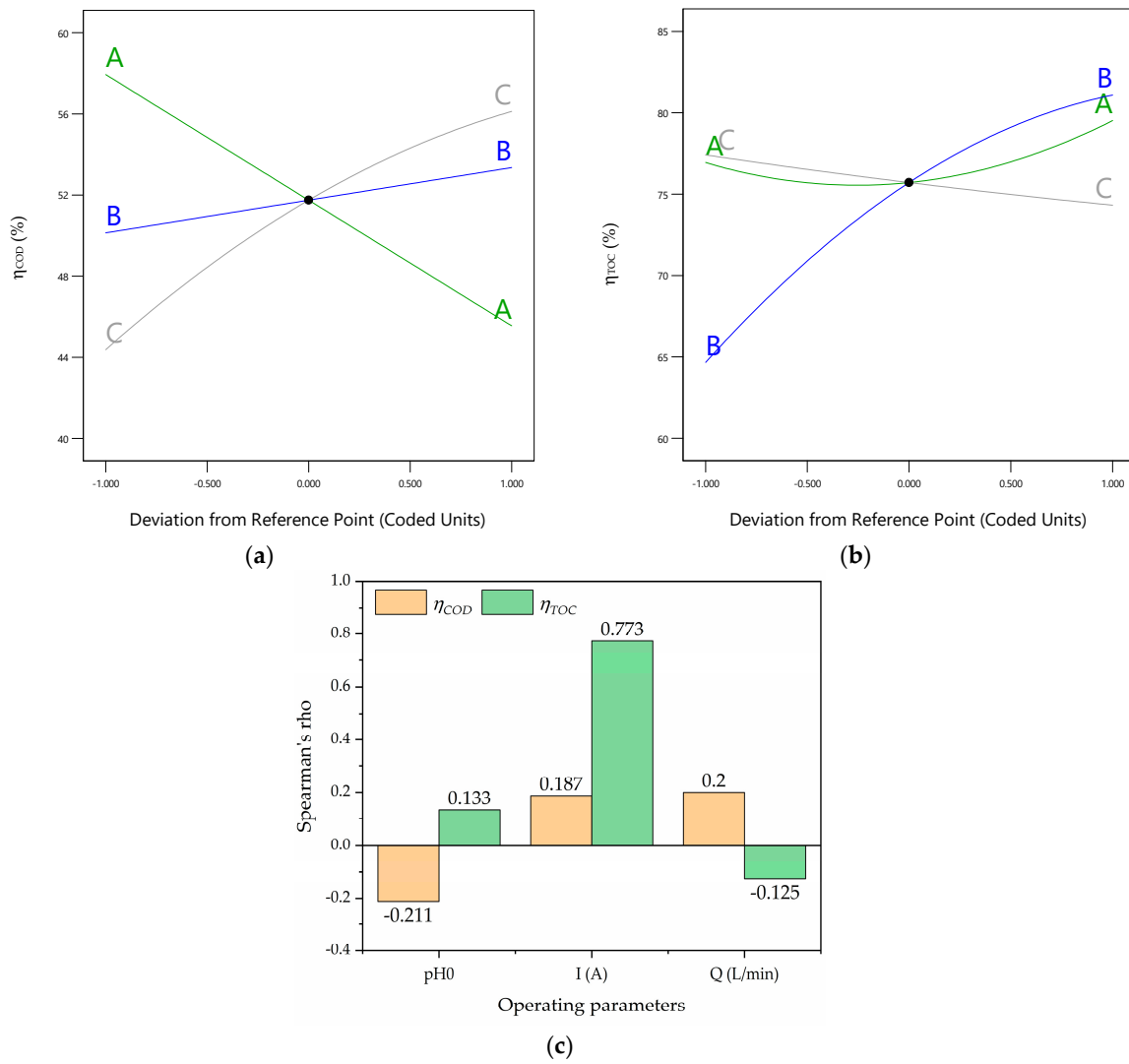


Figure 4. (a) Perturbation graphic for η_{COD} ; (b) perturbation graphic for η_{TOC} ; (c) Spearman's rho plot.

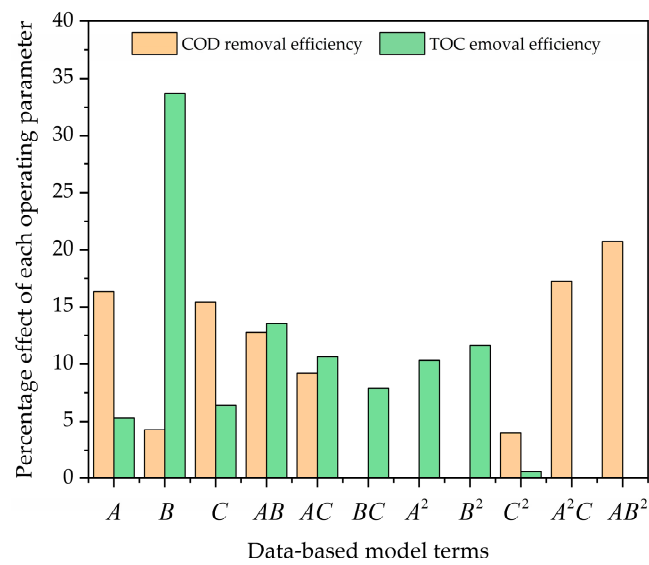


Figure 5. Pareto graphic for both responses (η_{COD} and η_{TOC}).

3.5. Optimization of the Electro-Oxidation of MET

A maximum COD removal efficiency (η_{COD}) of 60.81% and a TOC removal efficiency (η_{TOC}) of 90.07% were achieved at pH₀ 5.0, an I value of 3.84 A, and a Q value of 0.8 L/min within 7.5 h of the total operating time. The overall desirability was 0.9502 (see Figure 6g), indicating that both objective functions ($\max(\eta_{COD}(A, B, C))$ and $\max(\eta_{TOC}(A, B, C))$) were fully met because the desirability attained is classified as fully acceptable [67]. Additionally, both objective functions (responses η_{COD} and η_{TOC}) were simultaneously optimized. These functions were fitted through data-based modeling using Equation (9).

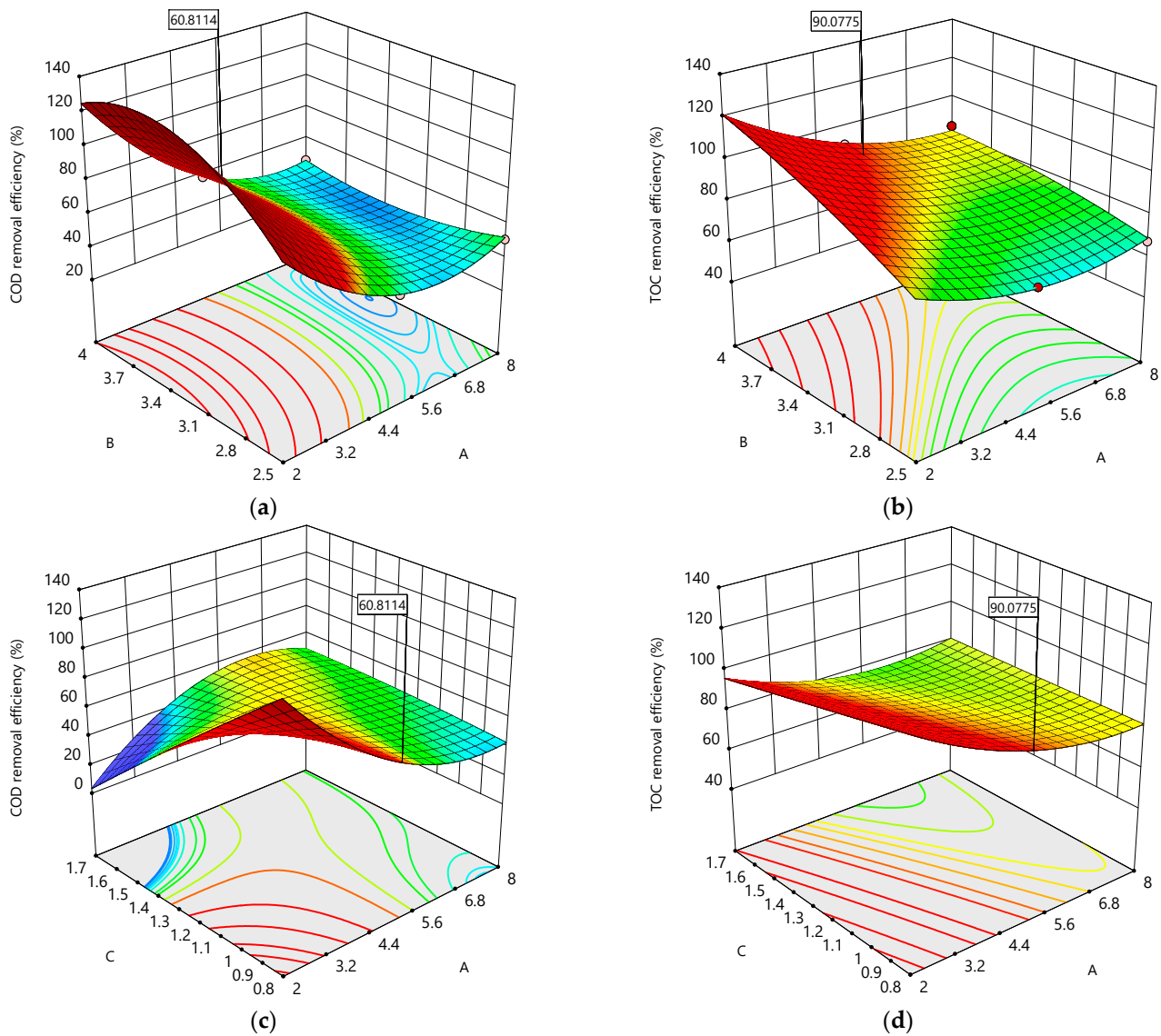


Figure 6. Cont.

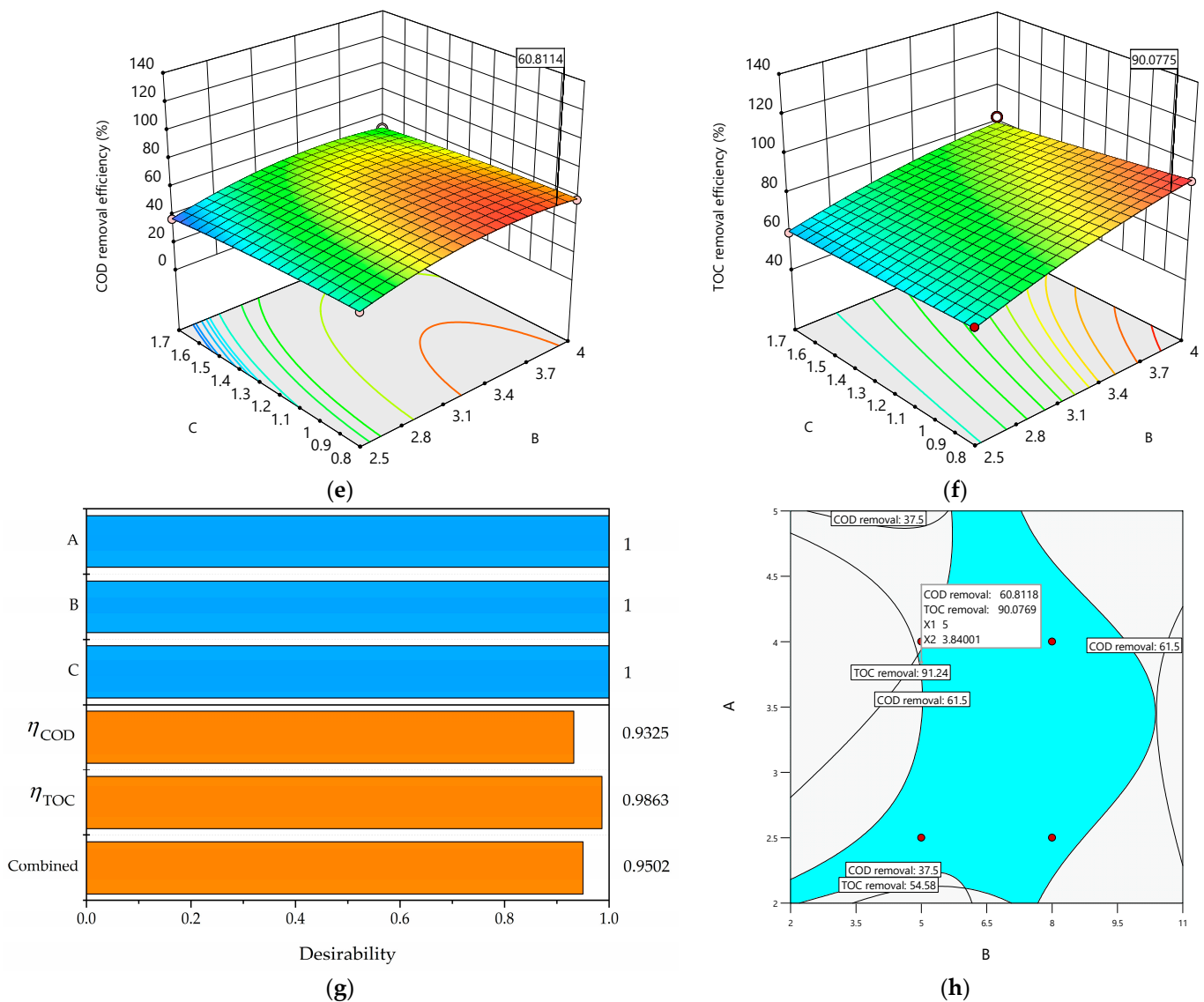


Figure 6. (a) The 3D plot of η_{COD} as a function of B and A ; (b) the 3D plot of η_{TOC} as a function of B and A ; (c) the 3D plot of η_{COD} as a function of C and A ; (d) the 3D plot of η_{TOC} as a function of C and A ; (e) the 3D plot of η_{COD} as a function of C and B ; (f) the 3D plot of η_{TOC} as a function of C and B ; (g) desirability bar plot; (h) overlaid contour plot for three design operating parameters (pH_0 , I , and Q) and two response constraints (η_{COD} and η_{TOC}) with the volumetric flow rate at the optimal level ($C = 0.8$ L/min).

The 3D surface plots (Figure 6a–f) show the interaction effects among the three operating parameters (A , B , and C) studied herein. All the 3D surface plots form saddle areas because the contour plots are hyperboles, indicating that all the interactions between the chosen operating parameters are significant. Figure 6a–f show increases in η_{COD} and η_{TOC} when pH_0 (A) increases because there is a large amount of hydroxyl radicals ($\bullet OH$) produced on the Nb/BDD electrode's surface. Also, Figure 6a,b,d reveal increases in η_{COD} and η_{TOC} when the current intensity (B) increases, indicating a large amount of $\bullet OH$ produced on the Nb/BDD anode's surface and that the main reactions are taking place on the BDD's surface. Similarly, Figure 6d–f display increases in η_{COD} and η_{TOC} when the volumetric flow rate (C) increases because the high mass transport of the pollutant from the bulk to the Nb/BDD anode electrode is achieved. Meanwhile, Figure 6c shows an increase in η_{COD} when the volumetric flow rate (C) decreases because of the low mass transport of the pollutant from the bulk to the Nb/BDD anode's surface. In addition, Figure 6a,b

demonstrate that the interaction between the *A* and *B* operating parameters exerts negative effects on both responses (η_{COD} and η_{TOC}) because the production of $\bullet OH$ at low *A* and *B* values is insufficient to remove the MET. Figure 6c,d reveal that the interaction between the *A* and *C* operating parameters produces positive effects on both responses (η_{COD} and η_{TOC}) because the amount of $\bullet OH$ is sufficient to remove the MET, and the contact time of the MET with the $\bullet OH$ is enough to react effectively when the electrochemical treatment operates at low *A* and *C* values. As well, Figure 6e,f show that the interaction between the *B* and *C* operating parameters has negative effects on η_{COD} and η_{TOC} because of the low amount of $\bullet OH$ produced and the low mass transport of the MET from the bulk to the Nb/BDD anode's surface.

Figure 6h depicts the overlaid graphic, shaped according to the constraint criteria described in Table 4 and both objective functions (η_{COD} and η_{TOC}). The black lines in the overlaid graphic are assembled based on the constraint criteria; the area between the constraints is in blue, and the region that is not convened based on the constraints is in gray. Also, the feasible region is enclosed by pH_0 (2–11), and $I = (2\text{--}5\text{ A})$ at $Q = 0.8\text{ L/min}$ within 7.5 h of the total operating time.

3.6. Data-Based Model Validation

Three complementary trials were conducted using the optimal operating parameters to validate the data-based models (η_{COD} and η_{TOC}). The modeled and experimental data for both responses (η_{COD} and η_{TOC}) are recorded in Table 8. The results in Table 8 indicate the high degree of effectiveness of the multi-objective optimization process employed in this study because the deviation between the modeled and experimental data is less than 10% for both responses (0.49 and 8.37% for η_{COD} and η_{TOC} , respectively). Hence, the maximum experimental COD and TOC removal efficiencies were 61.11 and 83.12%, respectively, when the electrochemical plant operates at an initial pH of 5, a current intensity of 3.84 A, and a volumetric flow rate of 0.8 L/min within 7.5 h of the total operating time. Additionally, the high mineralization percentages of the MET achieved indicate the excellent performance of the Nb/BDD electrodes in the degradation of organic pollutants, such as MET [68].

Table 8. Modeled and experimental responses for the optimal operating parameters.

Operating Parameter	Response	Values (%)		Error (%)
		Modeled	Experimental	
$pH_0 = 5.0$ $I = 3.84\text{ A}$ $Q = 0.8\text{ L/min}$	η_{COD}	60.81	61.11 ± 1.60	0.49
	η_{TOC}	90.07	83.12 ± 1.56	8.37

3.7. The Total Operating Cost of the Electro-Oxidation of MET

The total operating cost of the electrochemical treatment of MET was calculated based on the electricity consumed and the spent mass of the electrolyte support ($NaSO_4$), using Equations (4)–(8). Hence, the total energy consumed was 3.09 kWh, with a total operating cost of 0.19 USD/L when the electrochemical plant operates under the optimal operating conditions ($pH_0 = 5.0$, $I = 3.84\text{ A}$, and $Q = 0.8\text{ L/min}$) within 7.5 h of the electrochemical treatment. It is worth noting that this work includes the cost of the electricity consumed by the pump flow and heat exchanger and the costs of the used mass of the electrolyte support, unlike other studies in the literature.

3.8. Kinetic Models

The electro-oxidation of MET conducted in this work is mainly assumed to be driven by $\bullet OH$ on the Nb/BDD anode's surface [69]. The pseudo first-, pseudo second-, and pseudo first-order kinetic models used to fit the MET, COD, and TOC abatements, respectively,

were set to fit the experimental data according to Equations (20) and (21) for the MET, COD, and TOC abatements:

$$\frac{d[\text{MET}]}{dt} = -\underbrace{k[\bullet\text{OH}]}_{k_{app, \text{MET}}}[\text{MET}] = -k_{app}[\text{MET}] \quad (19)$$

$$\frac{d[\text{COD}]}{dt} = -\underbrace{k[\bullet\text{OH}]^2}_{k_{app, \text{COD}}}[\text{COD}]^2 = -k_{app}[\text{COD}]^2 \quad (20)$$

$$\frac{d[\text{TOC}]}{dt} = -\underbrace{k[\bullet\text{OH}]}_{k_{app, \text{TOC}}}[\text{TOC}] = -k_{app}[\text{TOC}] \quad (21)$$

Figure 7a illustrates the MET's depletion over time because of the treatment, reaching a 95.19% degradation efficiency of the MET after 5 h of the treatment (see Figure S1 in the Supplementary Materials). This behavior suggests that the electro-oxidation of MET forms other simple organic compounds (e.g., carboxylic acid and aromatic molecules) because $\bullet\text{OH}$ attacks the chemical structure of the MET. Additionally, Figure 7b–d show that pseudo first-, pseudo second-, and pseudo first-order kinetic models effectively describe the decays of the MET, COD, and TOC, with determination coefficients (R^2) of 0.9999, 0.9977, and 0.9945, respectively. These high R^2 values indicate good model fits because the R^2 values are close to 1 [61], and the low RMSEs (0.16, 1.49, and 1.00 for MET, COD, and TOC decays, respectively) further support this conclusion (in line with Table 9). Also, Figure 7b–d suggest the constant production of $\text{HO}\bullet$ on the Nb/BDD anode's surface. Furthermore, the high correlation between the experimental and modeled MET depletions suggests a large amount of $\bullet\text{OH}$ generated on the Nb/BDD anode's surface (see Figure S1), as stated in [70].

Table 9. Estimated kinetic coefficients, goodness-of-fit values, and determination coefficients.

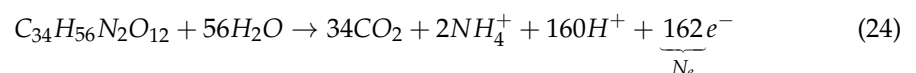
Variable	Order (n)	k_{app}	R^2	RMSE
[MET]	1	0.6072 1/h	0.9999	0.16
[COD]	2	0.0019 L/mg h	0.9973	1.49
[TOC]	1	0.2380 1/h	0.9945	1.00

The kinetic constants of the pseudo first-, pseudo second-, and pseudo first-order fitted models for the MET, COD, and TOC abatements are summarized in Table 9. Also, the pseudo first-order models for the MET and TOC abatements agree with references [35,36,71]. Additionally, the kinetic constants for the decays of the COD and TOC ($k_{app, \text{COD}}$ and $k_{app, \text{TOC}}$, respectively) can be considered as global because the electro-oxidation of the byproducts was not discriminated against for their calculations.

To understand whether the electro-oxidation of MET is controlled by the mass transport or current, the limit current intensity (I_{lim}) can be estimated using Equations (22)–(24) [72]:

$$\ln\left(\frac{[\text{TOC}]_t}{[\text{TOC}]_0}\right) = -\underbrace{\frac{A_A}{V_{treated}}k_m t}_{\text{slope}} \quad (22)$$

$$I_{lim} = N_e A_A F k_m [\text{MET}]_0 \quad (23)$$



where A_A (0.0032 m²) is the anode's surface area, N_e is the number of electrons, $V_{treated}$ (0.0025 m³) is the treated volume, F (96,487 C/mol) is the Faraday constant, k_m (m/s) is the mass transfer coefficient, I_{lim} (mA) is the current intensity, t (s) is the time, $[\text{MET}]_0$

(0.073 mol/m³) is the initial concentration of the MET, and [TOC]_t and [TOC]₀ are the final and initial TOC concentrations (mg/L), respectively.

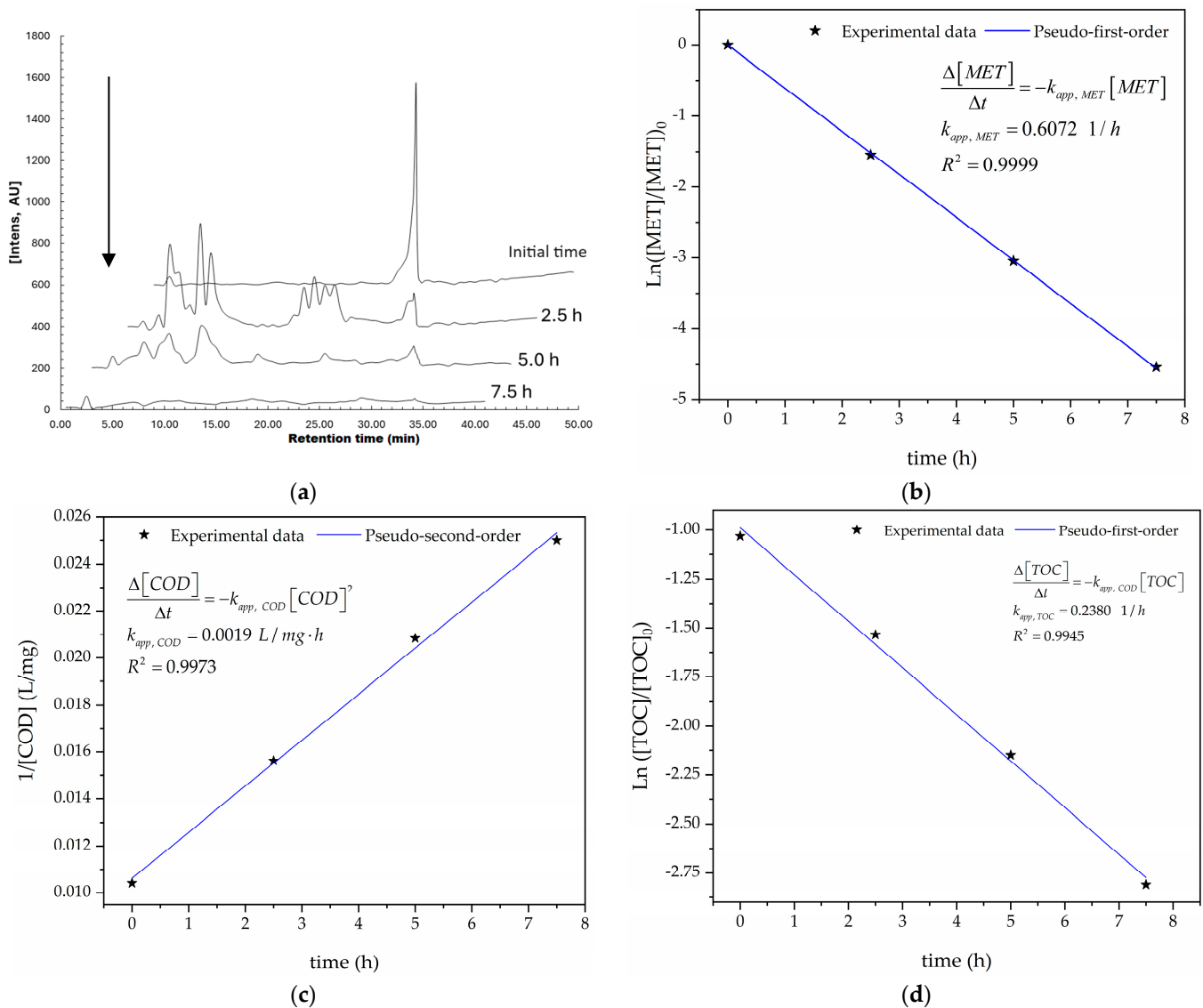


Figure 7. (a) Abatement of the concentrations of MET as functions of time, as determined using HPLC; (b) kinetic analysis for the pseudo first-order model for the MET decay; (c) kinetic analysis for the pseudo second-order model for the COD decay; (d) kinetic analysis for the pseudo first-order model for the TOC decay, performed under the optimal operating conditions: $T = 25 \text{ }^\circ\text{C}$, $\text{pH}_0 = 5$, $I = 3.84 \text{ A}$, and $Q = 0.8 \text{ L/min}$ within a 7.5 h reaction time.

After performing the linear regression analysis of the TOC data and using Equation (22), the mass transfer coefficient was calculated at $5.16 \times 10^{-5} \text{ m/s}$. This value is similar to those calculated for the electro-oxidations of ciprofloxacin [39] and 2-chlorophenol [73] using the same flow-by reactor. The electro-oxidation of MET is primarily controlled by the mass transfer because the applied current intensity ($I_{app} = 3.84 \text{ A}$) is greater than the limiting current intensity ($I_{lim} = 0.189 \text{ A}$), which is in line with what was reported in reference [68].

3.9. Byproduct Identification and Reaction Pathway

During the electro-oxidation of MET with two Nb/BDD electrodes under the optimal operating conditions ($\text{pH}_0 = 5$, $I = 3.84 \text{ A}$, and $Q = 0.8 \text{ L/min}$ within a 7.5 h reaction time), an LC/MS analysis was conducted to identify the byproducts. The LC/MS analysis revealed

11 intermediates (refer to Table 10). The electro-degradation of MET was observed to occur in three phases. In the first phase (indicated by the green line), the branched aromatic ring with an ether functional group in the MET was oxidized by $\bullet\text{OH}$ (see Figure 8 and Table 10). The second phase (marked by the orange line) involved the attack of the $\bullet\text{OH}$ on the aromatic ring, leading to the complete degradation of the compound, including the breakage of the aromatic ring and the degradation of the branch with the amine group (see Figure 8). These stages illustrate how MET is progressively broken down into simpler compounds through advanced oxidation processes. This observation aligns with reference [74]. It is important to note that because of the low abundance and complexity of the sample, there are additional unidentified intermediates, resulting in peak overlap and a lack of reference information.

Table 10. Byproducts identified after the electro-oxidation treatment of MET.

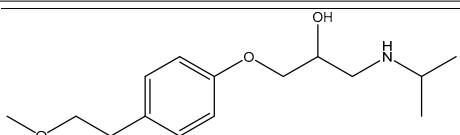
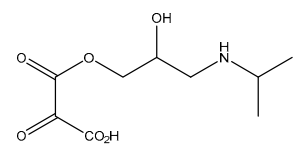
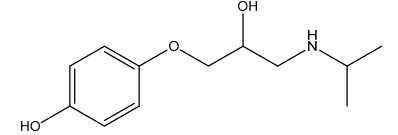
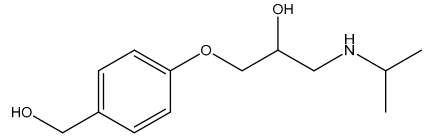
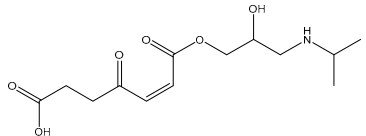
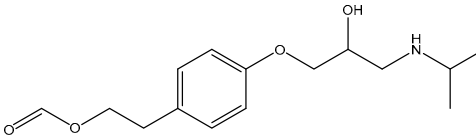
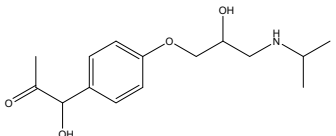
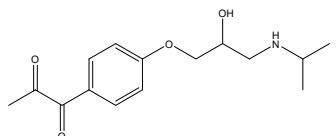
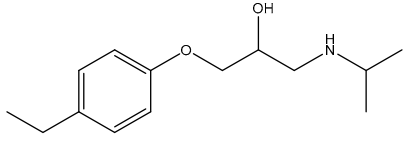
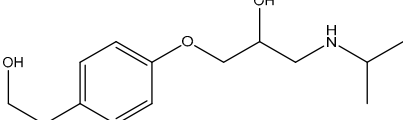
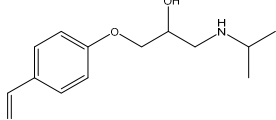
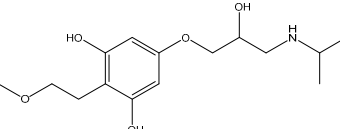
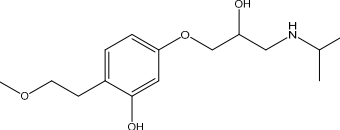
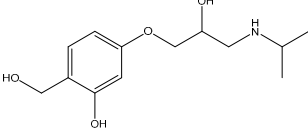
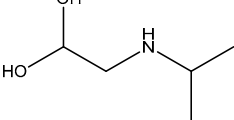
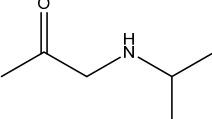
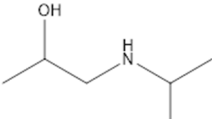
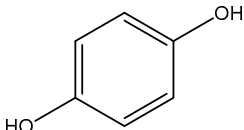
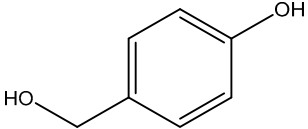
Stage	Retention Time (min)	[M + H] ⁺ m/z Experimental	Proposed Structure	Difference (%)
MET	34.59	268		0.05
I	2.56	206		0.02
II	4.78	226		0.04
III	5.68	240		0.03
IV	6.45	274		0.02
V	8.56	282		0.02
VI	9.35	284		0.03
VII	16.21	282		0.02

Table 10. Cont.

Stage	Retention Time (min)	[M + H] ⁺ m/z Experimental	Proposed Structure	Difference (%)
VIII	17.25	238		0.02
IX	19.54	254		0.02
X	20.47	226		0.02
XI	11.68	300		0.04
XII	14.26	284		0.05
XIII	18.08	256		0.02
XIV	5.26	134		0.03
XV	8.46	116		0.02
XVI	10.92	118		0.02
XVII	12.23	128		0.01
XVIII	18.43	140		0.01

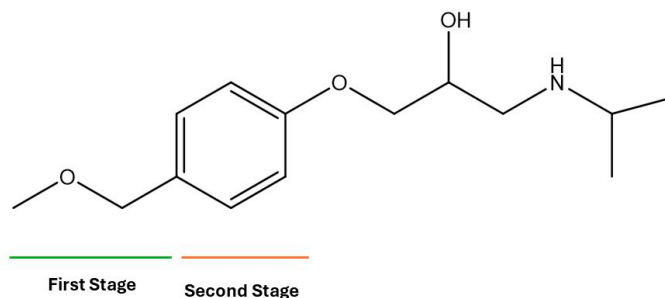


Figure 8. Electro-degradation reaction pathway of MET stages, performed under the optimal operating conditions: $T = 25\text{ }^{\circ}\text{C}$, $\text{pH}_0 = 5$, $I = 3.84\text{ A}$, and $Q = 0.8\text{ L/min}$ within a 7.5 h reaction time.

The low susceptibility of the MET to degradation in the part containing the alcohol group is because of the stability of the C-O bonds, the absence of conjugated systems that facilitate electron delocalization, and the general resistance of secondary alcohols to oxidation [74–76]. The amino group, with a pH below 9.7, becomes protonated (NH_3^+), causing the lone pair of electrons on the nitrogen atom to bond with the additional proton, making the nitrogen less available to react with electrophilic species; this reduces the availability of the lone pair of electrons, decreasing the capacity of the $\bullet\text{OH}$ to participate in chemical reactions [77].

The proposed reaction pathway for the electro-degradation of MET consists of three main stages of transformation, as shown in Figure 9. In the first stage, metoprolol decomposes into several intermediate products: one with a molecular weight of 226 m/z , resulting from the loss of the isopropylamine group; another intermediate with a weight of 240 m/z , formed by the oxidation of the secondary alcohol group to a ketone; and a third product with a weight of 282 m/z , indicating the formation of an oxidized dimer. In the second stage, these intermediates undergo further decomposition: The 226 m/z compound transforms to another with a weight of 238 m/z , suggesting the oxidative cleavage of the branch with the ether group and the aliphatic chain of the aromatic ring of the metoprolol; the 254 m/z intermediate transforms to a 256 m/z compound through the additional oxidation of the aromatic ring and degradation of the aliphatic chain, and the 282 m/z product decomposes into a compound with a weight of 274 m/z , indicating the breakage of the aromatic ring. Finally, in the final stage, the products from the second stage decompose into even simpler compounds: one product with a weight of 140 m/z , another with a weight of 118 m/z , and a compound with a weight of 116 m/z , reflecting the complete breakage of the aliphatic bonds and oxidation of the functional groups.

3.10. Comparison of Results with Those in the Literature

The results obtained in this study were compared with those found in the literature (refer to Table 11). This work is the only one reporting the associated operating cost, using two Nb/BDD electrodes, and carrying out the optimized electro-oxidation of MET. Although previous studies have reported that MET was degraded and mineralized to 100% [34,36], it is essential to consider that the treated volume in those studies was less (0.6 L and 1.0 L, respectively) compared to the volume used in this work (2.5 L). Additionally, the active areas reported in the literature [34,36] are 2.44 and 3.13 times larger (78 and 100 cm^2 , respectively) than the area used in this work (32 cm^2). Reference [34] and this work are unique studies that used a flow-by reactor for the electro-degradation of MET. However, the literature typically reports that four cells were used to achieve the 100% degradation and mineralization of MET. Also, reference [35] achieves a TOC removal efficiency (97%) higher than that in this study for a treated volume four times higher (10 L) than the volume used in this study (2.5 L). Nevertheless, this study uses five times the initial MET concentration (50 mg/L) than the initial MET concentration (10 mg/L) reported in the literature [35]. Based on the results presented in this study, the electrochemical method employed for the wastewater treatment effectively removes MET.

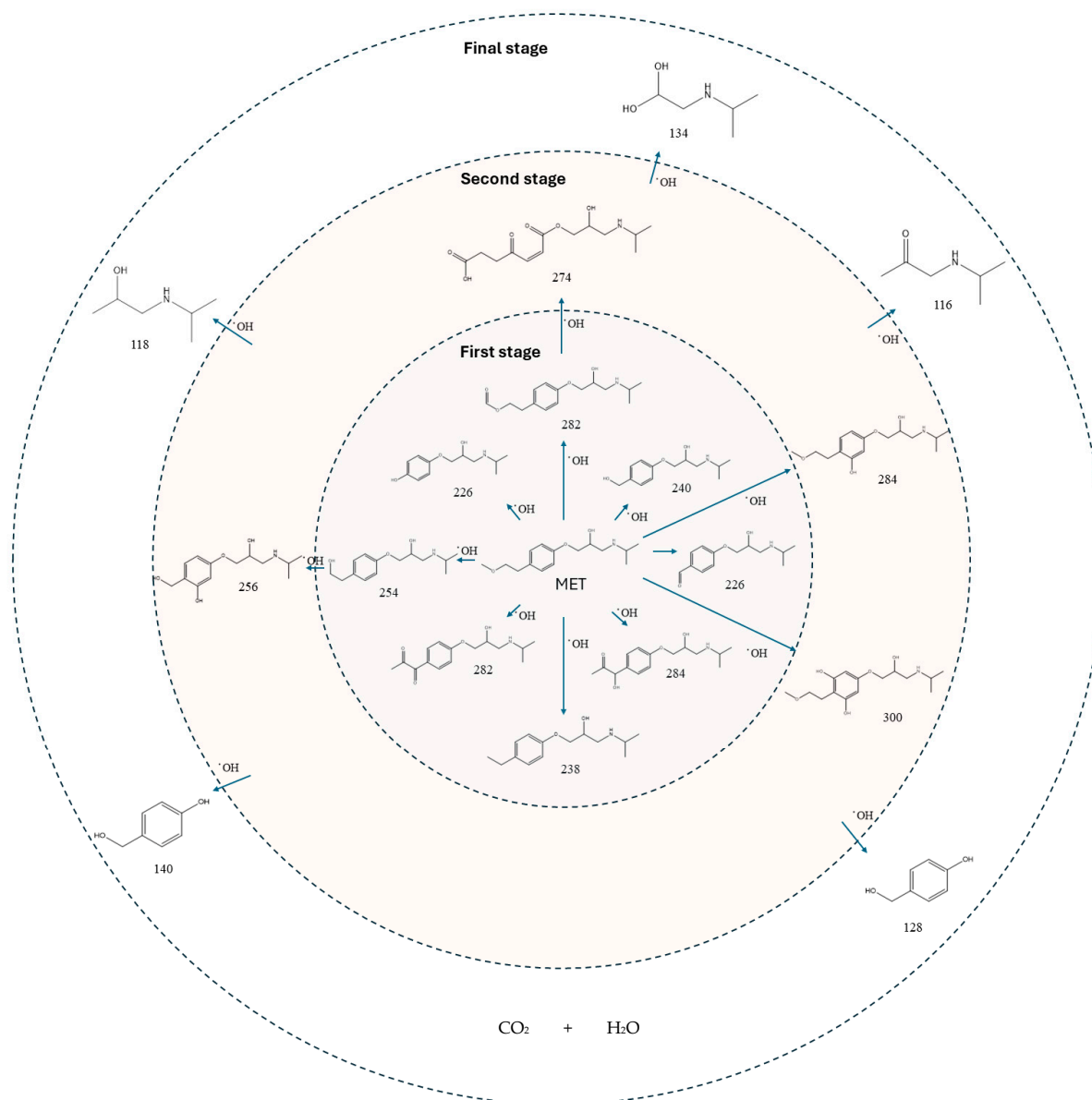


Figure 9. Reaction pathway of MET, performed under the optimal operating conditions: $T = 25\text{ }^\circ\text{C}$, $\text{pH}_0 = 5$, $I = 3.84\text{ A}$, and $Q = 0.8\text{ L/min}$ within a 7.5 h reaction time.

3.11. Toxicity Test

To complement the information about the hazardousness of the aqueous solution of MET at the beginning of and after the electro-oxidation treatment, a phytotoxicity test was conducted using mung bean seeds (*Vigna radiata*) according to Section 2.8.

Table 12 indicates that the germination percentage of *Vigna radiata* irrigated with the treated effluent (93.33%) equals those of *Vigna radiata* irrigated with the distilled water and electrolyte controls. However, the germination percentage of *Vigna radiata* drops to 73.33% for the solution containing 50 mg/L of MET, which is lower than the germination percentage of *Vigna radiata* (93.33%) irrigated with the treated effluent. Additionally, the germination percentage of *Vigna radiata* irrigated with tap water was 100%, with a shorter average root length (5.1 cm) compared to those of *Vigna radiata* irrigated with the distilled water, electrolyte support, and wastewater at the beginning and the end of the electro-oxidation process. Furthermore, the average shoot lengths of *Vigna radiata* (24.0 and 20.7 cm for irrigation with distilled water and tap water, respectively) are longer than those of *Vigna radiata* irrigated with the electrolyte support and wastewater, at the beginning and

the end of the electro-oxidation process, when irrigated with distilled water and tap water. Finally, the germination percentages of *Vigna radiata* irrigated with the electrolyte solution and treated solution of the MET are equal (93.33%), and the average shoot length of *Vigna radiata* treated with the solution of the MET (6.8 cm) is longer than those of *Vigna radiata* treated with the solution of the MET and without treatment (5.6 cm), which demonstrates the removal and degradation of the MET from the wastewater through the application of the electrochemical process using BDD electrodes as the cathode and anode, as assessed in this work.

Table 11. Comparison of various electro-oxidations of MET.

Process	Reaction Conditions	Electrodes		$V_{treated}$ (L)	Main Results				Ref.
		Ca.	An.		COD Removal (%)	TOC Removal (%)	MET Removal (%)	Cost (USD/L)	
E ^O	Flow-by reactor, $A_A = 32 \text{ cm}^2$, $[\text{MET}]_0 = 50 \text{ mg/L}$, $\text{pH}_0 = 5.0$, $I = 3.84 \text{ A}$, $[\text{Na}_2\text{SO}_4] = 0.1 \text{ M}$, $Q = 0.8 \text{ L/min}$, $T = 25 \text{ }^\circ\text{C}$, and $t = 7.5 \text{ h}$	BDD	BDD	2.5	60.81	90.07	95.19	0.19	This work
E ^{NO}	Batch cell, $A_A = 4 \text{ cm}^2$, $[\text{MET}]_0 = 2 \text{ mg/L}$, 5 V , $[\text{NaCl}] = 5000 \text{ mg/L}$, and $t = 1.33 \text{ h}$	Pt	Graphite-PVC	0.1	---	---	>95.00	---	[33]
US/UV-E ^{NO}	Batch cell, $A_A = 78 \text{ cm}^2$, Ultrasound at 200 W , UV at 250 nm , $I = 2.34 \text{ A}$, $[\text{Na}_2\text{SO}_4] = 0.035 \text{ M}$, $[\text{MET}]_0 = 100 \text{ mg/L}$, and $t = 10 \text{ h}$	SS	BDD	1.0	---	100.00	100.00	---	[34]
SPEF ^{NO}	Flow reactor, four cells, $A_A = 100 \text{ cm}^2$, $\text{pH}_0 = 3.0$, $I = 3 \text{ A}$, $[\text{Fe}^{2+}] = 0.5 \text{ mM}$, $[\text{Na}_2\text{SO}_4] = 0.1 \text{ M}$, $[\text{TOC}]_0 = 10 \text{ mg/L}$, $T = 35 \text{ }^\circ\text{C}$, $Q = 4.17 \text{ L/min}$, and $t = 6 \text{ h}$	ADE	BDD	10.0	---	97.00	---	---	[35]
E ^{NO}	Batch cell, $A_A = 78 \text{ cm}^2$, $I = 1.2 \text{ A}$, $[\text{Na}_2\text{SO}_4] = 5000 \text{ mg/L}$, $[\text{MET}]_0 = 10 \text{ mg/L}$, and $t = 10 \text{ h}$	SS	BDD	0.6	---	100.00	100.00	---	[36]

Ca.: cathode; An.: anode; O: optimized; NO: not optimized; $V_{treated}$: volume treated; Ref.: reference; E: electro-oxidation; U-UV-E: ultrasonic/UV-assisted electro-oxidation; SS: stainless steel; SPEF: solar photo-electro Fenton; ADE: air-diffusion electrode.

Table 12. Variables assessed in phytotoxicity test.

Variable	Controls			Wastewater Containing MET	
	Tap Water	Distilled Water	Electrolyte	Initial	Final
Germination (%)	100	93.33	93.33	73.33	93.33
Average shoot length (cm)	20.7	24.0	5.7	5.6	6.8
Average root length (cm)	5.1	11.2	10.8	10.4	10.3

Figure 10 shows the growth of *Vigna radiata* when irrigated with different water types. It demonstrates that treated effluent (see Figure 10e) is less toxic than the solution without the electrochemical treatment (see Figure 10d).

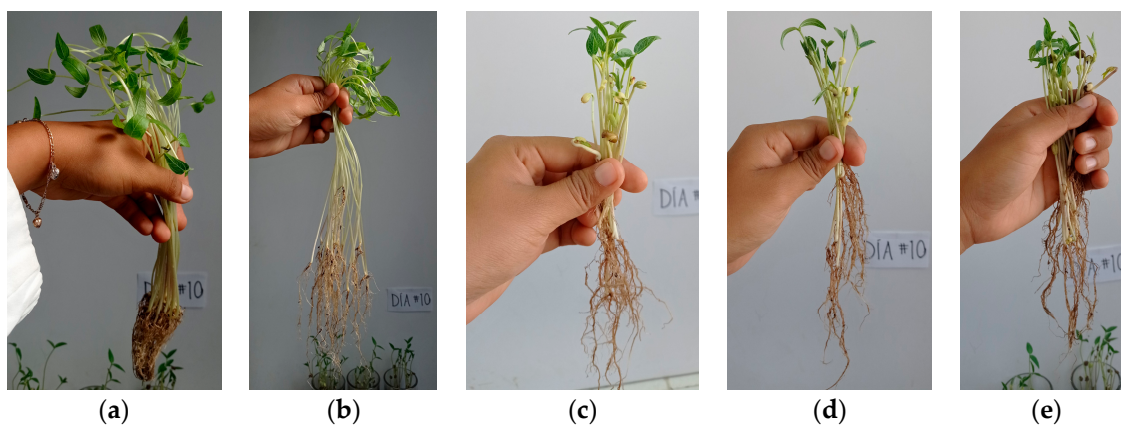


Figure 10. Growth of *Vigna radiata* in (a) tap water, (b) distilled water, (c) electrolyte support (Na_2SO_4 at 0.1 M), (d) solution of MET at 50 mg/L, and (e) wastewater treated using electro-oxidation. The electro-oxidation of the solution of the MET was carried out under the optimal operating conditions: $T = 25\text{ }^\circ\text{C}$, $\text{pH}_0 = 5$, $I = 3.84\text{ A}$, and $Q = 0.8\text{ L/min}$ within a 7.5 h reaction time.

4. Conclusions

The electro-oxidation of MET was successfully driven by $\bullet\text{OH}$ in a flow-by reactor equipped with two Nb/BDD electrodes and running in the batch recirculation mode.

The electrochemical treatment employed in this study effectively removed the MET (95.19%) from the wastewater when the flow-by electrochemical reactor operated under the optimal operating conditions ($\text{pH}_0 = 5$, $I = 3.84\text{ A}$, and $Q = 0.8\text{ L/min}$ within a 7.5 h reaction time).

The maximum COD and TOC removals were 60.8 and 90.1%, respectively, with an associated operating cost of 0.19 USD/L.

The polynomial equations adjusted to the experimental data (COD and TOC removal efficiencies) as functions of pH_0 , I , and Q (A, B, and C, respectively) are as follows:

$$\left(\begin{array}{l} \eta_{\text{COD}} = 51.75 - 6.19A + 1.61B + 5.87C - 4.83AB + 3.5AC - 1.5C^2 - 6.54A^2C + 7.86AB^2 \\ \eta_{\text{TOC}} = 75.71 + 1.29A + 8.22B - 1.55C - 3.31AB + 2.59AC - 1.93BC + 2.52A^2 - 2.84B^2 + 0.15C^2 \end{array} \right)$$

The electro-oxidation of MET, COD removal, and TOC removal follow pseudo first-order, pseudo second-order, and pseudo first-order kinetic models, respectively, when the flow-by electrochemical reactor operates under the optimal operating conditions.

The LC/MS analysis of the byproducts from the electro-degradation of the MET molecule revealed 11 intermediates. It was also confirmed that the process occurs in three phases, which is consistent with previous reports.

The phytotoxicity test indicates that the proposed electrochemical wastewater led to a diminution in the phytotoxicity of the effluent produced because of an increase in the percentage germination of *Vigna radiata*.

Supplementary Materials: The following supporting information can be downloaded at <https://www.mdpi.com/article/10.3390/pr12091958/s1>, Figure S1: Decay of MET over time during the oxidation treatment under the optimal operating conditions. Operating parameters: $[\text{Na}_2\text{SO}_4] = 0.1\text{ M}$; $[\text{MET}]_0 = 50\text{ mg/L}$; $I = 3.84\text{ A}$; $\text{pH}_0 = 5$; $Q = 0.8\text{ L/min}$; $T = 25\text{ }^\circ\text{C}$.

Author Contributions: Conceptualization, A.R.-M., R.N. and C.J.E.; methodology, A.R.-M. and D.V.-V.; software, A.R.-M.; validation, D.V.-V., E.P.-R. and A.R.-M.; formal analysis, A.R.-M. and E.E.R.-G.; investigation, D.V.-V., A.R.-M. and E.P.-R.; resources, A.R.-M., E.P.-R. and E.E.R.-G.; data curation, R.N., C.J.E. and A.R.-M.; writing—original draft preparation, A.R.-M.; writing—review

and editing, R.N., A.R.-M., E.E.R.-G., C.J.E. and E.P.-R.; visualization, A.R.-M., C.J.E. and E.P.-R.; supervision, A.R.-M.; project administration, A.R.-M.; funding acquisition, A.R.-M., E.P.-R. and E.E.R.-G. All authors have read and agreed to the published version of the manuscript.

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